

Metal-Organic Frameworks Based on Tetraphenylpyrazine Derived Tetracarboxylic
Acid for Electrocatalytic Hydrogen Evolution Reaction and NACs Sensing

Tao Wang, Kun Huang*, Mengni Peng, Xianglin Li, Defang Han, Linhai Jing, Dabin Qin*

*Key Laboratory of Chemical Synthesis and Pollution Control of Sichuan Province,
School of Chemistry and Chemical Engineering, China West Normal University, Nanchong
637002, P. R. China. E-mail: hkun2017@cwnu.edu.cn; qdbkyl@cwnu.edu.cn*

Content

Table S1. Selected bond lengths (Å) and angles (°) for 1 - 6
Table S2. The corresponding data of 1 - 6 evaluated from nitrogen adsorption
Table S3. Comparison of electrocatalytic activities of 1 - 6
Fig. S1- Fig. S6 Infrared spectrum of 1 - 6
Fig. S7 - Fig. S12 PXRD patterns of 1 - 6
Fig. S13 TGA plots of 1 - 6
Fig. S14 N ₂ adsorption isotherm of 1 - 6 at 77 K
Fig. S15 Solid-state absorption spectra of H ₄ TCPP and 1 - 6
Fig. S16 Solid-state emission spectra of H ₄ TCPP and 1 - 6
Fig. S17 The emission spectra of MOFs 1 - 6 titrated with TNP in water
Fig. S18 The Stern - Volmer plots of the fluorescent quenching of MOFs 1 - 6 with TNP, respectively
Fig. S19 PXRD patterns of 2 immersed in 1M KOH at different time
Fig. S20 PXRD patterns of 5 immersed in 1M KOH at different time
Fig. S21 Electrochemical impedance spectra (EIS) of 1 - 6 and H ₄ TCPP

Compound 1					
Cd1-O1	2.391(4)	Cd1-O2	2.376(4)	N1-C9	1.345(6)
Cd1-O3	2.240(4)	O1-C16	1.261(7)	C2-C3	1.387(7)
O2-C16	1.247(7)	O3-C1	1.253(7)	C1-C2	1.506(6)
O4-C1	1.252(7)	N1-C8	1.335(6)	C7-H7	0.93
O1-Cd1-O1	133.7(2)	O2-Cd1-O1	88.67(1)	C16-O1-Cd1	90.6(4)
O2-Cd1-O2	80.9(2)	O3-Cd1-O1	88.64(1)	C1-O3-Cd1	111.1(3)
Compound 2					
Pb1-O1	2.428(5)	Pb1-O7	2.283(5)	C1-C2	1.509(1)
Pb1-O4	2.436(7)	Pb1-O2	2.547(5)	Pb2-O3	2.324(6)
N1-C32	1.247(7)	O3-C1	1.253(7)	C1-O2	1.260(9)
O4-C1	1.340(9)	N1-C8	1.346(9)	Pb2-O5	2.398(5)
C1-O1-Pb1	133.7(2)	O7-Pb1-O1	76.1(2)	O1-Pb1-O2	137.2(2)
O7-Pb1-O4	87.5(2)	O7-Pb1-O7	78.8(3)	O1-C1-O2	127.0(7)
Compound 3					
La1-O1	2.565(4)	La1-O3	2.638(6)	N1-C8	1.338(4)
La1-O4	2.543(4)	La1-O5	2.673(5)	N2-O26	1.335(4)
La1-O6	2.499(4)	La1-O7	2.559(5)	N1-C9	1.329(4)
O3-C19	1.255(4)	O4-C19	1.274(5)	N2-O27	1.337(4)
O1-La1-O3	70.83(1)	O1-La1-O5	142.76(7)	O4-La1-O3	50.27(9)
O3-La1-O5	139.57(1)	O1-La1-O4	82.63(1)	O5-La1-O7	118.02(9)
Compound 4					
Sr1-O1	2.537(2)	Sr1-O3	2.717(2)	Sr2-O7	3.005(3)
Sr1-O10	2.508(2)	Sr1-O87	2.547(3)	Sr2-O12	2.651(2)
Sr1-O88	2.709(2)	Sr1-O89	2.647(2)	Sr2-O10	2.648(2)
Sr2-O3	2.539(2)	Sr2-O4	2.518(2)	Sr2-O18	2.689(4)
O1-Sr1-Sr2	78.46(4)	O1-Sr1-O2	24.29(7)	O1-Sr1-O89	70.96(7)
O1-Sr1-O3	48.60(6)	O1-Sr1-O88	128.63(7)	O10-Sr1-O89	79.21(7)
Compound 5					
Ce1-O1	2.518(2)	Ce1-O2	2.606(2)	N1-C6	1.335(4)
Ce1-O3	2.518(2)	Ce1-O4	2.657(2)	C12-O3	1.218(5)
Ce1-O5	2.541(2)	Ce1-O6	2.495(19)	N1-C7	1.343(4)
Ce1-O7	2.522(19)	Ce1-O9	2.562(2)	C12-O2	1.314(5)
O3-Ce1-O1	76.60(8)	O4-Ce1-O1	132.61(7)	O5-Ce1-O4	50.06(7)
O4-Ce1-O3	147.95(8)	O5-Ce1-O2	122.62(8)	O6-Ce1-O1	79.55(7)
Compound 6					

