Supplementary Information

Effect of Ligand Substituents on Nickel and Copper [N₄] Complexes: Electronic and Redox Behavior, and Reactivity towards Protons

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Contents

1. Electronic Spectrum for Complex 1-5

Figure S1. UV-visible absorption spectra of 1 in MeOH

Figure S2. UV-visible absorption spectra of 2 in MeOH

Figure S3. UV-visible absorption spectra of 3 in MeCN

Figure S4. UV-visible absorption spectra of 4 inMeOH

Figure S5. UV-visible absorption spectra of 5 inMeCN

- Electrochemical Data for Complexes 1-5
 Figure S6. Cyclic voltammograms of 1-5 in MeCN
- 3. Acid Titration Experiment

Figure S7. Cyclic voltammograms of **1-5** in MeCN, TBAPF₆ (0.1 M) with 0-10 equiv. of acetic acid

Figure S8. Cyclic voltammograms of **1-5** in MeCN with and without complexes upon addition of four equivalent of acid

4. Onset Overpotential Experiment

Figure S9.Charge vs. overpotential plot and charge accumulation (3 min.) vs. time plot for 1
Figure S10. Charge vs. overpotential plot and charge accumulation (3 min.) vs. time plot for 2
Figure S11. Charge vs. overpotential plot and charge accumulation (3 min.) vs. time plot for 3
Figure S12. Charge vs. overpotential plot and charge accumulation (3 min.) vs. time plot for 4
Figure S13. Charge vs. overpotential plot and charge accumulation (3 min.) vs. time plot for 5

- Pre and post catalysis UV-Vis spectra
 Figure S14. UV-Visible spectra for complex 1-5 before and after catalysis.
- DFT calculated energetics, structures, and spin density plots
 Figure S15. DFT calculated mechanistic pathway for the decomposition of Cu catalyst 3 in water.
 Figure S16. DFT calculated optimized geometries of the intermediates and catalytic complexes
 Figure S17. DFT calculated spin density plots of the intermediates and catalytic complexes
- 7. Table S1: Comparison of DFT and crystallographic structures of catalyst 3
- 8. Table S2: Comparison of DFT and crystallographic structures of catalyst 2
- 9. Table S3: Comparison of DFT-calculated structures of 3 and 5
- 10. Table S4: Comparison of DFT-calculated structures of 2 and 4
- 11. **Table S5:** Coordinates of the DFT-calculated geometries





Figure S1. UV-visible absorption spectra of 1 in MeOH.





Figure S2. UV-visible absorption spectra of 2 in MeOH.



Figure S3. UV-visible absorption spectra of 3 in MeCN.







Figure S4. UV-visible absorption spectra of 4 in MeOH.





Figure S5. UV-visible absorption spectra of 5 in MeCN.



Figure S6. Left: Cyclic voltammograms in acetonitrile for complexes 1 (black), 2 (red) and 4 (blue). Right: Cyclic voltammograms in acetonitrile for complexes 3 (orange) and 5 (green).



Figure S7. Cyclic voltammograms of **1-5** in MeCN, TBAPF₆ (0.1 M) with 0-10 equiv. of acetic acid. (a) Complex **1**, (b) complex **2**, (c) complex **4**, (d) complex **3**, (e) complex **5**. Electrodes: C_{glassy} (W), Ag/AgCl (R), Pt wire (A). The spike observed at -590mV for complexes **3** and **5** is attributed to Cu¹/Cu⁰ couple.¹

¹ M. Kugler, J. Scholz, A. Kronz, I. Siewert, *Dalton Trans.* 2016, **45**(16), 6974-6982. D. M. Ekanayake, K. M. Kulesa, J. Singh, K. K. Kpogo, S. Mazumder, H. B. Schlegel, C. N. Verani, *Dalton Trans.* 2017, **46**, 16812-16820.



Figure S8. Cyclic Voltammograms of **1-5** in MeCN with and without complexes upon addition of four equivalent of acid. (a) Complex **1**, (b) complex **2**, (c) complex **4**, (d) complex **3**, (e) complex **5**. Electrodes: C_{glassy} (W), Ag/AgCl (R), Pt wire (A).



Figure S9. (a) Charge vs. overpotential plot at various overpotentials over 3 min intervals. (b) Charge accumulation (3 min.) vs. time plot; 0.1 µmol of 1; Electrodes: Hg pool (W), Ag/AgCl (R), Pt wire (A). Onset Overpotential = 78 mV.



Figure S10. (a) Charge *vs.* overpotential plot at various overpotentials over 3 min intervals. (b) Charge accumulation (3 min.) *vs.* time plot; 0.1 µmol of **2**; Electrodes: Hg pool (W), Ag/AgCl (R), Pt wire (A). Onset Overpotential = 833 mV.



Figure S11. (a) Charge *vs.* overpotential plot at various overpotentials over 3 min intervals. (b) Charge accumulation (3 min.) *vs.* time plot; 0.1 µmol of **3**; Electrodes: Hg pool (W), Ag/AgCl (R), Pt wire (A). Onset overpotential = 883mV.



Figure S12. (a) Charge *vs.* overpotential plot at various overpotentials over 3 min intervals. (b) Charge accumulation (3 min.) *vs.* time plot; 0.1 µmol of **4** Electrodes: Hg pool (W), Ag/AgCl (R), Pt wire (A). Onset Overpotential = 983 mV.



