

Supporting Informations

A highly selective “turn-on” fluorescent chemosensor based on 8-aminoquinoline for Zn²⁺

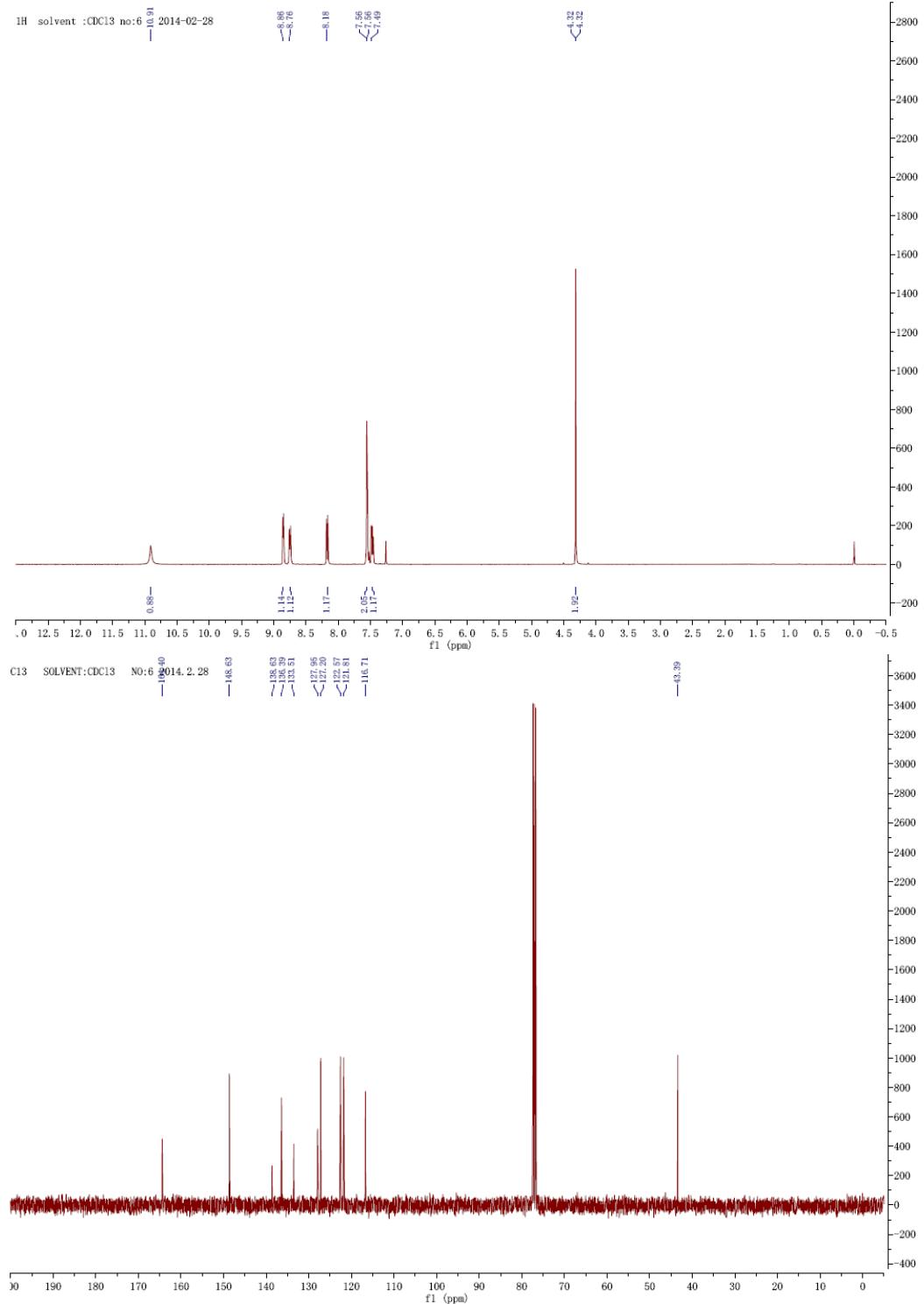
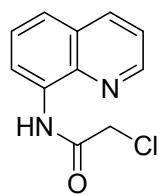
