

Supporting Information

Pyridine–Triazole Ligands for Copper–Catalyzed Aerobic Alcohol Oxidation

*Pech Thongkam,^{†,§} Sudarat Jindabot,^{†,§} Samran Prabpai,[§] Palangpon Kongsaeree,[§] Taveechai Wititsuwannakul,[§] Panida Surawatanawong,[§] Preeyanuch Sangtrirutnugul**^{†,§}

* Corresponding author. Email: preeyanuch.san@mahidol.ac.th

[†] Center for Catalysis, Department of Chemistry, Faculty of Science, Mahidol University, 272 Rama VI Rd., Bangkok 10400, Thailand

[§] Department of Chemistry and Center of Excellence for Innovation in Chemistry, Faculty of Science, Mahidol University, Rama VI Rd., Bangkok 10400, Thailand.

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Crystallographic Data and Structures

Table S1. Crystallographic data

	1·2H₂O	3	4
empirical formula	C ₂₈ H ₂₀ CuF ₆ N ₈ O ₆ S ₂ •2(H ₂ O)	C ₂₆ H ₁₈ CuF ₆ N ₁₀ O ₆ S ₂	C ₂₂ H ₁₆ CuF ₆ N ₄ O ₆ S ₂
formula weight	842.22	808.16	674.05
T (K)	293	293	293
crystal system	orthorhombic	triclinic	monoclinic
space group	Pcba	P-1	C2/c
<i>a</i> (Å)	17.9015(3)	8.6034(3)	10.0099(3)
<i>b</i> (Å)	9.3530(2)	9.2137(4)	13.9540(5)
<i>c</i> (Å)	19.7128(4)	11.0129(5)	18.8133(8)
α (deg)	90	111.653(2)	90
β (deg)	90	92.753(2)	100.1196(14)
γ (deg)	90	102.528(2)	90
<i>V</i> (Å ³)	3300.57(11)	784.13(6)	2586.93
Z	4	1	4
ρ _{calc} (mg m ⁻³)	1.695	1.711	1.731
μ (mm ⁻¹)	0.885	0.925	1.096
crystal size (mm)	0.28x0.20x0.10	0.15x0.12x0.10	0.20x0.18x0.15
color, habit	green, plate	blue, block	blue, block
reflections collected	14453	8677	4368
independent reflections (<i>R</i> _{int})	3375 (0.055)	2708 (0.062)	1871 (0.050)
data/restraints/parameters	3373/3/248	2708/0/233	1868/0/187
goodness of fit	1.05	1.14	1.05
final <i>R</i> indices [<i>I</i> > 2σ(<i>I</i>)]	<i>R</i> ₁ = 0.0574 <i>wR</i> ₂ = 0.1429	<i>R</i> ₁ = 0.0393 <i>wR</i> ₂ = 0.1289	<i>R</i> ₁ = 0.0543 <i>wR</i> ₂ = 0.1427
final <i>R</i> indices (all data)	<i>R</i> ₁ = 0.0722 <i>wR</i> ₂ = 0.1510	<i>R</i> ₁ = 0.0422 <i>wR</i> ₂ = 0.1342	<i>R</i> ₁ = 0.0631 <i>wR</i> ₂ = 0.1553
largest diff. peak/hole	0.607, -0.606	0.388, -0.378	0.423, -0.742

