

*Electronic supplementary information (ESI)*

**Synthesis and stabilization of a hypothetical porous framework  
based on a classic flexible metal carboxylate cluster**

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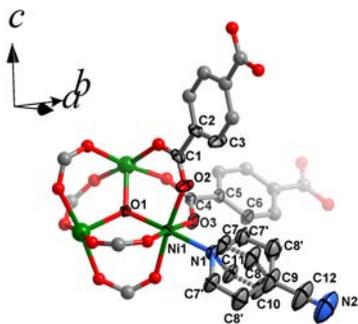
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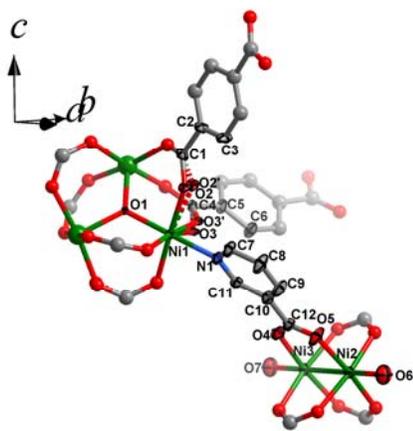
**Table S1.** Crystallographic Data and Structural Refinements.

Complex	<b>1</b>	<b>2</b>	<b>3</b>
Formula	C <sub>84</sub> H <sub>50</sub> N <sub>12</sub> Ni <sub>6</sub> O <sub>26</sub>	C <sub>84</sub> H <sub>51.5</sub> N <sub>6</sub> Ni <sub>9</sub> O <sub>41</sub>	C <sub>84</sub> H <sub>51.5</sub> Fe <sub>7.5</sub> N <sub>6</sub> O <sub>39.5</sub>
Formula weight	1995.62	2329.20	2191.62
Temperature (K)	150	213	213
Crystal system	Cubic	Cubic	Cubic
Space group	<i>Fm-3m</i>	<i>Fm-3m</i>	<i>Fm-3m</i>
<i>a</i> /Å	41.5216(3)	41.286(2)	41.258(3)
<i>V</i> /Å <sup>3</sup>	71585.0(16)	70375(12)	70230(15)
<i>Z</i>	16	16	16
$\rho_{\text{calc}}$ (g cm <sup>-3</sup> )	0.741	0.879	0.829
$\mu$ (mm <sup>-1</sup> )	1.031	1.453	1.273
$R_1^a$ ( <i>I</i> > 2 $\sigma$ )	0.0656	0.0628	0.1120
$wR_2^b$ (all data)	0.2421	0.2235	0.3774
GOOF	1.036	1.001	1.065

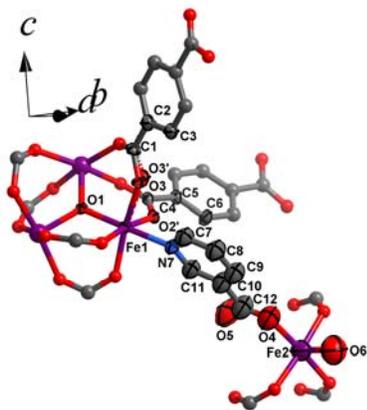
<sup>a</sup> $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ . <sup>b</sup> $wR_2 = [\sum w(F_o^2 - F_c^2)^2 / \sum w(F_o^2)^2]^{1/2}$ .