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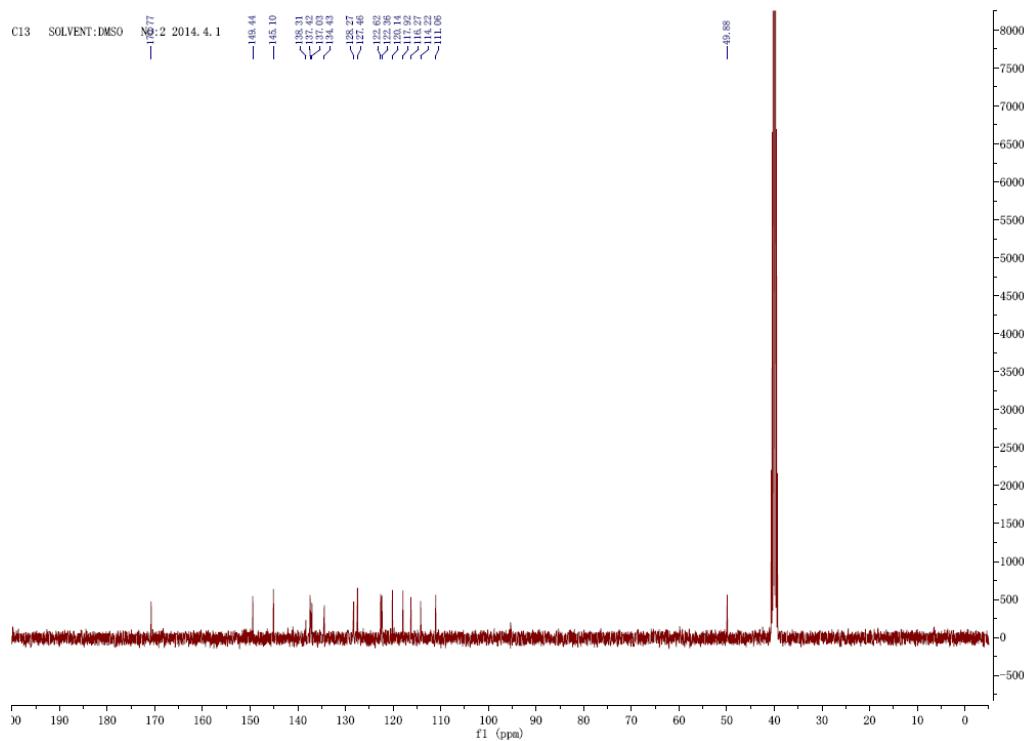
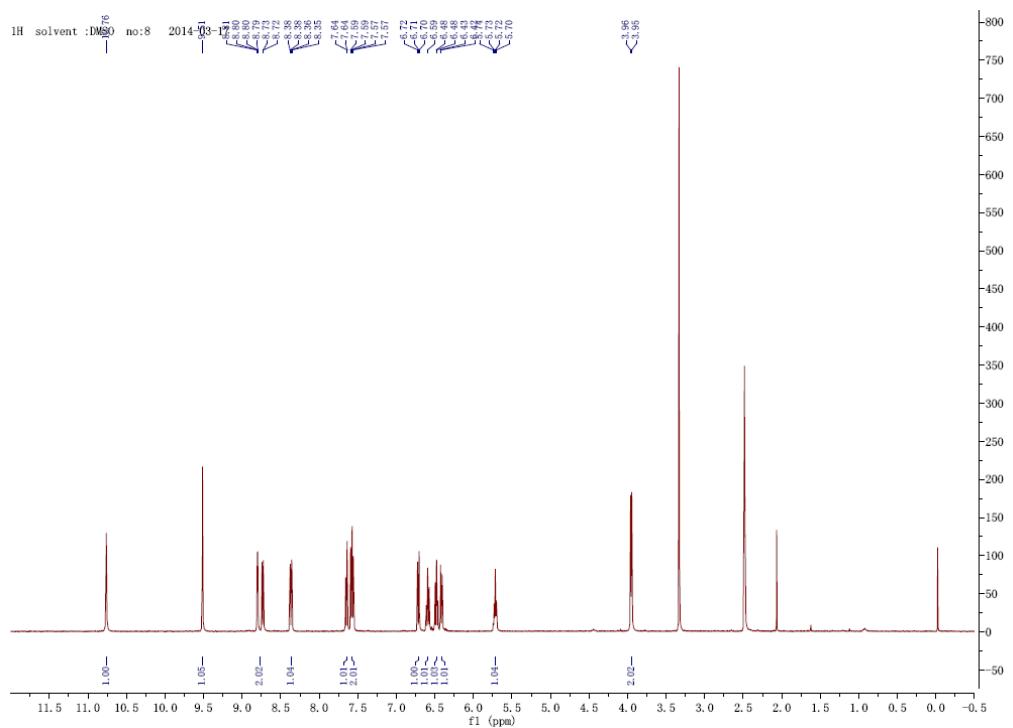
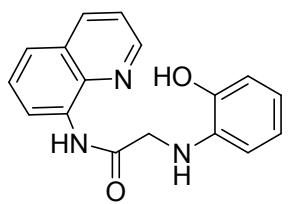