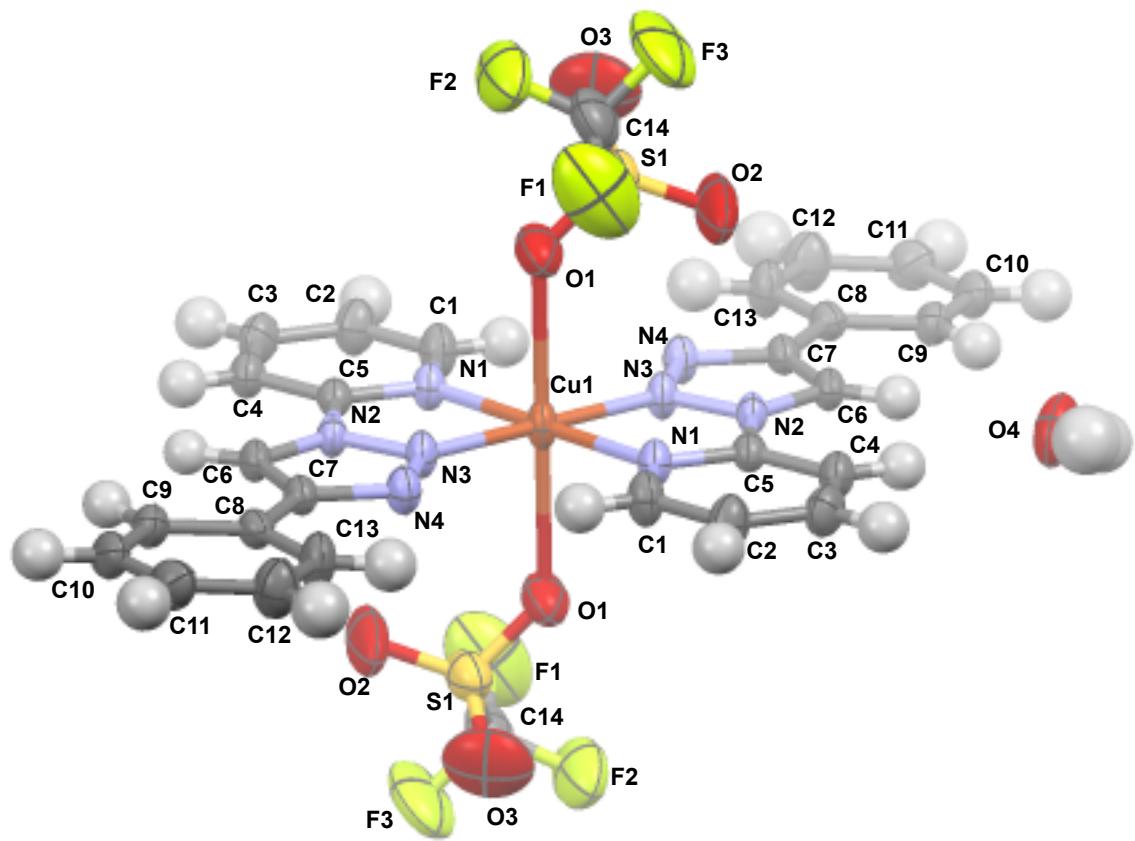


Figure S1. ORTEP diagram of **1**·2H₂O. Solvent molecules were omitted for clarity. The thermal ellipsoids were shown at the 50% probability level.