Mn4-O5	2.319(3)	Mn4-O4	2.027(4)	Mn3-O9	2.201(4)
Mn4-O5	2.219(3)	Mn4-O4	2.027(4)	Mn3-O9	2.201(4)
Mn4-O6	2.217(3)	Mn3-O3	2.143(4)	Mn3-O3	2.143(4)
Mn4-O6	2.217(3)	Mn3-O5	2.208(4)	Mn3-O5	2.208(4)
O5-Mn4 -O5	96.13(17)	O4-Mn4-	97.12(15)	O5-Mn3-O5	85.83(19)
O6-Mn4 -O5	86.79(13)	O4-Mn4-	127.24(16)	O3-Mn3-O3	162.4(2)

Symmetry code: 1. Cd = x, y, z; O2 = 1-x, 0.25+y, 0.25+z.; 2. Pb1= x, y, z; O4= -x, 1-y, 1-z.; 3. La1= x, y, z; O5= x, y, z; 4. Sr1= x, 1+ y, z; Sr2=1- x, 1-y, 1-z; 5. Cel = x, y, z; O1 = x, y, 1 + z; 6. Mn3= x, y, z; Mn4= x, y, z.

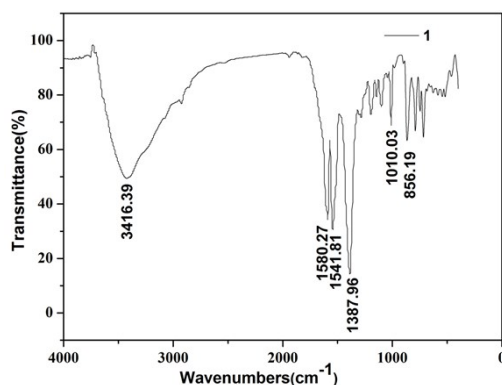


Fig. S1 Infrared spectrum of 1.

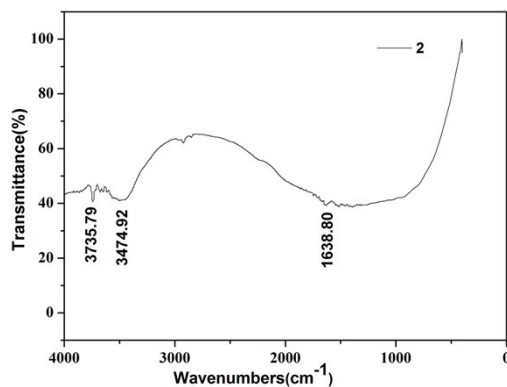


Fig. S2 Infrared spectrum of 2.

