

Supporting Information for

**Sn-Co nano alloy embedded in porous N-doped carbon microboxes as a
stable anode material for lithium-ion batteries**

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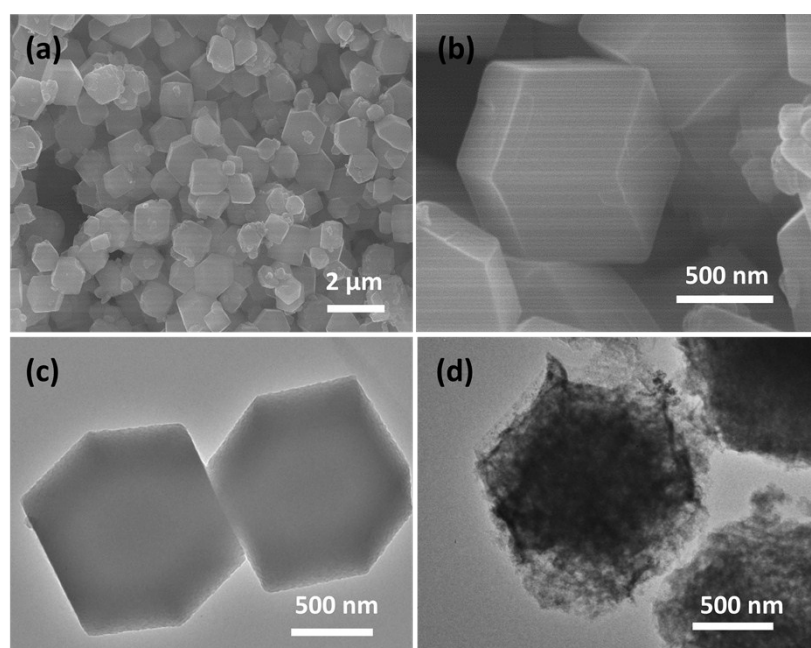


Fig. S1. SEM images of ZIF67 (a) and (b); TEM image of ZIF-67 (c); TEM image of concentrated HCl washed Sn-Co@C-2 –Porous carbon (d).

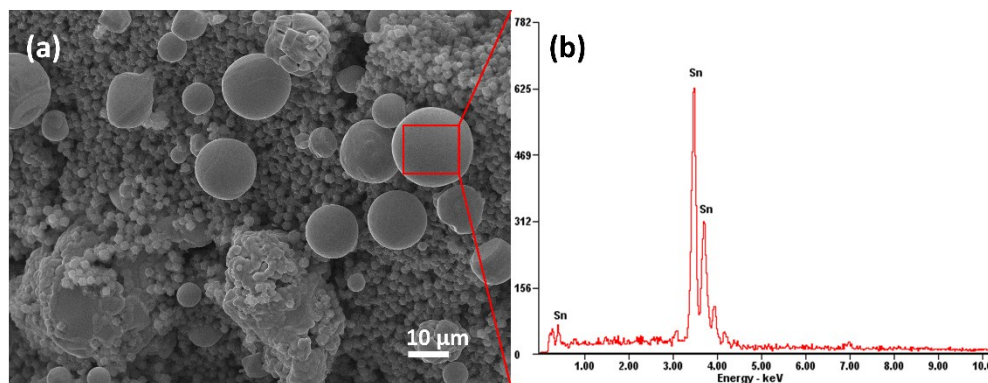


Fig. S2. SEM image (a) and the EDX spectra (b) of the selected area of the sample Sn-

Co@C-650 (carbonization at 650°C).

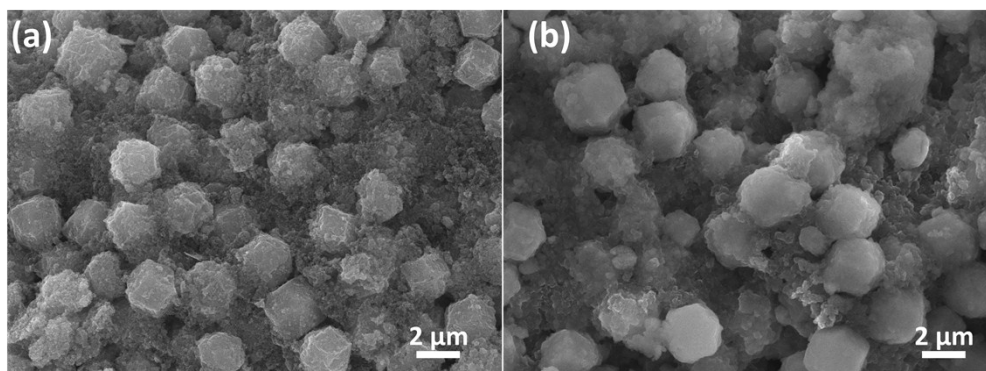


Fig. S3. SEM image of the Sn-Co@C-2 electrode before (a) and after 100 cycles (b).

