

Supporting Information

**A dual-function photoelectrochemical solar cell which assimilates light-harvesting, charge-transport and photoelectrochromic nanomaterials in a tandem design**

Ankita Kolay,<sup>a</sup> Nathan Potts,<sup>b</sup> Kripasindhu Sardar,<sup>b</sup> Elizabeth A. Gibson,<sup>b</sup> and Melepurath Deepa<sup>a,\*</sup>

<sup>a</sup>Department of Chemistry, Indian Institute of Technology Hyderabad, Kandi, Sangareddy, Telangana 502285, India. E-mail: mdeepa@iith.ac.in

<sup>b</sup>Energy Materials Laboratory, Chemistry, School of Natural and Environmental Science, Newcastle University, Newcastle upon Tyne NE1 7RU, United Kingdom.

Brunauer-Emmett-Teller (BET) surface area analysis and Barrett-Joyner-Halenda (BJH) pore size and volume analysis were performed under Nitrogen at 77.3 K, after degassing at 300 °C for 3 h. This analysis was performed on a Nova 2200e Quantachrome instrument.

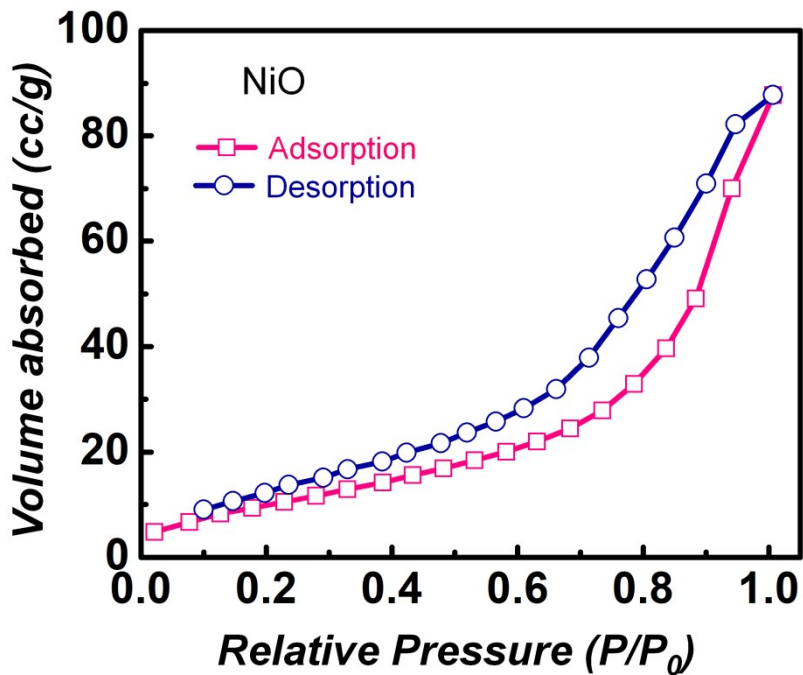


Figure S1 Nitrogen adsorption-desorption isotherm for NiO.

