

Photo- and Electro-luminescence of cuprous complexes with Sterically Demanding and Hole Transmitting Diimine Ligands

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Experimental section

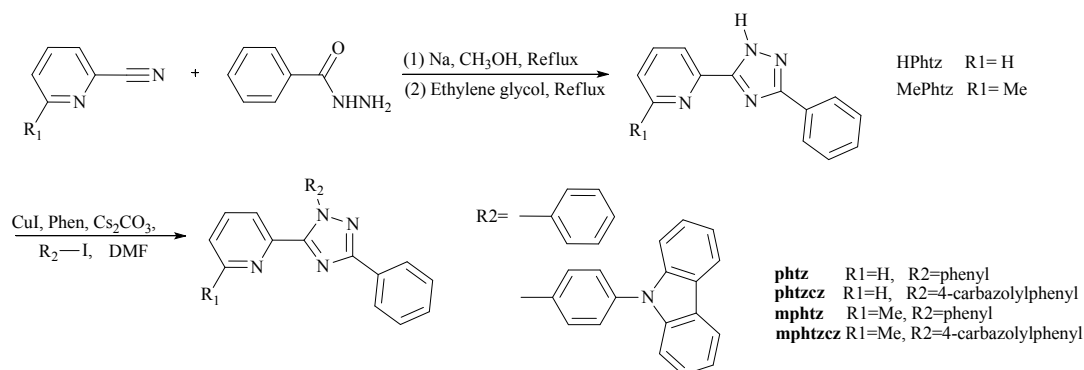


Fig. S1 The simplified synthetic route for functionalized diimine ligands.

Table S1 Crystallographic data and Structural Refinements for **1–4**

	1	2	3	4
Formula	C ₅₅ H ₄₂ BCuF ₄ N ₄ OP ₂	C ₆₈ H ₅₂ BCuF ₄ N ₅ O _{1.5} P ₂	C ₅₇ H ₄₆ BCl ₂ CuF ₄ N ₄ OP ₂	C ₆₈ H ₅₁ BCuF ₄ N ₅ OP ₂
<i>M_r</i> (g mol ⁻¹)	987.22	1175.44	1086.17	1166.43
Space group	C2/c	P2 ₁ /c	P-1	P2 ₁ /c
<i>a</i> /Å	20.846(5)	22.9549(10)	11.2117(6)	12.527(4)
<i>b</i> /Å	17.594(5)	21.3720(15)	11.6277(8)	20.392(6)
<i>c</i> /Å	29.817(7)	24.8432(9)	20.2682(12)	22.743(7)
α /°	90	90	92.924(5)	90
β /°	90.721(5)	111.602(4)	90.208(5)	101.445(5)
γ /°	90	90	102.716(5)	90
<i>V</i> /Å ³	10935(5)	11331.8(10)	2573.8(3)	5695(3)
<i>Z</i>	8	8	2	4
<i>D_c</i> /g cm ⁻³	1.199	1.378	1.402	1.361
μ /mm ⁻¹	0.511	1.605	2.634	0.503
<i>F</i> (000)	4064	4856	1116	2408
total reflns	39802	43988	16781	39413
unique reflns	9524	20725	9426	10003
<i>R</i> _{int}	0.0332	0.0663	0.0375	0.0442
GOF	1.162	0.865	0.998	1.050
<i>R</i> ₁ ^a [<i>I</i> > 2σ(<i>I</i>)]	0.0946	0.0652	0.0757	0.0852
<i>wR</i> ₂ ^b (all data)	0.3088	0.1956	0.2241	0.2470
CCDC	1034650	1034651	1034652	1034653

^a $R_1 = \sum(F_o - F_c) / \sum F_o$; ^b $wR_2 = [\sum w(F_o^2 - F_c^2)^2 / \sum w(F_o^2)^2]^{1/2}$.

Table S2 Selected bond lengths (Å) and angles (°) for **1–4**.

