

Supporting Information

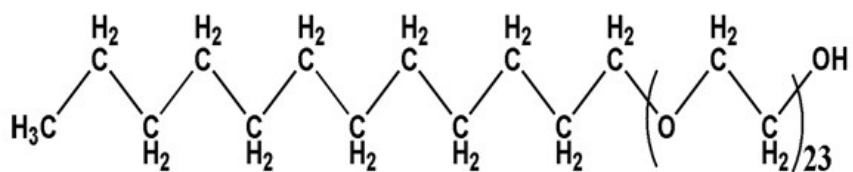
Solar Assisted Photocatalytic reduction of Methyl Orange Azo Dye over Porous TiO₂ nanostructures

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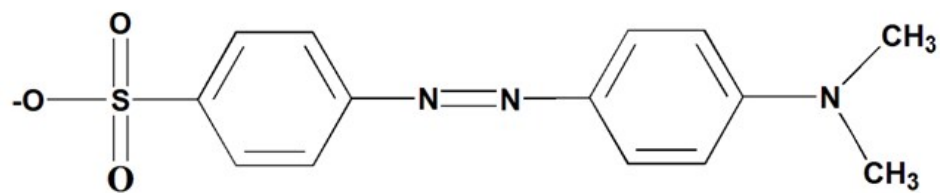
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a

Brij-35

Molecular Weight 1225



b

Methyl Orange

Molecular Weight

304

Figure S1 (a) Structure of Surfactant Brij-35 (b) Structure of Methyl orange dye.

