

## Supporting Information

### **Suppressed polarization by epitaxial growth of SrTiO<sub>3</sub> on BaTiO<sub>3</sub> nanoparticles for high discharged energy density and efficiency nanocomposites**

Yupeng Ma, <sup>†a</sup> Hang Luo, <sup>†\*a</sup> Xuefan Zhou, <sup>a</sup> Ru Guo, <sup>a</sup> Feng Dang, <sup>\*b</sup> Kechao Zhou <sup>a</sup> and Dou Zhang <sup>\*a</sup>

a. State Key Laboratory of Powder Metallurgy, Central South University, Changsha, Hunan 410083, China.

b. Key Laboratory for Liquid-Solid Structural Evolution and Processing of Materials (Ministry of Education), Shandong University, Jinan 250061, China.

<sup>†</sup>These authors contributed equally to this work.

## Materials

Barium hydroxide octahydrate ( $\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$ ), strontium hydroxide octahydrate ( $\text{Sr}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$ ), sodium hydroxide ( $\text{NaOH}$ ) and titanium tetrachloride ( $\text{TiCl}_4$ ). All chemicals were of analytical grade and were purchased from Sigma-Aldrich Co. LLC, China. They were used in the experiments without any further purification.

## Synthesis of $\text{BaTiO}_3@\text{SrTiO}_3$ core-shell nanoparticles

A typical synthetic route by two-step hydrothermal reaction was used for  $\text{BaTiO}_3@\text{SrTiO}_3$ . Briefly, 2.1809g  $\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$  was added into 24 mL 1M ethanoic acid solution to obtain solution A. 0.8 mL  $\text{TiCl}_4$  was added into 16 mL alcohol to obtain solution B. 0.0967g  $\text{Sr}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$  was dissolved in 16 mL 1M ethanoic acid solution named solution C. The precursors containing  $\text{BaTiO}_3$  nanoparticles and unreacted gels were obtained at the first stage by adding an excess of solution B and 5.0 g  $\text{NaOH}$  to solution A and mechanically stirring for 10 min. Then the solution C was added into the precursors and mechanically stirred to get a uniform mixture. To proceed to the second stage, the mixture was loaded into autoclave and hydrothermally synthesized at  $200^\circ\text{C}$  for 24 h to prepare core-shell nanoparticles. After the synthesis, the particles were washed with ethyl alcohol absolute twice and dried at  $80^\circ\text{C}$  for 24 h.

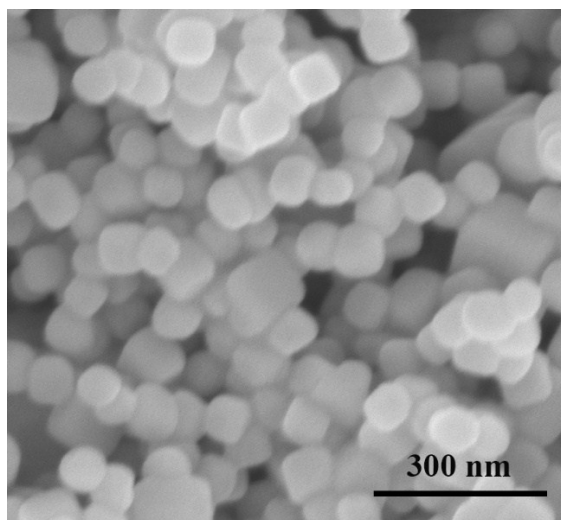


Fig.S1 SEM image of BaTiO<sub>3</sub> NPs.