1			
Cu(1)–N(4)	2.127(4)	Cu(1)–P(2)	2.2706(14)
Cu(1)–N(3)	2.156(4)	Cu(1)–P(3)	2.2967(15)
N(4)–Cu(1)–N(3)	79.27(16)	N(4)–Cu(1)–P(3)	110.32(13)
N(4)–Cu(1)–P(2)	119.13(13)	N(3)–Cu(1)–P(3)	106.36(12)
N(3)–Cu(1)–P(2)	122.64(12)	P(2)–Cu(1)–P(3)	114.12(6)
2			
Cu(1)–N(4)	2.067(3)	Cu(2)–N(8)	2.082(3)
Cu(1)–N(3)	2.110(3)	Cu(2)–N(7)	2.080(3)
Cu(1)–P(2)	2.2453(11)	Cu(2)–P(4)	2.2235(11)
Cu(1)–P(3)	2.2872(11)	Cu(2)–P(5)	2.2969(11)
N(4)–Cu(1)–N(3)	79.60(11)	N(8)–Cu(2)–N(7)	79.71(12)
N(4)–Cu(1)–P(2)	118.07(10)	N(8)–Cu(2)–P(4)	122.18(10)
N(3)–Cu(1)–P(2)	120.47(9)	N(7)–Cu(2)–P(4)	127.44(9)
N(4)–Cu(1)–P(3)	109.45(9)	N(8)–Cu(2)–P(5)	100.59(9)
N(3)–Cu(1)–P(3)	110.29(9)	N(7)–Cu(2)–P(5)	105.73(9)
P(2)–Cu(1)–P(3)	114.18(4)	P(4)–Cu(2)–P(5)	114.28(4)
3			
Cu(1)–N(4)	2.093(3)	Cu(1)–P(3)	2.2463(10)
Cu(1)–N(3)	2.117(3)	Cu(1)–P(2)	2.3010(12)
N(4)–Cu(1)–N(3)	80.70(12)	N(4)–Cu(1)–P(2)	104.62(9)
N(4)–Cu(1)–P(3)	118.79(9)	N(3)–Cu(1)–P(2)	108.68(9)
N(3)–Cu(1)–P(3)	121.78(9)	P(3)–Cu(1)–P(2)	116.29(4)
4			
Cu(1)–N(3)	2.085(4)	Cu(1)–P(2)	2.2470(14)
Cu(1)–N(4)	2.118(4)	Cu(1)–P(3)	2.3148(15)
N(3)–Cu(1)–N(4)	80.62(15)	N(3)–Cu(1)–P(3)	106.64(12)
N(3)–Cu(1)–P(2)	120.44(12)	N(4)–Cu(1)–P(3)	102.28(11)
N(4)–Cu(1)–P(2)	121.71(11)	P(2)–Cu(1)–P(3)	118.27(5)

