

**Electronic Supporting Information**

**Selenium Containing Imidazolium Salt In Designing Single Source  
Precursors for Silver Bromide and Selenide Nano-Particles**

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## 1. Experimental Section

### General

$^1\text{H}$ ,  $^{13}\text{C}\{^1\text{H}\}$  and  $^{77}\text{Se}\{^1\text{H}\}$  NMR spectra were recorded on a Bruker Spectrospin DPX 300 NMR spectrometer at 300.13, 75.47 and 57.24 MHz respectively. The chemical shifts are reported in ppm relative to the residual deuterated solvent or the internal standard (tetramethylsilane in case of  $^1\text{H}$ ,  $^{13}\text{C}\{^1\text{H}\}$  NMR and  $\text{Me}_2\text{Se}$  for  $^{77}\text{Se}\{^1\text{H}\}$  NMR). All reactions were carried out in glassware dried in an oven and under  $\text{N}_2$  atmosphere at room temperature.

X-ray diffraction data for crystals of **1** was collected on a BRUKER AXS SMART-APEX CCD diffractometer using  $\text{Mo-K}\alpha = 0.71073 \text{ \AA}$ ; radiations monochromator, graphite). Frames were collected at  $T = 298 \text{ K}$  by  $\omega$ ,  $\varphi$ , and  $2\theta$ -rotations with full quadrant data collection strategy (four domains each with 600 frames) at 10s per frame with SMART. The measured intensities were reduced to  $F^2$  and corrected for absorption with SADABS.<sup>1</sup> Structure solution, refinement, and data output were carried out with the SHELXTL package by direct methods.<sup>2</sup> Non-hydrogen atoms were refined anisotropically.

Elemental composition of these phases has been obtained by a Carl ZEISS EVO50 scanning electron microscope (SEM). Sample was mounted on a circular metallic sample holder with a sticky carbon tape, associated EDX system Model QuanTax 200, which is based on the SDD technology and provides an energy resolution of 127 eV at Mn K alpha.

Powder X-ray diffraction (PXRD) studies were carried out on a Bruker D8 Advance diffractometer using Ni-filtered  $\text{CuK}\alpha$  radiation, scan speed of 1 s and scan step of  $0.05^\circ$ . Transmission electron microscopic (TEM) studies were carried out using a JEOL JEM 200CX TEM instrument operated at 200 kV. The specimens for these studies were prepared by dispersing the powdered sample in ethanol by ultrasonic treatment. Few drops of resulting solution were put on a porous carbon film supported on a copper grid and dried in air. High Resolution Mass Spectral (HR-MS) measurements were performed with electron spray ionization (10 eV, 180 °C source temperature) and using sodium formate as calibrant on a Bruker MicroTOF-Q II, taking sample in  $\text{CH}_3\text{CN}$ .

### Chemicals

Oleic acid (OA), octadecylamine (ODA), oleylamine (OM), 1-octadecene (ODE), trioctylphosphine (TOP),  $\text{Ag}_2\text{O}$ ,  $\text{AgBF}_4$ , benzyl bromide procured from Sigma-Aldrich (USA) were used as received. All the solvents of AR grade i.e. 1,2-dichloroethane, dichloromethane,

acetonitrile and ethanol were dried and distilled before use by known standard procedures.<sup>3</sup> The 1-(2-chloroethyl)-1*H*-imidazole was synthesized by reported procedure.<sup>4</sup>

### 1-(2-(Phenylseleno)ethyl)-1*H*-imidazole (A)

Diphenyldiselenide (0.624 g, 2.0 mmol) dissolved in 30 mL of EtOH was stirred under N<sub>2</sub> atmosphere at room temperature. Sodium borohydride (0.151 g, 4.0 mmol) was added to it as solid and the mixture further stirred for 15 min. When it became colorless due to the formation of PhSeNa, 1-(2-chloroethyl)-1*H*-imidazole (0.522g, 4.0 mmol) dissolved in 10 cm<sup>3</sup> of ethanol was added with constant stirring and the mixture stirred further for 3h. It was extracted with chloroform (4 × 25 mL). The extract was washed with water (3 × 40 mL) and dried over anhydrous sodium sulphate. The solvent was evaporated off under reduced pressure on a rotary evaporator to get 1-(2-(phenylseleno)ethyl)-1*H*-imidazole (A) as yellow oil.

Yield 0.909 g, 90%. <sup>1</sup>H NMR (CDCl<sub>3</sub>, 25 °C, vs Me<sub>4</sub>Si): δ (ppm): 3.07 (t, <sup>3</sup>J<sub>H-H</sub> = 7.2 Hz, 2H, H<sub>5</sub>), 4.06 (t, <sup>3</sup>J<sub>H-H</sub> = 7.2 Hz, 2H, H<sub>6</sub>), 6.84 (s, 1H, H<sub>7</sub>), 6.99 (s, 1H, H<sub>8</sub>), 7.17–7.24 (m, 3H, H<sub>1</sub>& H<sub>2</sub>), 7.38 (s, 1H, H<sub>9</sub>), 7.41–7.45 (m, 2H, H<sub>3</sub>). <sup>13</sup>C{<sup>1</sup>H} NMR (CDCl<sub>3</sub>, 25 °C, vs Me<sub>4</sub>Si): δ (ppm): 27.4 (C<sub>5</sub>), 46.2 (C<sub>6</sub>), 127.1 (C<sub>1</sub>) 127.1 (C<sub>7</sub>), 128.6 (C<sub>8</sub>), 128.8 (C<sub>2</sub>), 130.8 (C<sub>9</sub>), 132.6 (C<sub>3</sub>), 136.4 (C<sub>4</sub>); <sup>77</sup>Se{<sup>1</sup>H} NMR (CDCl<sub>3</sub>, 25 °C, vs Me<sub>2</sub>Se); (δ, ppm): 282.60.

### 3-Benzyl-1-(2-phenylselanyl-ethyl)-3*H*-imidazolium bromide (L)

A (0.502 g, 2.0 mmol) was placed in a Schlenk tube equipped with a magnetic stirrer and benzyl bromide (0.342 g, 2.0 mmol) was added. The mixture was heated for 8 h at 85°C under N<sub>2</sub> atm and thereafter allowed to cool to room temperature. The dark red oil obtained was washed with dry CH<sub>3</sub>CN (2 × 40 mL) and dried in vacuo.

Yield 0.802 g, 95%. <sup>1</sup>H NMR (CDCl<sub>3</sub>, 25 °C, TMS); (δ, ppm): 3.40 (t, <sup>3</sup>J<sub>H-H</sub> = 6.3 Hz, 2H, H<sub>5</sub>), 4.55 (t, <sup>3</sup>J<sub>H-H</sub> = 6.0 Hz, 2H, H<sub>6</sub>), 5.51(s, 2H, H<sub>10</sub>), 7.20 - 7.46 (m, 11H, H<sub>1</sub>,H<sub>2</sub>,H<sub>3</sub>,H<sub>7</sub>,H<sub>12</sub>,H<sub>13</sub>&H<sub>14</sub>), 7.58 (s, 1H, H<sub>8</sub>), 10.09 (s, 1H, H<sub>9</sub>). <sup>13</sup>C{<sup>1</sup>H} NMR (CDCl<sub>3</sub>, 25 °C, TMS); (δ, ppm): 26.9 (C<sub>5</sub>), 49.6 (C<sub>6</sub>), 52.8 (C<sub>10</sub>), 121.5 (C<sub>7</sub>), 122.6 (C<sub>8</sub>), 127.4 (C<sub>1</sub>), 128.7 (C<sub>2</sub>), 128.9 (C<sub>4</sub>), 129.0 (C<sub>13</sub>&C<sub>14</sub>), 129.2 (C<sub>12</sub>), 131.1 (C<sub>11</sub>), 132.8 (C<sub>3</sub>), 136.1 (C<sub>9</sub>). <sup>77</sup>Se {<sup>1</sup>H} NMR (CDCl<sub>3</sub>, 25 °C, Me<sub>2</sub>Se); (δ, ppm): 277.83. HRMS: M-Br (m/z) 343.0723.

### Complex 1

**L** (0.211 g, 0.5 mmol) was dissolved in dry CH<sub>2</sub>Cl<sub>2</sub> (20 mL) and solid Ag<sub>2</sub>O (0.058 g, 0.25 mmol) was added under N<sub>2</sub> atm. The reaction mixture was stirred overnight in the dark at room temperature. Then the suspension was filtered through a celite pad and the solvent was evaporated under reduced pressure to give a light sensitive white solid **1**.

Yield 0.026 g, 43%. <sup>1</sup>H NMR (DMSO, 25 °C, TMS); (δ, ppm): 3.28 (t, <sup>3</sup>J<sub>H-H</sub> = 6.3 Hz, 4H, H<sub>5</sub>), 4.38 (t, <sup>3</sup>J<sub>H-H</sub> = 6.3 Hz, 4H, H<sub>6</sub>), 5.24 (s, 4H, H<sub>10</sub>), 7.09 (s, 2H, H<sub>7</sub>), 7.21–7.27 (m, 12H, H<sub>1</sub>,H<sub>2</sub>,H<sub>13</sub>&H<sub>14</sub>), 7.36–7.39 (m, 10H, H<sub>3</sub>, H<sub>8</sub> &H<sub>11</sub>), 8.80 (s, 2H, H<sub>9</sub>). <sup>13</sup>C{<sup>1</sup>H} NMR (CDCl<sub>3</sub>, 25 °C, TMS); (δ, ppm): 27.5 (C<sub>5</sub>), 51.0 (C<sub>6</sub>), 54.1 (C<sub>10</sub>), 121.9 (C<sub>7</sub>), 122.3 (C<sub>8</sub>), 126.7 (C<sub>1</sub>), 127.5 (C<sub>2</sub>), 127.9 (C<sub>4</sub>), 128.7 (C<sub>14</sub>), 129.0 (C<sub>13</sub>), 129.3 (C<sub>12</sub>), 131.2 (C<sub>3</sub>), 137.1 (C<sub>11</sub>), 179.2 (C<sub>9</sub>). <sup>77</sup>Se{<sup>1</sup>H} NMR (DMSO, 25 °C, Me<sub>2</sub>Se); (δ, ppm): 264.70. HR-MS: 343.0729 [PhCH<sub>2</sub>-Imi-2CSePh]<sup>+</sup>, and 451.0856 [{Ag(PhCH<sub>2</sub>-Imi-2CSePh)}+H]<sup>+</sup>.