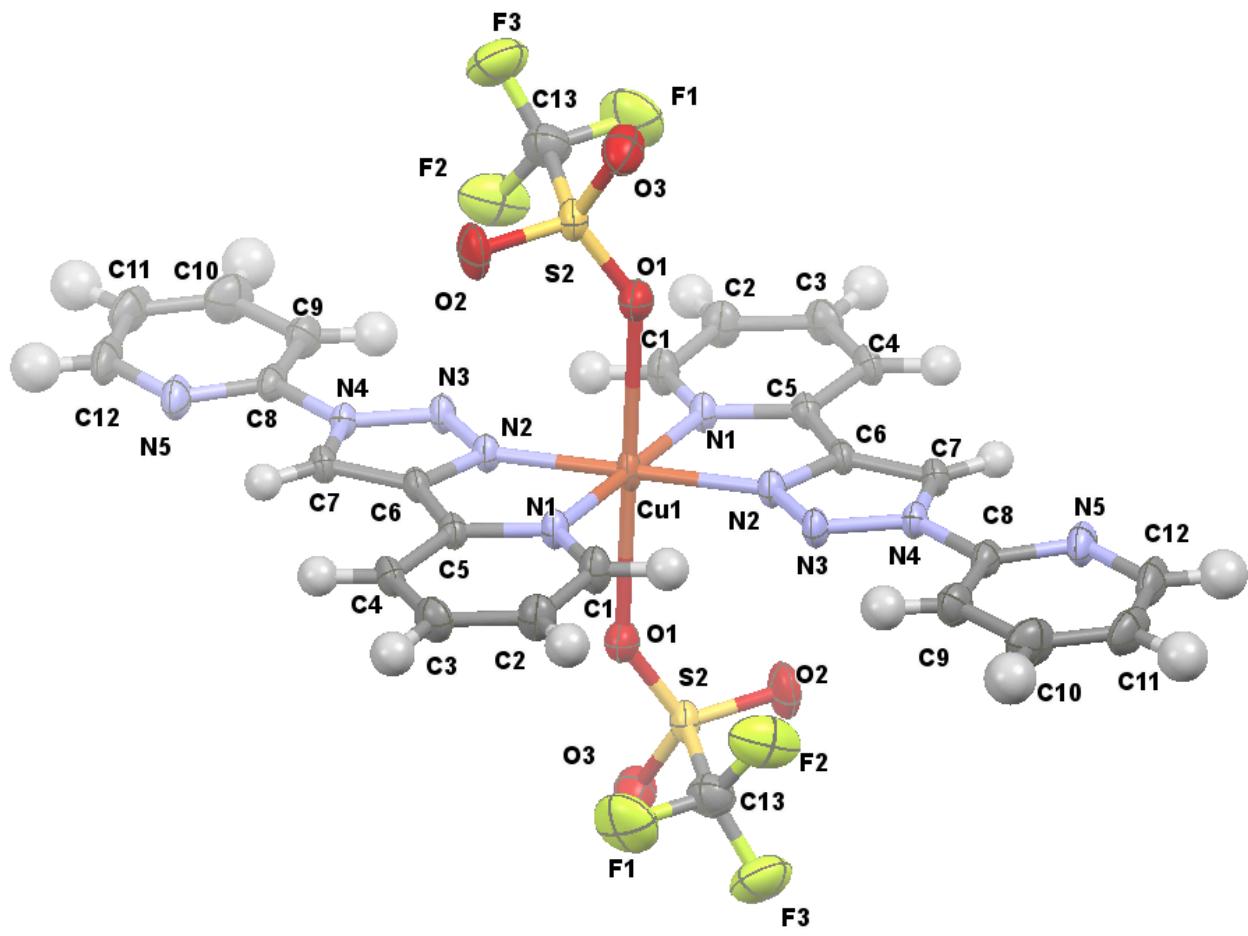


Figure S2. ORTEP diagram of 3. The thermal ellipsoids were shown at the 50% probability level.

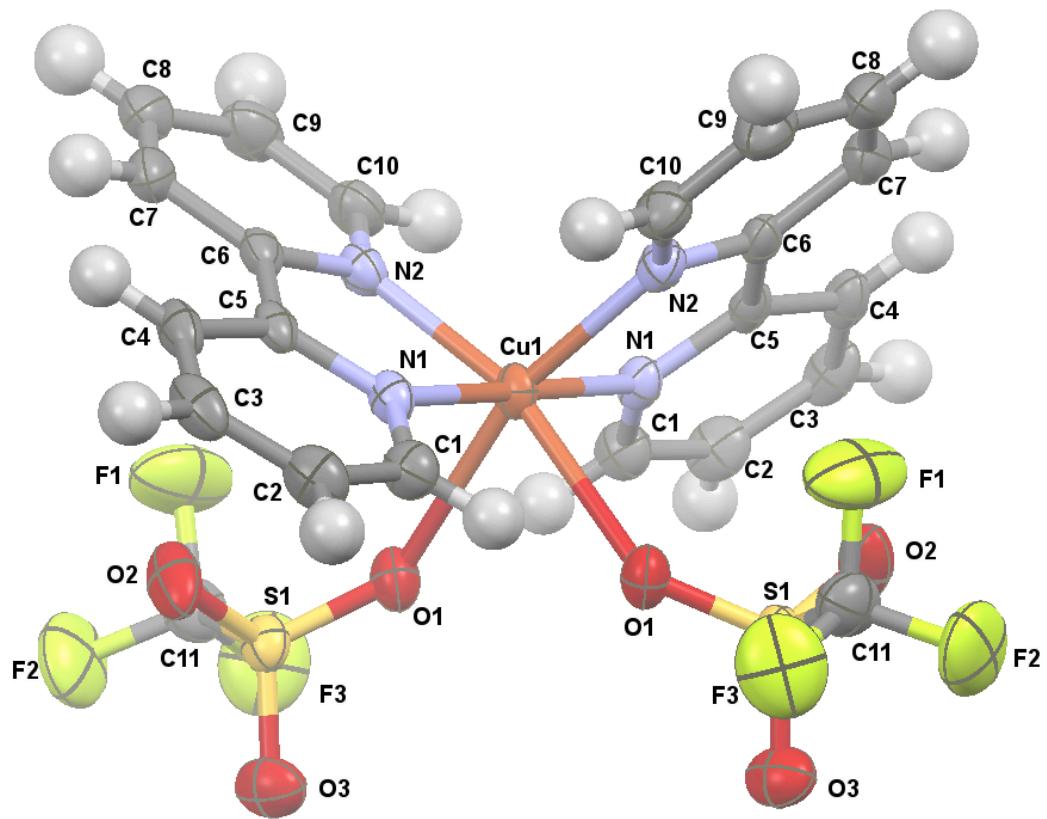
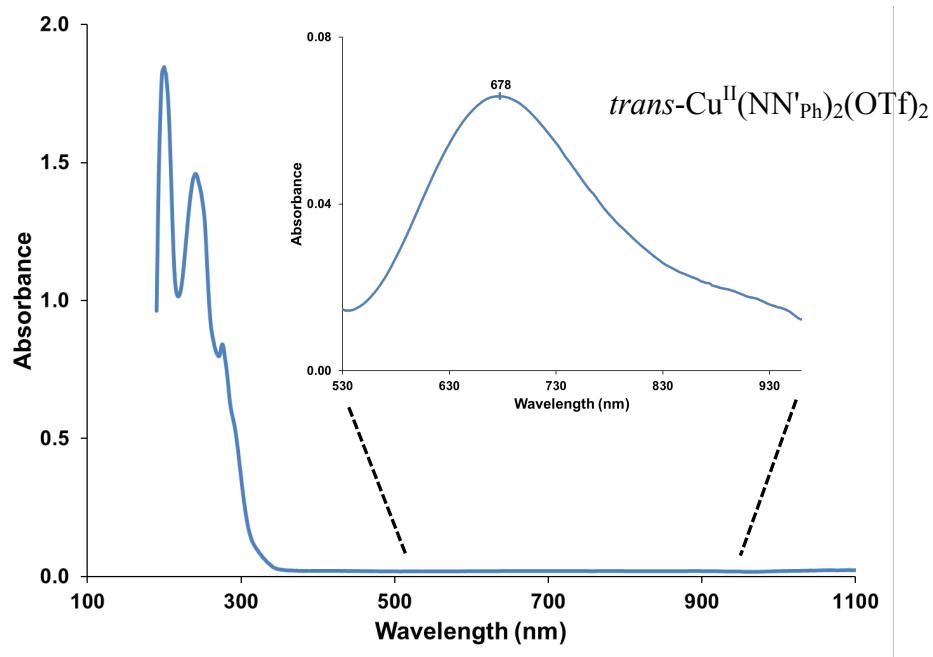