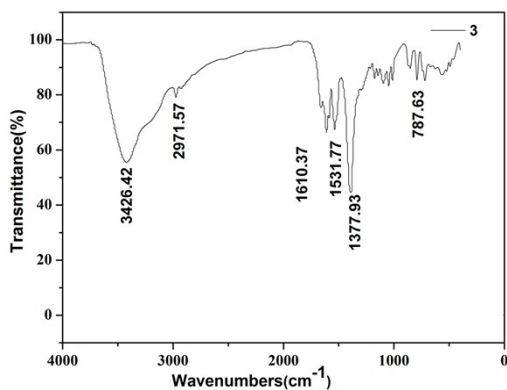


Fig. S3 Infrared spectrum of 3.

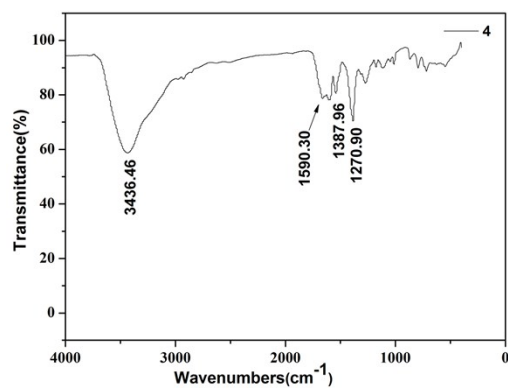


Fig. S4 Infrared spectrum of 4.

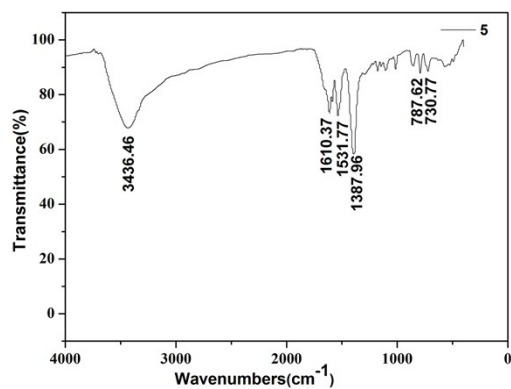


Fig. S5 Infrared spectrum of 5.

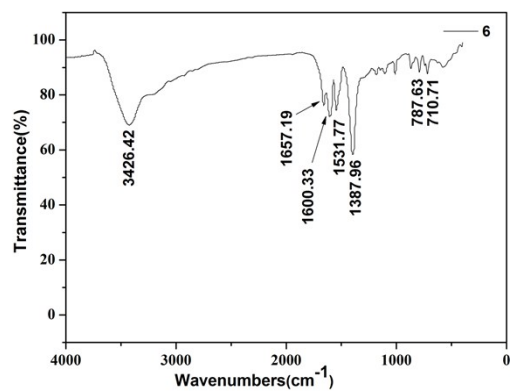


Fig. S6 Infrared spectrum of 6.

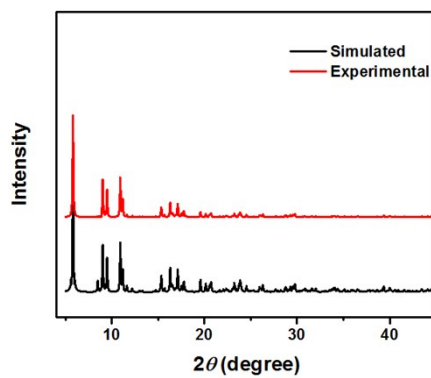


Fig. S7 PXRD patterns of 1

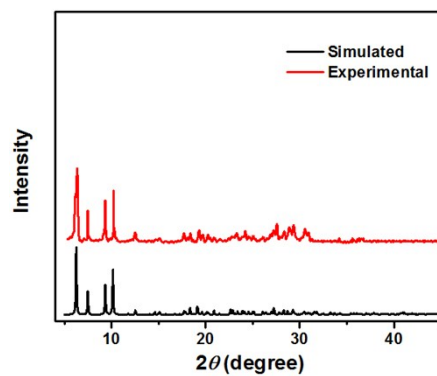


Fig. S8 PXRD patterns of 2.

