Electronic Supporting Information for:

Considering the chemical energy requirements of the tri-*n*-propylamine co-reactant pathways for the judicious design of new electrogenerated chemiluminescence detection systems

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Table S1. Data for Figures 1-6: metal complexes in acetonitrile solution.[1, 2]

Complex	E_{ox} / V (*vs* $Fc^{0/+}$)	E_{red} / V (*vs* $Fc^{0/+}$)	λ_{max} / nm
$[Ru(bpy)_3]^{2+}$	0.89	-1.75	620
$[Ir(ppy)_2(phen)]^+$	0.77[a]	-1.85[a]	606
$[Ir(ppy)_3]$	0.33	-2.67	530
$[Ir(df\text{-}ppy)_3]$	0.68	-2.53	492
$[Ir(df\text{-}ppy)_2(ptb)]^+$	1.17	-2.14	455
$[Ir(pmi)_3]$	0.22	[b]	384
$[Ir(pq)_2(tmd)]$	0.53[a]	-2.01[a]	615
$[Ir(pq)_2(acac)]$	0.61[a]	-2.05[a]	609

[a]Reported in original work[1] in V *vs* SCE and referenced to $Fc^{0/+}$ using a conversion factor of -0.38 V.[3] [b]Beyond the potential window of the solvent.

Table S2. Data for Figure 7: metal complexes in aqueous solution.[4]

Complex	E_{ox} / V (*vs* Ag/AgCl)	E_{red} / V (*vs* Ag/AgCl)	λ_{max} / nm
$[Ir(df\text{-}ppy)_2(pt\text{-}PEG)]^+$	1.43	-1.69	456
$[Ir(df\text{-}ppy\text{-}SO_3)_2(ptb)]^-$	1.47	-1.69	462

References

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