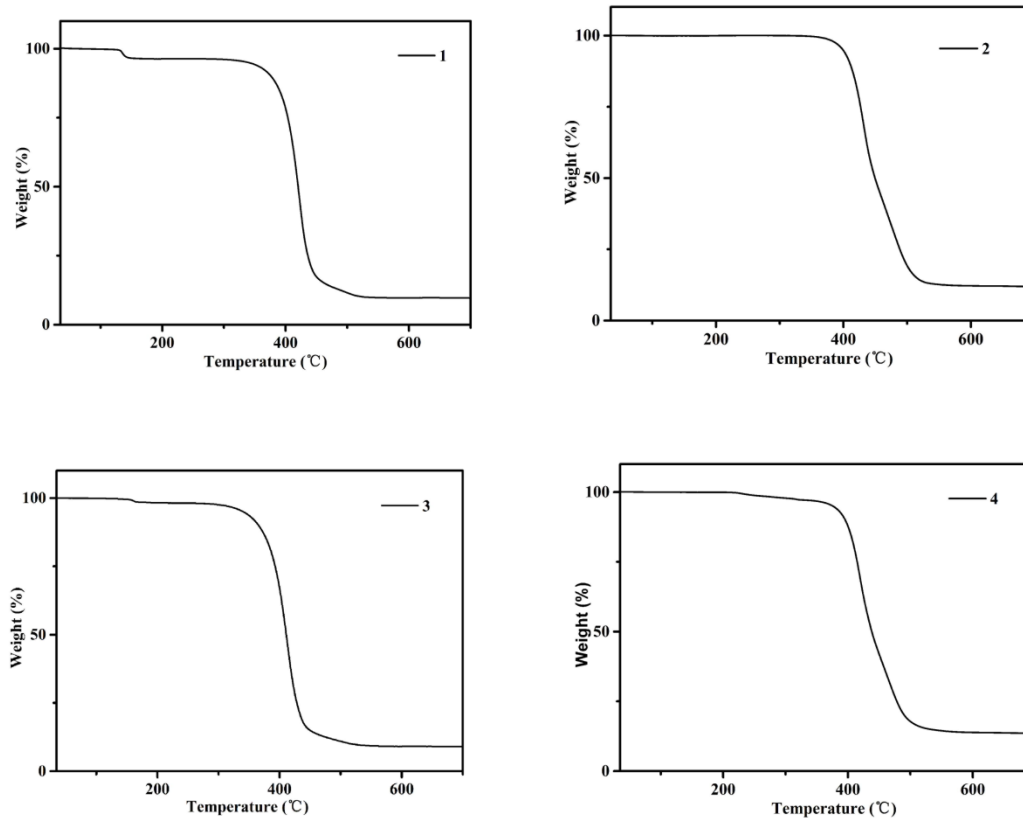


Fig. S2 TGA curves for 1–4.

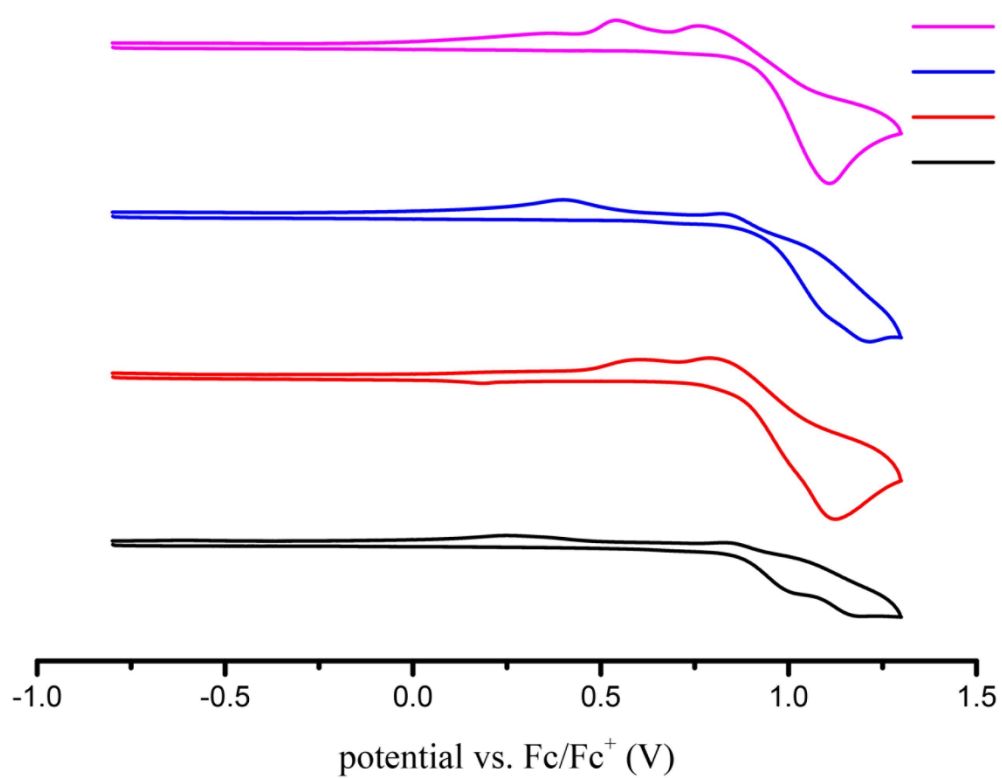


Fig. S3 Oxidation waves of **1-4** in CH_2Cl_2 solution at room temperature. Scan rate 100mVs^{-1} in 0.1M TBAP.