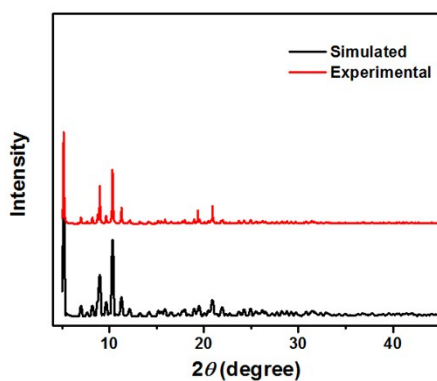


Fig. S9 PXRD patterns of 3.

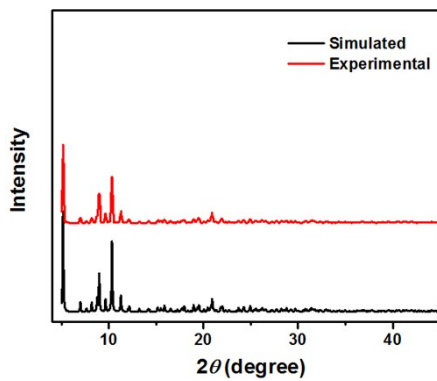


Fig. S10 PXRD patterns of 4.

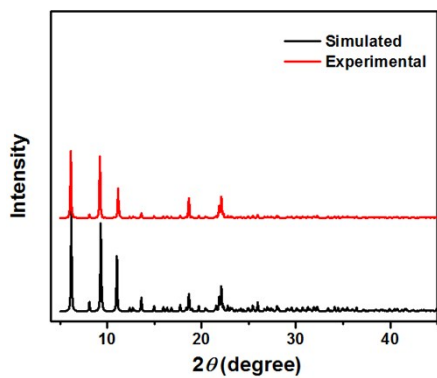


Fig. S11 PXRD patterns of 5.

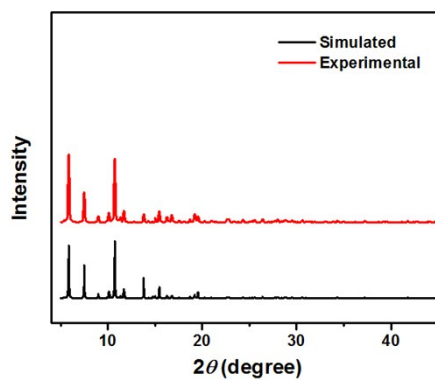


Fig. S12 PXRD patterns of 6.

