

Electronic Supplementary Information

Unsolvated Al(C₆F₅)₃: Structural Features and Electronic Interaction with Ferrocene

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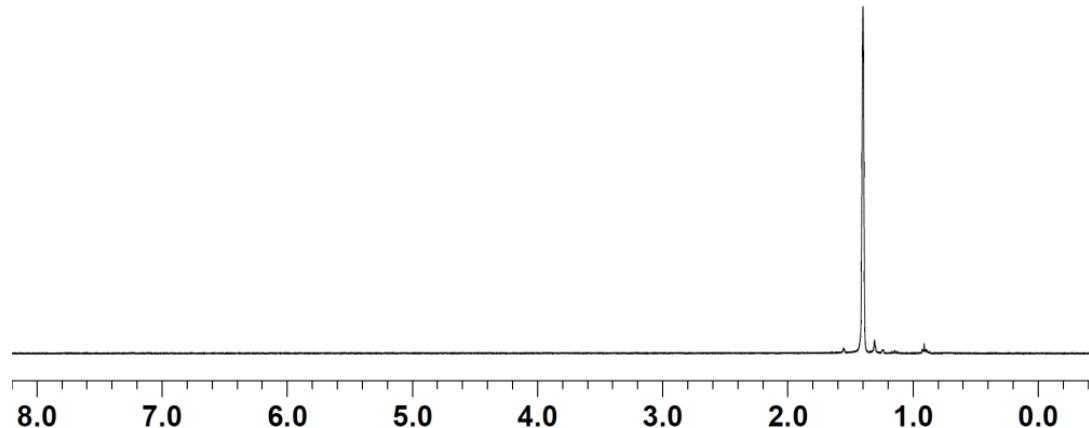


Figure S1. ¹H NMR (C₆D₁₂, 25 °C) spectrum of [Al(C₆F₅)₃]₂.

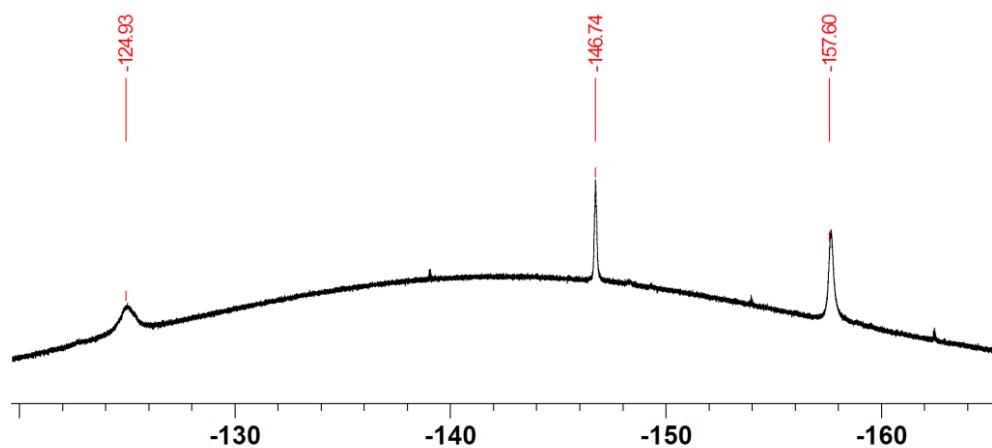


Figure S2. ¹⁹F NMR (C₆D₁₂, 25 °C) spectrum of [Al(C₆F₅)₃]₂. A trace amount of C₆F₅H (-139.2, -154.2, -162.50) was also present.

