Electronic Supplementary Material (ESI) for ChemComm. This journal is © The Royal Society of Chemistry 2015

**Supporting Information:** 

Facile Synthesis of nanostructured TiNb<sub>2</sub>O<sub>7</sub> anode materials

with superior performance for high-rate lithium ion

**batteries** 

Shuaifeng Lou<sup>a</sup>, Yulin Ma<sup>a,b</sup>, Xinqun Cheng<sup>a</sup>, Jinlong Gao<sup>a</sup>, Yunzhi Gao<sup>a</sup>, Pengjian Zuo<sup>a</sup>,

Chunyu Du<sup>a</sup> and Geping Yin<sup>a</sup>\*

a Institute of Advanced Chemical Power Sources, School of Chemical Engineering and Technology, Harbin

Institute of Technology, Harbin 150001, China.

b Energy Storage & Distributed Resources Division, Lawrence Berkeley National laboratory, Berkeley CA 94720,

USA

E-mail: yingeping@hit.edu.cn

Tel: +86-0451-86403807

**Experimental Section** 

**Preparation of TiNb<sub>2</sub>O<sub>7</sub>:** TiNb<sub>2</sub>O<sub>7</sub> nanorods were prepared through a simple sol-gel method. First,

Nb<sub>2</sub>O<sub>5</sub> (Sinopharm Chemical Reagent, analytically pure) was added to a hydrofluoric acid solution

(Sinopharm Chemical Reagent, 40 wt.%) and stirred for about 4 hours at 70°C. After the solution

became transparent, excess ammonium hydroxide was added and then a white Nb(OH)<sub>5</sub>

precipitated. The solid was filtered, washed and dried. 1.78g of the obtained Nb(OH)<sub>5</sub> was added

in an oxalic acid solution at 80°C to form transparent solution including Nb5+, with the pH value

adjusted by controlling the amount of oxalic acid. Meanwhile, 1.7 ml tetrabutyltitanate

(Sinopharm Chemical Reagent, analytically pure) was added into absolute ethanol with a little

oxalic acid and then stirred vigorously. Subsequently, The resulted light yellow solution was

added to the Nb5+-contained oxalic acid solution slowly. The mixed solution was then stirred and

heated at 90°C to completely remove ethanol and water. The obtained white powder was calcined

at 900°C in air to obtain the slightly yellow TiNb<sub>2</sub>O<sub>7</sub> product.

*Materials characterization:* The size and morphology of the TiNb<sub>2</sub>O<sub>7</sub> powders were characterized

using a Helios Nanolab 600i high-resolution Field emission Scanning Electron Microscope (FESEM) operated at 10 kV. X-ray powder diffraction (XRD) was carried out on a Rigaku D/max- $\gamma$ A X-ray diffractometer with Cuk $\alpha$  radiation ( $\lambda$ =1.54178 Å). The FTIR spectra was collected using a ThermoNicolet iS10 FT-IR spectrometer with an attenuated total reflectance(ATR) unit. Transmission electron microscopy (TEM) and high resolution TEM (HR-TEM) images were taken with a TecnaiG2F30 transmission electron microscope with an acceleration voltage of 300 kV. X-ray photoelectron spectra (XPS) was recorded on a PHI 5700 ESCA system fitted with Al  $\kappa\alpha$  radiation. The energy calibration and deconvolution of XPS peaks were calculated using a XPSPEAK software.

Electrode fabrication and electrochemical measurements: The working electrodes were fabricated as follows. First, the TiNb<sub>2</sub>O<sub>7</sub> nanostructured materials (70 wt.%), acetylene black (20 wt.%), and polyvinylidene fluoride (PVDF, 10 wt.%) were mixed and then stirred in N-methyl-2-pyrrolidone (NMP) to form a black slurry. Then, the black slurry was spread onto a Cu foil by a doctor blade method, followed by drying in vacuum at 120°C for 12 h. The average loading weight of the active material in the coin cell fabrication is ~2 mg. Lithium metal was used as both the counter electrode and the reference electrode, and a porous polypropylene membrane (Celgard 2500) was used as the separator. CR2025-type coin cells were assembled in an argon-filled glove box with both moisture and oxygen contents below 0.5 ppm. The electrolyte was 1M LiPF<sub>6</sub> in ethylene carbonate/dimethyl carbonate/ethyl methyl carbonate (EC/DMC/EMC, 1:1:1 by volume). The galvanostatic charge/discharge test (1.0-3.0V) was performed on a Neware-CT3008 test system under ambient temperature.