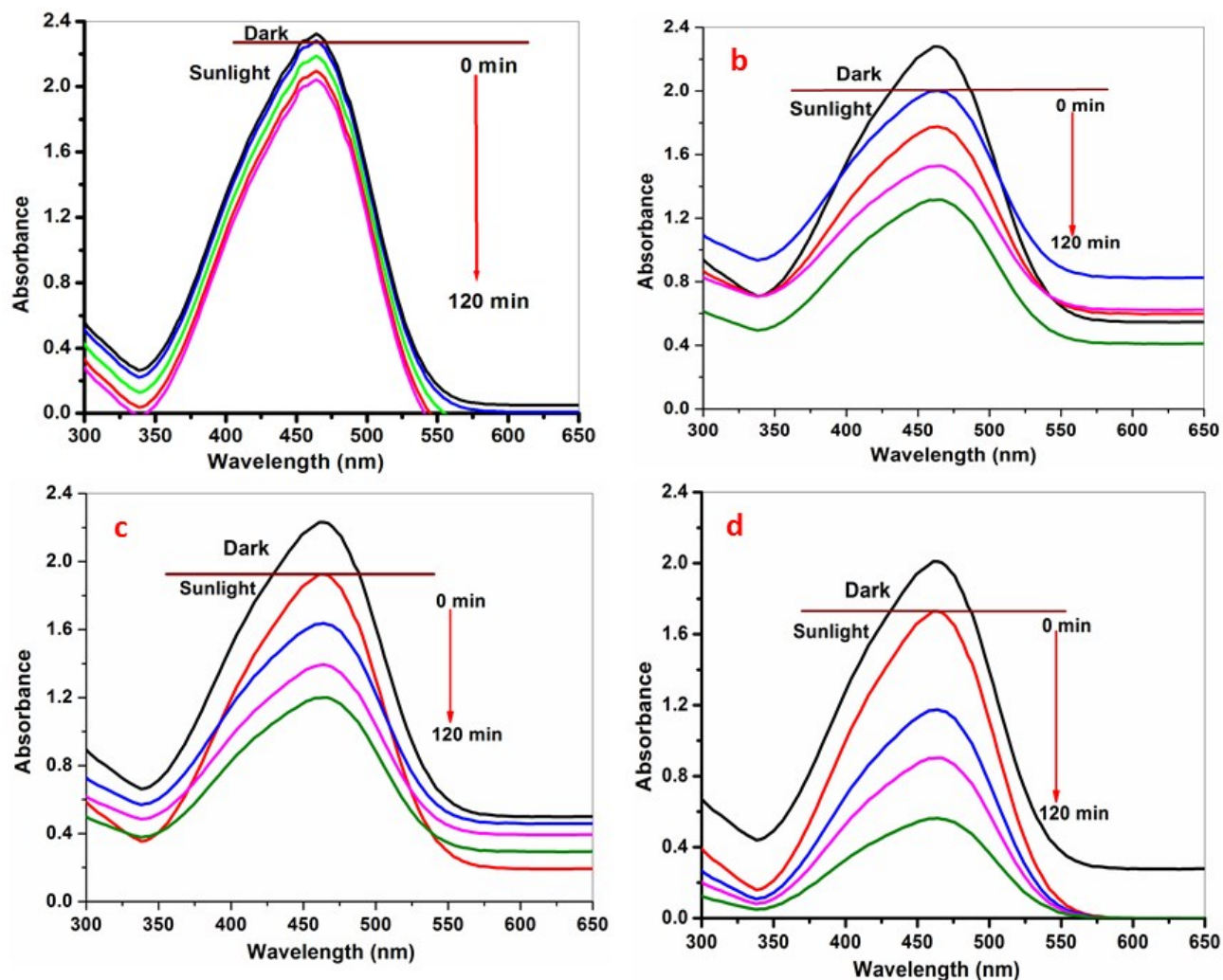


Figure S2 UV-VIS absorption plots of (a) methyl orange dye without catalyst. UV-VIS absorption plots of methyl orange dye in dark for 30 min and then solar light irradiation up to 120 min in presence of catalysts (b) $\text{TiO}_2/\text{Brij-35}$ (c) $\text{TiO}_2/\text{Brij-35}/\text{dextran}$ (d) $\text{TiO}_2/\text{Brij-35}/\text{Silica-Nps}$. (MO concentration $\square$ 20 ppm, catalyst dose $\square$ 12 mg, irradiation time=120 min)

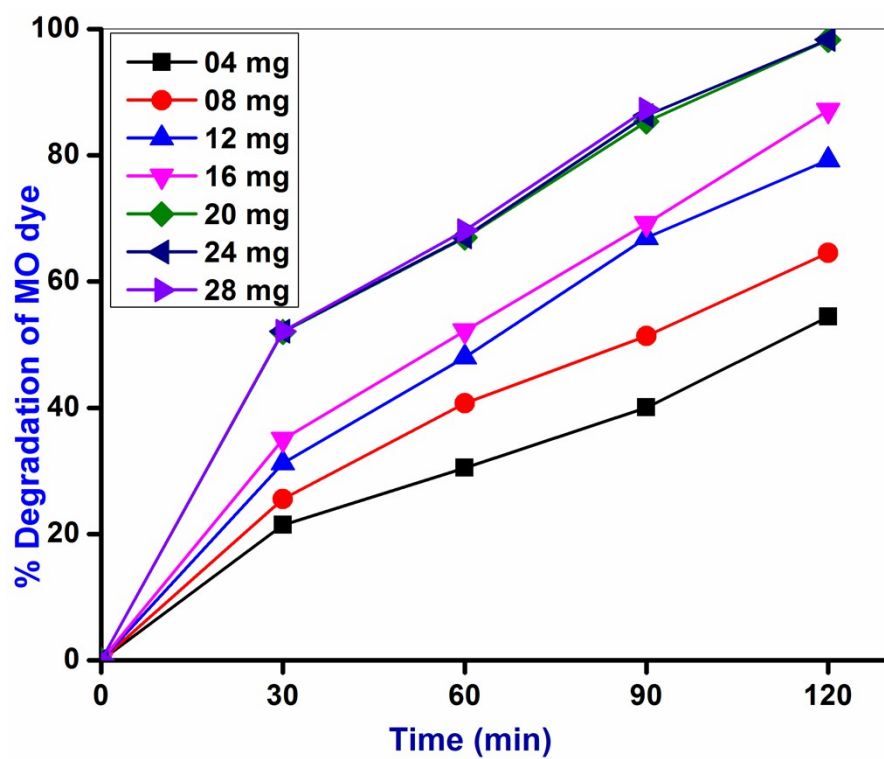


Figure S3 Degradation of MO using different catalyst doses of $\text{TiO}_2/\text{Brij-35}/\text{Silica-Nps}$ with time. (MO concentration $\approx$ 20 ppm, irradiation time=120 min)