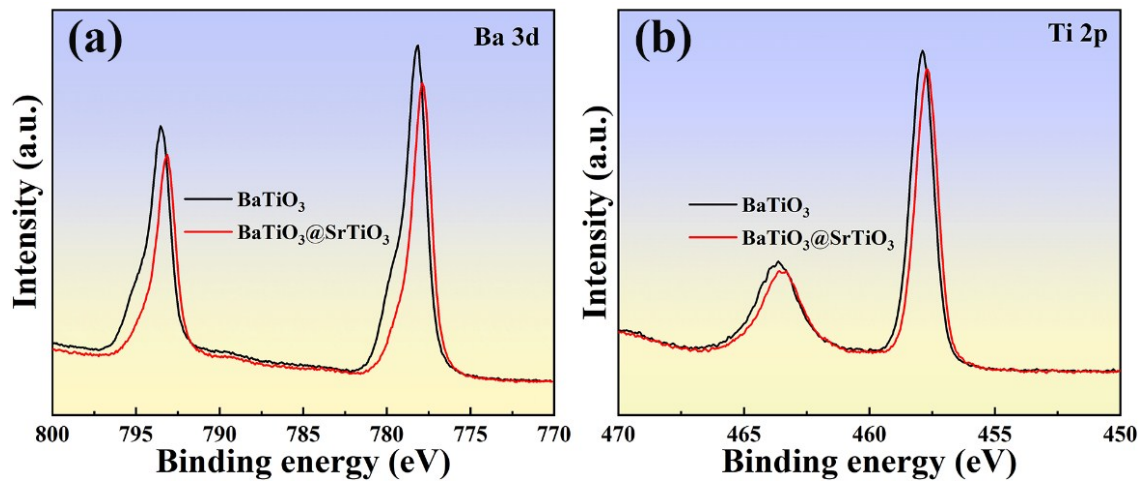


Fig.S2 XPS spectra of BaTiO<sub>3</sub>@SrTiO<sub>3</sub> NPs and BaTiO<sub>3</sub> NPs: (a) Ba and (b) Ti.

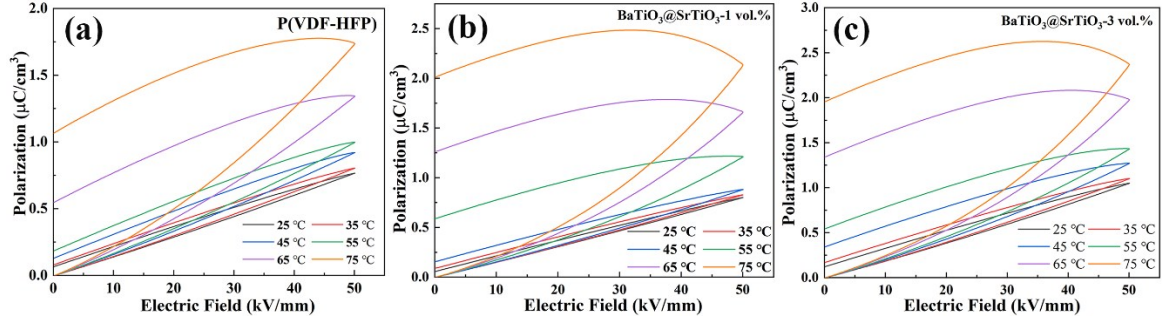


Fig.S3 P-E loops of (a) P(VDF-HFP), (b)(c) BaTiO<sub>3</sub>@SrTiO<sub>3</sub>/P(VDF-HFP) nanocomposites (filler content: 1 vol.%, 3 vol.%) in different temperature (25 °C-75 °C).