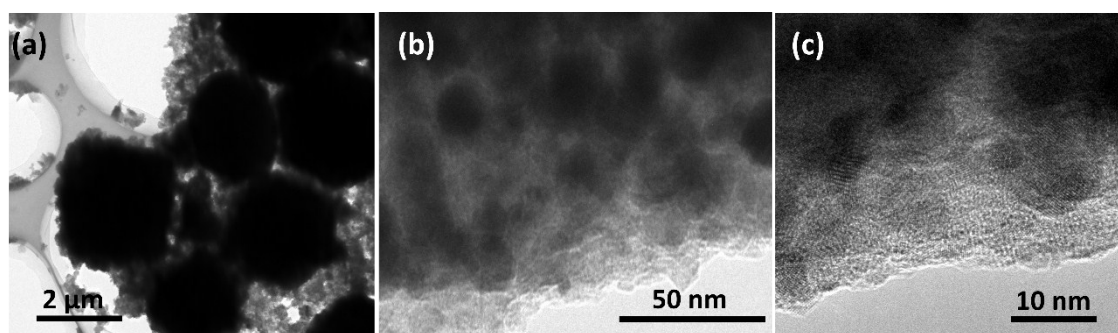


Fig. S1. TEM and HRTEM images of Sn-Co@C-2 after 100 charge/discharge cycles.

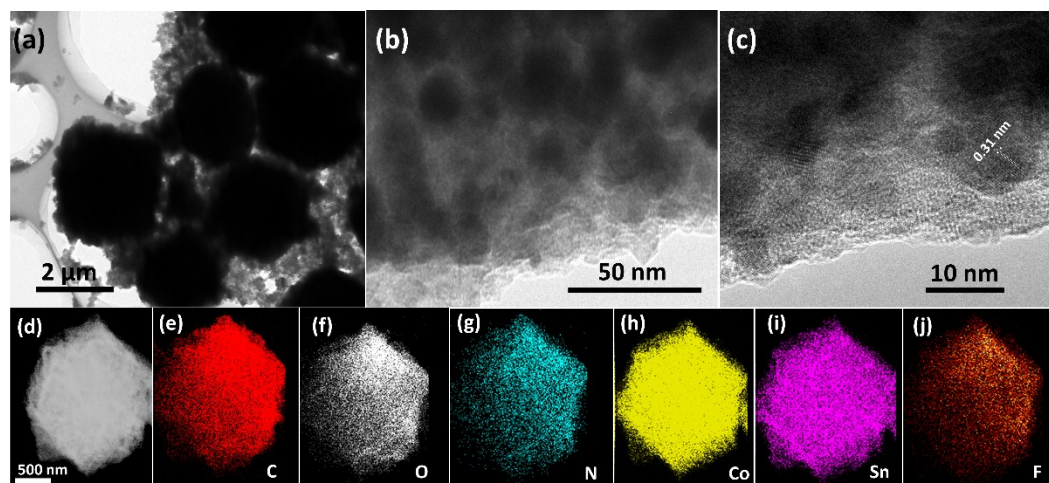


Fig. S4. TEM images (a,b), HR-TEM image (c), STEM image (d), and corresponding elemental mapping images (e-j) of Sn-Co@C-2 after 100 charge-discharge cycles.

Table S1. Electrochemical performances of various Sn-contained anode materials for lithium-ion batteries.

Material	Voltage range (V)	Current density (mA g ⁻¹)	First charge capacity (mAh g ⁻¹)	Capacity after (x) cycles ^a	Capacity at high current density (mAh g ⁻¹) ^b	Reference
Sn-Co@C	0.01-3	100	945	818 (100)	472 (2 A g⁻¹)	This work
Sn-Co@graphene	0.01-3	72	1117	571 (60)	720 (922 mA g ⁻¹)	1
Sn-Co@graphene	0.01-3	500	644	560 (60)	~500 (800 mA g ⁻¹)	2
rod-shaped Cu/SnCo	0.01-2	249	1342	660 (80)	~450 (3976 mA g ⁻¹)	3
SnCo film on nickel	0.02-2	148	~ 900	663 (60)	—	4
porous Sn-Co alloy	0.02-1.5	100	528.7	511 (70)	—	5
SnCo@ordered mesoporous carbon	0.01-3	100	1338	562 (60)	351 (1 A g ⁻¹)	6
SnO ₂ @C	0.01-3	100	1208	880 (200)	450 (1 A g ⁻¹)	7
SnO ₂ @carbon hierarchical tubular structures	0.05-1.5	200	—	700 (50)	505 (1 A g ⁻¹)	8

^aCapacity after a certain cycle number (in brackets)

^bCapacity at a certain current density (in brackets)

Reference

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