Figure S13. (a) Charge *vs.* overpotential plot at various overpotentials over 3 min intervals. (b) Charge accumulation (3 min.) *vs.* time plot; 0.1 µmol of **5**; Electrodes: Hg pool (W), Ag/AgCl (R), Pt wire (A). Onset Overpotential = 933 mV.



Figure S14. UV-Visible spectra for complex 1-5 before and after catalysis: (a) Complex 1, (b) complex 2, (c) complex 4, (d) complex 3, (e) complex 5.



Figure S15. DFT calculated mechanistic pathway for the decomposition of Cu catalyst **3** in water. The formation of 3B has been calculated from four monomeric water molecules. The ΔG for this process is calculated to be +20.57 kcal/mol. However, in solution, it is expected that the four water molecules will not be monomeric. Rather they are known to form a single tetramer complex, (H₂O)₄. Therefore, it is more appropriate that the formation of 3B should be considered from a single tetrameric complex, (H₂O)₄, rather than four monomeric water molecules, 4H₂O. When considered the formation of 3B from (H₂O)₄, the process will not be considerably affected by entropy. Therefore, the DFT-calculated ΔG (= +20.57 kcal/mol) is only an upper limit for the formation of 3B and, in actual case, ΔG will be considerably less than this value.



Figure S16. DFT calculated optimized geometries of the intermediates and catalytic complexes.



Figure S17. DFT calculated spin density plots (isodensity = 0.004 au) with Mulliken Spin Density (MSD) values on the metal atoms of the intermediates and catalytic complexes.

Bond	Bond Distance (Å) DFT	Bond Distance (Å) Crystal	Difference (Å)
Cu-N1	2.041	2.000	0.041
Cu-N2	2.063	2.028	0.035
Cu-N3	2.059	2.031	0.028
Cu-N4	2.060	2.000	0.060

 Table S1. Comparison of DFT and crystallographic structures of catalyst 3.

Table S2. Comparison of DFT and crystallographic structures of catalysis	t 2.
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Bond	Bond Distance (Å) DFT	Bond Distance (Å) Crystal	Difference (Å)
Ni-Cl1	2.484	2.433	0.051
Ni-N2	2.225	2.160	0.065
Ni-N3	2.141	2.094	0.047
Ni-N4	2.226	2.160	0.066
Ni-Cl5	2.487	2.433	0.054
Ni-N6	2.142	2.094	0.048



н

3

Table S3. Comparison of DFT-calculated structures of 3 and 5.

Bond	Bond Distance (Å) Complex 3	Bond Distance (Å) Complex 5
Cu-N1	2.041	2.038
Cu-N2	2.063	2.058
Cu-N3	2.059	2.055
Cu-N4	2.060	2.056



* Comparison of the structures of **3** and **5** shows that ligand L3 in complex **5** has better coordination ability compared to the parent ligand L2 in complex **3**. The metal-ligand bond distances in case of L3 are shorter than the distances in case of L2 (Table S3). However, the differences in the metal-ligand bond distances are not very prominent between L3 and L2. Ligand L3 has larger steric impedance than L2, this factor tends to increase the metal-ligand bond distances in case of L3 than L2. On the other hand, L3 being electron richer has better coordination ability than L2, this factor tends to decrease the metal-ligand bond distances in case L3 than L2. Therefore, overall there is only a slight decrease of the metal-ligand bond distances from L2 to L3.

Table S4. Comparison of DFT-calculated structures of 2 and 4.

Bond	Bond Distance (Å) Complex 3	Bond Distance (Å) Complex 5
Ni-N2	2.225	2.208
Ni-N3	2.141	2.143
Ni-N4	2.226	2.213
Ni-N6	2.142	2.139



Table S5. Coordinates of the DFT-calculated geometries.

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Н

Н

Н

Н

н

н

Н

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Cu N	0.875505000 0.998824000	4.227930000 4.950476000	10.329342000 8.510997000
С	-0.081280000	5.595949000	8.018531000
Н	-0.940632000	5.661326000	8.676265000
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Н	-1.004406000	6.670200000	6.408013000
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	2 143660000	6.495494000 5 307158000	4.955744000
ц	2.143000000	5.397 130000	5 864837000
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č	3.290514000	4.106855000	8.313925000
Ĥ	4.198015000	4.381839000	7.751001000
Н	3.423717000	4.462355000	9.337013000
Ν	3.085862000	2.649816000	8.336001000
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C	0.993764000	3 415798000	13 097493000
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C	0.546305000	3.211074000	14.402477000
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С	-0.710710000	3.678654000	14.779957000
Н	-1.071731000	3.527625000	15.792386000
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Н	-2.482500000	4.717239000	14.079353000
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С	3.323322000	2.077302000	7.008877000
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Н	-1.816653000	-0.859821000	1.033925000
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Н	-1.644059000	-3.241681000	-1.758204000
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н	-2.687375000	1.197387000	1.023774000
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Ν	-1.402177000	0.580620000	-0.498106000
Ν	0.340800000	-2.150089000	0.069514000
Ν	0.105623000	1.924421000	1.347102000
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Н	2.581485000	-1.390955000	-1.942476000
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Ċ	1 29762/000	1 227193000	3 587899000
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Ν	1.107679000	4.667675000	8.574273000
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н	4.327526000	3.491298000	10.885935000
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С	2.485658000	3.050118000	12.629453000
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5

С	-1.393991000	5.738913000	15.670302000
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н	-0.343522000	5.965598000	15.878550000
С	1.545376000	8.925920000	6.116548000
Н	2.619193000	8.821911000	6.301272000
н	1.381248000	9.604611000	5.278602000
Н	1.056666000	9.321107000	7.012685000