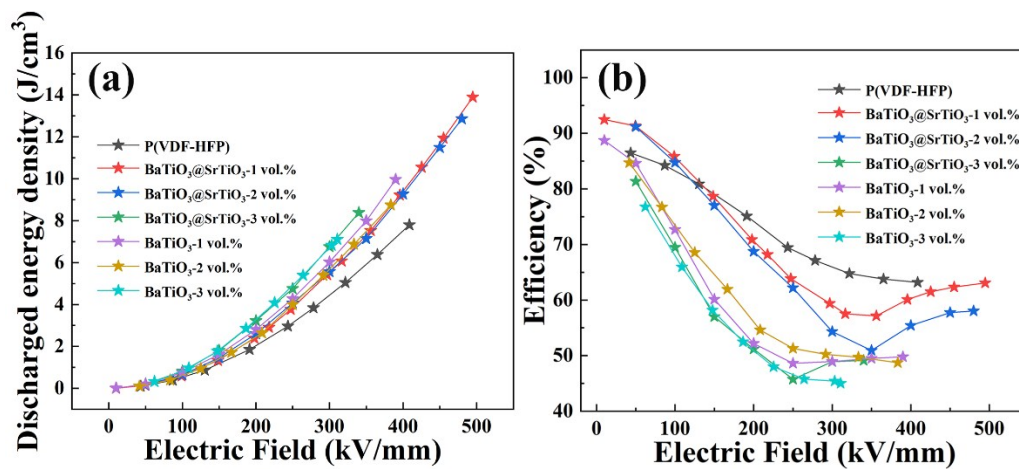


Fig. S4 (a) Discharged energy density and (b) efficiency of P(VDF-HFP), BaTiO<sub>3</sub>@SrTiO<sub>3</sub>/P(VDF-HFP) and BaTiO<sub>3</sub>/P(VDF-HFP) nanocomposites (1 vol.%, 2 vol.%, 3 vol.%).

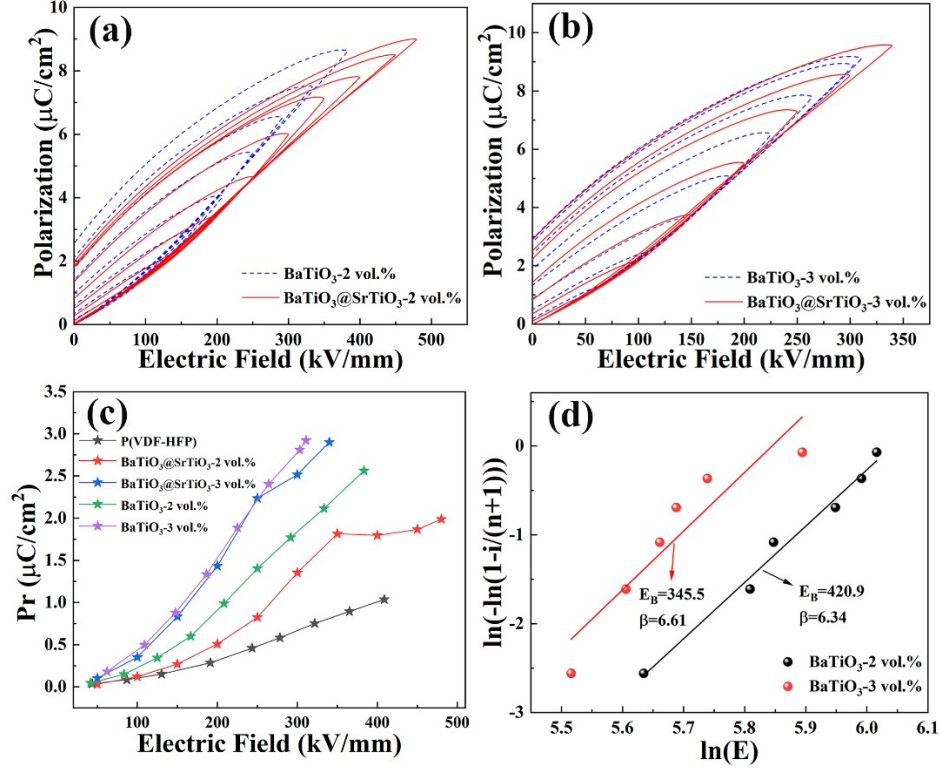


Fig. S5 (a)(b) P-E loops, (c) remanent polarization and (d) Weibull plots of BaTiO<sub>3</sub>@SrTiO<sub>3</sub>/P(VDF-HFP) and BaTiO<sub>3</sub>/P(VDF-HFP) nanocomposites (2 vol.%, 3 vol.%).