### Complex 2

Solid AgBF<sub>4</sub> (0.097 g, 0.5 mmol) was added to **L** (0.211 g, 0.5 mmol, dissolved in 20 mL of CH<sub>2</sub>Cl<sub>2</sub>). The reaction mixture was stirred for 4h at room temperature. In the resulting mixture, AgBr appeared as white precipitate was filtered off through celite. Solid Ag<sub>2</sub>O (0.116 g, 0.5 mmol) was added to colorless filtrate under N<sub>2</sub> atm. in the dark. The reaction mixture was further stirred overnight at room temperature. It was filtered through a celite pad and solvent from filtrate was evaporated under reduced pressure to give a light sensitive white solid **2**.

Yield 0.33 g, 75%. <sup>1</sup>H NMR (DMSO, 25 °C, TMS); (δ, ppm): 3.36 (t, <sup>3</sup>J<sub>H-H</sub> = 6.0 Hz, 4H, H<sub>5</sub>), 4.35 (t, <sup>3</sup>J<sub>H-H</sub> = 6.0 Hz, 4H, H<sub>6</sub>), 5.24(s, 4H, H<sub>10</sub>), 7.19–7.35 (m, 20H, H<sub>1</sub>,H<sub>2</sub>,H<sub>3</sub>,H<sub>12</sub>,H<sub>13</sub>&H<sub>14</sub>), 7.49 (d, 4H, H<sub>7</sub>&H<sub>8</sub>), <sup>13</sup>C{<sup>1</sup>H} NMR (DMSO, 25 °C, TMS); (δ, ppm): 28.0 (C<sub>5</sub>), 51.9 (C<sub>6</sub>), 54.9 (C<sub>10</sub>), 122.4 (C<sub>7</sub>), 122.8 (C<sub>8</sub>), 128.7 (C<sub>1</sub>), 128.8 (C<sub>2</sub>), 129.0 (C<sub>4</sub>), 129.4 (C<sub>14</sub>), 131.3(C<sub>13</sub>), 131.9 (C<sub>12</sub>), 134.7 (C<sub>11</sub>), 137.1 (C<sub>3</sub>) 179.6 (C<sub>9</sub>). <sup>77</sup>Se{<sup>1</sup>H} NMR (DMSO, 25 °C, Me<sub>2</sub>Se); (δ, ppm): 262.07. HRMS: M-BF<sub>4</sub><sup>-</sup> (m/z) 791.0313.

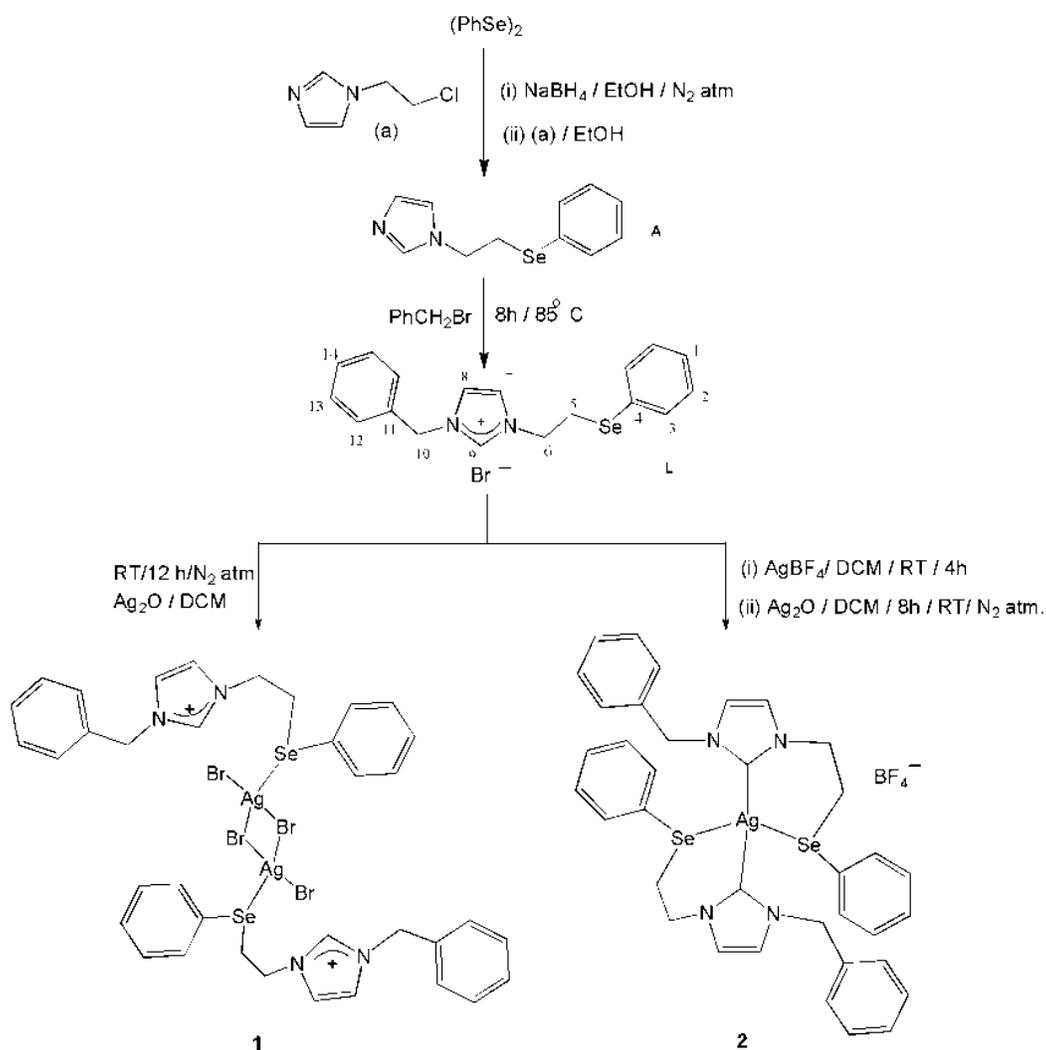
### Synthesis of Silver Bromide Nano-Particles

A slurry containing 0.1 mmol of complex **1**, 10 mmol of octadecylamine (ODA), 10 mmol of 1-octadecane (ODE) and 20 mmol of oleic acid (OA), prepared in a three necked flask (100 mL) was heated to 100 °C to remove water and oxygen, The resulting homogeneous brown solution was heated to 200 °C under N<sub>2</sub> and kept at the same temperature for 30 min, affording a dark

colloidal solution. The solution was air-cooled and NPs of AgBr were precipitated with excess ethanol. They were washed with ethanol and dried in air at 60 °C.

### Synthesis of Silver Selenide Nano-Particles

A slurry containing 0.1 mmol of complex **2**, 10 mmol of trioctylphosphine (TOP) and 20 mmol of oleic acid (OA) was prepared and treated by a procedure described above to get the NPs of Ag<sub>2</sub>Se.



**Figure S1.** Synthesis of **L**, complexes **1** and **2**

## 2. NMR Spectra

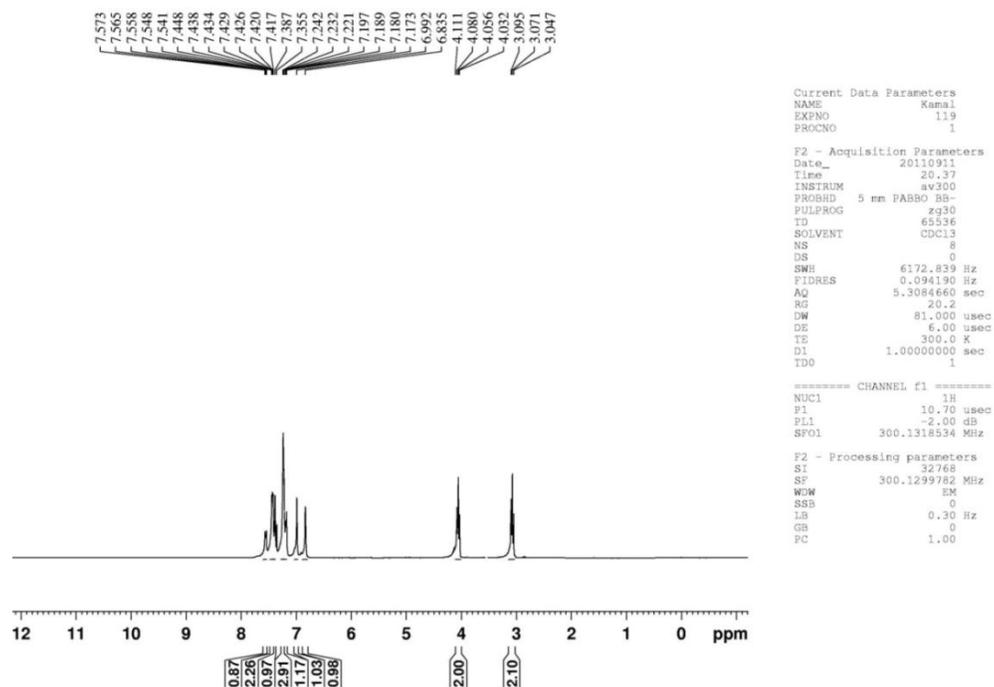


Figure S2.  $^1\text{H}$  NMR of (A).