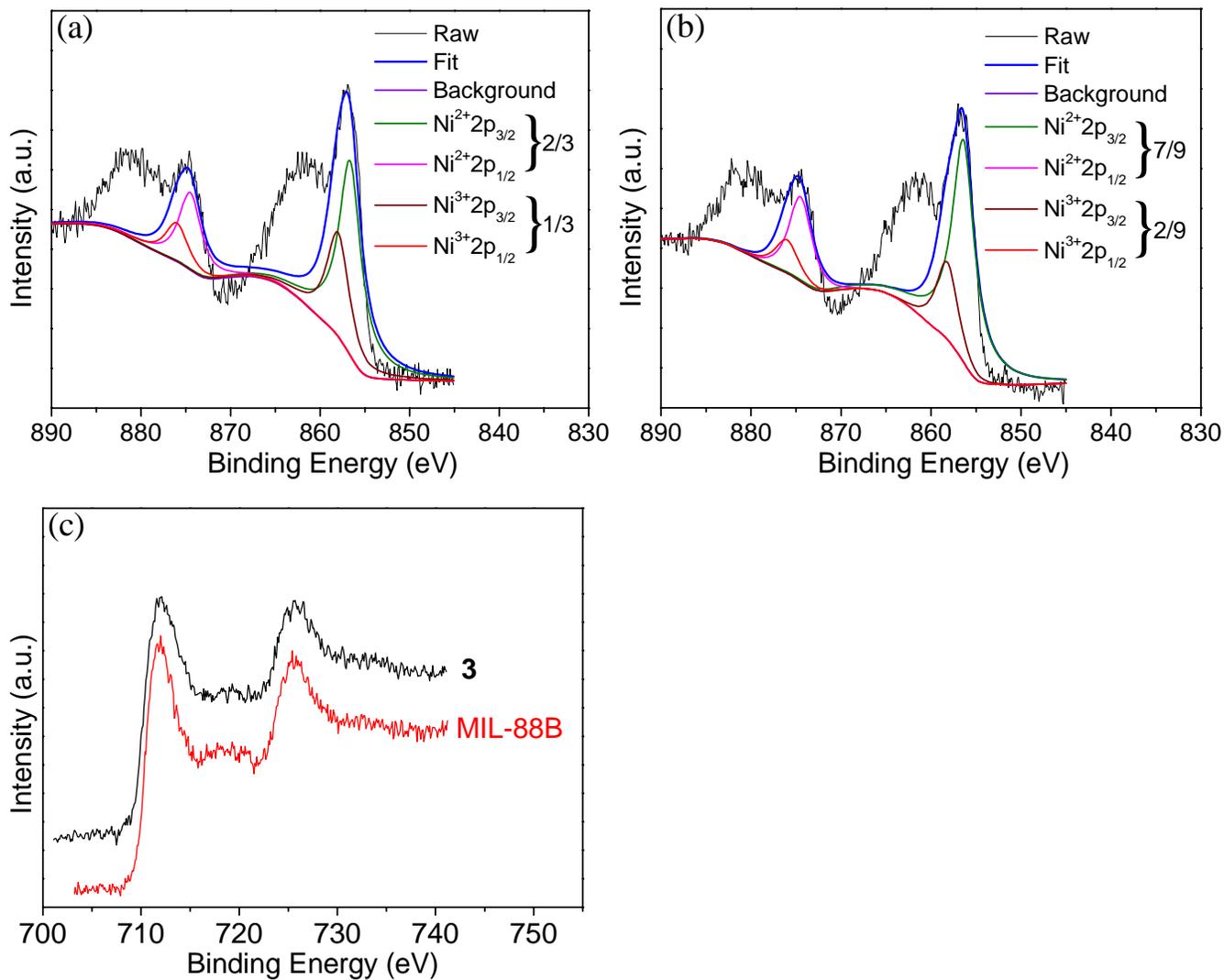
**Figure S1(a).** Perspective views of the coordination environments in **1**. Hydrogen atoms are omitted for clarity (Thermal ellipsoids are drawn for the asymmetric units with probability 30%). Dashed bonds represent another part of the 2-fold disordered pyridyl group.



**Fig. S1(b).** Perspective views of the coordination environments in **2**. Hydrogen atoms are omitted for clarity (Thermal ellipsoids are drawn for the asymmetric units with probability 30%). Dashed bonds represent another part of the 2-fold disordered carboxylate groups.