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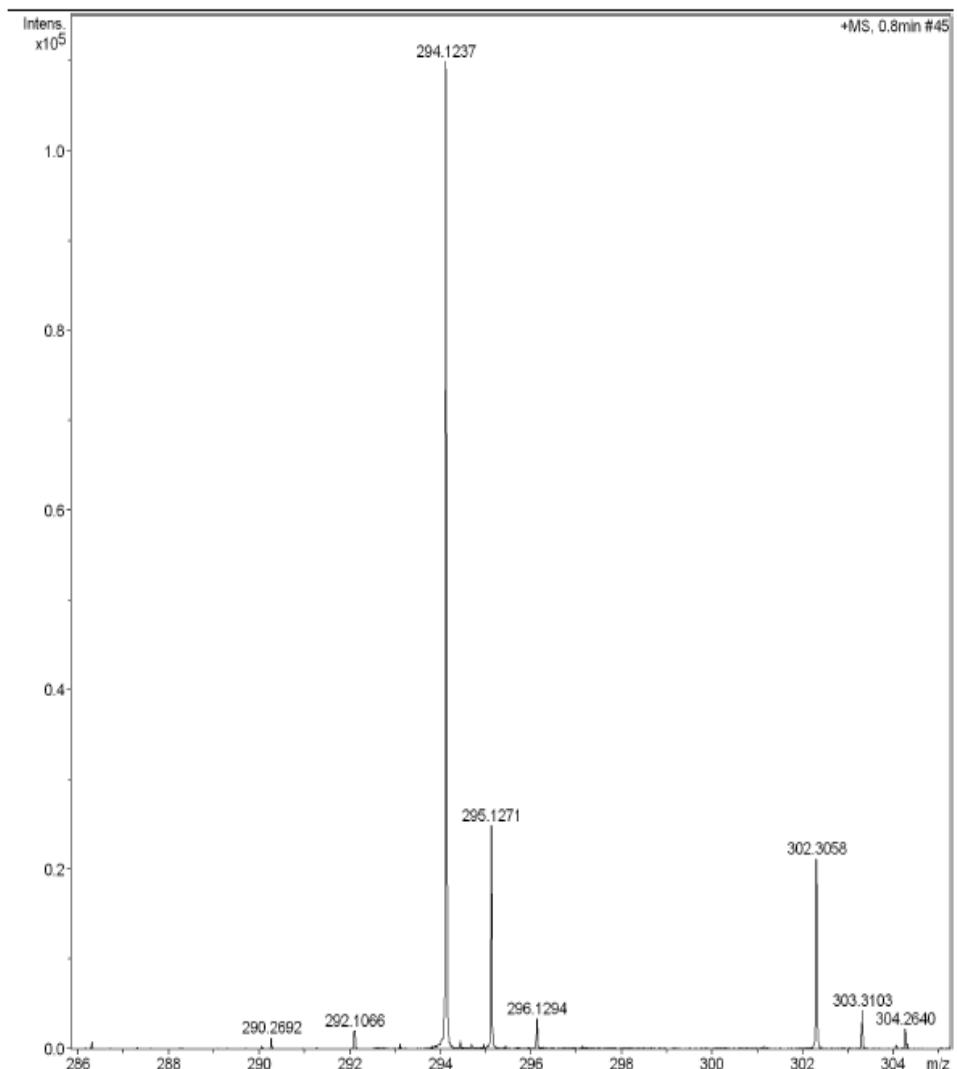
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1. ^1H NMR and ^{13}C NMR copies of products