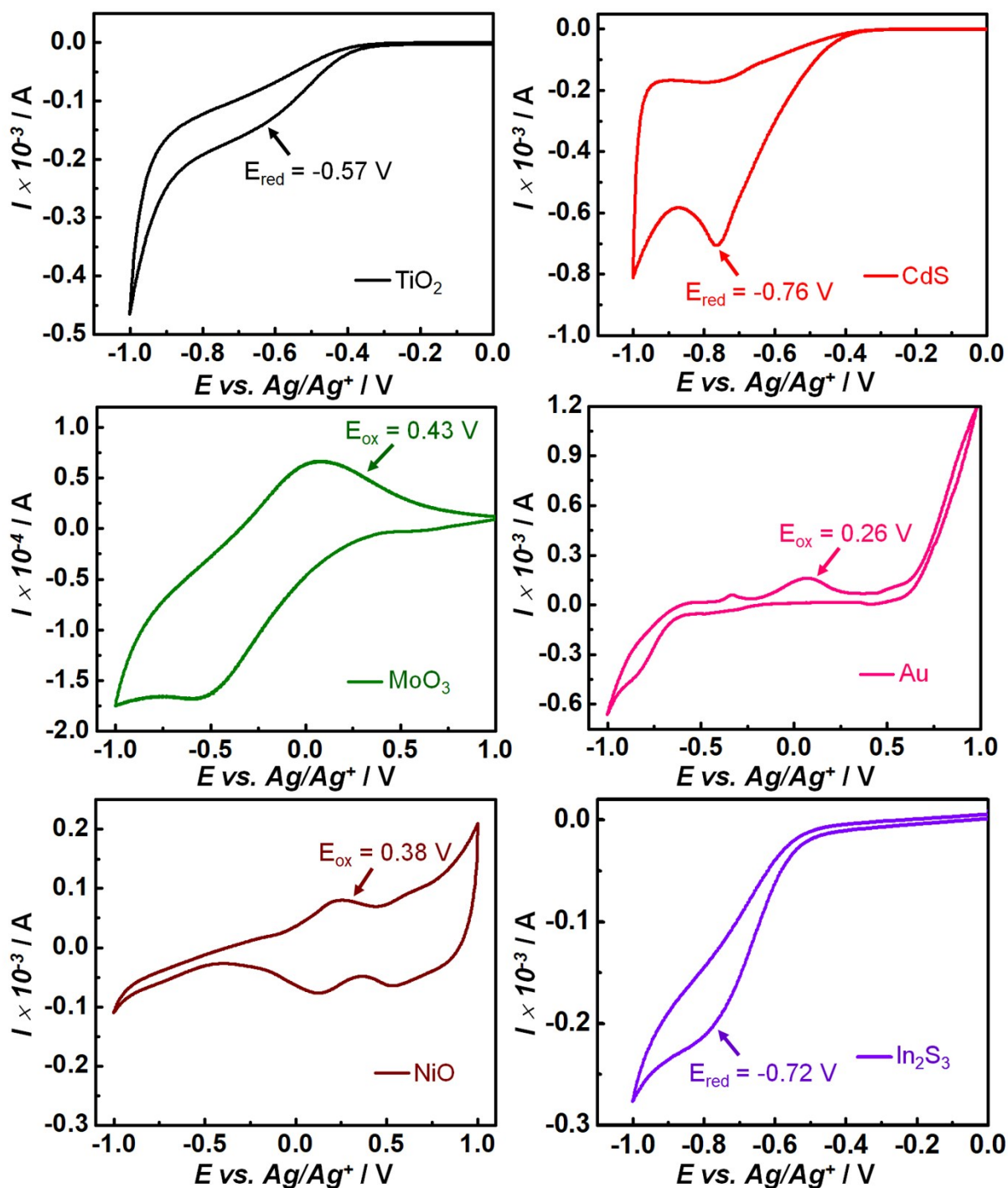


Figure S2 Cyclic voltammograms of TiO<sub>2</sub>, CdS, MoO<sub>3</sub>, Au, NiO and In<sub>2</sub>S<sub>3</sub> films deposited over FTO substrates. All plots were recorded in a 0.1 M KCl solution as electrolyte, with a Pt rod as the counter electrode and Ag/AgCl/KCl as the reference electrode, at a scan rate of 20 mVs<sup>-1</sup>.

The HOMO (VB), LUMO (CB) and band gap values of different materials used in PECVD fabrication were acquired from cyclic voltammograms and absorbance spectra using the equations provided below.

$$E_{\text{red}} = -4.5 \text{ eV} (\equiv 0 \text{ V versus NHE}) - (\text{Red. Peak (V) vs. Ag/AgCl/KCl} + 0.197 \text{ V}) \quad (1)$$

$$E_{\text{ox}} = -4.5 \text{ eV} (\equiv 0 \text{ V versus NHE}) - (\text{Ox. Peak (V) vs. Ag/AgCl/KCl} + 0.197 \text{ V}) \quad (2)$$

Table S1 Energy level positions of the photoanode and photocathode components.

Material	Reduction peak / V vs. Ag/AgCl	Oxidation peak / V vs. Ag/AgCl	$E_{\text{red}}$ (versus NHE) / eV $\equiv$ LUMO	Band Gap / eV	$E_{\text{ox}}$ (versus NHE) / eV $\equiv$ HOMO
TiO <sub>2</sub>	-0.571	---	-4.126	3.16	-7.286
CdS	-0.758	---	-3.939	2.29	-6.229
NiO	---	0.377	1.484	3.59	-5.074
In <sub>2</sub> S <sub>3</sub>	-0.723	---	-3.974	2.22	-6.194

Material	Peak / V vs. Ag/AgCl	E (versus NHE) / eV
Au	0.257	-4.954
MoO <sub>3</sub>	0.434	-5.131