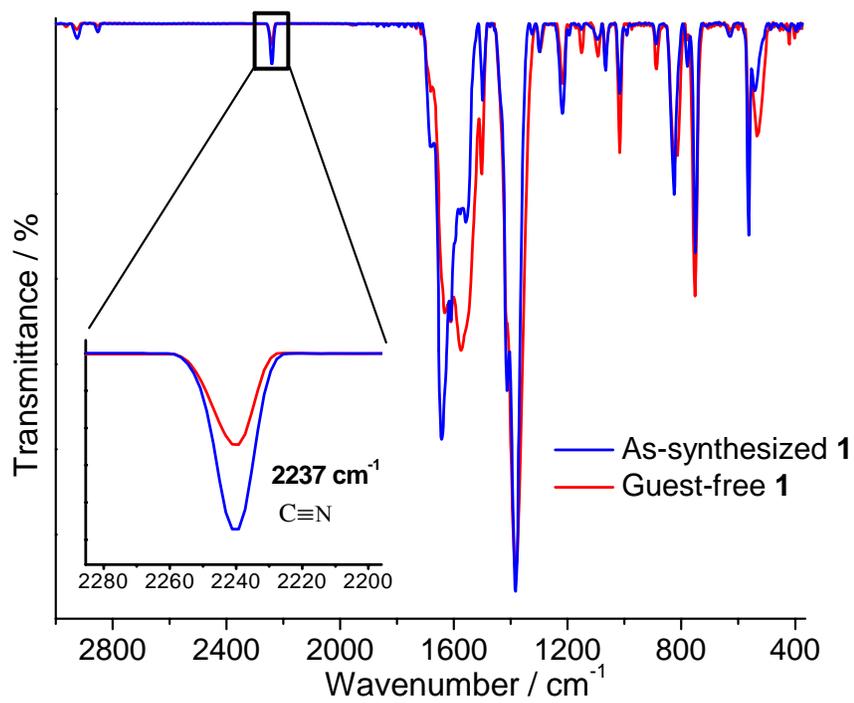
**Fig. S1(c).** Perspective views of the coordination environments in **3**. Hydrogen atoms are omitted for clarity (Thermal ellipsoids are drawn for the asymmetric units with probability 20%). Dashed bonds represent another part of the 2-fold disordered carboxyl groups.



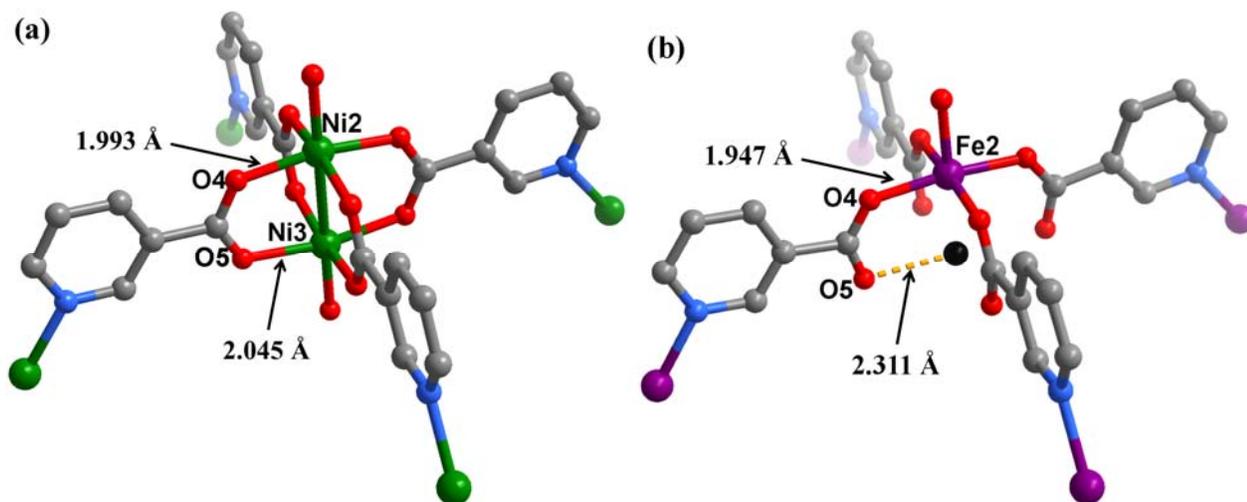
**Fig. S2.** XPS spectra of (a) **1**, (b) **2** and (c) **3**.