2. HRMS spectra of HAQT



3. Benesi-Hildebrand plot of HAQT-Zn²⁺

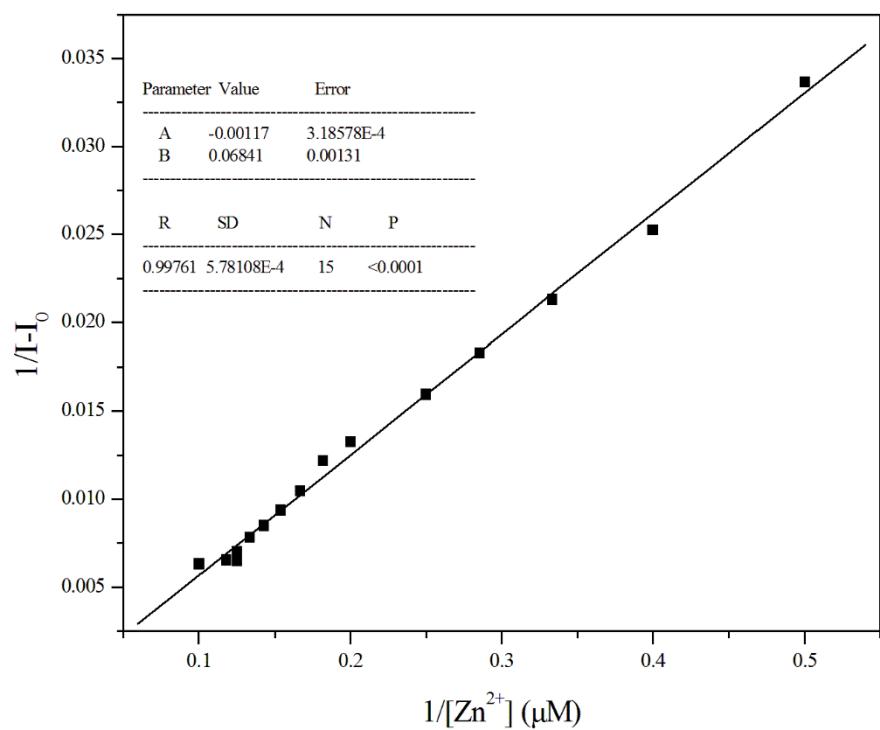


Fig. S1 Benesi-Hildebrand plot of HAQT-Zn²⁺ in buffered CH₃OH-H₂O solution (4:1,

v/v,

Tris-HCl

pH

7.40).

