## Electronic supplementary information (ESI) for

## Fe<sub>2</sub>O<sub>3</sub>@Au Core@Shell Nanoparticle-Graphene Nanocomposites as Theranostic Agents for Bioimaging and Chemo-Photothermal Synergistic Therapy

Hongda Chen,<sup>a</sup> Fuyao Liu,<sup>a,b</sup> Zhen Lei,<sup>a,b</sup> Lina Ma<sup>\*,a</sup> and Zhenxin Wang<sup>\*,a</sup>

<sup>a</sup>State Key Laboratory of Electroanalytical Chemistry, Changchun Institute of Applied Chemistry, Chinese Academy of Sciences, Changchun, 130022, P. R. China <sup>b</sup>University of Chinese Academy of Sciences, Beijing, 100039, P. R. China E-mail: malina@ciac.ac.cn (LM) and wangzx@ciac.ac.cn (ZW).

Nanocomposites	Zeta potential (mv)	hydrodynamic diameter (nm)
NH <sub>2</sub> -PEG-Fe <sub>2</sub> O <sub>3</sub> @Au NPs	12.9	129.4
rGO-COOH	-29.5	616.0
rGO-Fe <sub>2</sub> O <sub>3</sub> @Au NPs	-21.1	611.7

 Table S1
 The Zeta potentials and hydrodynamic diameters of as-prepared nanocomposites.

The PEG can significantly improve the colloidal stability and dispersibility of nanocomposites, resulting in that hydrodynamic diameter of rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs is slightly smaller than that of rGO-COOH.

Concentration (µg mL <sup>-1</sup> )	$f_{add}$	$f_{chem-PTT}$
10	0.786	0.756
20	0.566	0.482
30	0.313	0.254
40	0.123	0.111
50	0.04	0.03

**Table S2** Comparison of calculated  $f_{add}$  from  $f_{chem}$  and  $f_{PTT}$  with  $f_{chem-PTT}$ .



**Fig. S1** UV-visible spectra of NH<sub>2</sub>-PEG-Fe<sub>2</sub>O<sub>3</sub>@Au NPs (blue line), rGO-COOH (red line) and rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs (black line), respectively.



**Fig. S2** Raman spectra of NH<sub>2</sub>-PEG-Fe<sub>2</sub>O<sub>3</sub>@Au NPs (black line), rGO-COOH (red line) and rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs (blue line), respectively.



Fig. S3 (a) Hydrodynamic size and (b) UV-visible spectrum of rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs which were incubated in fresh DMEM supplemented with 10% (v/v) FBS for 5 h.



**Fig. S4**  $R_2$  relaxivity of rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs as a function of the molar concentration of Fe<sup>3+</sup> in the solution.



**Fig. S5** UV-visible spectra of DOX (black line) and DOX-rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs (red line), respectively.



**Fig. S6** DOX release from rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs over time in PBS. In the NIRtriggered release of DOX from rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs, which were irradiated with an 808 nm NIR laser (2 W cm<sup>-2</sup>) for 5 min at time points indicated by arrows. Error bars mean standard deviations (n = 5).



Fig. S7 Cellular Uptake of DOX-rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs or rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs. The concentrations of DOX-rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs and rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs are 50  $\mu$ g mL<sup>-1</sup> at a rGO concentration in cell culture medium, respectively.



**Fig. S8** The viabilities HeLa cells which incubated with different concentrations of nanomaterials for 24 h, respectively. The concentrations of rGO-COOH, rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs and DOX-rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs are defined by rGO content, and the concentration of NH<sub>2</sub>-PEG-Fe<sub>2</sub>O<sub>3</sub>@Au NPs is defined by Fe<sub>2</sub>O<sub>3</sub>@Au NPs content, respectively. The error bars mean standard deviations (n = 5).



**Fig. S9** Digital photos of HeLa cells incubated with DOX-rGO-Fe<sub>2</sub>O<sub>3</sub>@Au NPs and a magnet. Position 1, right above the magnet; position 2, at the edge of the magnet; and position 3, far from the magnet.