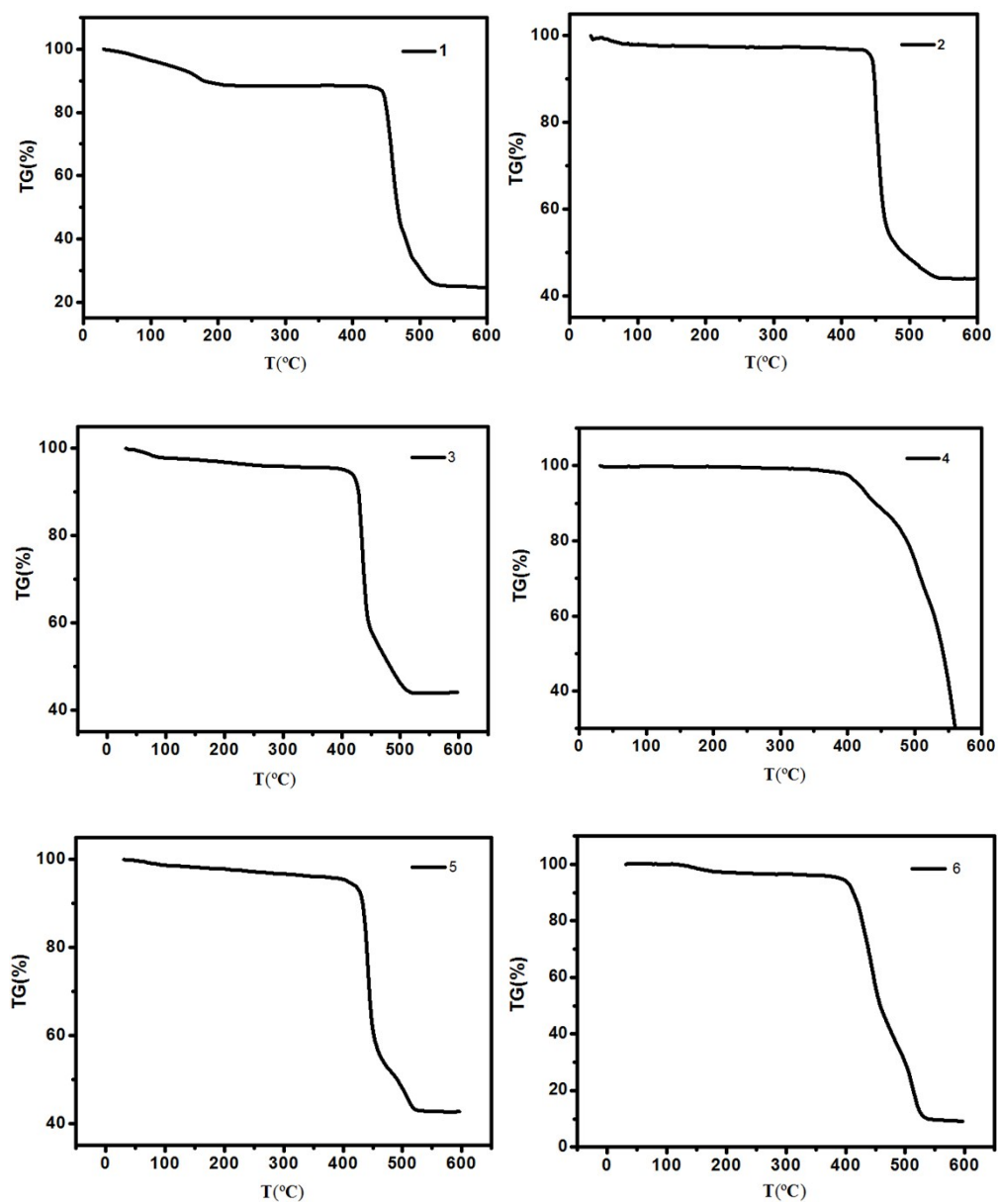


Fig. S13 TGA plots of 1 - 6.

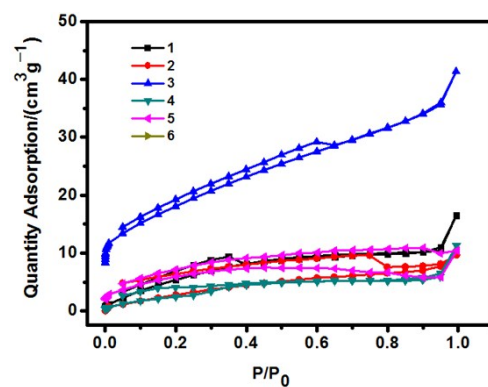


Fig. S14 N₂ adsorption isotherm of 1 - 6 at 77 K.

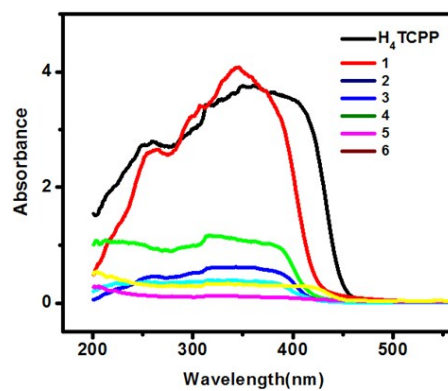


Fig. S15 Solid-state absorption spectra of H₄TCPP and 1 – 6.