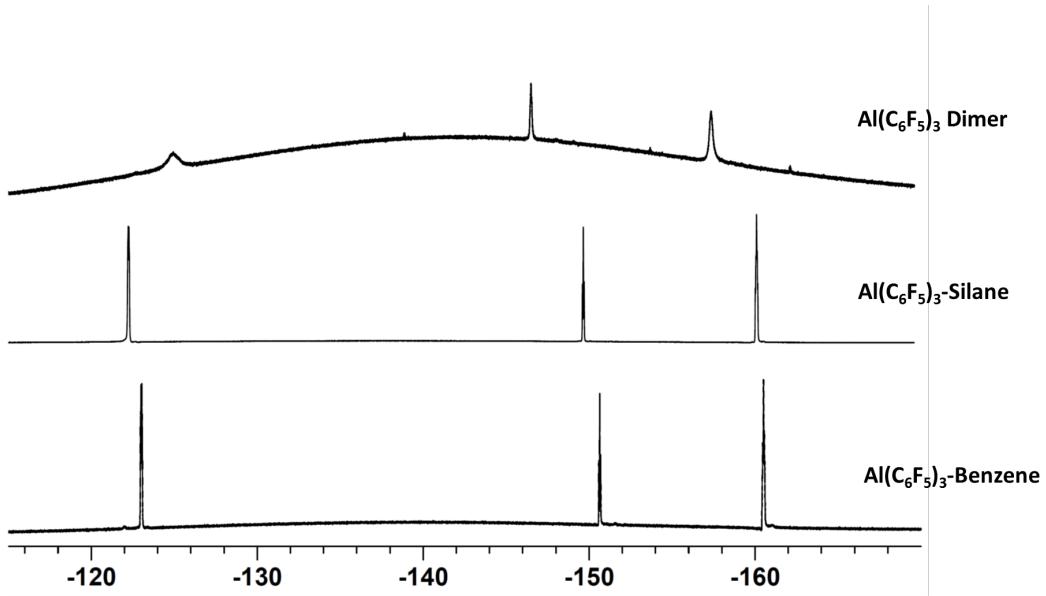


Figure S3. ^{19}F NMR (25°C) spectra of $[\text{Al}(\text{C}_6\text{F}_5)_3]_2$ and $\text{Et}_3\text{SiH}\cdot\text{Al}(\text{C}_6\text{F}_5)_3$ in C_6D_{12} , and $\text{C}_6\text{D}_6\cdot\text{Al}(\text{C}_6\text{F}_5)_3$ in C_6D_6 .

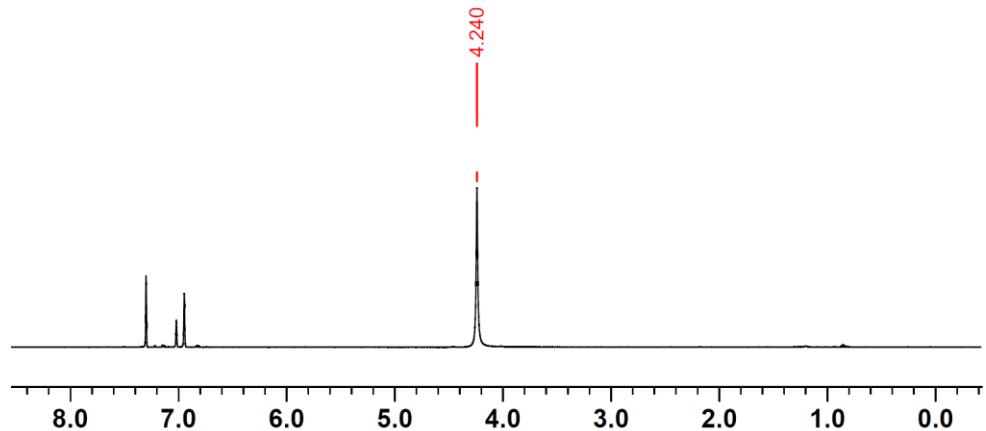


Figure S4. ^1H NMR ($\text{C}_6\text{D}_5\text{Br}$, 25°C) spectrum of $\text{Cp}_2\text{Fe}\cdot\text{Al}(\text{C}_6\text{F}_5)_3$.