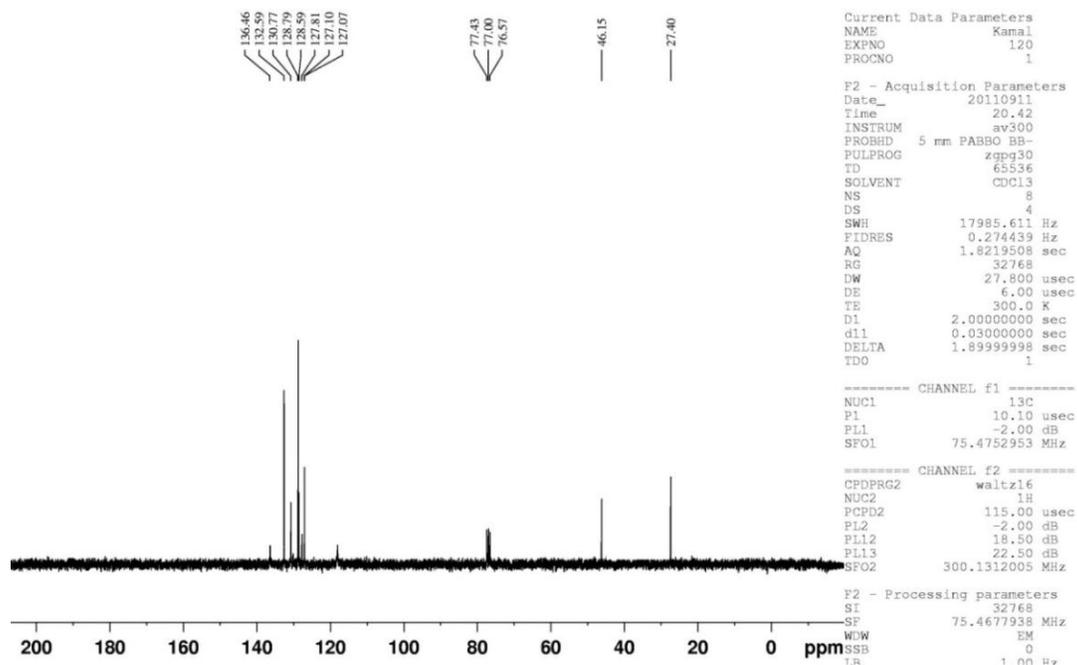


Figure S3.  $^{13}\text{C}\{^1\text{H}\}$  NMR of (A).

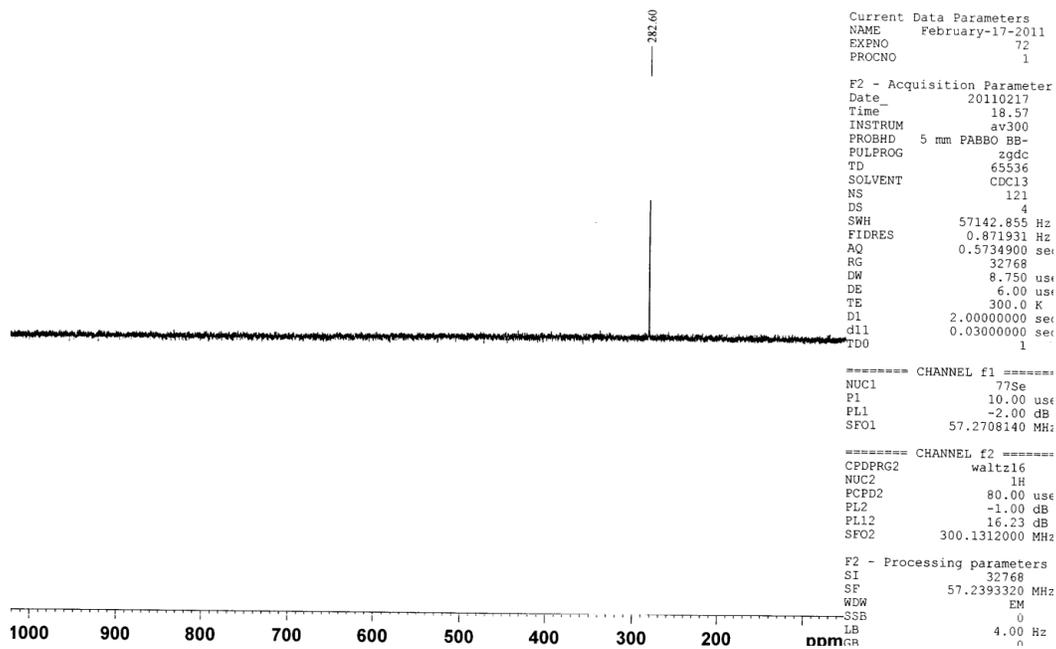


Figure S4.  $^{77}\text{Se}\{^1\text{H}\}$  NMR of (A).

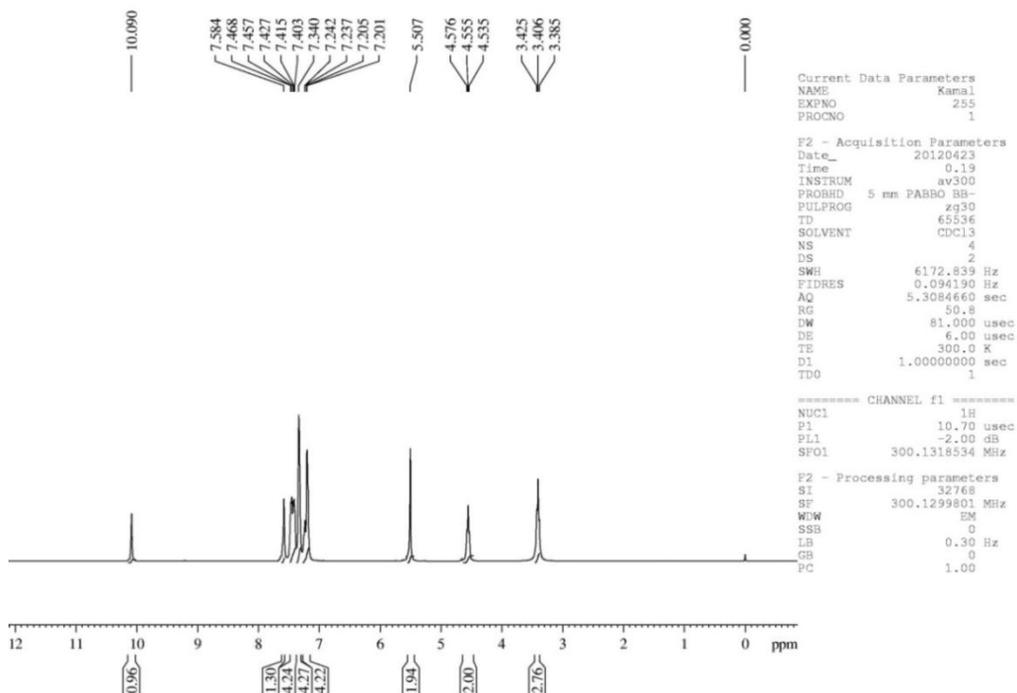


Figure S5.  $^1\text{H}$  NMR of (L).

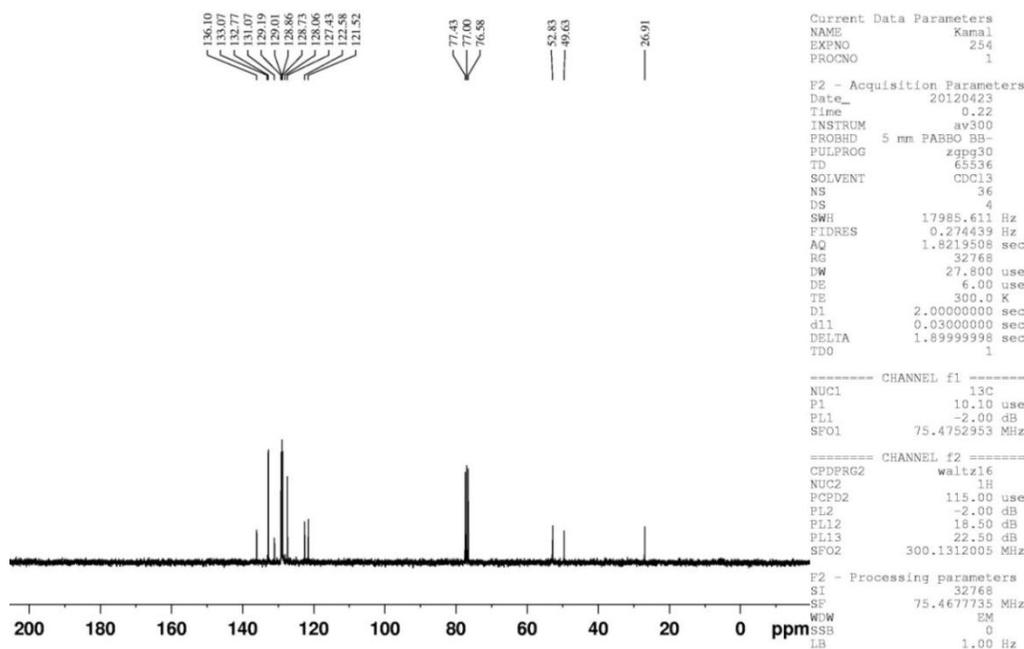


Figure S6.  $^{13}\text{C}\{^1\text{H}\}$  NMR of (L).

