Electronic Supplementary Material

Rational design of Co-S-P nanosheet arrays as bifunctional electrocatalysts for both ethanol oxidation reaction and hydrogen

evolution reaction

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Fig. S1. SEM images of bare carbon cloth (a) and CoS₂/CC electrode (b).



Fig. S2. Nitrogen adsorption-desorption isotherms of CoS₂ and Co-S-P.



Fig. S3. LSV curves of Co-S-P/CC electrode for OER, EOR and collected using various

contents of ethanol as an electrolyte.



Fig. S4. LSV polarization curves of the Co-S-P_{1.0}/CC, Co-S-P/CC, Co-S-P_{2.0} and CoP/CC





Fig. S5. (a) LSV polarization curves of the CoS_2/CC and Co-S-P/CC in 1.0 M KOH electrolyte. (b) The corresponding Tafel plots.



Fig. S6. Nyquist plots of the bare CC, CoS₂/CC, Co-S-P/CC catalysts for EOR process at 1.38 V vs. RHE.



Fig. S7. CV curves for (a) bare CC, (b) CoS_2/CC , (c) Co-S-P/CC and (d) Pt-C/CC electrodes in a potential range of 0.83~0.93 V vs. RHE with scan rates from 20 to 100 mV s⁻¹ upon EOR catalysis.



Fig. S8. (a) Chronoamperometric curves of Co-S-P/CC electrode at a potential of 1.5 V vs. RHE for 2 h. (b) 13 C NMR spectra of electrolyte before and after oxidation reaction.



Fig. S9. LSV polarization curves of the Co-S-P/CC electrode at a scan rate of 5 mV s⁻¹

in 1.0 M KOH with and without 1.0 M ethanol for HER.



Fig. S10. Nyquist plots of the bare CC, CoS_2/CC , Co-S-P/CC catalysts at an overpotential of 200 mV toward HER.



Fig. S11. CV curves for (a) bare CC, (b) CoS_2/CC , (c) Co-S-P/CC and (d) Pt-C/CC electrodes in a potential range of -0.1^{-0} V vs. RHE with various scan rates for HER catalysis.



Fig. S12. Full water splitting performance of the CC, CoS₂/CC and Co-S-P/CC in a twoelectrode system.





electrolyser.

Table S1. Comparison with the performance of electrocatalytic ethanol oxidationcatalysts reported in the previous literature

Catalysts	Electrolyte	Current density at 1.5 V vs. RHE (mA cm ⁻²)	Potential at 10 mA cm ⁻² (V)	Electrolyte product	Ref.
Perforated CoNi hydroxide nanosheets	1.0 M KOH with 1.0 M ethanol	57	<mark>1.39</mark>	acetic acid	1
NGr–NiO/ITO	0.5 M NaOH with 1.0 M ethanol	3	<mark>-</mark>	_	2
MgFe layered double hydroxide	1.0 M NaOH with 1.0 M ethanol	2.2	ł	-	3
NiAl layered double hydroxide	1.0 M NaOH with 1.0 M ethanol	40	<mark>1.72</mark>	Acetaldehy de acetic acid	4
NiFe layered double hydroxide	1.0 M KOH with 1.0 M ethanol	1.2	ł	-	5
NiFe LDH@MnO ₂ spheres	1.0 M KOH with 1.0 M ethanol	4	<mark>>1.70</mark>	_	5
Co-S-P/CC	1.0 M KOH with 1.0 M ethanol	70	<mark>1.38</mark>	acetic acid	This work

Table S2. Comparison with the performance of electrocatalytic HER catalystsreported in the previous literature

Catalysts	Electrolyte	Overpotential at	Tafel slope	Ref.
		10 mA cm ⁻²	(mV dec ⁻¹)	
PdCu Alloy Nanosheets	1.0 M KOH	106 mV	124	6
PdCu Alloy Nanoparticles	1.0 М КОН	182 mV	170	6
Ni_3S_2	1.0 M KOH	335 mV	97	7
CoNi ₂ S ₄ @CoS ₂ /NF	1.0 M KOH	173 mV	51	8
CoNi ₂ O ₄ @Co ₃ O ₄ /NF	1.0 M KOH	206 mV	92	8
NiS ₂ /CFC	1.0 M KOH	210 mV	114	9
Carbon Paper/Carbon Tubes/C-S Sheets	1.0 М КОН	190 mV	131	10
Co-S-P/CC	1.0 M KOH with 1.0 M ethanol	167 mV	86	This work

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