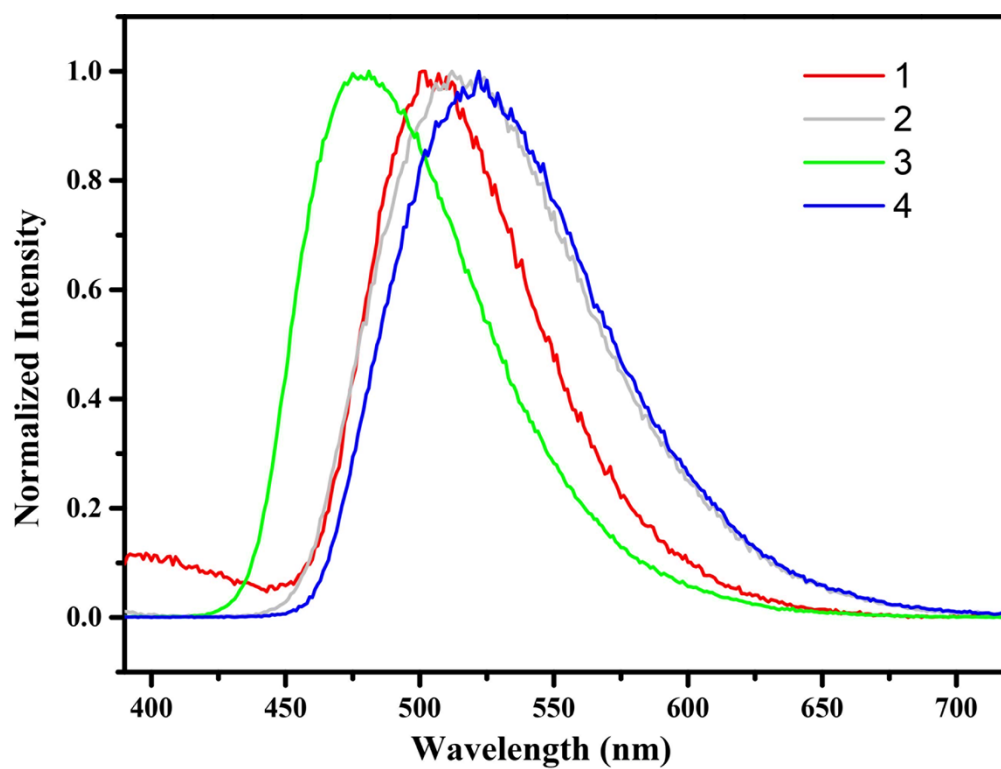


Fig. S4 Emission spectra of **1-4** in the solid state.

Table S3 PL properties in the crystal powder of **1–4**.

	$\lambda_{\text{em}}(\text{nm})$	$\tau_{\text{ave}}(\mu\text{s})$	$\Phi(\%)$
1	503	25	14
2	513	12	59
3	477	26	80
4	522	17	36

Table S4. The calculated HOMO from the oxidation curves of cyclic voltammetry, the ΔE_{gap} calculated from onset of the absorption spectra and the energy level of LUMO obtained by add the values of ΔE_{gap} to values of HOMO.

	HOMO	LUMO	ΔE_{gap}
1	-5.63ev	-2.85ev	2.78
2	-5.63ev	-2.86ev	2.77
3	-5.66ev	-2.81ev	2.85
4	-5.66ev	-2.81ev	2.85

Appendix 1