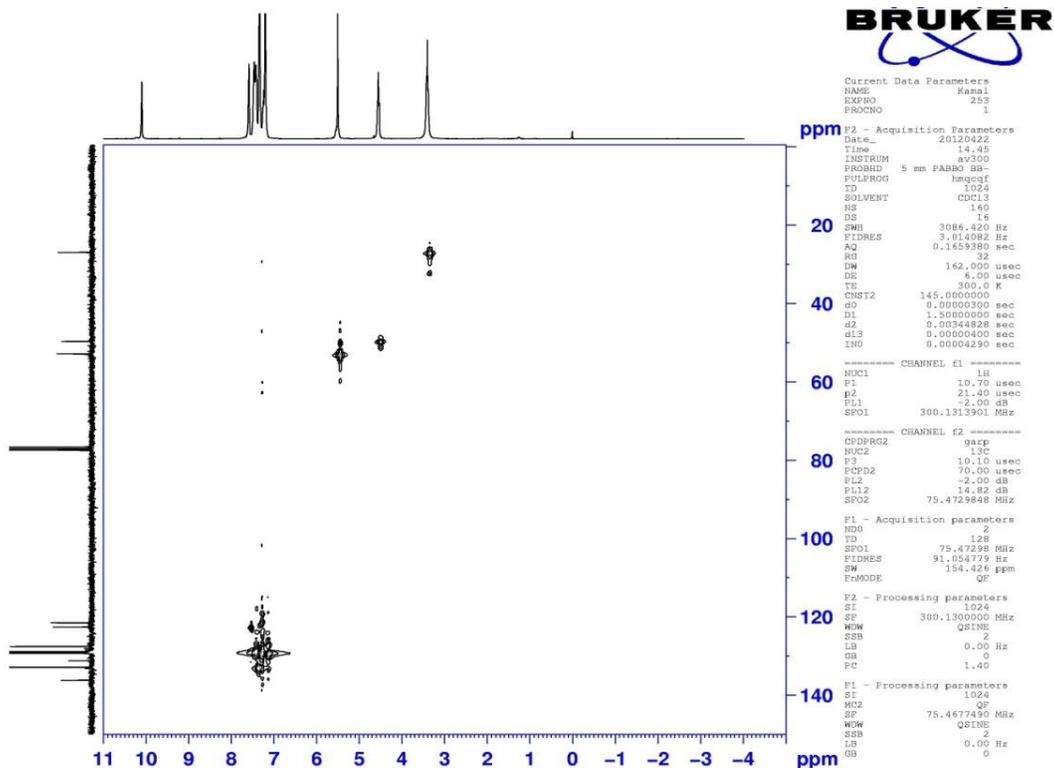


Figure S7. HMQC NMR of (L).

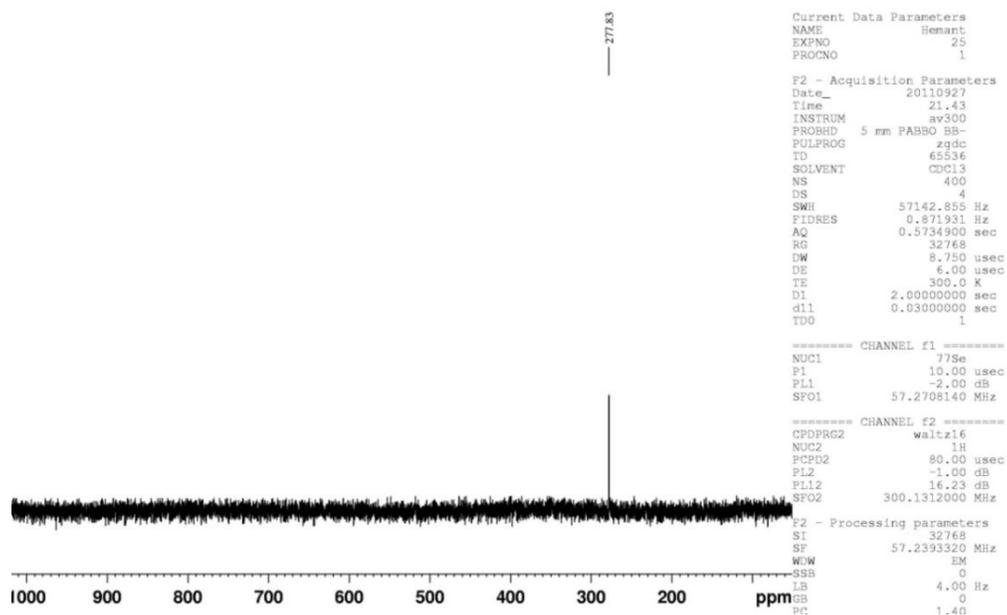


Figure S8. <sup>77</sup>Se{<sup>1</sup>H} NMR of (L).

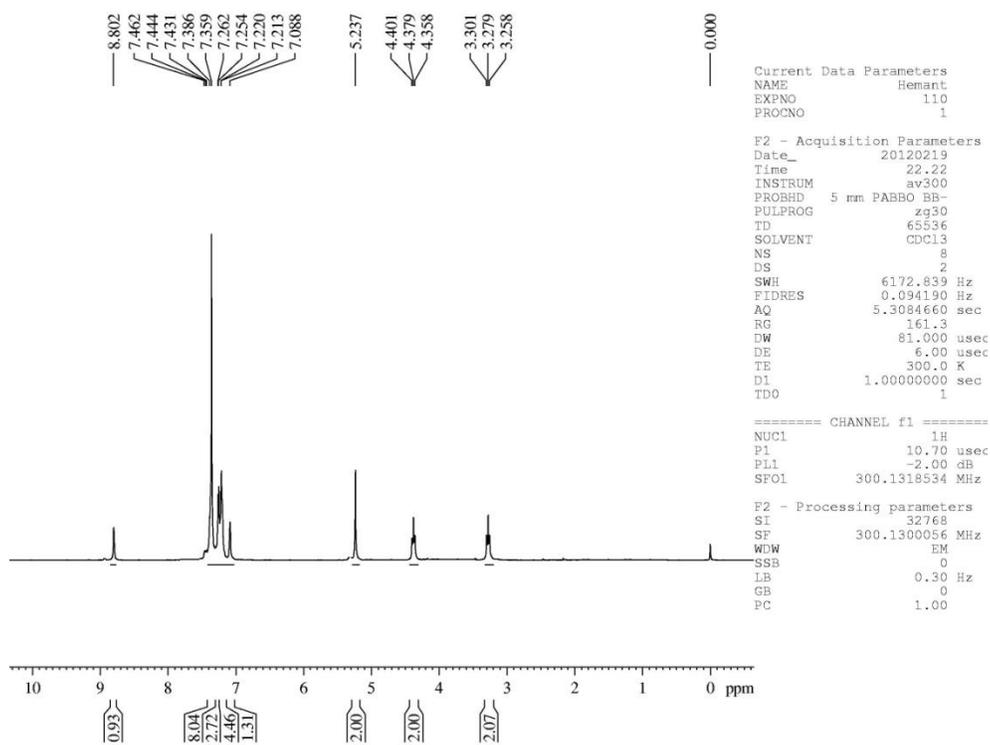


Figure S9. <sup>1</sup>H NMR of 1.

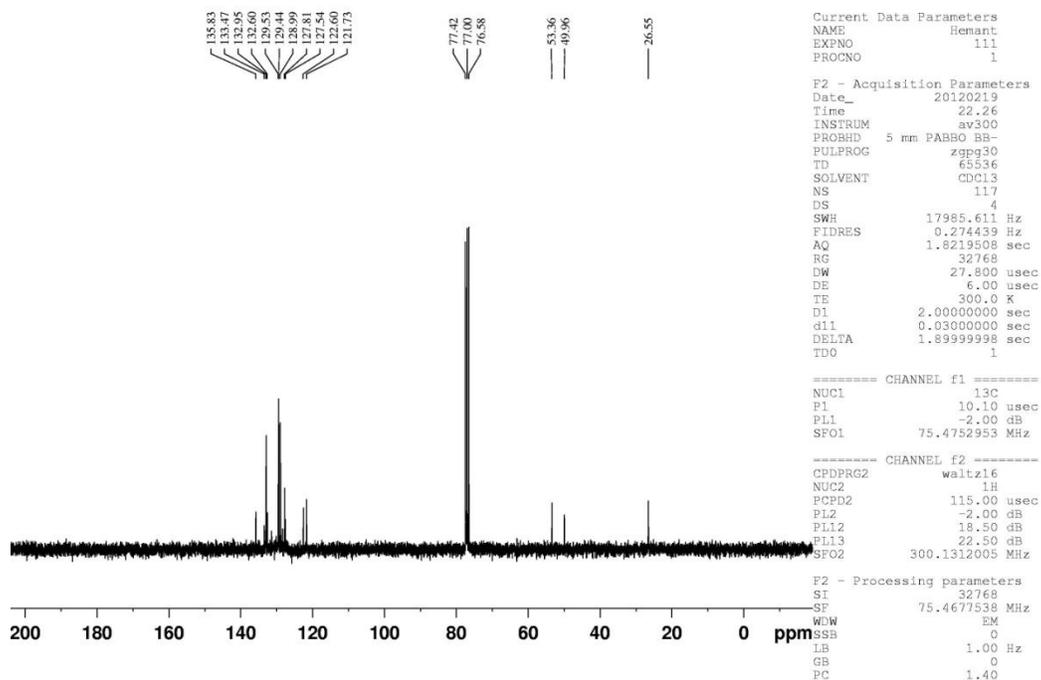


Figure S10.  $^{13}\text{C}\{^1\text{H}\}$  NMR of **1**.