**Table S2.** Results of the fitting of the Ni 2p X-ray Photoelectron Spectra.

Compound	Spectral line	Ni(II)		Ni(III)	
		2p <sub>1/2</sub>	2p <sub>3/2</sub>	2p <sub>1/2</sub>	2p <sub>3/2</sub>
<b>1</b>	Binding Energy (eV)	856.6	874.5	858.0	876.0
	Area (%)	2/3		1/3	
<b>2</b>	Binding Energy (eV)	856.4	874.5	858.2	876.0
	Area (%)	7/9		2/9	



**Fig. S3.** IR spectrum of as-synthesized and guest-free **1**. The characteristic carbonitrile stretching bands of the 4-pyCN at  $2237\text{ cm}^{-1}$  are highlighted, which indicate that this terminal ligand tend to leave the coordination framework during activation.



**Fig. S4.** Local coordination structures of (a)  $\{\text{Ni}_2(\text{na})_4(\text{H}_2\text{O})_2\}$  in **2** and (b)  $\{\text{Fe}(\text{na})_4(\text{H}_2\text{O})\}^-$  in **3**. Note that because the pyridyl ends of the  $\text{na}^-$  ligands in these metalloligands are fixed by coordination with the metal ions in the  $\text{M}_3(\mu_3\text{-O/OH})(\text{bdc})_3$  networks, the four carboxylate ends cannot adopt the ideal  $D_{4d}$  symmetry of the  $\text{M}_2(\text{RCOO})_4$  paddle wheel structure, which can be judged by the unequal bonding distances of Ni2-O4 and Ni3-O5. Because the high-valence cation Fe(III) has a smaller radius than Ni(II), the Fe2-O4 bonds in  $\{\text{Fe}(\text{na})_4(\text{H}_2\text{O})\}^-$  (1.947 Å) are shorter than the Ni2-O4 bonds in  $\{\text{Ni}_2(\text{na})_4(\text{H}_2\text{O})_2\}$  (1.993 Å). To approach the position for binding Fe(III), the  $\text{na}^-$  ligands need to bend toward the 4-fold symmetry axis, which push O5 away from the 4-fold symmetry axis. The shortest bonding distance will be 2.311 Å if four O5 atoms coordinate with a metal ion (the black sphere at the centre of the square defined by four O5 atoms), which is not suitable for an Fe(III) ion.