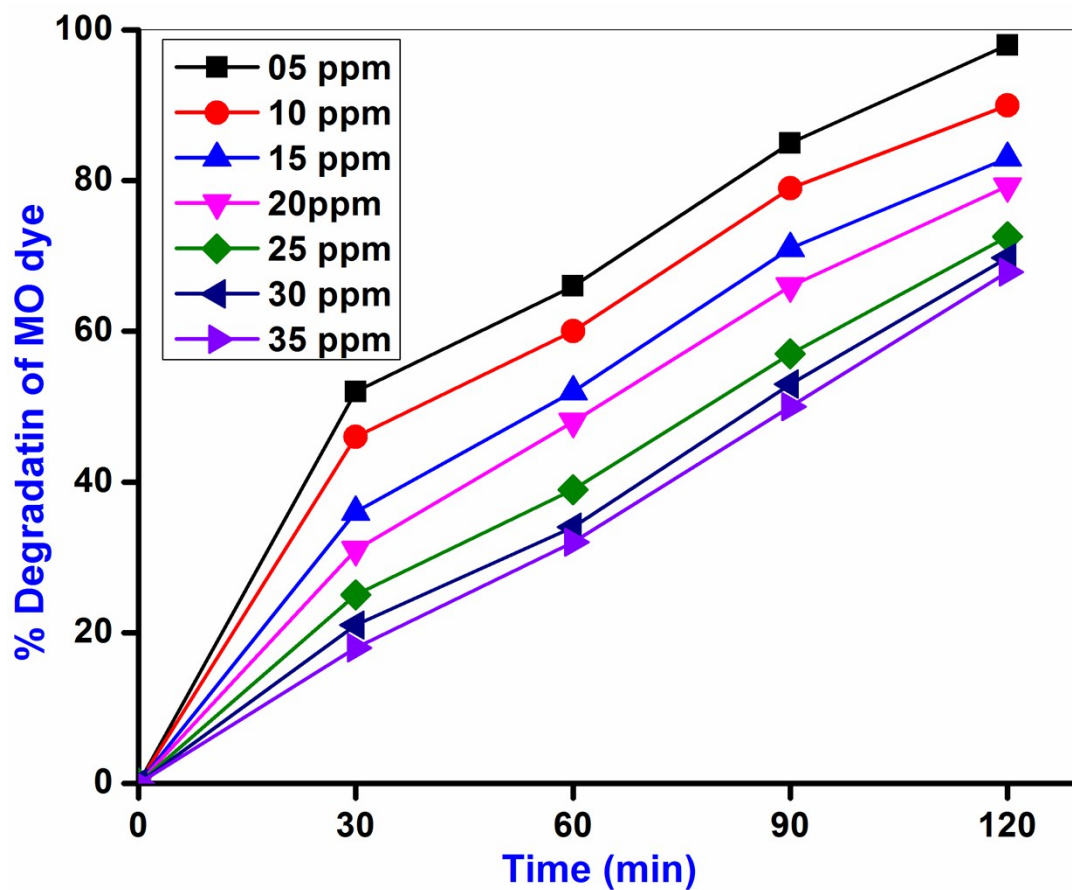


Figure S4 Degradation of MO using different concentrations of MO dye in presence of $\text{TiO}_2/\text{Brij-35}/\text{Silica-Nps}$ catalyst with time. (catalyst dose= 12 mg, irradiation time=120 min)