Figure S3. ORTEP diagram of **4**. The thermal ellipsoids were shown at the 50% probability level.

UV-Vis Spectra

(a)



(b)

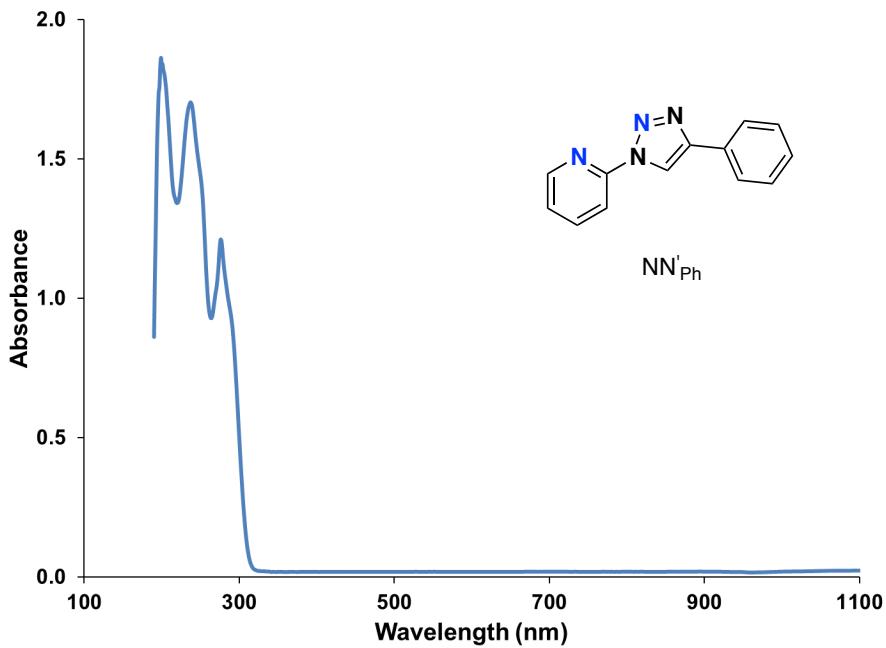
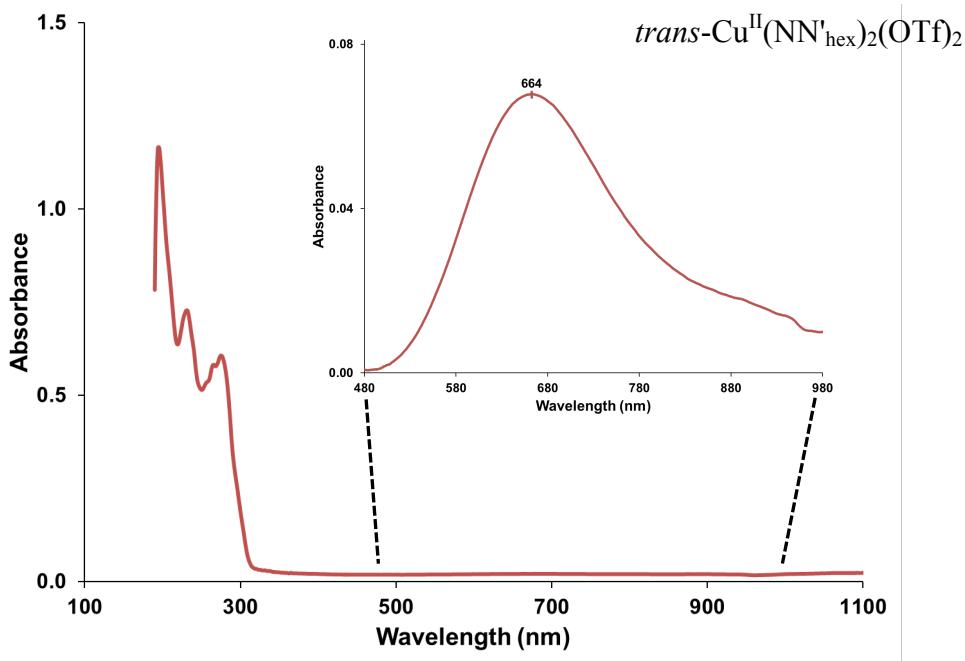