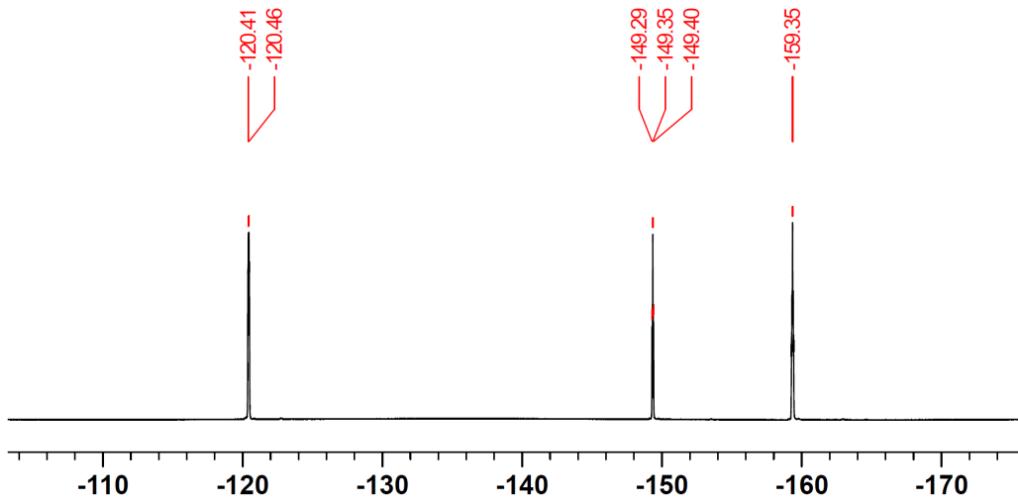


Figure S5. ^{19}F NMR ($\text{C}_6\text{D}_5\text{Br}$, 25 °C) spectrum of $\text{Cp}_2\text{Fe}\cdot\text{Al}(\text{C}_6\text{F}_5)_3$.

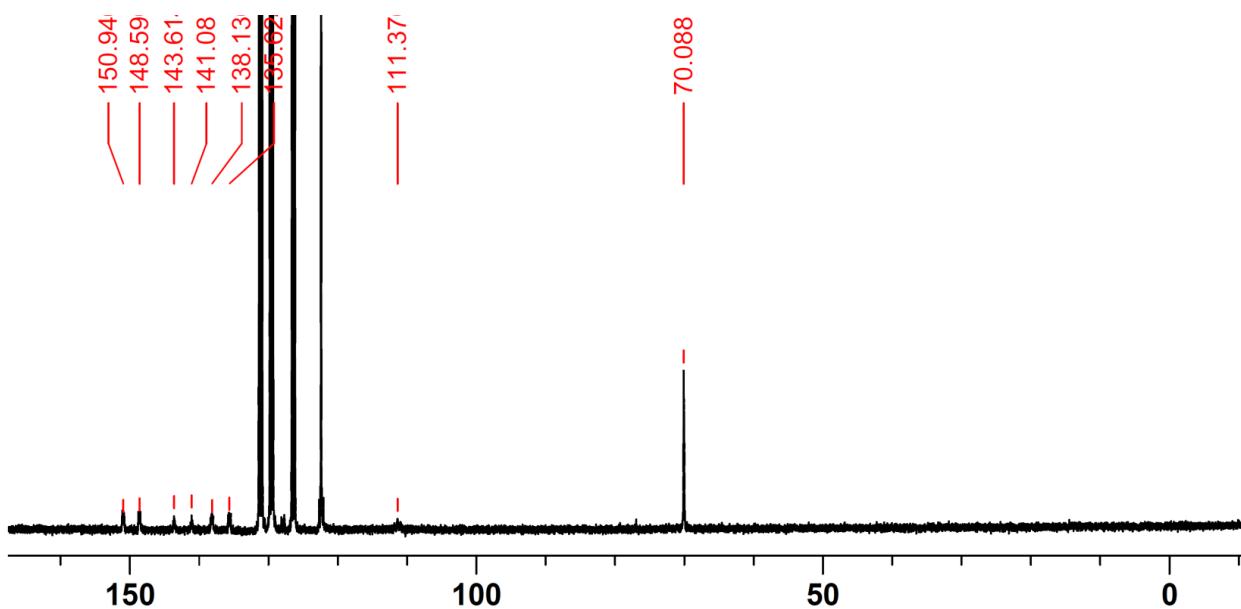


Figure S6. ^{13}C NMR ($\text{C}_6\text{D}_5\text{Br}$, 25 °C) spectrum of $\text{Cp}_2\text{Fe}\cdot\text{Al}(\text{C}_6\text{F}_5)_3$.

