Supplementary Material
**Figure SI1**

**Fig. SI1** The pH dependence of the $^1$H chemical shifts of $V^{V}$-trihyat complexes formed in solution. $C_v = 4.0 \times 10^{-3}$ M, $L : M = 1$, $l = 0.20$ M KCl. The samples contain 10% D$_2$O. Type 1 hydrogen marked with ○ (left axis) and type 2 with ▲ (right axis).
Fig. SI2 A) Concentration distribution curves of $V^{V}$-trihyat complexes formed in solutions with $c_V = 2.6 \times 10^{-4}$ M and L:M=1, calculated by using the stability constants listed in Table 1 together with the observed absorbance at $\lambda = 360$ nm. B) The molar absorbances at the same conditions, in the pH range 10-12.