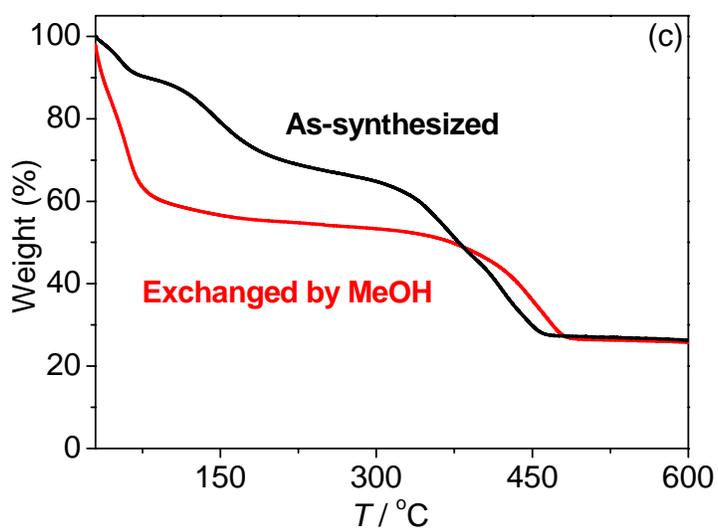
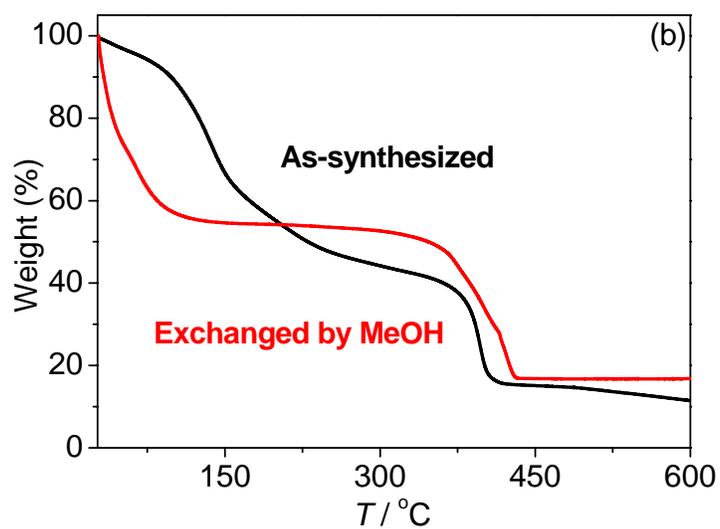
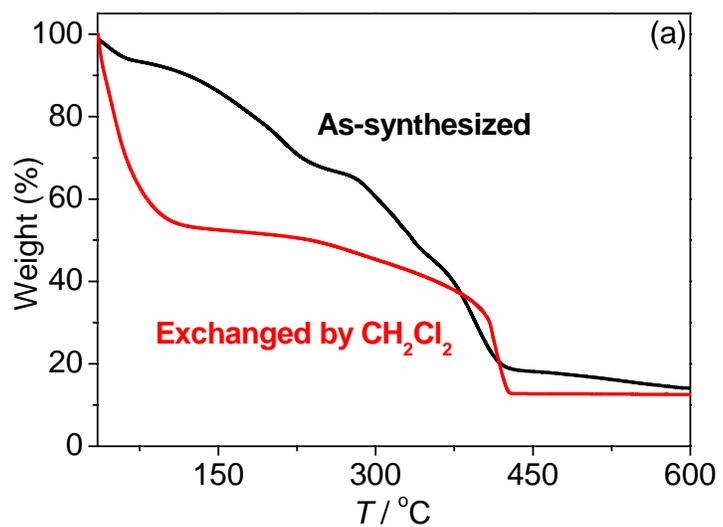
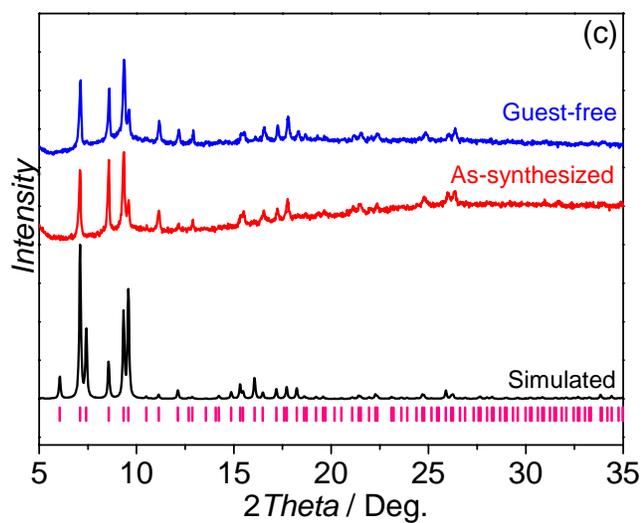
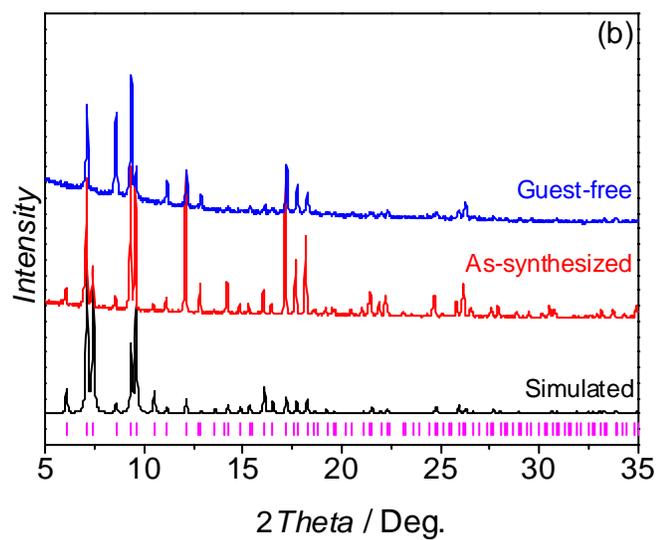