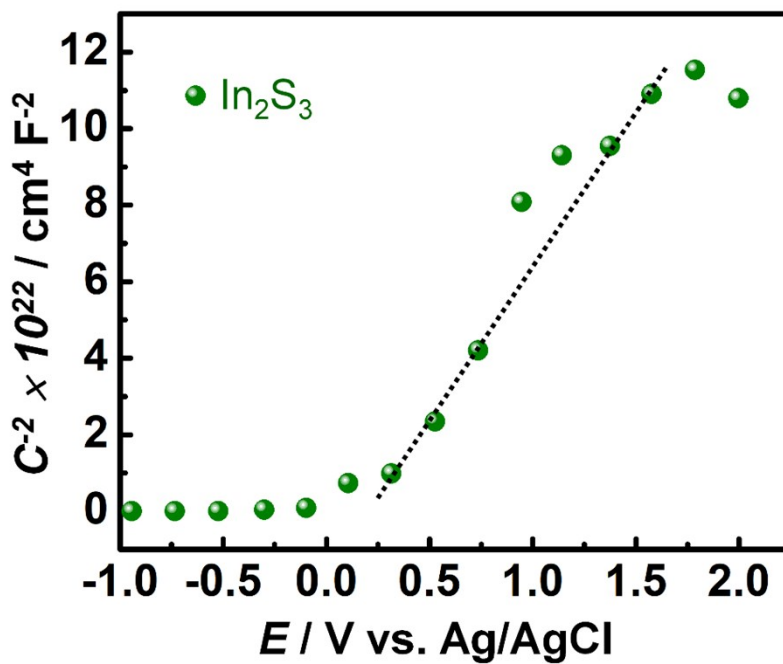


Figure S3 Mott-Schottky plot of the  $\text{In}_2\text{S}_3$  film.

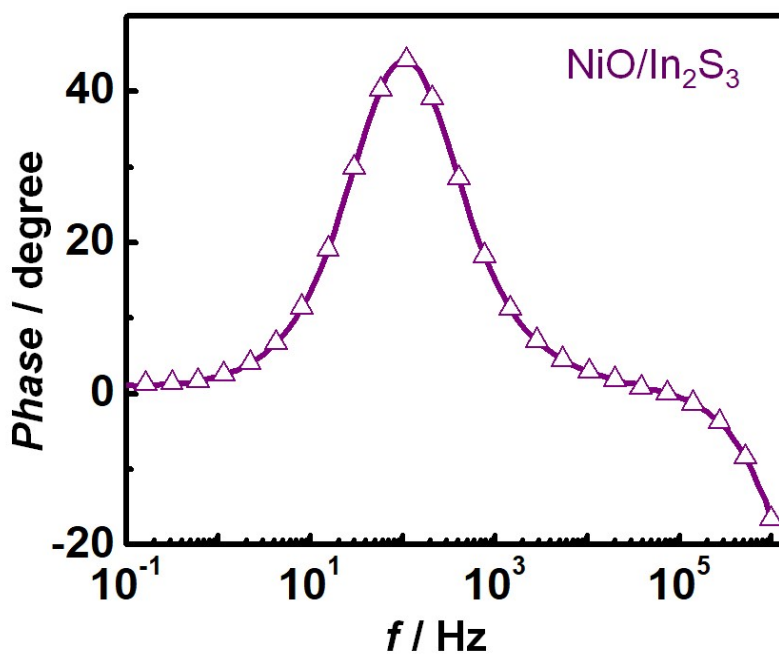


Figure S4 Bode plot of the  $\text{NiO}/\text{In}_2\text{S}_3\text{-nS}_2^-/\text{S}_n^{2-}$  C-fabric half-cell, with  $\text{NiO}/\text{In}_2\text{S}_3$  as the photoanode.

Table S2 Solar cell parameters of the 3-cells in a 1 M polysulfide gel electrolyte, exposed cell area: 0.12-0.15 cm<sup>2</sup>, under 1 sun illumination (AM 1.5, 100 mW cm<sup>-2</sup>).

Cells	V <sub>oc</sub> (V)	J <sub>sc</sub> (mA cm <sup>-2</sup> )	FF (%)	Efficiency (η %)
<b>TiO<sub>2</sub>/CdS/Au/MoO<sub>3</sub> -In<sub>2</sub>S<sub>3</sub>/NiO/C-fabric</b>				
Cell 1	0.785	17.99	56.54	7.987
Cell 2	0.787	17.67	52.60	7.314
Cell 3	0.783	17.32	51.63	7.002
<b>Average</b>	0.785	17.66	53.59	7.434

