Solvent modulated assembly of two Zn Metal-Organic Frameworks: syntheses, luminescent, and gas adsorption properties

Liangliang Zhang^{†,}*, Yu Zhang[†], Rongming Wang[†], Zixi Kang[†], Xiaobin Liu[†], Daofeng Sun[†] and Qingguo Meng^{‡,}*

[†]College of Science, China University of Petroleum (East China), Qingdao, Shandong, 266580, People's Republic of China.

[‡]Chemistry & Chemical and Environmental Engineering College, Weifang University, Weifang 261061, People's Republic of China.

complex 1				
Zn1—O2	2.1185 (13)	Zn1—O4	2.3708 (15)	
Zn1—N4	2.2099 (16)	Zn1—N6	2.1851 (15)	
Zn1—N9	2.0941 (14)	Zn1—N10	2.0855 (13)	
02—Zn1—O4	58.36 (5)	O2—Zn1—N4	150.85 (5)	
O2—Zn1—N6	88.02 (5)	N4—Zn1—O4	97.48 (5)	
N6—Zn1—O4	88.59 (5)	N6—Zn1—N4	74.31 (6)	
N9—Zn1—O2	100.40 (5)	N9—Zn1—O4	88.24 (5)	
N9—Zn1—N4	94.32 (6)	N9—Zn1—N6	89.16(11)	
N16—Zn1—N7 ii	103.24(9)	N27—Zn3—N12	167.68 (6)	
N10—Zn1—O2	97.53 (5)	N10—Zn1—O4	155.72 (5)	
N10—Zn1—N4	106.42 (6)	N10—Zn1—N6	93.70 (5)	
N10—Zn1—N9	94.08 (5)			
Symmetry codes: (i) $-x+1/2$, $y-1/2$, z ; (ii) $-x+1/2$, $y+1/2$, z ; (iii) $x-1/2$, y , $-z+1/2$; (iv) $x+1/2$, y , $-z+1/2$				

Table S1. Selected bond length (Å) and angles (°) for complex 1.

Table S2. Selected bond length (\AA) and angles $(^{\circ})$ for complex 2.

complex 2				
Zn1—01	1.910 (2)	Zn3—N12	2.026(2)	
Zn1—N3 ⁱ	1.969(2)	Zn3—O1W	2.087(3)	
Zn1—N7 ⁱⁱ	2.018(2)	Zn3—N23	2.078(2)	
Zn1—N16	1.988(2)	Zn3—N27	2.189(3)	
Zn3—O4 ³	1.959(2)			
N3 ⁱ —Zn1—O1	115.11(9)	01W—Zn3—N12	98.12(10)	
N7 ⁱⁱ —Zn1—O1	97.05(8)	N23—Zn3—O4 ⁱⁱⁱ	152.85(10)	
N7 ⁱⁱ —Zn1—N3 ⁱ	115.26(8)	N23—Zn3—N12	103.18(10)	
N16—Zn1—O1	117.59(9)	N23—Zn3—O1W	91.78(10)	
N16—Zn1—N3 ⁱ	107.87(9)	N27—Zn3—O4 ⁱⁱⁱ	89.16(11)	
N16—Zn1—N7 ⁱⁱ	103.24(9)	N27—Zn3—N12	101.85(10)	
N12—Zn3—O4 ⁱⁱⁱ	102.01(9)	N27—Zn3—O1W	158.46(10)	
O1W—Zn3—O4 ⁱⁱⁱ	94.53(11)	N27—Zn3—N23	75.94(11)	
Symmetry codes: (i) -1/2+X,+Y,1/2-Z; (ii)1/2-X,1/2+Y,+Z; (iii)-X,-Y,-Z.				

Table S3.	The	photop	hysical	data for	complex	1, 2	and	2 a.
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Complex	τ, ns, (Weight)	χ^2	$\Phi(\%)$
1	$\tau_1 = 1.29(70.12\%); \tau_2 = 4.78(26.37\%); \tau_3 = 23.45(3.50\%)$	1.218	2.72
2	$\tau_1 = 0.60(61.88\%); \tau_2 = 2.66(27.71\%); \tau_3 = 27.01(10.41\%)$	1.380	2.85
2a	$\tau_1 = 0.83(73.13\%); \tau_2 = 3.03(20.53\%); \tau_3 = 20.79(6.35\%)$	1.287	4.35



Figure S1. The simulated and experimental PXRD of complex 1.



Figure S2. The simulated and experimental PXRD of complex 2 and experimental PXRD of



Figure S3. The TGA curves of complexes 1, 2 and 2a.



Figure S4. The Ar sorption isotherms at 77K for 2a.



Figure S5. The CO₂ sorption isotherms at 273 and 295K for 2a.



Figure S6. The IR spectrum of complex 1.



Figure S7. The IR spectrum of complex 2.



Figure S8. The IR spectrum of complex 2a.



Figure S9. The photoluminescence excitation spectrum of H₂btca at 298 K



Figure S10. The photoluminescence excitation spectrum of complex 1 at 298 K



Figure S11. The photoluminescence excitation spectrum of complex 2 at 298 K



Figure S12. The photoluminescence excitation spectrum of complex 2a at 298 K