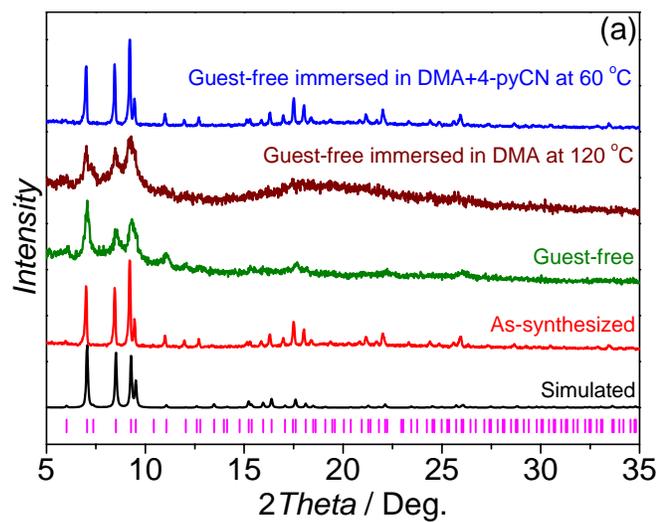
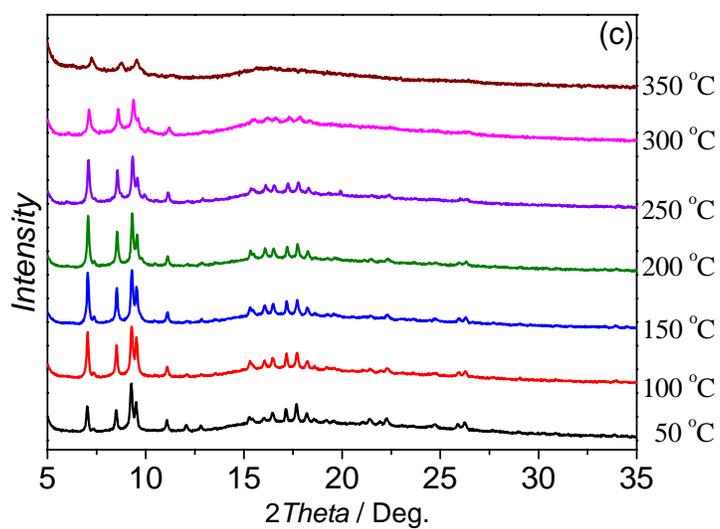
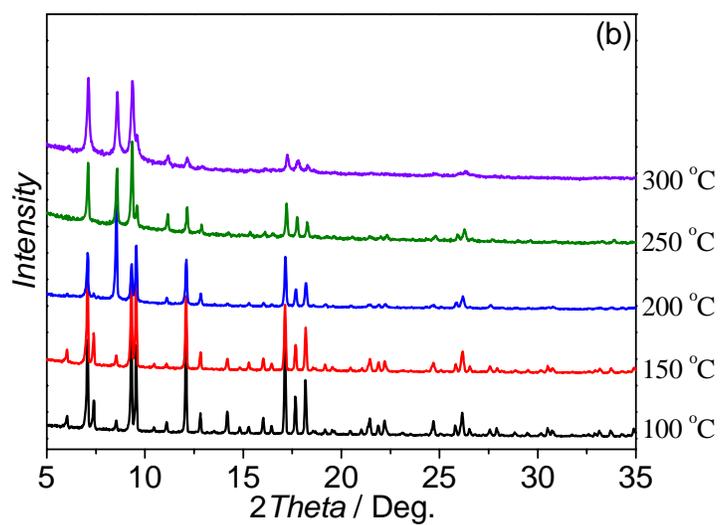
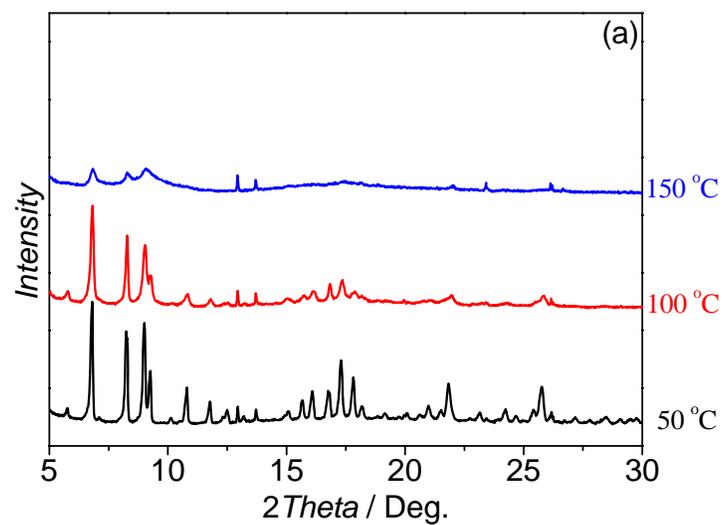