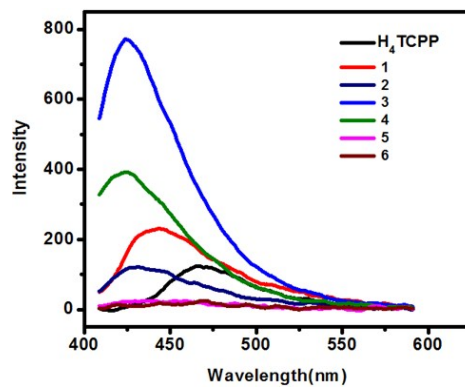


Fig. S16 Solid-state emission spectra of H₄TCPP and 1 - 6.

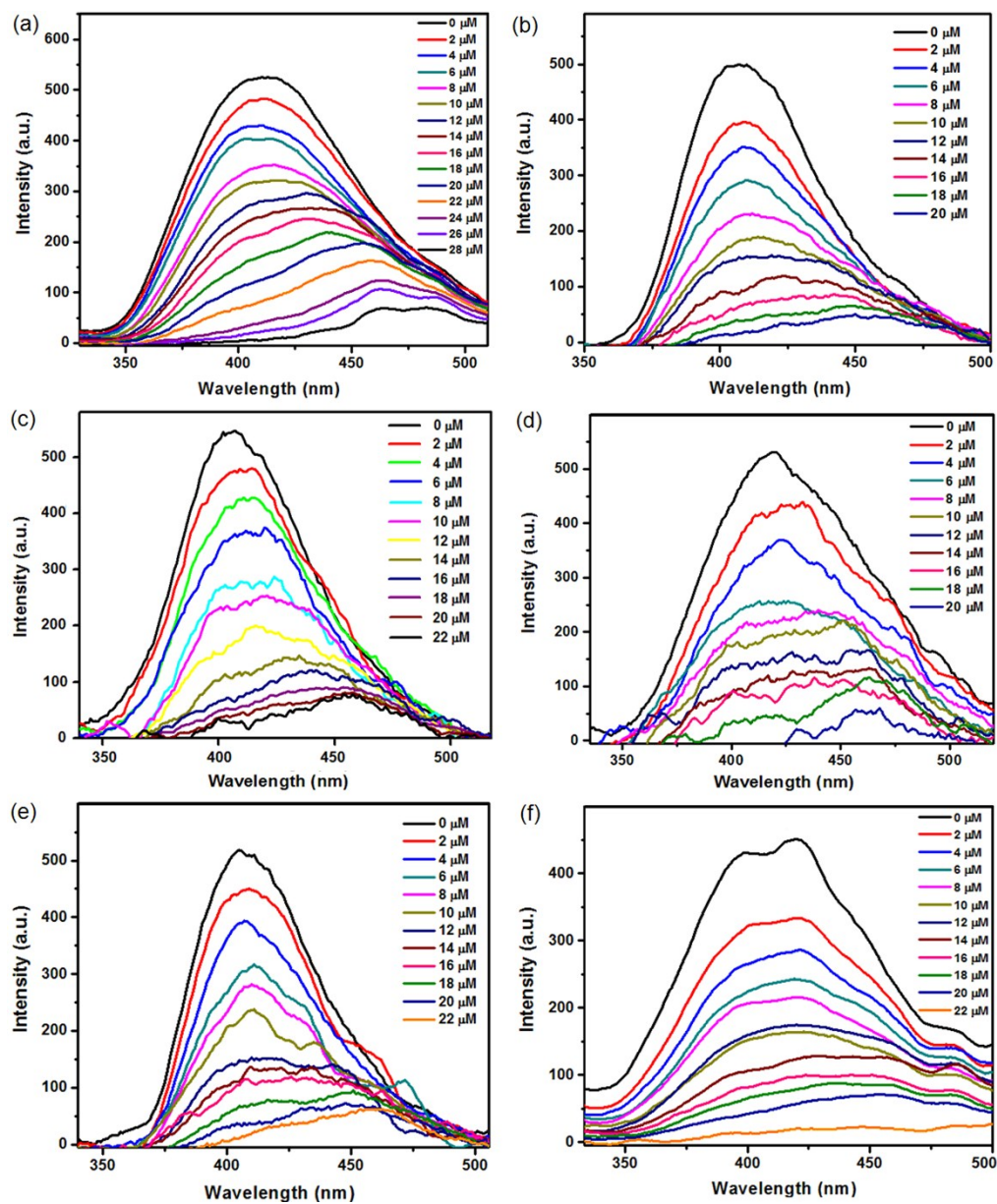


Fig. S17 The emission spectra of MOFs 1 - 6 (0.5 mg in 2 mL H₂O) titrated with TNP in water.

