Supplementary Information

Porous MnFe₂O₄ Microrods as Advanced Anodes for Li-ion Batteries with Long Cycle Lifespan

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Fig. S1 (a) FTIR spectra and (b) TGA curve of the precursor.

Synthesis of MnFe₂O₄ nanorods: the synthesis of the FeOOH nanorods was according to our previous report.^[1] First, FeCl₃·H₂O (2 mmol) and KNO₃ (15 mmol) were dissolved in deionized water (40 ml) under magnetic stirring. The resulting solution was then transferred into Teflon-lined autoclave and heated to 100 °C for 12 h. The precipitate was collected, washed with water and then dried in vacuum at 60 °C overnight. The obtained FeOOH nanorods (2 mmol) were mixed with MnC₂O₄•2H₂O (1 mmol) and several drops of ethanol using ball-milling for 2 h. The homogenous slurry was dried in vacuum, followed by annealing in Ar atmosphere at 500 °C for 2 h.



Figure S2. (a) XRD patterns and (b) cycling performance at a current density of 1 A g^{-1} (c, d) TEM images of MnFe₂O₄ nanorods.



Figure S3. Nitrogen adsorption and desorption isotherms and pore size distribution (inset) of MnFe₂O₄ nanorods.

MnFe₂O₄ nanorods show diameters of ~ 50 nm and lengths of 200 nm (Figure S2c, S2d) which is shorter than porous MnFe₂O₄ microrods. The BET specific area and pore volume of the MnFe₂O₄ nanorods is 16.17 m² g⁻¹ and the pore volume is 0.068 cm³ g⁻¹, respectively (Figure S3). It is clearly observed that the size distribution is broad in the range of 5 ~ 15 nm, confirming the irregular pores observed in the TEM images (Figure S2d).