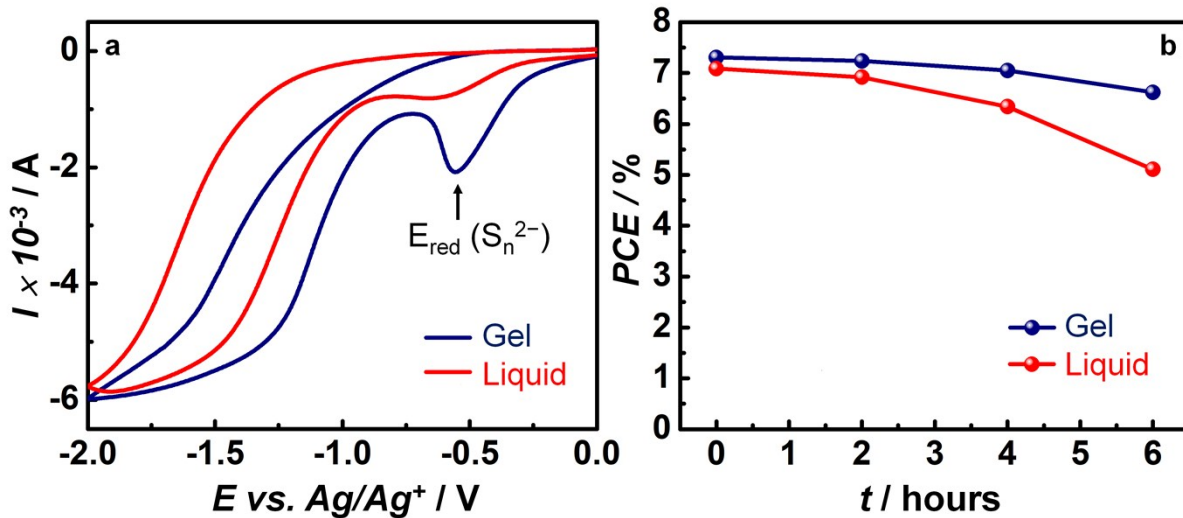


Figure S5 (a) Cyclic voltammograms of aqueous 1 M Na<sub>2</sub>S/1 M S/SiO<sub>2</sub> gel and the same liquid electrolyte without SiO<sub>2</sub>, recorded between two Pt electrodes at a scan rate of 20 mV s<sup>-1</sup>. (b) Variation of PCE as a function of exposure time to 1 sun (100 mW cm<sup>-2</sup>) for tandem cells with liquid and gel electrolytes.

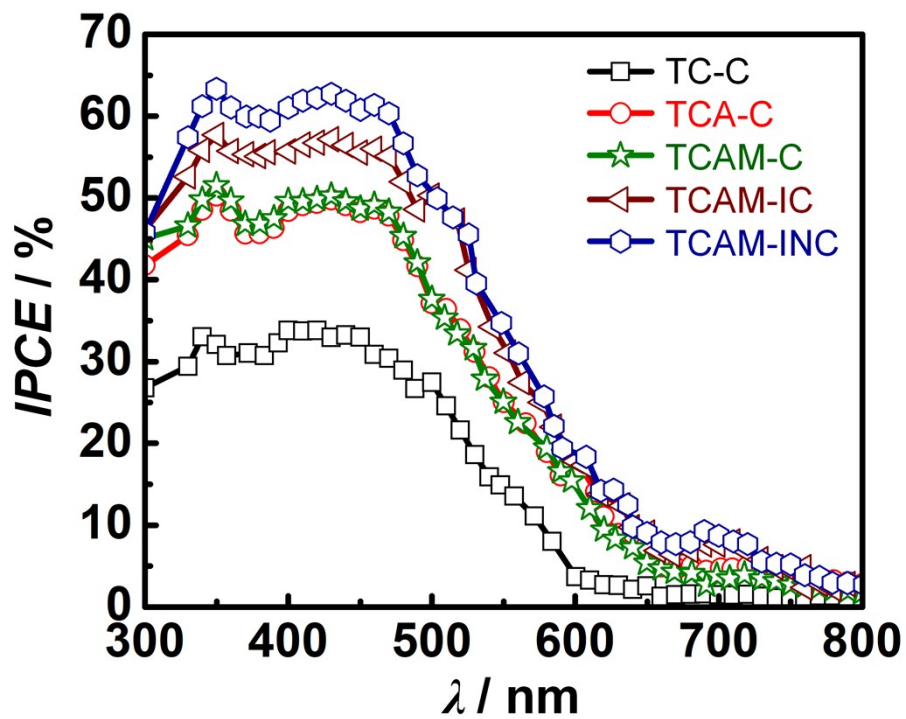


Figure S6 IPCE versus wavelength curves for  $\text{TiO}_2/\text{CdS}$  (TC),  $\text{TiO}_2/\text{CdS}/\text{Au}$  (TCA) and  $\text{TiO}_2/\text{CdS}/\text{Au}/\text{MoO}_3$  (TCAM) photoanode-based cells each with three different counter electrodes: C-fabric (C),  $\text{In}_2\text{S}_3/\text{C}$ -fabric (IC), and  $\text{In}_2\text{S}_3/\text{NiO}/\text{C}$ -fabric (INC) (polysulfide gel electrolyte, under 1 sun illumination, AM 1.5).