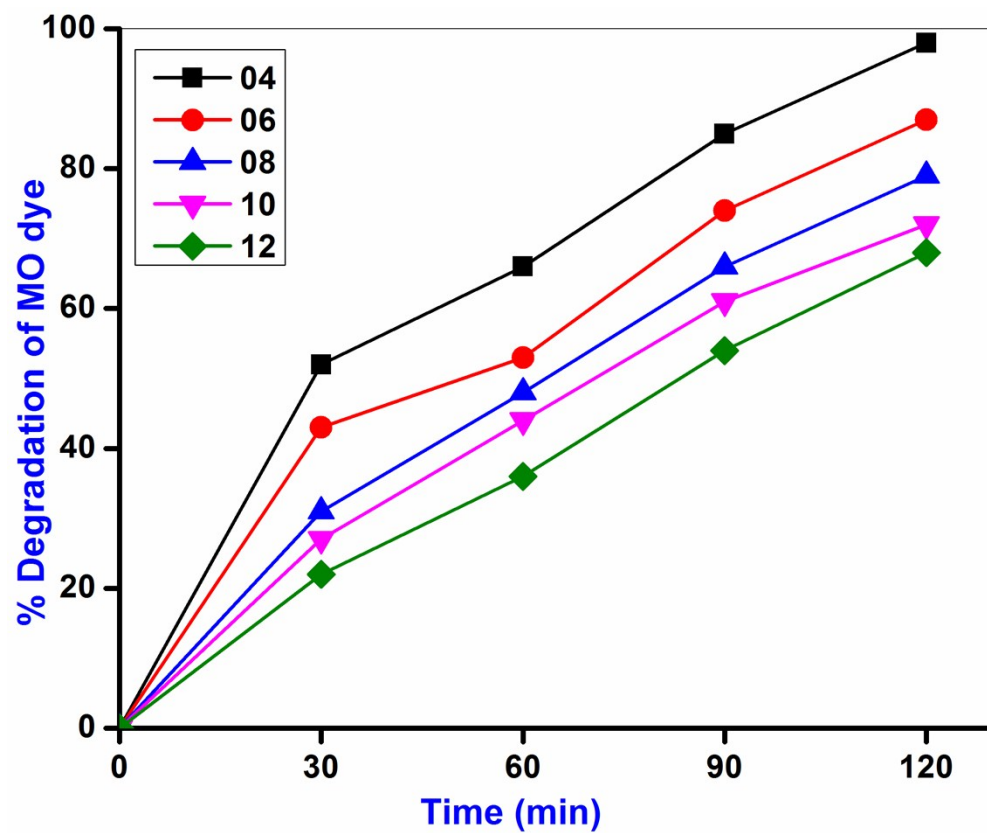
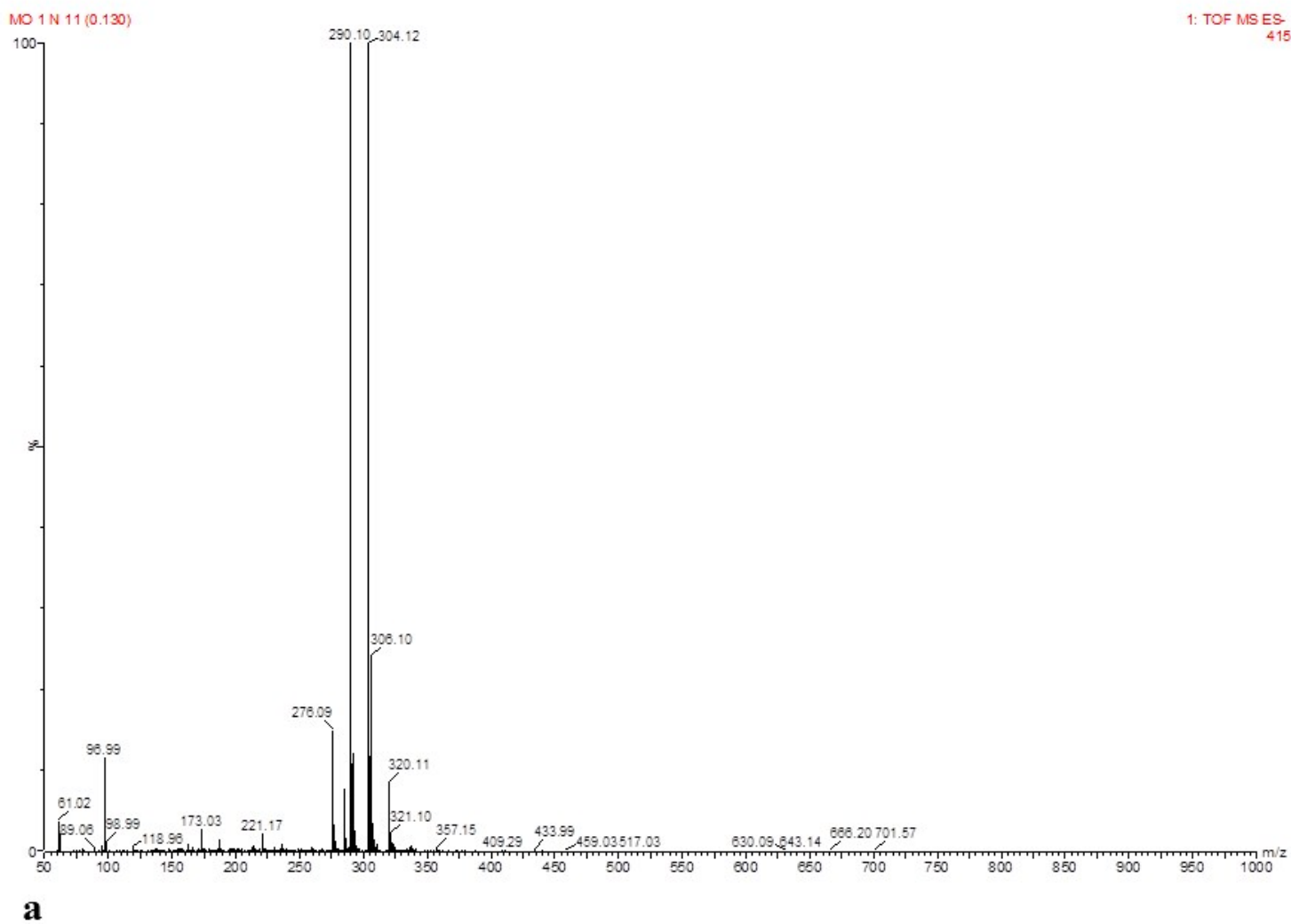