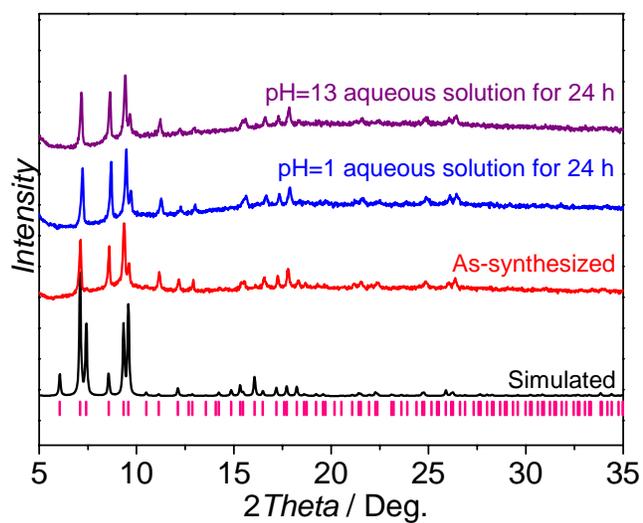
Fig. S5. TGA curves of (a) 1, (b) 2 and (c) 3.



**Fig. S6.** Room-temperature PXRd patterns of (a) **1**, (b) **2** and (c) **3**.



**Fig. S7.** PXR D patterns of (a) **1**, (b) **2**, and (c) **3** after heated at different temperatures under N<sub>2</sub> for 30 min.



**Fig. S8.** Room-temperature PXRD patterns of **3** after different treatments.