SUPPORTING INFORMATION

Synthesis of magnetic cobalt ferrite nanoparticles with controlled morphology, monodispersity and composition: the influence of solvent, surfactant, reductant and synthetic condition

Le T. Lu^a*, Ngo T. Dung^a, Le D. Tung^b, Cao T. Thanh^c, Ong K. Quy^d, Nguyen V. Chuc^c, Shinya Maenosono^e and Nguyen T. K. Thanh^b*

^aInstitute for Tropical Technology -Vietnam Academy of Science and Technology ,18 Hoang Quoc Viet, Cau Giay, Hanoi, Vietnam. ^bBiophysics Group, Department of Physics and Astronomy, University College London, Gower Street, London WC1E 6BT, UK. ^cInstitute of Materials Science, Vietnam Academy of Science and Technology, 18 Hoang Quoc Viet, Cau Giay, Hanoi, Vietnam. ^dInstitute of Materials, Ecole Polytechnique Federale de Lausanne, Lausanne, Switzerland. ^eSchool of Materials Science, Japan Advanced Institute of Science and Technology, 1-1 Asahidai, Nomi, Japan.



Figure S1: TEM images of the cobalt ferrite NPs synthesised in di-octyl ether with the precursor ratio Co^{2+} : $Fe^{3+} = 1 : 2$ under equimolar amount of OA/OLA at total concentration of 248 (a, b) and 620 mM (c,d); and different reaction times: 5 (a,c) and 120 min (b,d). Scale bar: 200 nm.



Figure S2: TEM images of NPs synthesised in dioctyl ether with a total concentration of equimolar amount of OA and OLA at 240 mM at ratio of Co: Fe = 1: 1 (left) and 1.5: 1 (right) and different reaction times. Scale bar: 200 nm.



Figure S3: TEM images of the cobalt ferrite NPs synthesised in the presence of equimolar amount of OA/OLA at the total concentration of 2.4 M. The starting precursor ratio Co^{2+} : Fe³⁺ = 1 : 1.5 and the reaction time is 30 min. Scale bar: 200 nm.



Figure S4: TGA measurements: **a**) metal acetylacetonate compounds $(1 \text{ Co}(\text{acac})_2 + 2 \text{ Fe}(\text{acac})_3 \text{ and } \text{b})$ metal-acac compounds in the presence of OA and OLA $(1 \text{ Co}(\text{acac})_2 + 2 \text{ Fe}(\text{acac})_3 + \text{OA} + \text{OLA})$. [OA] = [OLA] = 310 mM. TGA plots were recorded at a constant heating rate of 10 °C/min under nitrogen atmosphere.



Figure S5: TEM images of the cobalt ferrite NPs synthesised in the presence of equimolar amount of OA/OLA concentration of 372 mM (186 mM each) at different starting precursor Co^{2+} : Fe³⁺ ratios: a) 1: 2, b) 1: 1.5 and 1.5: 1. The reaction time is 60 min. Scale bar: 20 nm.



Figure S6: TEM images and size distribution histograms of the NPs synthesised in the presence of 496 mM OA/OLA (248 mM each) with HDD (a-d) and without (e-h) for 60 min (a,b,e,f) and 120 min (c,d,g,h). Scale bar: 20 nm.



Figure S7: TEM images and size distribution histograms of the NPs synthesised in the absence of HDD and at 372mM OA/OLA (186 mM ecah) (a,b) and 744 mM (372 mM (c,d) for 60 min. Scale bar: 20 nm.



Figure S8: TEM image of cobalt ferrite NPs synthesised in the presence of sole 620 mM OA for 30 min. Scale bar: 100 nm



Figure S9: XRD patterns of some cobalt ferrite nanoparticles synthesised under different synthetic conditions: a,b) in the presence of HDD with 248 mM equimlar OA/OLA for 60 min (a) and 120 min (b); c) in the absence of HDD/OCD-ol and 744 mM OA/OLA (372 mM each) for 60 min; d) in the presence of OCD-ol and 186 mM OLA for 30 min.



Figure S10: The plot of blocking temperature Tb (a) and magnetocrystalline anistrotropy constant K of them samples prepared at different Co: Fe ratios (1: 2, 1: 1.5 and 1: 1) vs. the average volume of the nanoparticles.



Figure S11: Hysteresis curves of $CoFe_2O_4$ NPs with different size synthesized under different reaction conditions: a,b) in the presence of OCD-ol in octadecene with 186 mM OA for 30 min (a) and 248 mM equimlar OA/OLA for 60 min (b); c) the presence of HDD and 124 mM equimolar OA/OLA for 30 min in dioctyl ether; d) in the absence of HDD/OCD-ol and 248 mM equimolar OA/OLA for 60 min.



IMS-NKL ×100k SE(M,LA0)







Figure S12: High magnification SEM (a) and TEM images of CNTs grown on the 4.9 nm cobalt ferrite NPs coated Si/SiO₂.