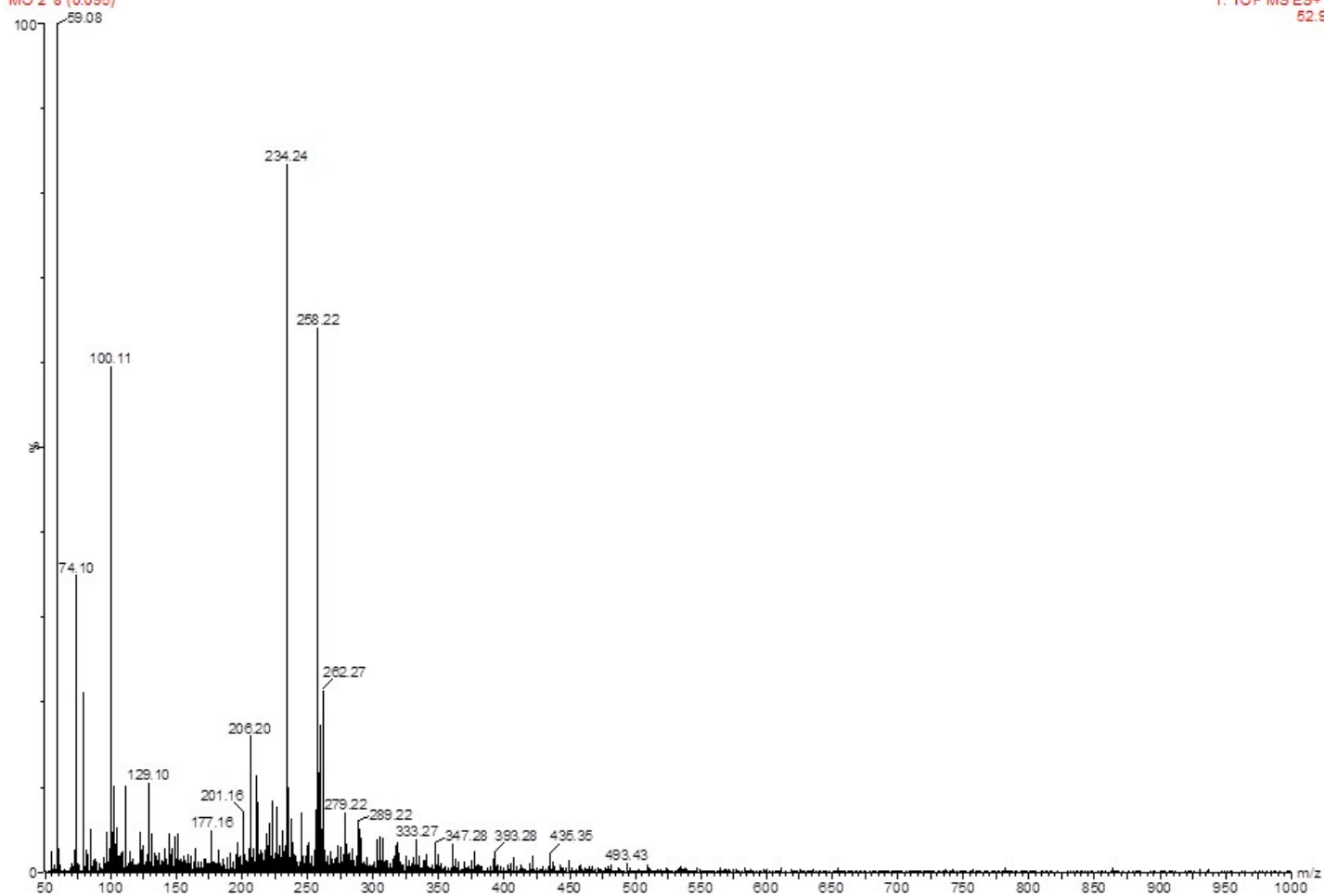
Figure S5 Effect of pH on the degradation of MO dye with time (MO concentration = 20 ppm, catalyst dose= 12 mg, irradiation time=120 min)