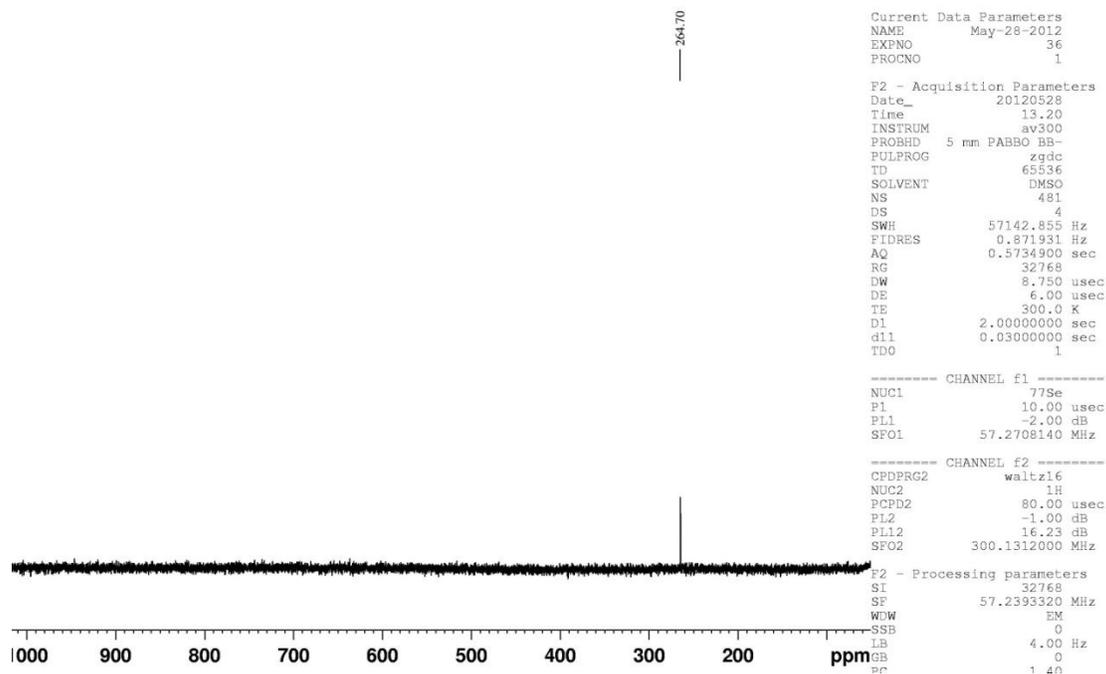


Figure S11.  $^{77}\text{Se}\{^1\text{H}\}$  NMR of **1**.

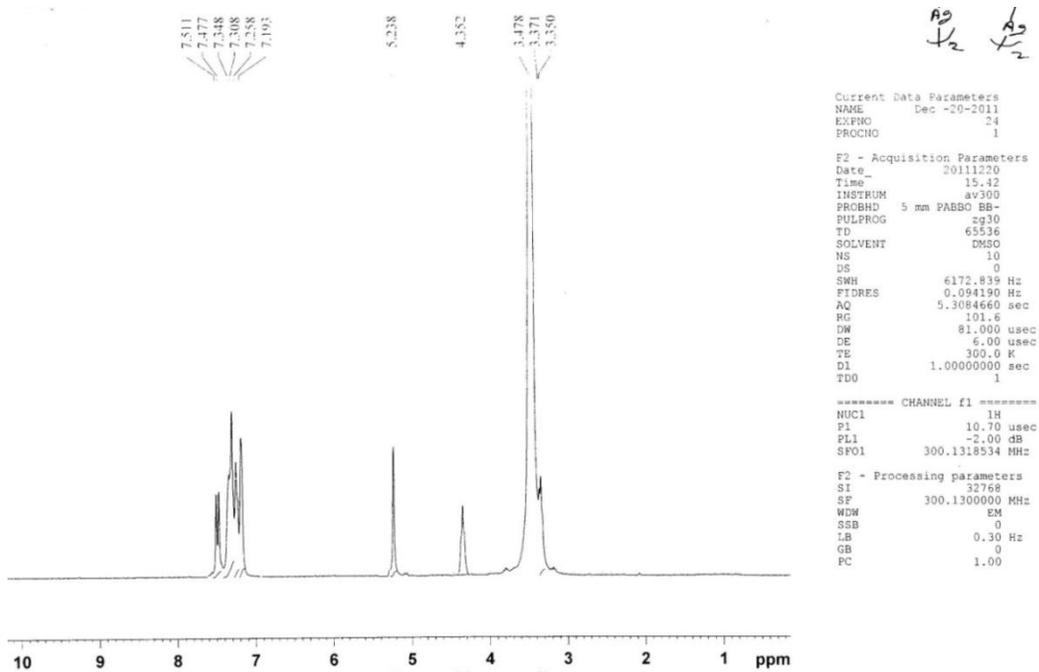


Figure S12.  $^1\text{H}$  NMR of 2.