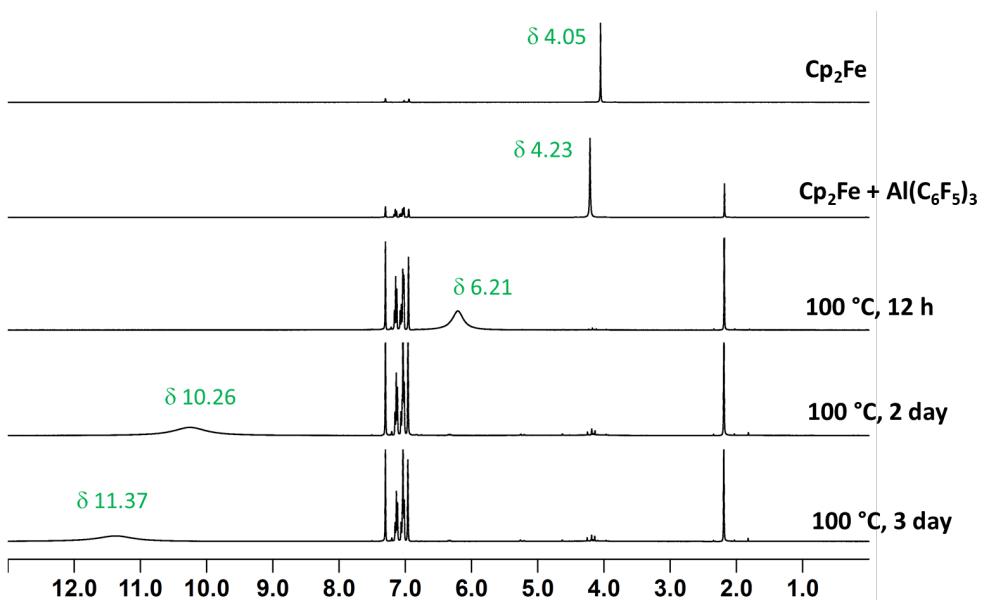


Figure S7. ^1H NMR ($\text{C}_6\text{D}_5\text{Br}$, $25\text{ }^\circ\text{C}$) investigation of coordination and electronic interaction between ferrocene and $\text{Al}(\text{C}_6\text{F}_5)_3$.

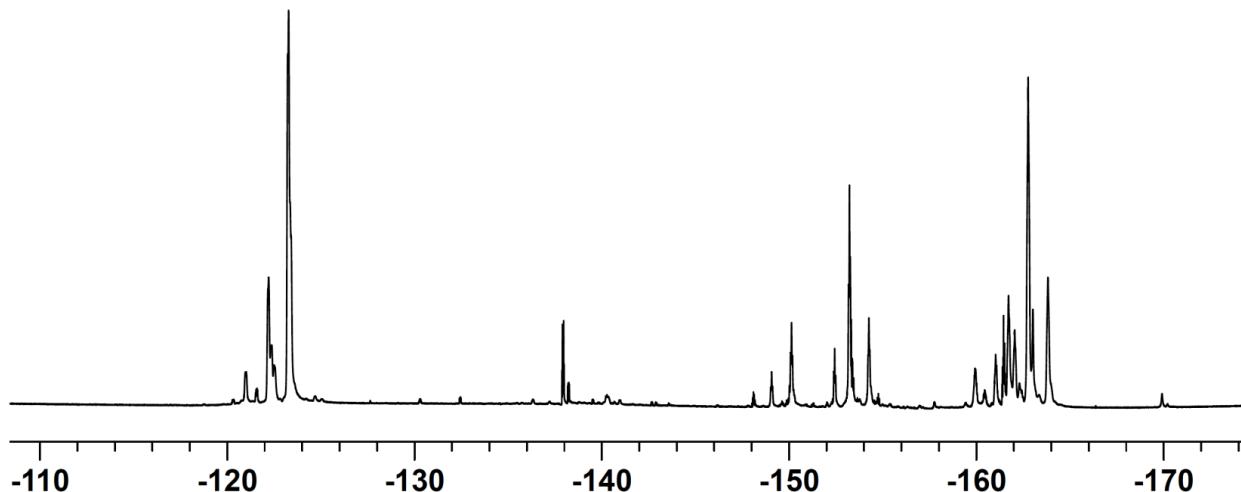


Figure S8. ^{19}F NMR ($\text{C}_6\text{D}_5\text{Br}$, $25\text{ }^\circ\text{C}$) of a mixture of ferrocene and $\text{Al}(\text{C}_6\text{F}_5)_3$ heated at $100\text{ }^\circ\text{C}$ for 3 days, showing the decomposition of $\text{Al}(\text{C}_6\text{F}_5)_3$.