Figure S4. Electronic absorption spectra of **1** (a) and NN'Ph (b).

(a)



(b)

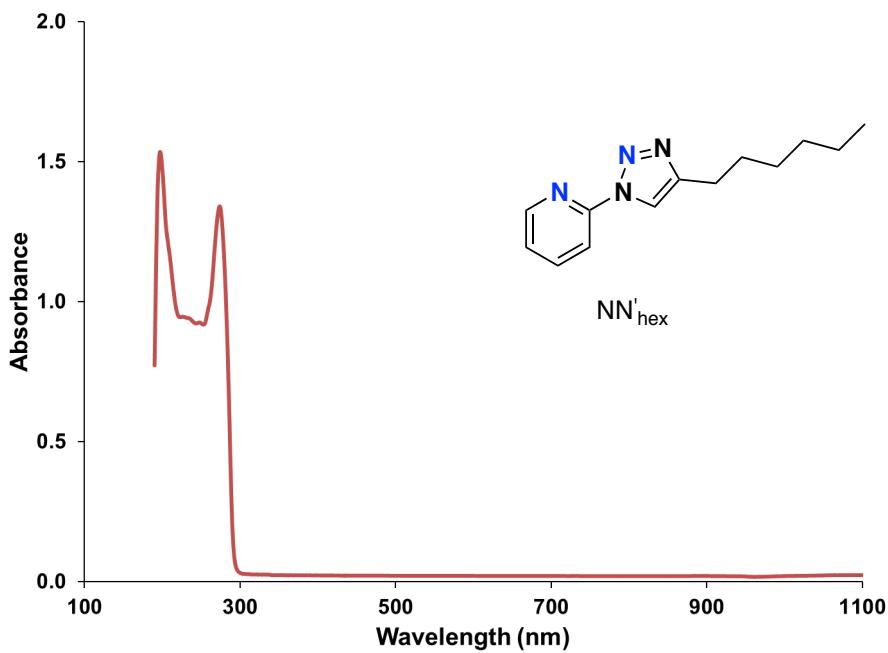
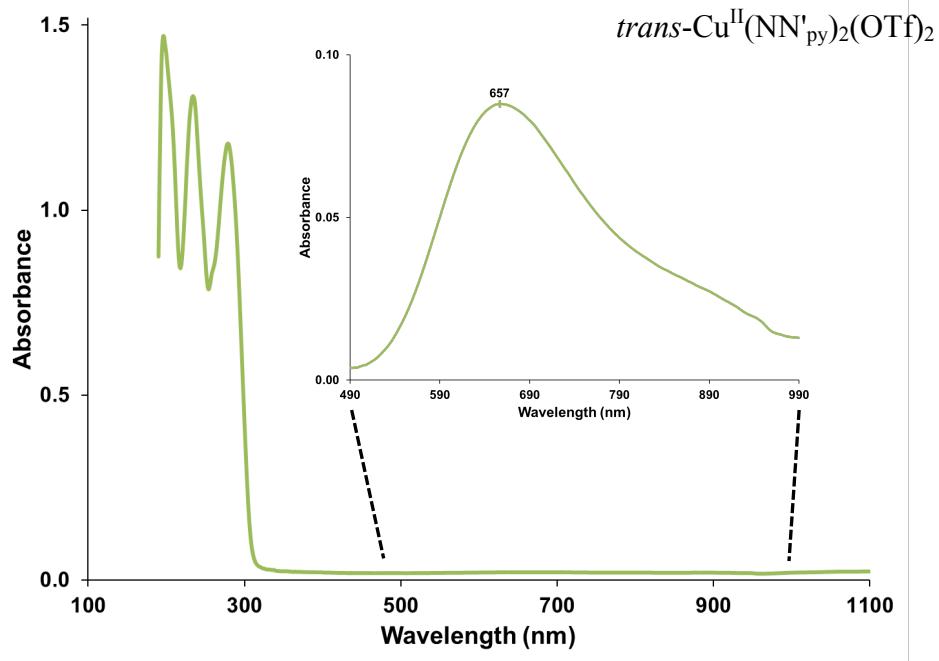