6. ESI-Mass Spectra of Methyl Orange (MO) dye degradation.



MO 2 8 (0.095)

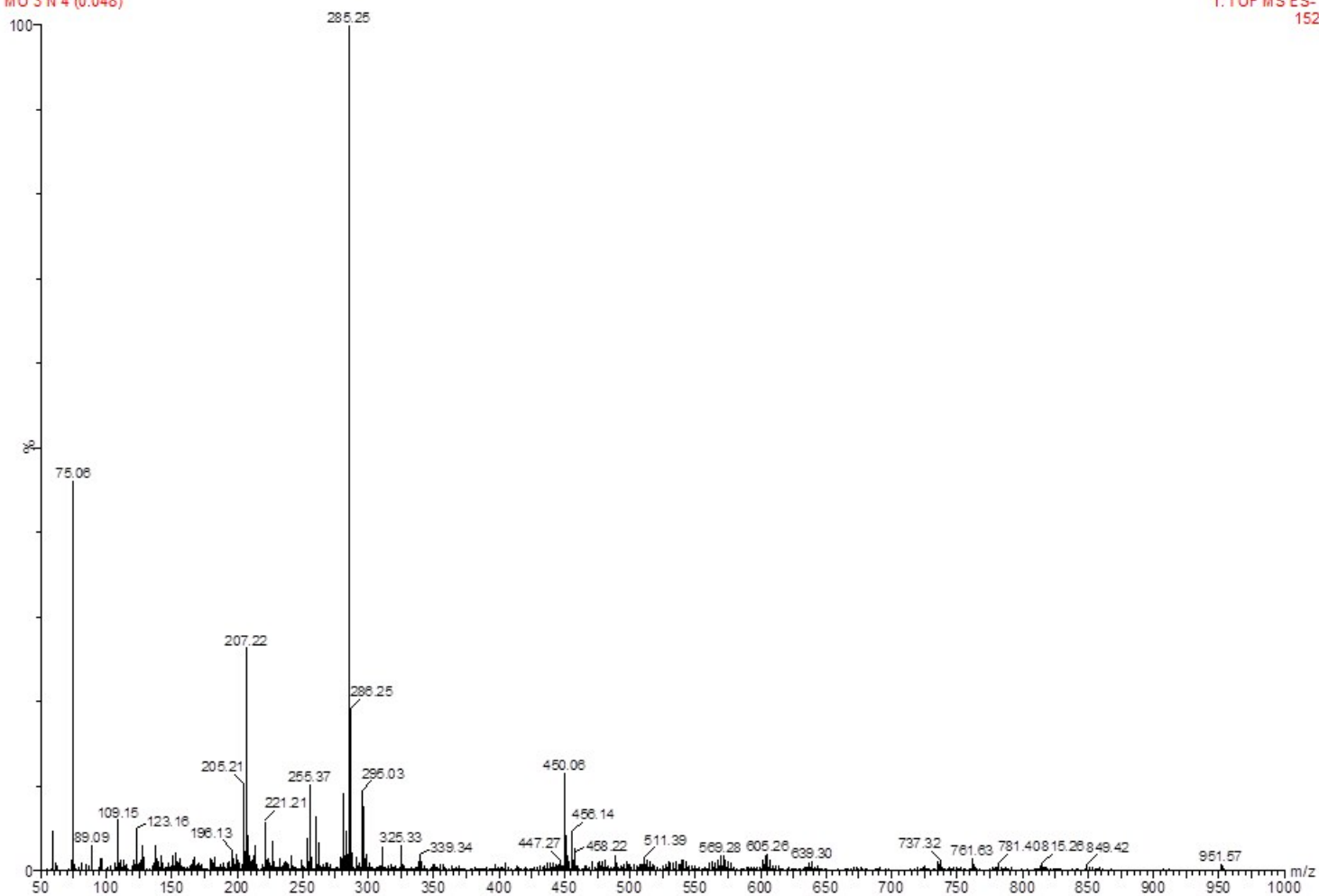
1: TOF MS ES+
52.9



b

MO 3 N 4 (0.048)