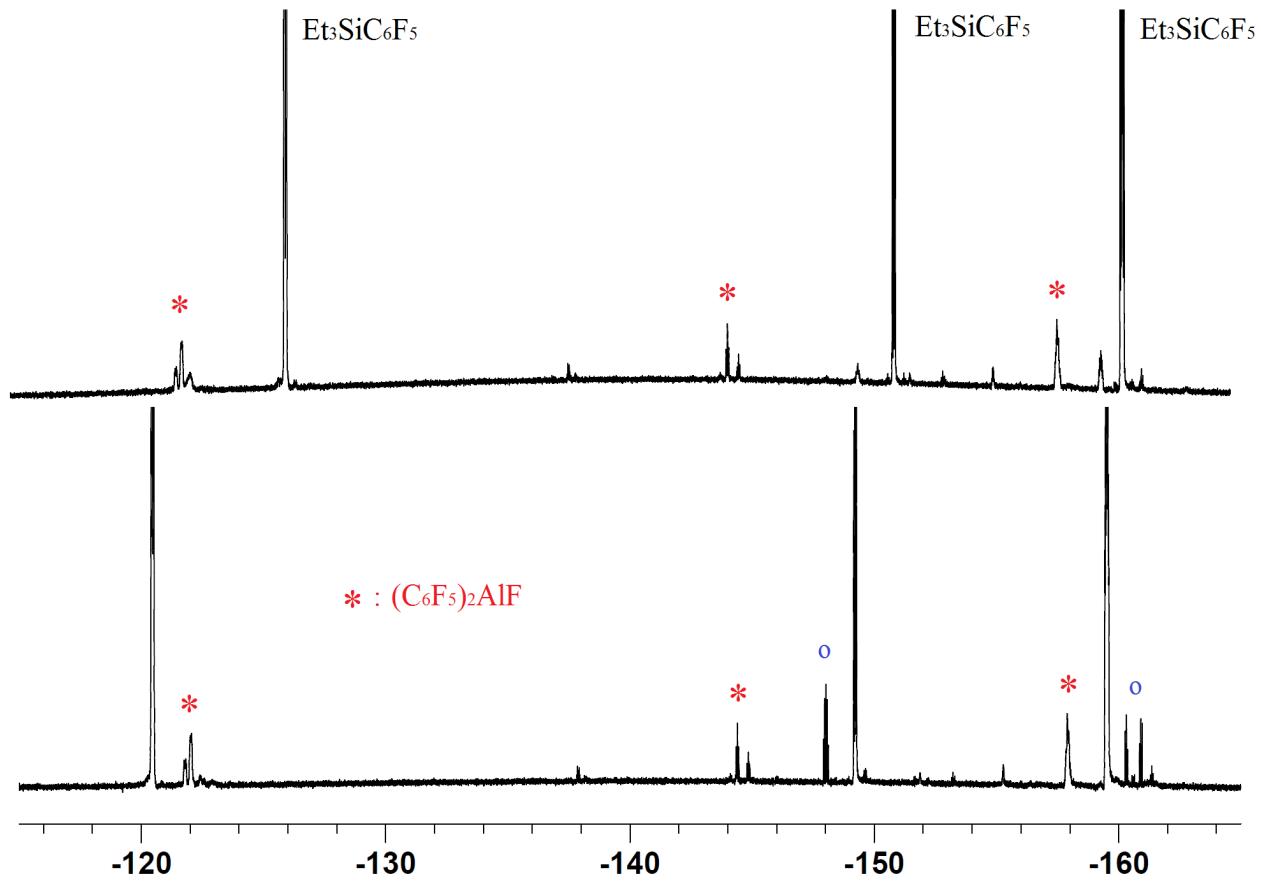


Figure S9. ^{19}F NMR ($\text{C}_6\text{D}_5\text{Br}$, 25 °C) spectra of metathesis between Et_3SiF and $\text{Al}(\text{C}_6\text{F}_5)_3$ (top spectrum, the resulting $(\text{C}_6\text{F}_5)_2\text{AlF}$ has limited solubility hence the stoichiometry in the spectrum was off), and the decomposition (about 10%) of $\text{Al}(\text{C}_6\text{F}_5)_3$ in $\text{C}_6\text{D}_5\text{Br}$ at 100 °C for 3 days (bottom spectrum). Peaks