Figure S5. Electronic absorption spectra of **2** (a) and NN'_{hex} (b).

(a)



(b)

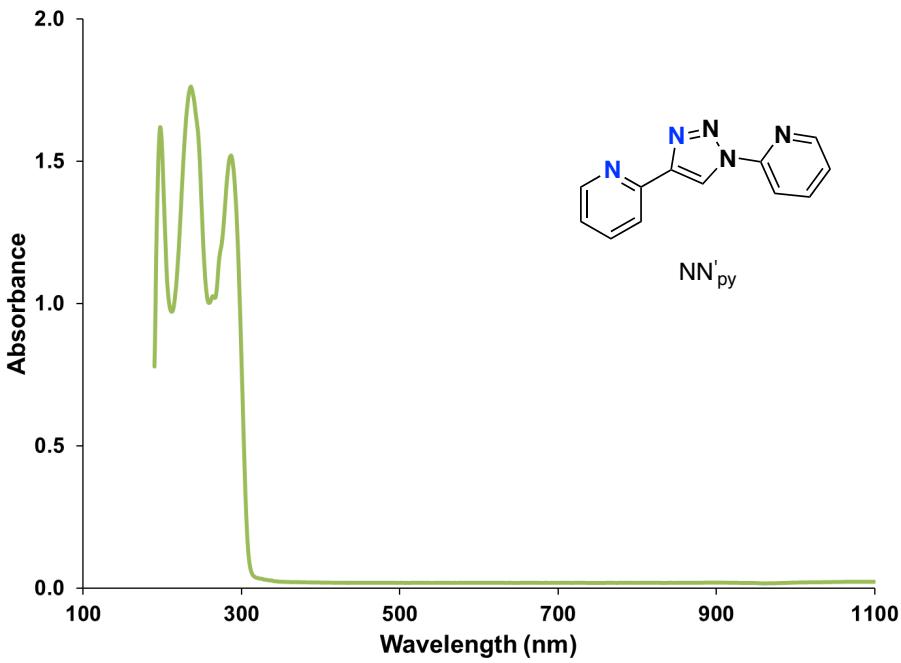
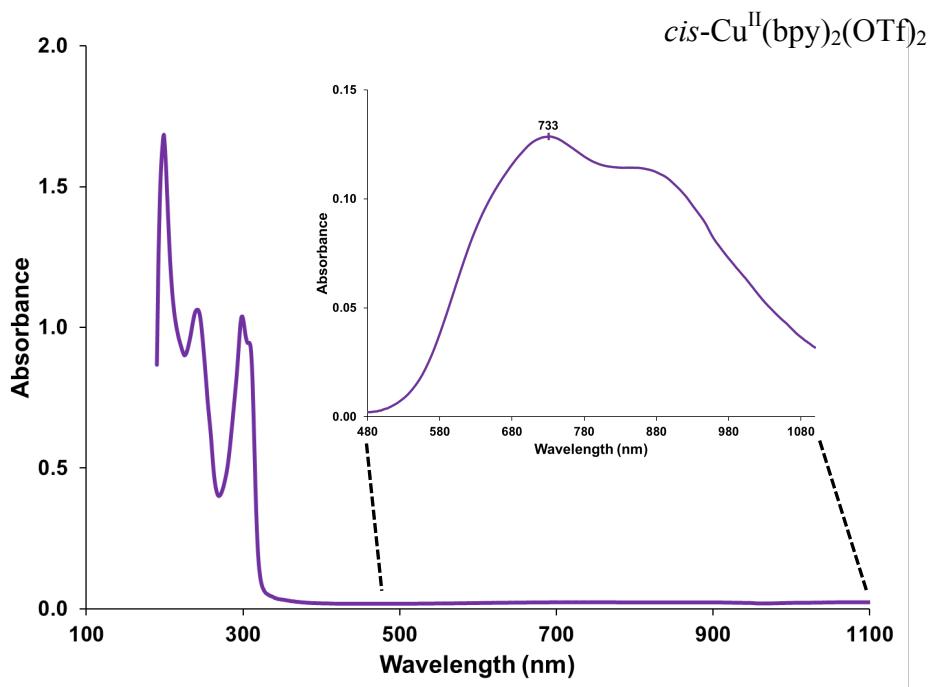