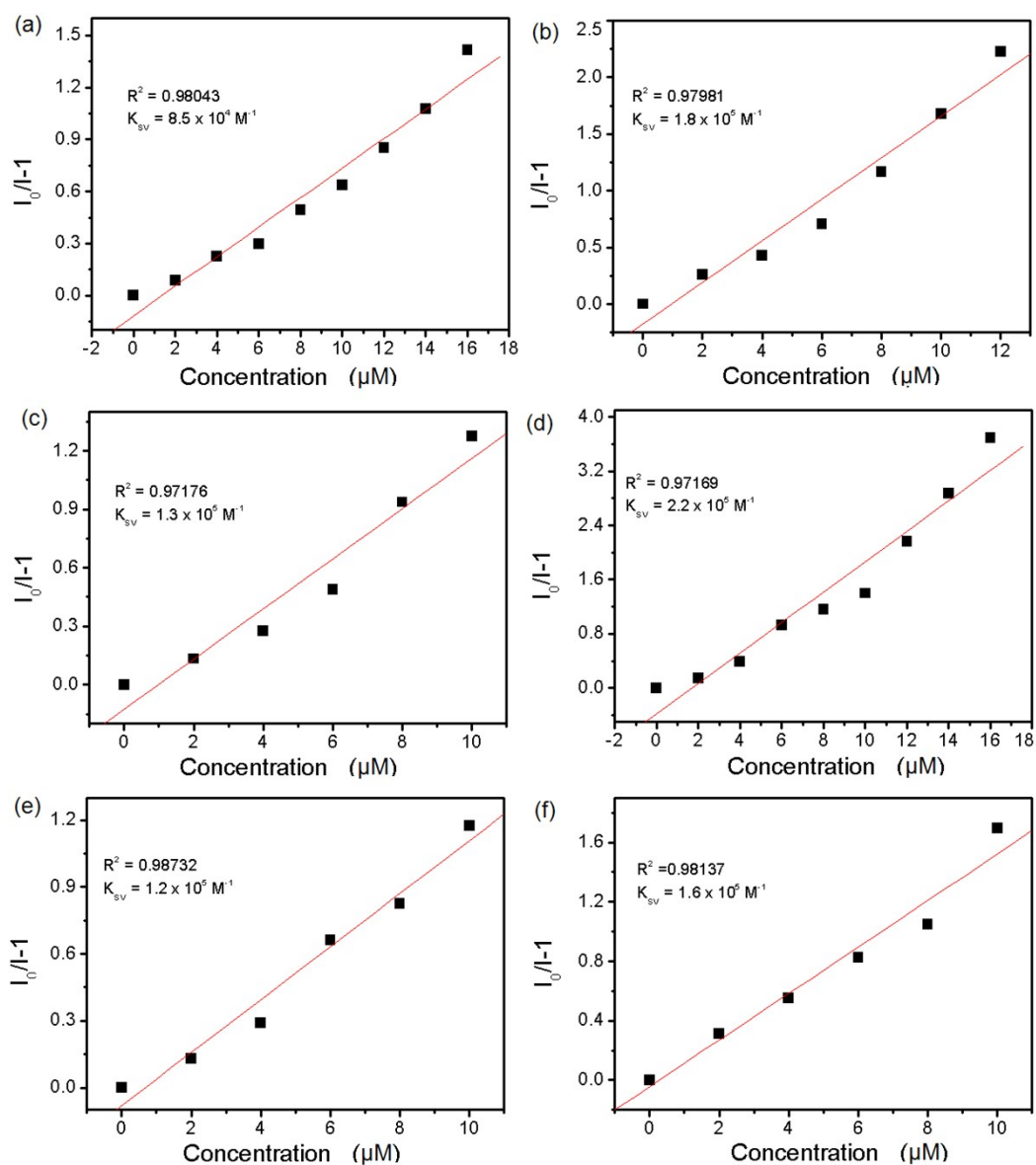


Fig. S18 Stern - Volmer plots of the fluorescent quenching of MOFs 1 - 6 with TNP, respectively. The solid lines represent fit to the concentration-resolved data using the Stern - Volmer equation. The fluorescence intensity was monitored at 432 nm.

Table S2. The corresponding data of 1 – 6 evaluated from nitrogen adsorption.

Complex	BET(m^2/g)	Pore volumes (cm^3/g)	pore size(μm)
1	4.493	0.0255	3.819
2	4.193	0.01491	3.408
3	18.325	0.06396	3.406
4	1.675	0.01745	3.405
5	3.642	0.01612	4.304
6	31.128	0.07064	3.405

Table S3. Comparison of electrocatalytic activities of 1 – 6.