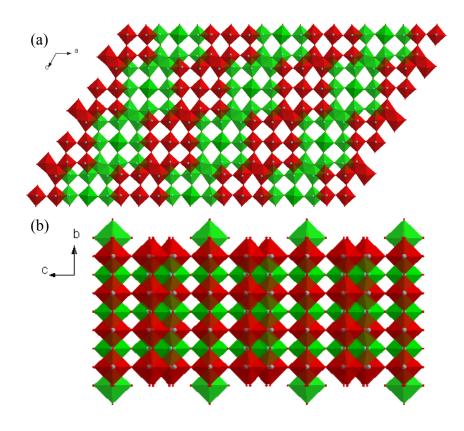


Fig. S1 Crystal structure of  $TiNb_2O_7$  oxide with a  $ReO_3$ -type structure, (a) ac plane and (b) bc plane.

The red blocks and green blocks represent the  $MO_6$  (M=Ti, Nb) octahedra (TiO<sub>6</sub> and NbO<sub>6</sub> octahedra show a random arrangement) on the adjoining layers, and each of these blocks contains nine  $MO_6$  (3×3) octahedra.

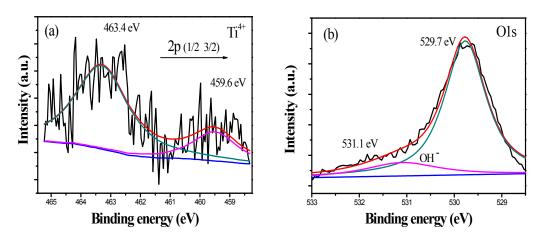


Fig. S2 XPS spectra of Ti and O elements in nanostructured TiNb<sub>2</sub>O<sub>7</sub>.

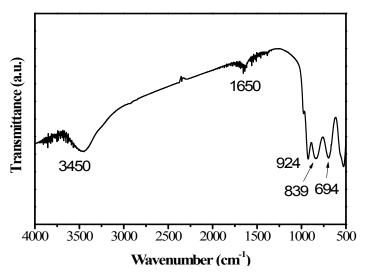


Fig. S3 FTIR of TiNb<sub>2</sub>O<sub>7</sub> nanostructured materials.

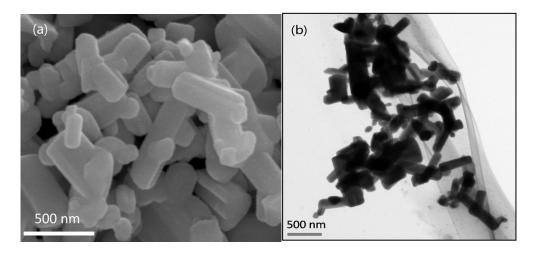


Fig. S4 (a) The SEM and (b) TEM image of TiNb<sub>2</sub>O<sub>7</sub> nanostructured materials.

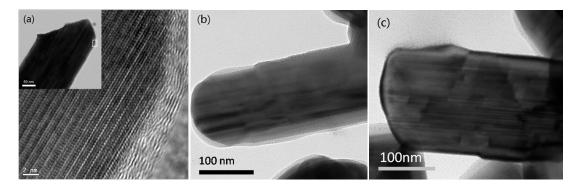


Fig. S5 The orientation of TiNb<sub>2</sub>O<sub>7</sub> nanorod materials.

The structural orientation shows a consistency results that crystal lattice fringes are parallel to the axial direction of the  $TiNb_2O_7$  rod. Fig. S5(a) is the enlarge image of the box in the inset image. Fig. S5 (b) and (c) are the representational TEM images in the products.

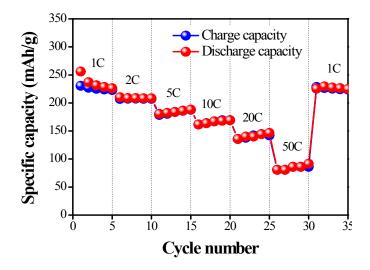


Fig. S6 Rate performance of  $TiNb_2O_7$  electrodes at various rates (1 C, 2 C, 5 C, 10 C, 20 C, 50 C)