A novel fluorescent-colorimetric probe for Al³⁺ and Zn²⁺ ions with different response and applications in F⁻ detection and cell imaging Jiaxin Fu¹, Yongxin Chang¹¹, Bai Li¹, Huihui Mei¹, Li Yang¹, Kuoxi Xu^{1,2}

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Fig. s2 ¹³C NMR of HL



Fig. s3 HRMS of HL



Fig. s4 FT-IR of HL



Fig. s5. The detection limit (DL) of probe **HL** for Zn^{2+} in EtOH-H₂O solution (v/v=4/1, 0.01 M, Tris-HCl buffer, pH = 7.30).



Fig. s6. The detection limit (DL) of probe **HL** for Al^{3+} in EtOH-H₂O solution (v/v=4/1, 0.01 M, Tris-HCl buffer, pH = 7.30).



Fig. s7. Metal-ions selectivity of **HL**; The blue and red bars represent in the fluorescence intensity of HL and metal ions. The green bars show the fluorescence intensity after the addition of 2.0 equiv. of (A) Zn^{2+} or (B) Al^{3+} to the solution containing **HL** and different 2.0 equiv. of other metal ions.



Fig. s8. The HRMS HL+Al(NO₃)₃ (1.0 equiv.)+NaF (3.0 equiv.).



Fig. s9. Recovery experiment of HL for Zn²⁺ and Al³⁺



Fig. s10. Job's plot for determining the stoichiometry of probe HL and (A) Zn^{2+} , (B) Al^{3+} in EtOH-H₂O solution (v/v=4/1, 0.01 M, Tris-HCl buffer, pH=7.30) the total concentration of HL and Zn^{2+} or Al^{3+} was 1.00×10^{-4} M.



Fig. s11. The HRMR of $[HL+Zn^{2+}-H]^+$.



Fig. s12. The HRMR of $[HL+Al^{3+}+2H_2O-H]^+$.



Fig. s13. Comparison between the FT-IR spectral data for HL-Zn²⁺ complex



Fig. s14. Comparison between the FT-IR spectral data for HL-Al³⁺ complex



Fig. s15. Dependence of fluorescence response of HL towards $Al^{3\scriptscriptstyle +}$ and $Zn^{2\scriptscriptstyle +}$ over variable pH medium



Fig. s16. The cell viability of PC12 at different concentration (0, 30, 60, 90 120, 150 μ M) of HL assessed through MTT assay