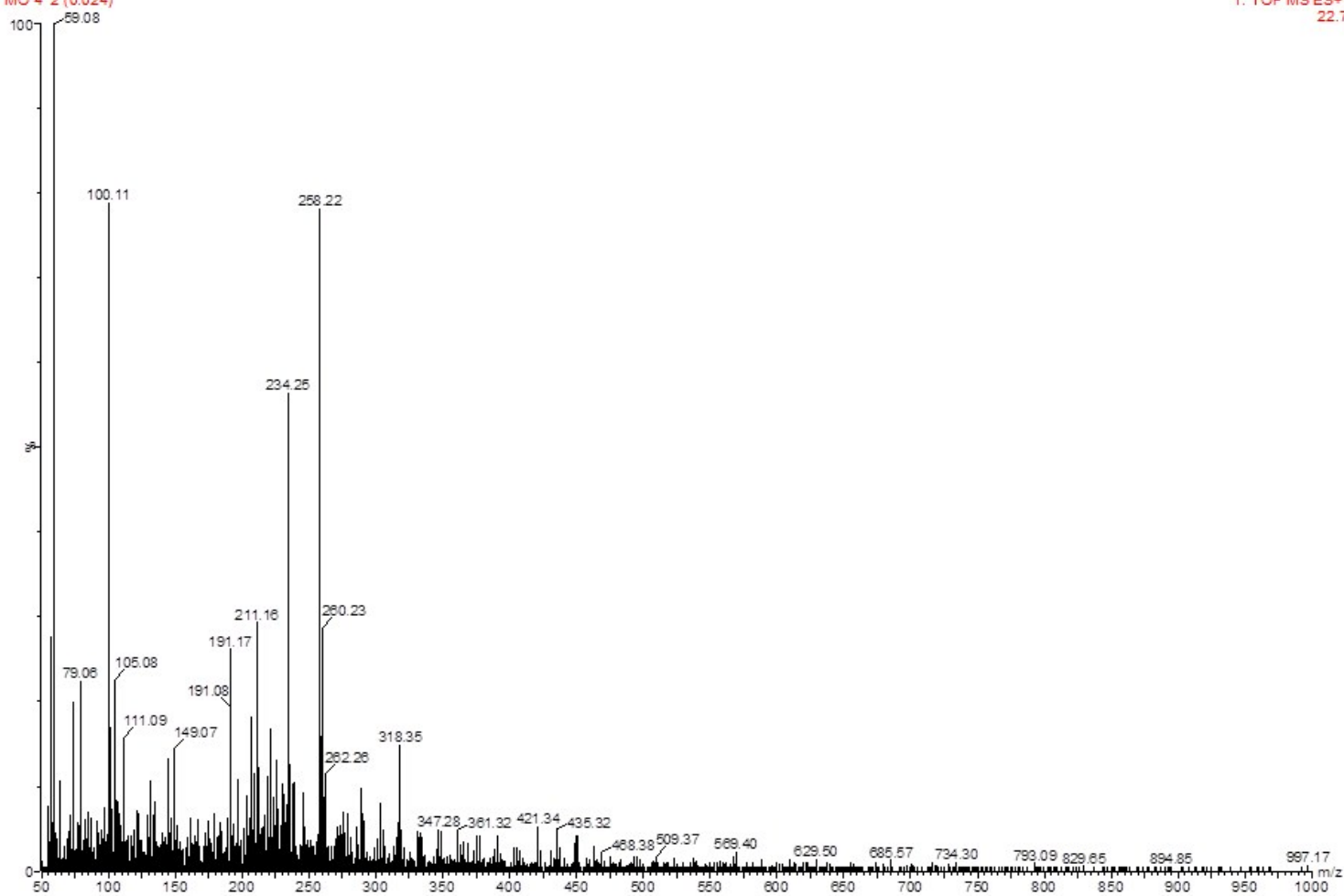
1: TOF MS ES-
152



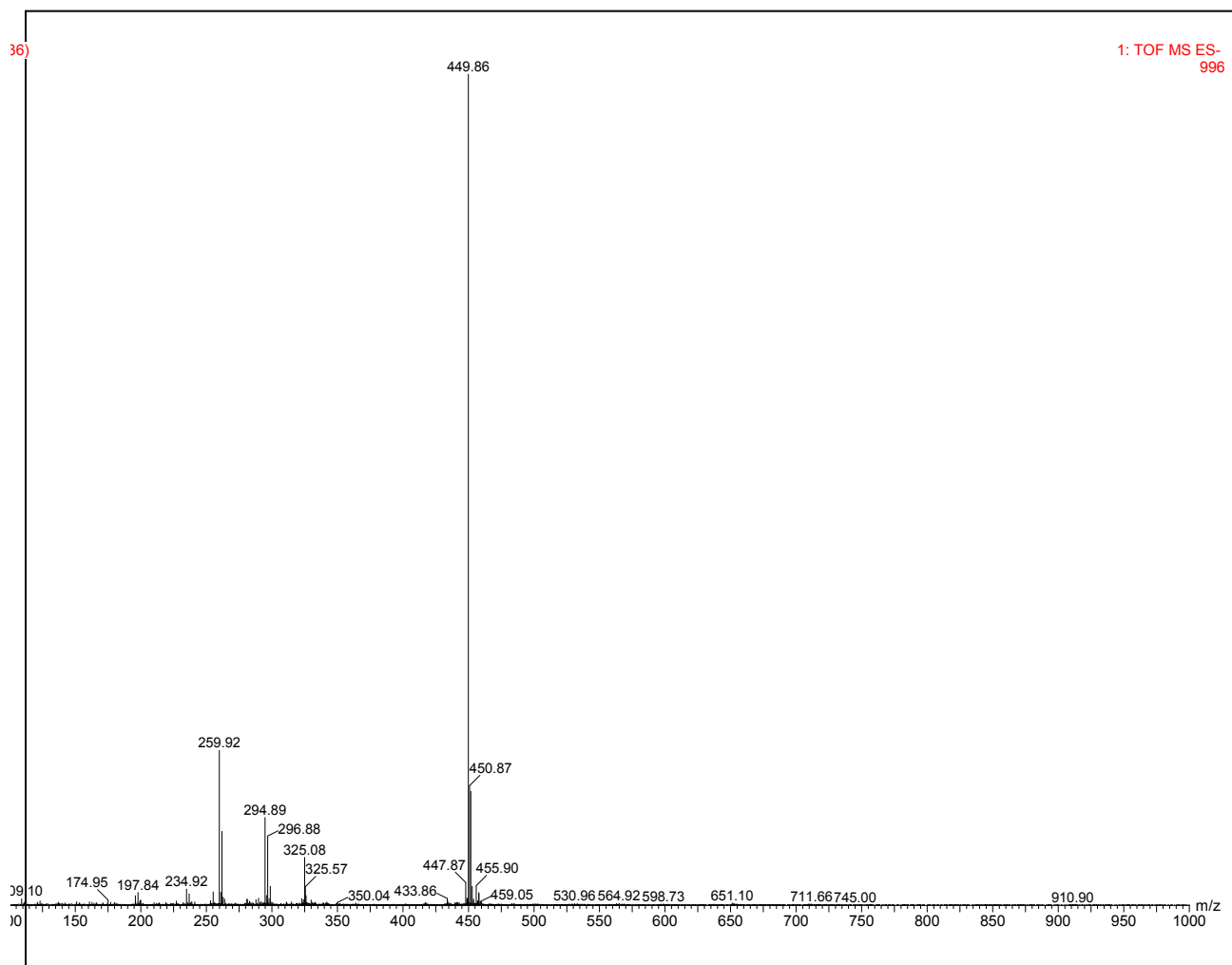
c

MO 4 2 (0.024)

1: TOF MS ES+
22.7



d



e

Figure S 6. ESI-Mass spectra of MO dye degradation at different time intervals (a) 30 min (b) 60 min (c) 90 min (d) 120 min (e) Original solution of MO (20 ppm)