Figure S6. Electronic absorption spectra of **3** (a) and NN'py (b).

(a)



(b)

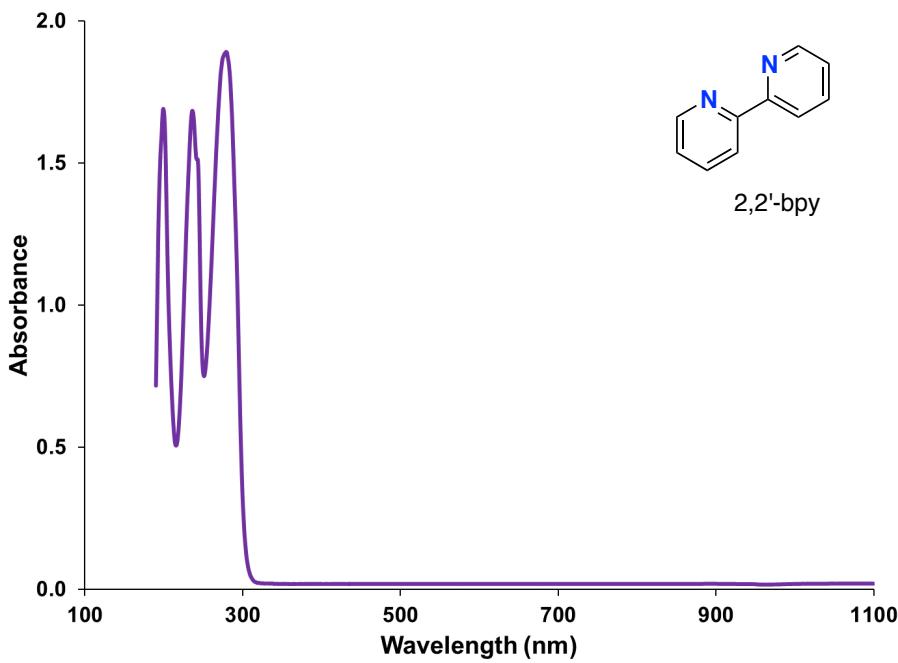


Figure S7. Electronic absorption spectra of **4** (a) and 2,2'-bpy (b).