marked as “*” were assigned to $(\text{C}_6\text{F}_5)_2\text{AlF}$, while those with “o” were assigned to species (such as dimer or trimer) presumably derived from the transient tetrafluorobenzyne.

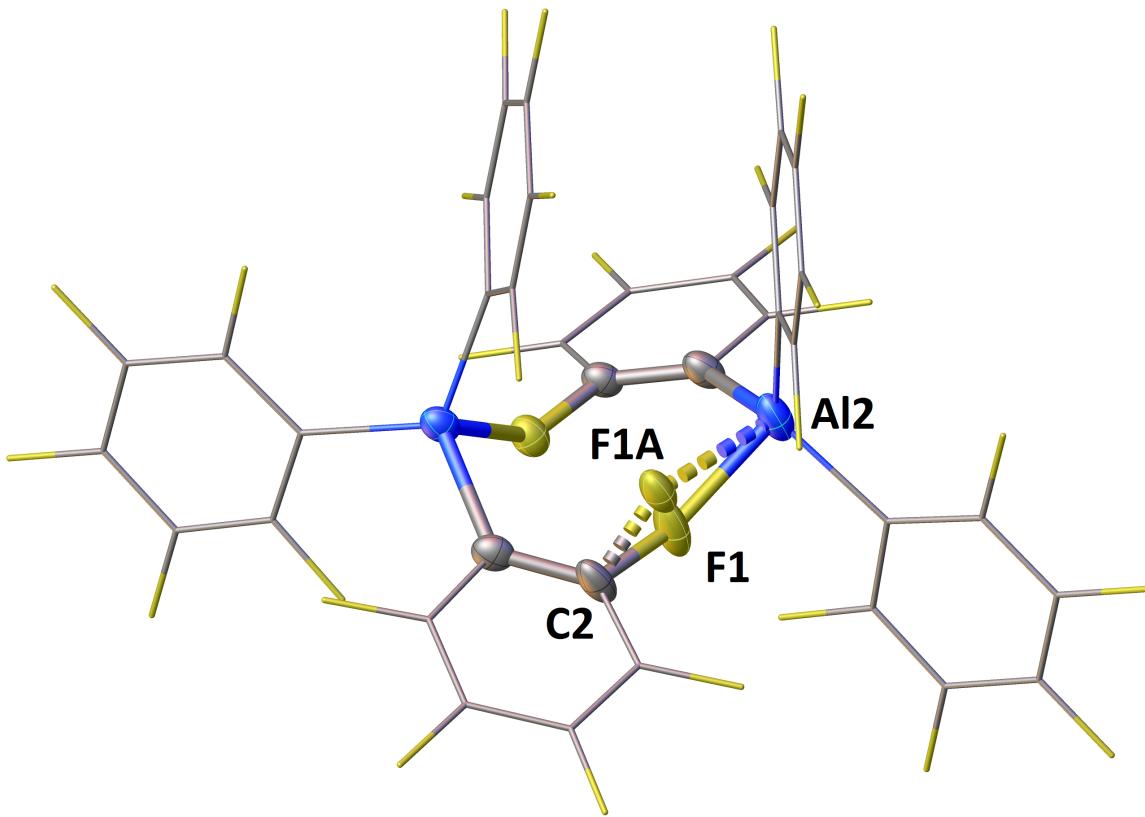


Figure S10. Modeling of the disorder issue for one of the bridging fluorines. The Site Occupancy Factors for F1 and F1A are 0.690 and 0.310, respectively. Other fluorines with high ADP max/min ratio are difficult to model and left as is.

