Tuning the emission of a water-soluble 3-hydroxyflavone derivative by host-guest complexation

Supporting Information

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Figure S1 (a) $^1$H NMR spectrum of 3HF-EO in CDCl$_3$ and (b) Q-TOF MS of 3HF-EO in methanol.

Figure S2 The PL spectra of 3HF-EO in the presence of 1, 5 and 8 equivalents of (a) $\alpha$-CD, (b) $\beta$-CD or (c) $\gamma$-CD in water.

Figure S3 (a) The NMR titration of 3HF-EO upon addition of $\beta$-CD. The total concentration of 3HF-EO and $\beta$-CD were kept constant, and the molar ratio of 3HF-EO was varied. (b) The aromatic region of (a). The chemical shifts of proton H-e on 3HF-EO were applied to draw the Job’s Plot. $R$ is the molar ratio of 3HF-EO in 3HF-EO and $\beta$-CD. $\Delta\delta = \delta$(3HF-EO/$\beta$-CD) - $\delta$(3HF-EO).
Figure S4 (a) The NMR titration of 3HF-EO upon addition of γ-CD. The total concentration of 3HF-EO and γ-CD were kept constant, and the molar ratio of 3HF-EO was varied. (b) The aromatic region of (a). The chemical shifts of proton H-e on 3HF-EO were applied to draw the Job’s Plot. $R$ is the molar ratio of 3HF-EO in 3HF-EO and γ-CD. $\Delta \delta = \delta (3HF-EO/\gamma-CD) - \delta (3HF-EO)$. 