

Electronic Supplementary Information

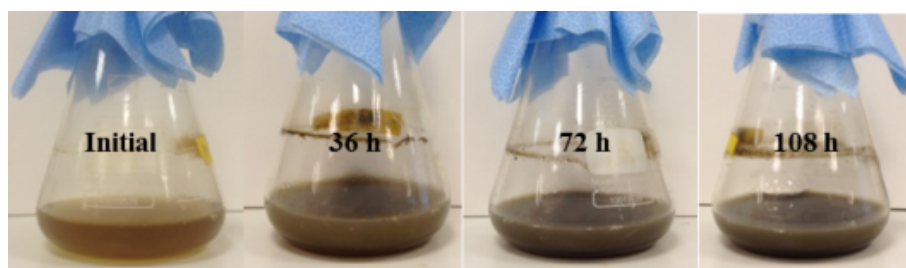


Fig. S1 Optical photographs of GO-bacteria mixture in culture medium at different cultivation time.

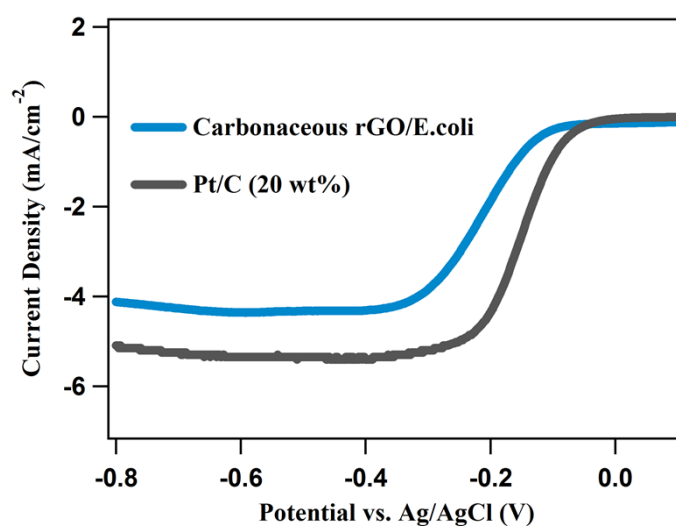


Fig. S2 LSV of carbonaceous rGO/ *E. coli* and commercial Pt/C (20 wt%) at roatation speed of 1600 rpm

Derivation of electron transfer number

The electron transfer number (n) of ORR is determined based on K-L equation: ^{1,2}

$$1/I = 1/I_L + 1/I_K = 1/(B\omega^{1/2}) + 1/I_K$$

$$B=0.2nF(D_0)^{2/3}\nu^{-1/6}$$

Where I , I_L and I_K are the measured current density, diffusion limiting current density and kinetic-limiting current density, respectively; ω is the rotation speed in rpm; F is the Faraday constant (96485 C mol⁻¹); D_0 is the diffusion coefficient of oxygen in 0.1 M KOH (1.9×10^{-5} cm² s⁻¹); ν is the kinetic viscosity (0.01 cm² s⁻¹); and C_0 is the bulk concentration of oxygen in the solution (1.2×10^{-6} mol cm⁻³).

Tafel Plot

For the Tafel plot, kinetic-limiting current density (I_K) is calculated based on mass-transport correction: $I_K = I \times I_L / (I_L - I)$. And the Tafel equation is given as $\eta = a + b \log I_k$, where η is the measured potential vs. Ag/AgCl.

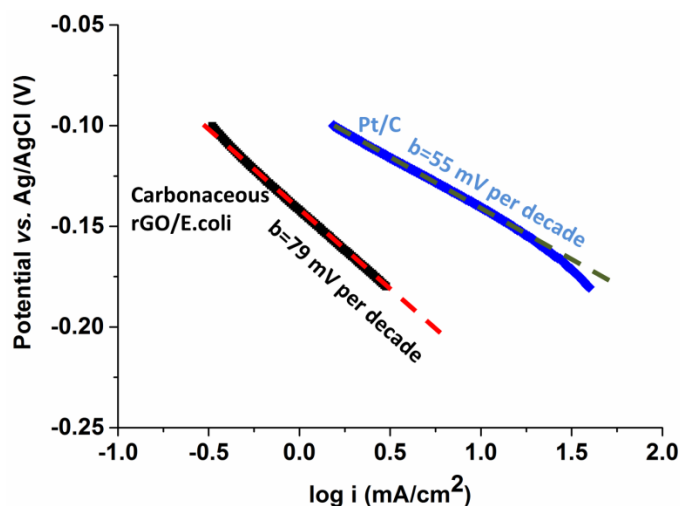


Fig. S3 Tafel plots of carbonaceous rGO/*E.coli* and Pt/C obtained at low currents and 1600 rpm, in 0.1 M KOH.

As shown in Figure S3, the Tafel slopes of carbonaceous graphene/*E.coli* and Pt /C are 79 and 55 mV per decade, respectively. Both values are comparable to the ideal theoretical value of 60 mV per decade for ORR catalyzed by Pt at room temperature.^{3,4}

References

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3. Y. G. Li, W. Zhou, H. L. Wang, L. M. Xie, Y. Y. Liang, F. Wei, J. C. Idrobo, S. J. Pennycook and H. J. Dai, *Nat Nanotechnol*, 2012, **7**, 394-400.
4. Y. Y. Liang, Y. G. Li, H. L. Wang, J. G. Zhou, J. Wang, T. Regier and H. J. Dai, *Nat Mater*, 2011, **10**, 780-786.