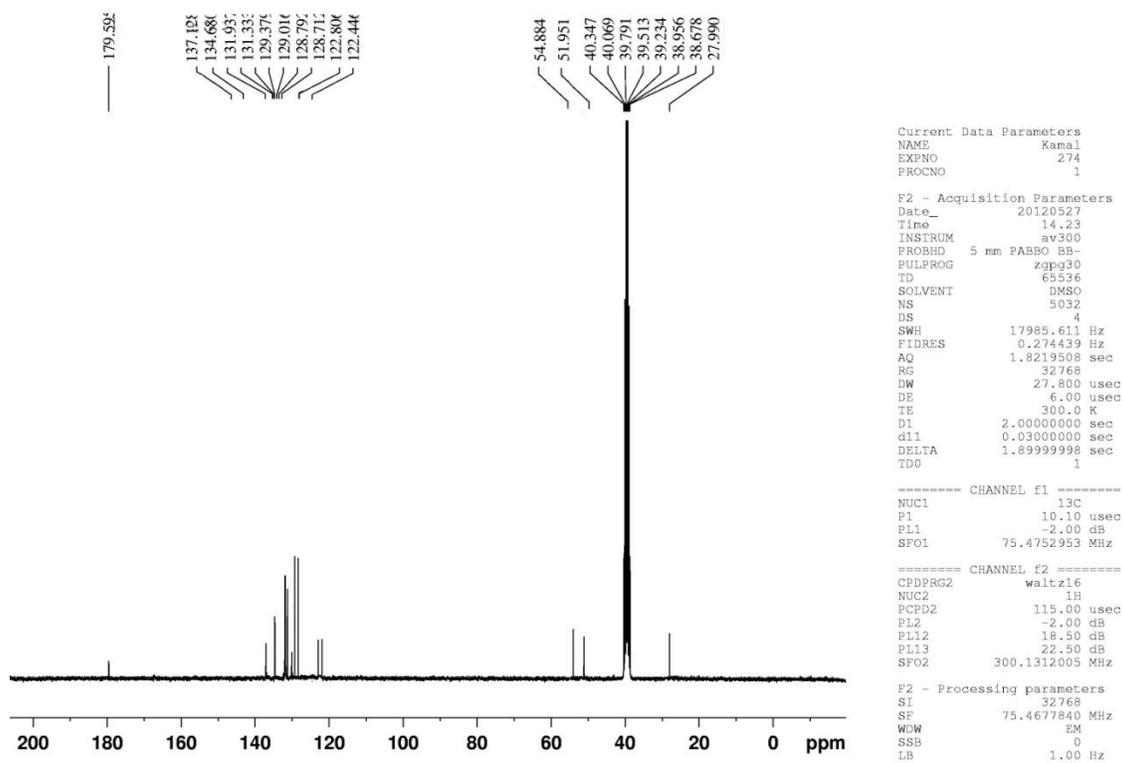


Figure S13.  $^{13}\text{C}\{^1\text{H}\}$  NMR of 2

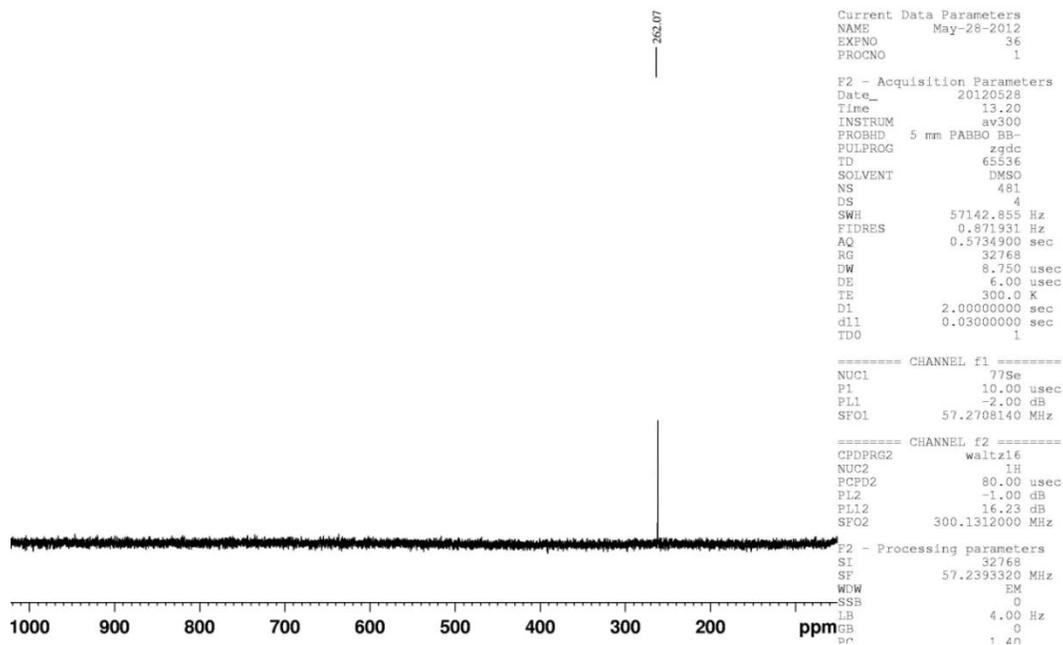
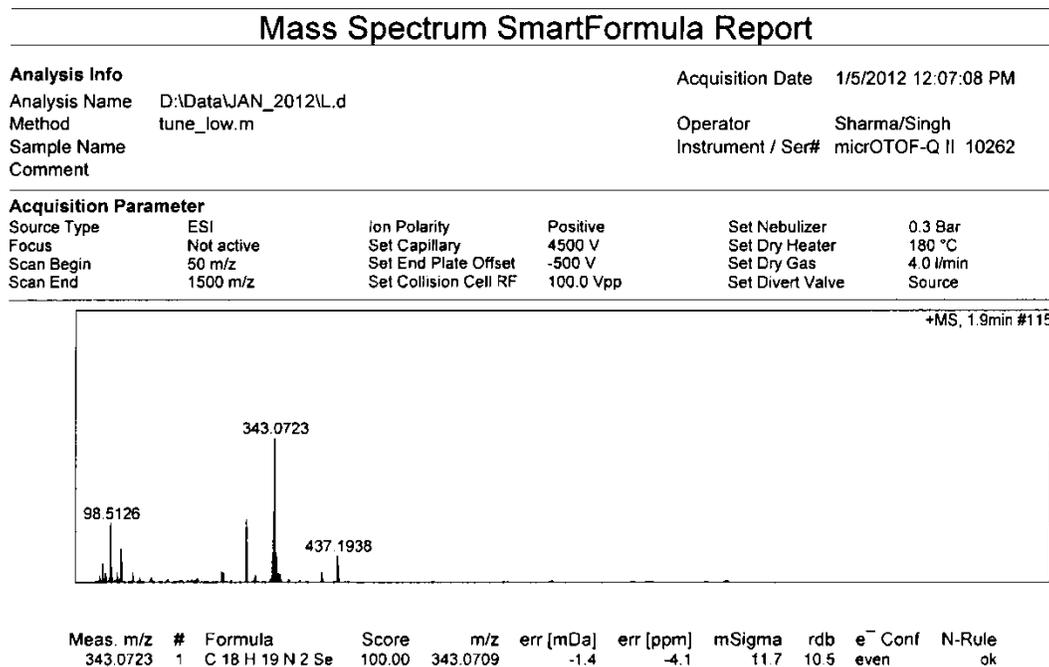


Figure S14.  $^{77}\text{Se}\{^1\text{H}\}$  NMR of 2

### 3. Mass Spectra

In mass spectrum of ligand **L** the peak appearing at 343.0723 corresponds to  $[\mathbf{L}\text{-Br}]^+$  cation. The mass spectrum of complex **1** shows two characteristic peaks at 343.0729 and 451.0856 which correspond to  $[\mathbf{L}\text{-Br}]^+$  and  $[\text{Ag}\{\mathbf{L}\text{-Br}\}+\text{H}]^+$  cationic fragments respectively. In mass spectrum of complex **2**, peak of  $[\mathbf{2}\text{-BF}_4]^+$  cation appears at 791.0313.



**Figure S15.** Mass spectra of (**L**)

## Display Report

### Analysis Info

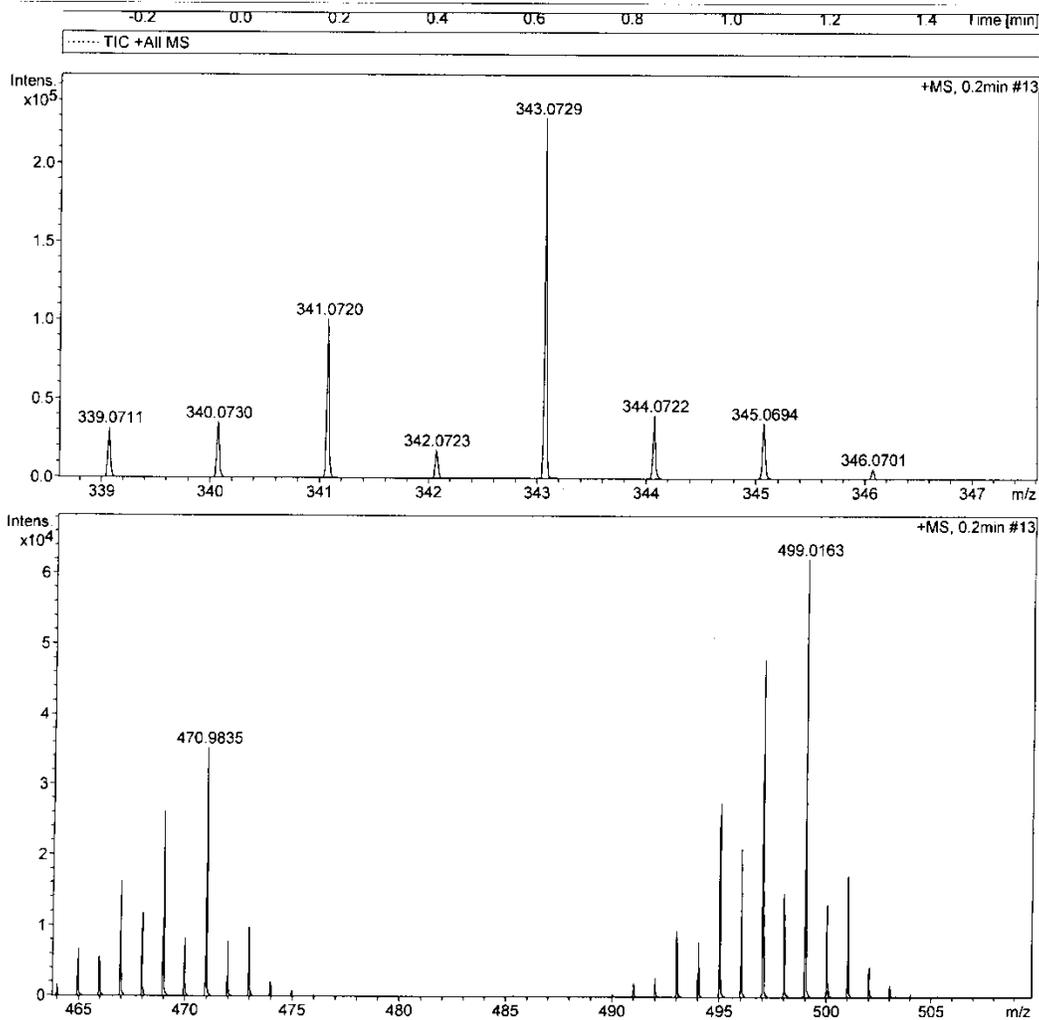
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Method tune\_wide.m  
Sample Name  
Comment

Acquisition Date 4/16/2012 12:18:21 PM

Operator Sharma/Singh  
Instrument micrOTOF-Q II 10262

### Acquisition Parameter

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Scan Begin	50 m/z	Set End Plate Offset	-500 V	Set Dry Gas	4.0 l/min
Scan End	3000 m/z	Set Collision Cell RF	600.0 Vpp	Set Divert Valve	Source



**Figure S16.** Mass spectra of **1**.

## Display Report

### Analysis Info

Analysis Name D:\Data\APRIL\_2012\h2.d  
Method tune\_wide.m  
Sample Name  
Comment

Acquisition Date 4/16/2012 12:18:21 PM

Operator Sharma/Singh  
Instrument micrOTOF-Q II 10262

### Acquisition Parameter

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Scan End	3000 m/z	Set Collision Cell RF	600.0 Vpp	Set Divert Valve	Source