catalyst	dimensionality	η_{10}^* (mV)	V_{onset} (mV)	Solution (1M)	Tafel Slope
H ₄ TCPP	2D	486	380	KOH	435
1	3D	423	348	KOH	400
2	3D	449	362	KOH	253
3	3D	545	336	KOH	485
4	3D	487	319	KOH	362
5	3D	457	244	KOH	273
6	3D	452	350	KOH	360

* η_{10} is defined as overpotential measured for a current of 10 mA cm⁻².

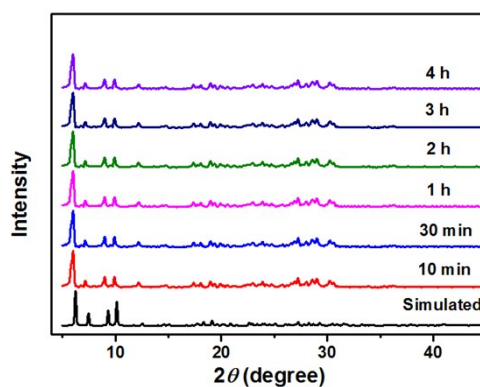


Fig. S19 PXRD patterns of 2 immersed in 1M KOH at different time.

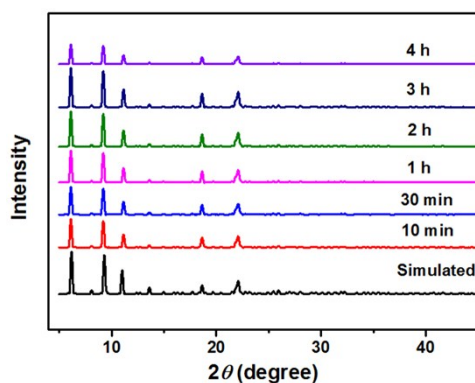


Fig. S20 PXRD patterns of 5 immersed in 1M KOH at different time.

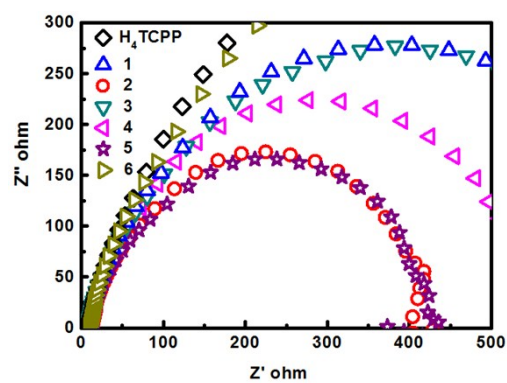


Fig. S21 Electrochemical impedance spectra (EIS) of 1 - 6 and H₄TCPP