4. Relative fluorescence intensity between HAQT and Zn²⁺

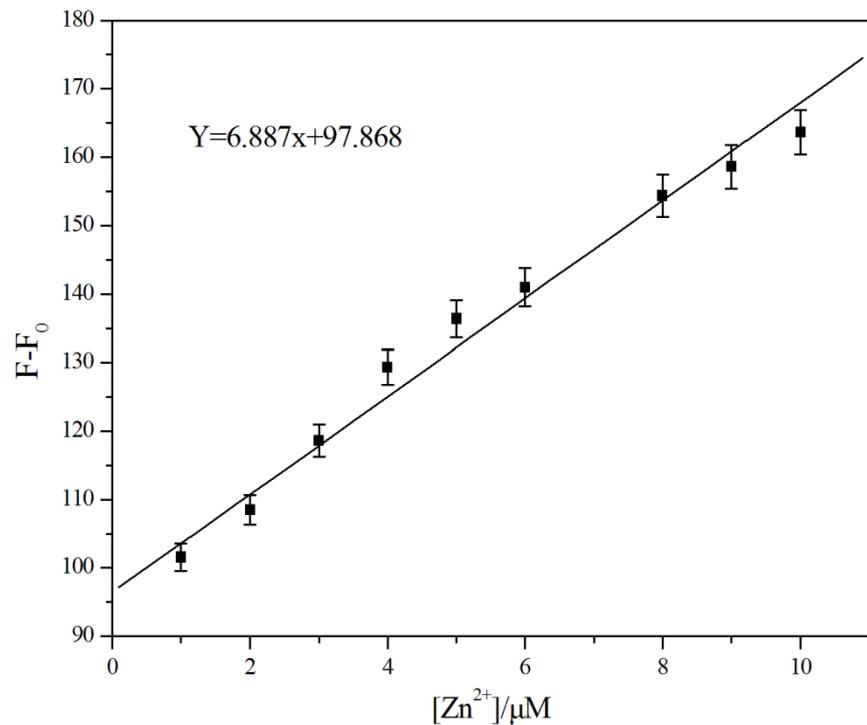


Fig. S2 Relative fluorescence intensity of HAQT at different concentrations of Zn²⁺

added in buffered CH₃OH-H₂O solution (4:1, v/v, Tris-HCl pH 7.40).

5. The effect of EDTA on HAQT-Zn²⁺ complexes

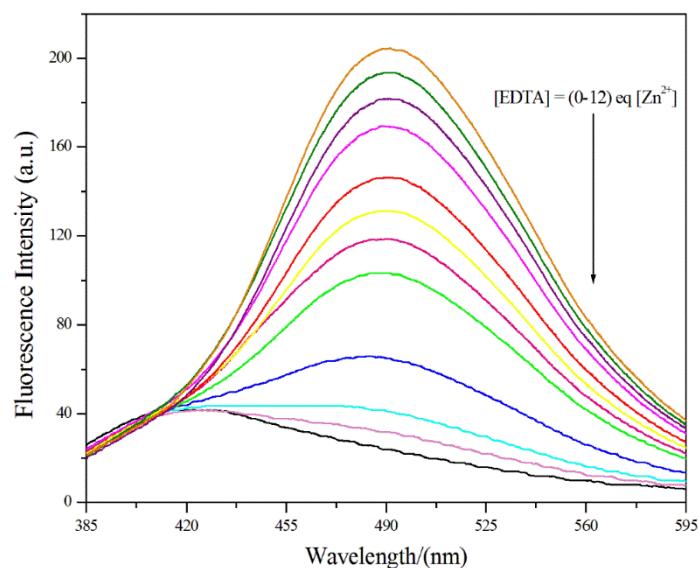


Fig. S3 The fluorescence spectra of HAQT+Zn²⁺ mixture upon gradual addition of EDTA showed the reversible binding nature of Zn²⁺ with HAQT.