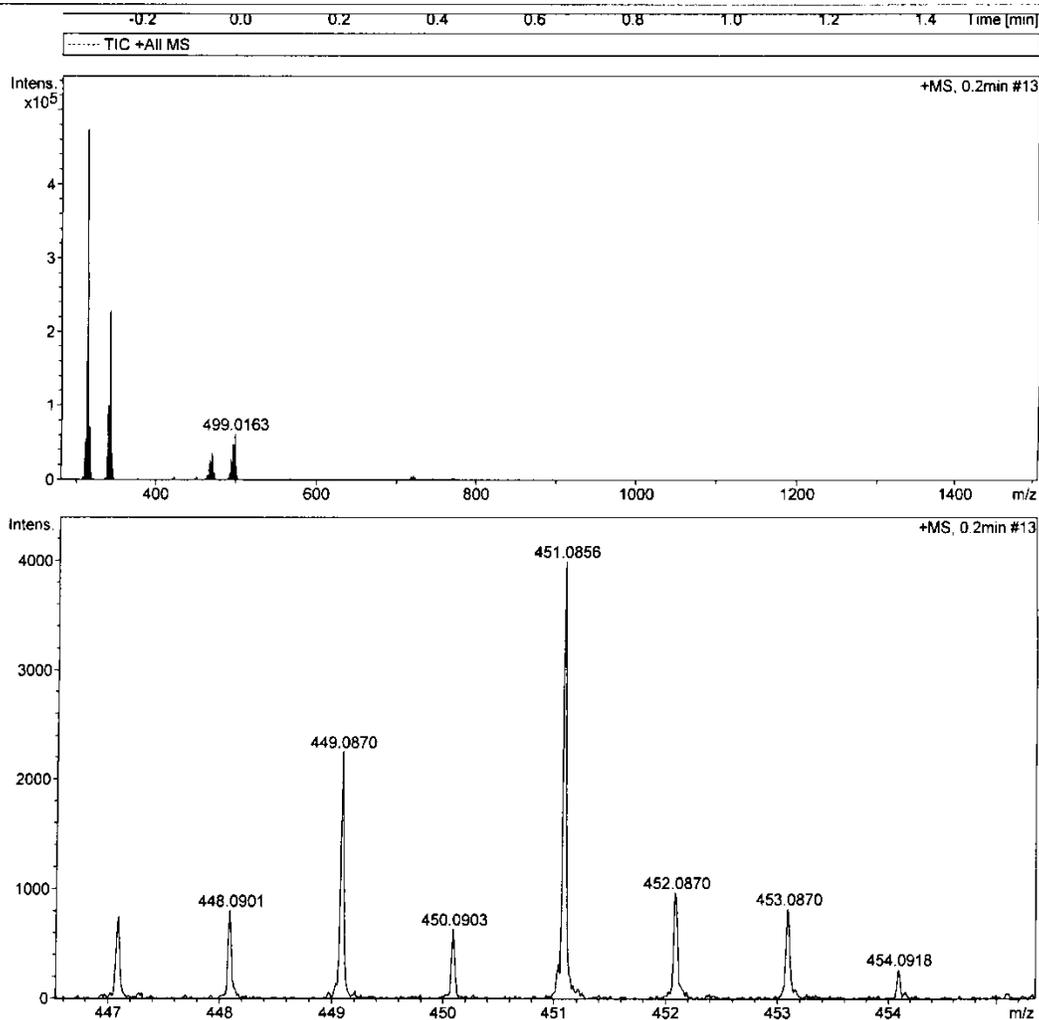
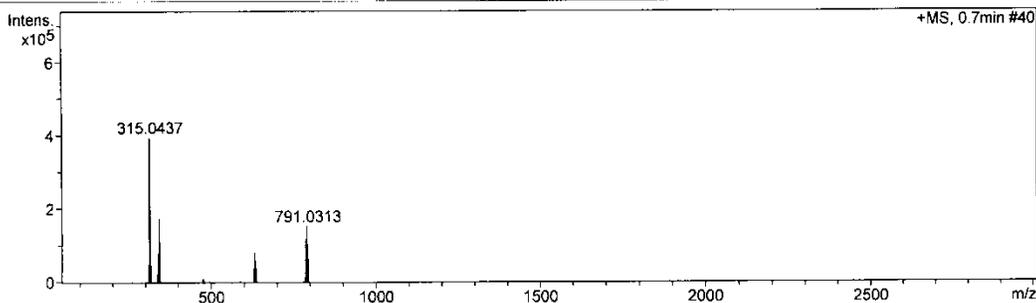


Figure S17. Mass spectra of 1.

## Mass Spectrum SmartFormula Report

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Method	tune_wide.m	Instrument / Ser#	micrOTOF-Q II 10262
Sample Name			
Comment			

<b>Acquisition Parameter</b>					
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Focus	Active	Set Capillary	4500 V	Set Dry Heater	180 °C
Scan Begin	50 m/z	Set End Plate Offset	-500 V	Set Dry Gas	4.0 l/min
Scan End	3000 m/z	Set Collision Cell RF	600.0 Vpp	Set Divert Valve	Source



Meas. m/z	#	Formula	m/z	err [ppm]	Mean err [ppm]	rdb	N-Rule	e <sup>-</sup> Conf	mSi gma	Std l	Std Me an m/z	Std l Va rN or m	Std m/ z Diff	Std Com b Dev
791.0313	1	C <sub>36</sub> H <sub>36</sub> AgN <sub>4</sub> Se <sub>2</sub>	791.0326	1.6	4.7	20.5	ok	even	38.6	36.8	4.3	5.4	2.7	842.7

**Figure S18. Mass spectra of 2.**

#### 4. Crystal data

**Table T1.** Structural refinement parameters of **1**.

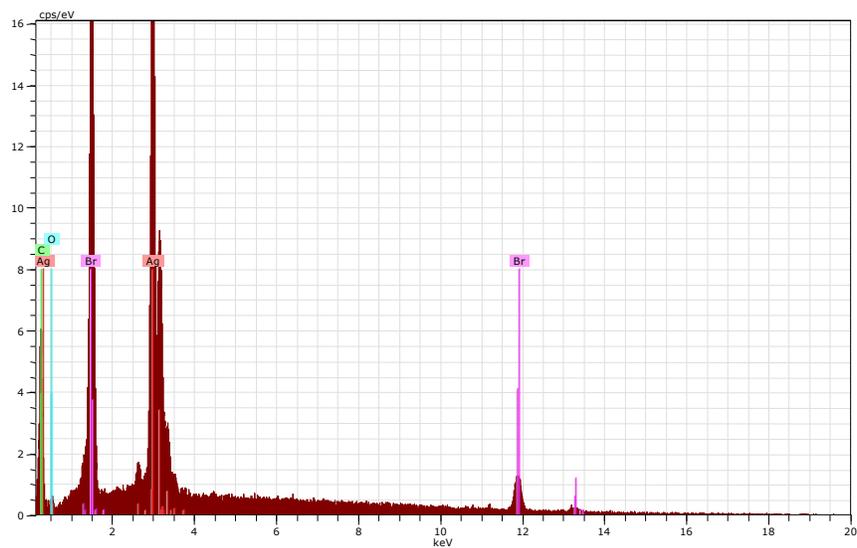
Compounds	<b>1</b>
Empirical formula	$C_{36}H_{38}Ag_2Br_4N_4Se_2$
Formula wt.	1219.96
Crystal size [mm]	0.29×0.25×0.21
Crystal system	Monoclinic
Space group	$P 21/c$
Unit Cell dimension	$a = 11.534(2)\text{Å}$ $b = 13.521(2)\text{Å}$ $c = 12.778(2)\text{Å}$ $\alpha = 90.00^\circ$ $\beta = 97.081(3)^\circ$ $\gamma = 90.00^\circ$
Volume [ $\text{Å}^3$ ]	1977.6(5)
$Z$	2
Density (Calc.) [ $\text{Mg m}^{-3}$ ]	2.049
Absorption coeff. [ $\text{mm}^{-1}$ ]	6.903
$F(000)$	1168.0
$\theta$ range [ $^\circ$ ]	2.20–26.97

Index ranges	$-13 \leq h \leq 13$ $-16 \leq k \leq 16$ $-15 \leq l \leq 15$
Reflections collected	18117
Independent reflections ( $R_{\text{int.}}$ )	3476 (0.0430)
Max./min. Transmission	0.025/0.004
Data/restraints/parameters	3476/0/217
Goodness-of-fit on $F^2$	1.052
Final R indices [ $I > 2\sigma(I)$ ]	$R_1 = 0.0271,$ $wR_2 = 0.0652$
R indices (all data)	$R_1 = 0.0354,$ $wR_2 = 0.0683$
Largest diff. peak/hole [ $\text{e.}\text{\AA}^{-3}$ ]	0.526/−0.625

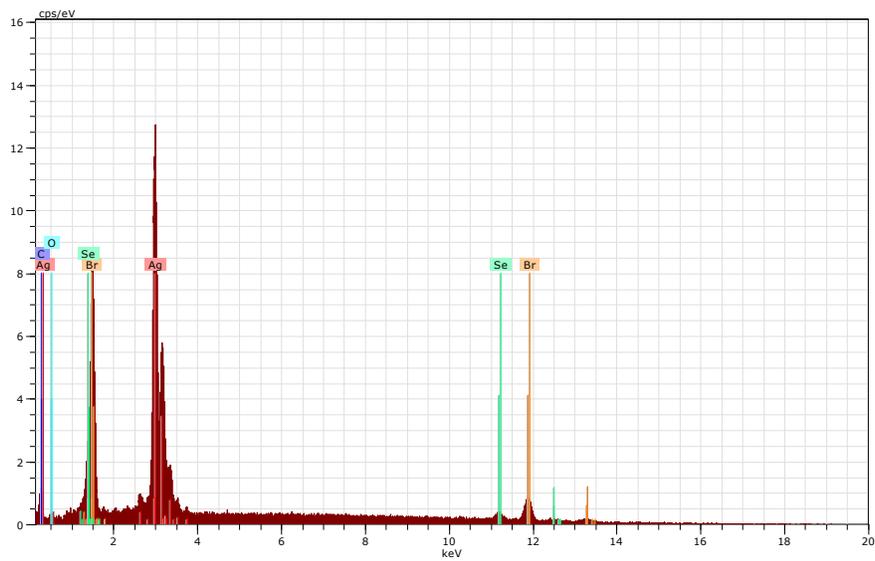
**Table T2.** Selected bond lengths [ $\text{\AA}$ ] and bond angles [ $^\circ$ ] of complex **1**