Powder X-ray diffraction data

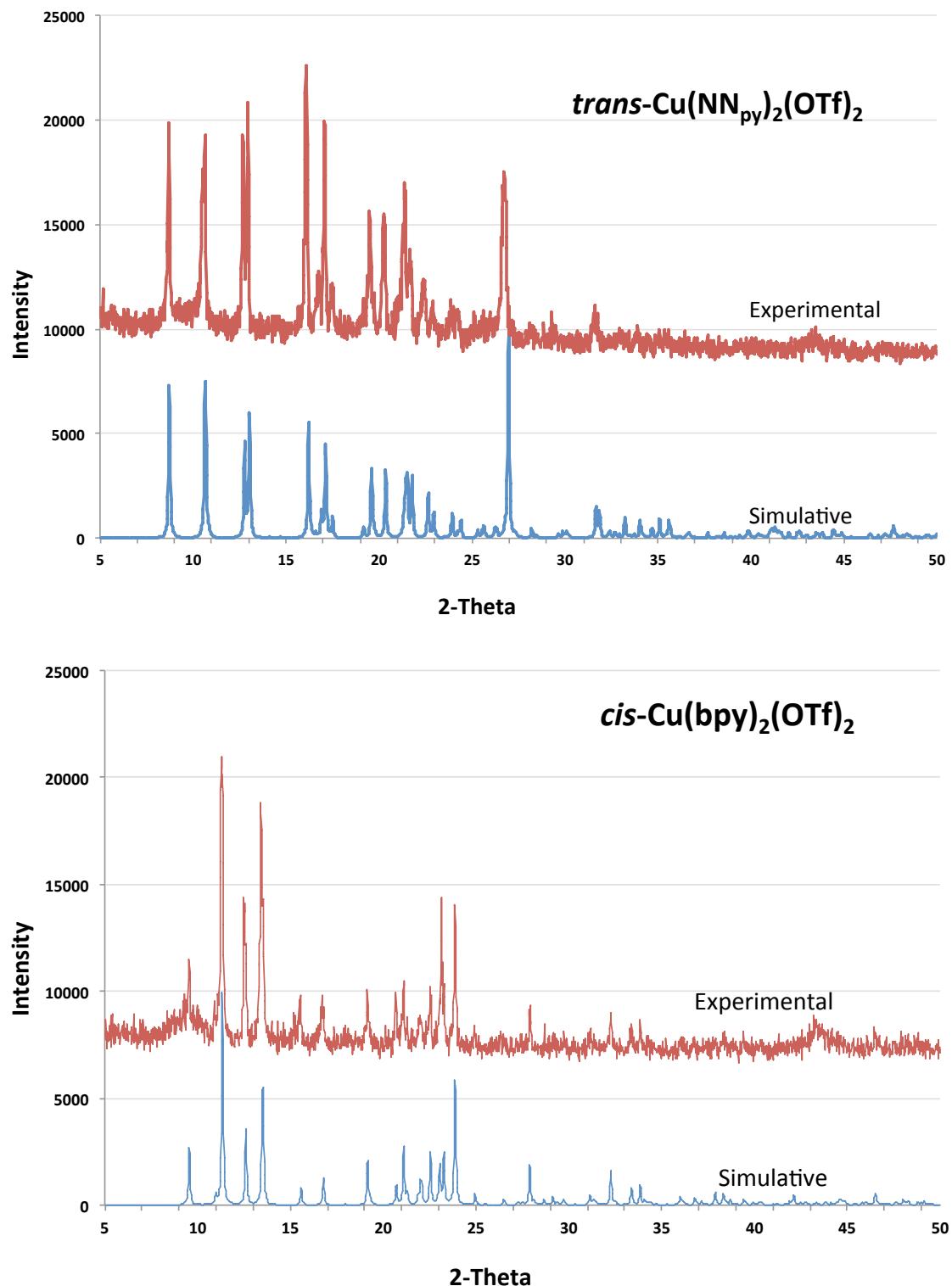
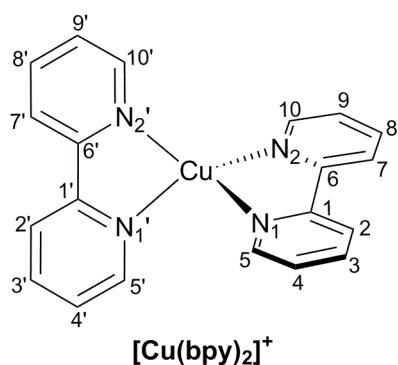


Figure S8. Powder X-ray diffraction patterns of Cu(II) complexes **3** and **4**.

DFT Calculations

Table S2. Selected geometrical parameters (bond lengths in Å and bond angles in °) of the X-ray crystal structure^a and the calculated structure of $[\text{Cu}(\text{bpy})_2]^+$.



	Cu $[(\text{bpy})_2]^+$	
	Calculation	X-ray ^a
Cu-N1	2.073	2.020(9)
Cu-N2	2.073	2.021(11)
C1-C6	1.490	1.440(15)
C1-N1	1.354	1.325(14)
C6-N2	1.354	1.385(15)
N1-Cu-N1'	125.9	127.1(4)
N1-Cu-N2	80.1	81.5(4)
N1-Cu-N2'	125.9	131.3(4)
N2-Cu-N2'	125.8	109.7(4)
C6-C1-N1	115.8	115.9(9)
C1-C6-N2	115.8	116.5(9)
Cu-N1-C1	114.2	114.2(8)
Cu-N1-C5	126.7	127.3(8)

^aMunakata, M.; Kitagawa, S.; Asahara, A.; Masuda, H.

Bull. Chem. Soc. Jpn **1987**, 60, 1927-1929.

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