6. The FTIR spectra HAQT-Zn²⁺ complexes

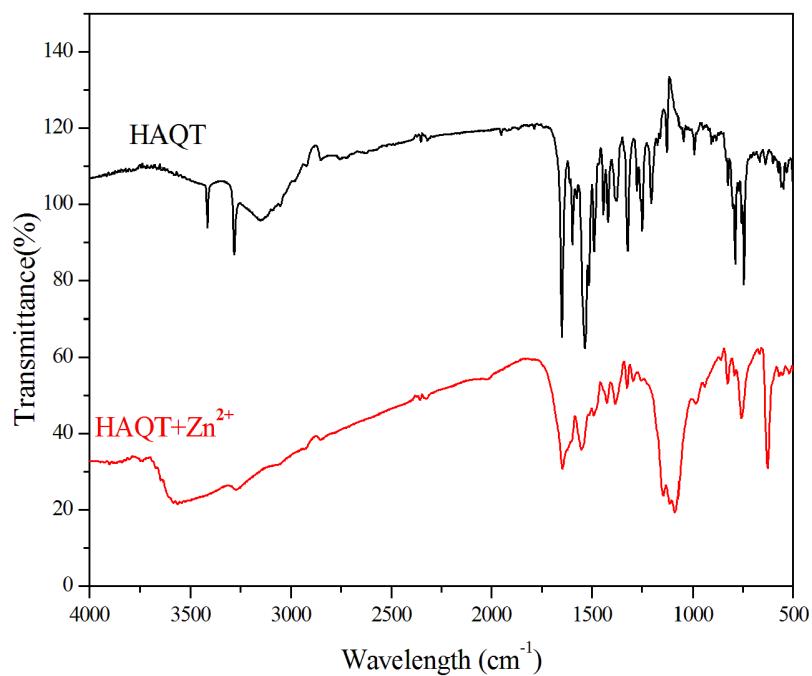
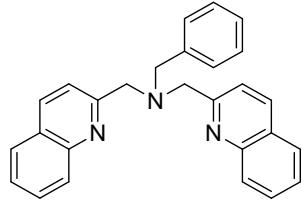
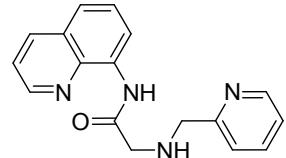
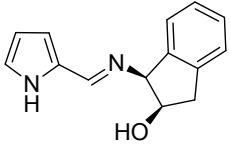
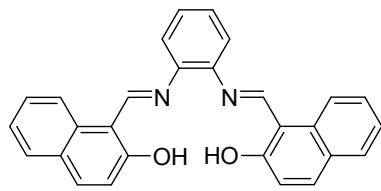
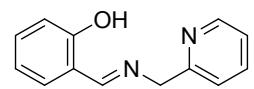
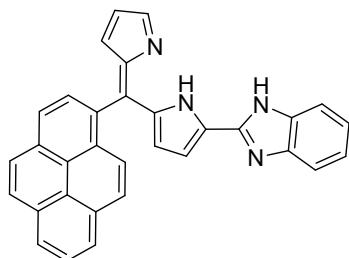


Fig. S4 The FT-IR spectrum of the chemosensor HAQT (black line) and the HAQT-Zn²⁺ complexes (red line).

7. A comparision to other reported chemosensors for detecting Zn²⁺

Table S1 A comparision to other reported chemosensors for the detection of Zn²⁺ based on quinoline and other fluorophores.

Fluorescence chemosensor	Analytes	K _a /M ⁻¹	LOD/μM	Ref.
	Zn ²⁺	3.31×10 ⁴	1.2	1
	Zn ²⁺	1.64×10 ⁵	3.2	2
	Zn ²⁺	3.00×10 ⁶	1.0	3
	Zn ²⁺	9.12×10 ⁶	0.15	4
	Zn ²⁺	8.99×10 ⁵	9.0	5
	Zn ²⁺	6.12×10 ⁴	0.24	6

8. Supporting references

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