Bond length [ $\text{\AA}$ ]	Bond angle [ $^\circ$ ]
Se(1)—C(6) 1.925(4)	Br(1)—Ag(1)—Se(1) 124.65(18)
Se(1)—C(7) 1.952(4)	Br(1)—Ag(1)—Br(2) 111.91(2)
N(1)—C(8) 1.464(4)	Se(1)—Ag(1)—Br(2) 104.86(17)
N(1)—C(9) 1.325(5)	C(8)—N(1)—C(9) 124.3(3)
N(1)—C(10) 1.371(5)	C(9)—N(1)—C(10) 108.8(3)
N(2)—C(9) 1.323(5)	C(8)—N(1)—C(10) 126.9(3)
N(2)—C(11) 1.364(5)	C(9)—N(2)—C(11) 108.3(3)
N(2)—C(12) 1.480(5)	C(9)—N(2)—C(12) 122.8(4)
Se(1)—Ag(1) 2.6633(6)	C(11)—N(2)—C(12) 128.7(4)
Br(1)—Ag(1) 2.6618(6)	C(6)—Se(1)—C(7) 100.17(15)
Br(2)—Ag(1) 2.6501(6)	C(6)—Se(1)—Ag(1) 111.36(11)
Ag(1)—Ag(1) 3.1172(8)	C(7)—Se(1)—Ag(1) 109.69(12)

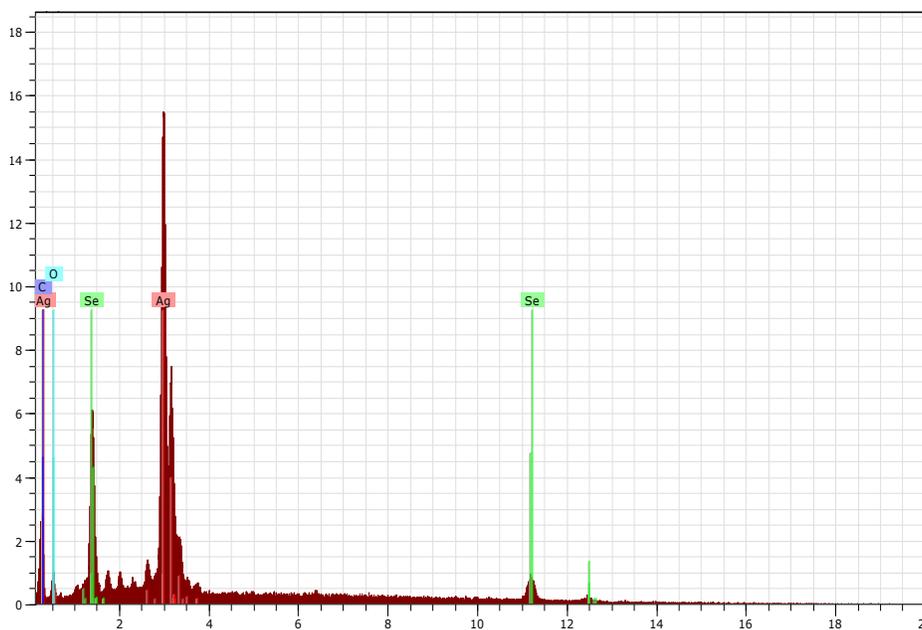
## 5. SEM-EDX data



**Figure S19.** SEM - EDX of AgBr NPs.



**Figure S20.** SEM - EDX of (Ag + AgBr + Ag<sub>2</sub>Se) NPs.



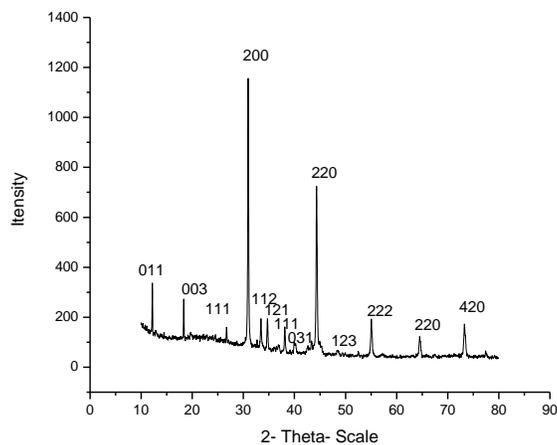
**Figure S21.** SEM - EDX of  $\text{Ag}_2\text{Se}$  NPs.

## 6. Powder XRD Data

Each of the powder XRD patterns of both types of NPs was found matching with that of the corresponding known standard phase of the same composition. The powder X-ray diffraction pattern of AgBr NPs was indexed on the basis of a cubic unit cell (JCPDS #79-0149) with the d values (*hkl*): 3.33 (111), 2.88 (200), 2.04 (220), 1.74 (311), 1.66 (222), 1.44 (400), 1.32 (331), 1.29 (420), 1.17 (422), 1.11 (511)

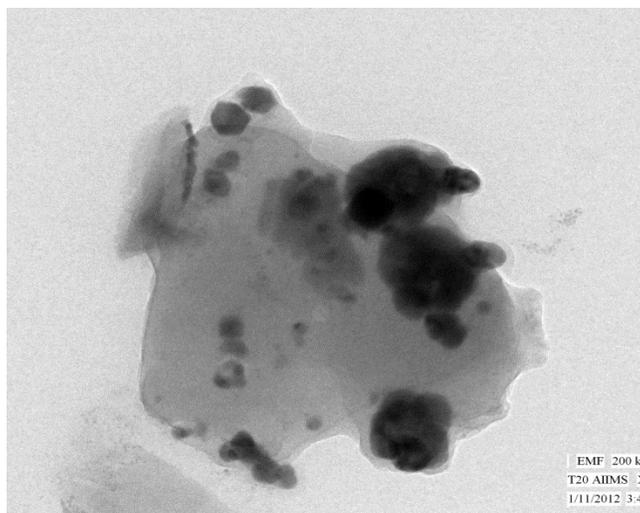
For (Ag + AgBr +  $\text{Ag}_2\text{Se}$ ) NPs a cubic unit cell (JCPDS #04-0783) with the d values (*hkl*): 2.35 (111), 2.04 (200), 1.44 (220), 1.23 (311) for Ag NPs, a cubic unit cell (JCPDS #06-0438) with the d values (*hkl*): 3.33 (111), 2.88 (200), 2.04 (220), 1.74 (311), 1.60 (222), 1.44 (400), 1.32 (331), 1.29 (420), 1.29 (420), 1.77 (422) for AgBr NPs, a orthorhombic unit cell (JCPDS #24-1041) with the d values (*hkl*): 3.33 (111), 2.89 (102), 2.73 (120), 2.67 (112), 2.58 (121), 2.42 (013), 2.25 (031), 2.23 (122), 2.16 (200), 2.11 (113), 2.08 (201), 2.06 (130), 2.01 (032), 2.00 (211), 1.94 (004), 1.88 (123), 1.87 (014), 1.84 (220), 1.82 (212), 1.76 (221), 1.76 (104), 1.71 (114), 1.66 (222), 1.63 (140), 1.61 (213), 1.60 (230), 1.59 (230), 1.56 (231), 1.50 (223), 1.47 (232), 1.45 (043), 1.44 (204) for  $\text{Ag}_2\text{Se}$  NPs,

For  $\text{Ag}_2\text{Se}$  NPs orthorhombic unit cell (JCPDS # 24-1041) with  $d$  values ( $hkl$ ): 3.88 (002), 3.77 (101), 3.69 (110), 3.52 (020), 3.33 (111), 2.89 (102), 2.73 (120), 2.67 (112), 2.58 (121), 2.42 (013), 2.25 (031), 2.23 (122), 2.16 (200), 2.11 (113), 2.08 (201), 2.06 (130), 2.01 (032), 2.00 (211), 1.94 (004), 1.88 (123), 1.87 (014), 1.84 (220), 1.82 (212), 1.79 (221), 1.76 (104), 1.71 (114), 1.66 (222), 1.61 (213), 1.56 (231), 1.50 (223), 1.44 (204), 1.42 (301), 1.41 (214) .



**Figure S22.** PXRD of (Ag + AgBr +  $\text{Ag}_2\text{Se}$ ) NPs.

## 7. TEM images



**Figure S23.** TEM image of (Ag + AgBr +  $\text{Ag}_2\text{Se}$ ) NPs.

## References

- 1 SADABS V2.10 (Sheldrick, G. M. **2003**).
- 2 (a) Sheldrick, G. M. *Acta Crystallogr., Sect. A: Found. Crystallogr.*, **1990**, 46, 467; (b) Sheldrick, G. M. *SHELXL-NT Version 6.12*, University of Gottingen, Germany, **2000**.
- 3 Furniss, B. S.; Hannaford, A. J.; Smith, P. W. G.; Tatchell, A. R. *Vogel's Textbook of Practical Organic Chemistry*, 5<sup>th</sup> Eds. ELBS, Longman Group U K Ltd., **1989**.
- 4 Lee, H. M.; Zeng, J. Y.; Hee, C. H.; Lee, M. T. *Inorg. Chem.*, **2004**, 43, 6822–6829.