

Supporting information

Methyl ketone bridged molecule as multi-stimuli-responsive color switch for electrochromic device

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1. Structure identification of M5

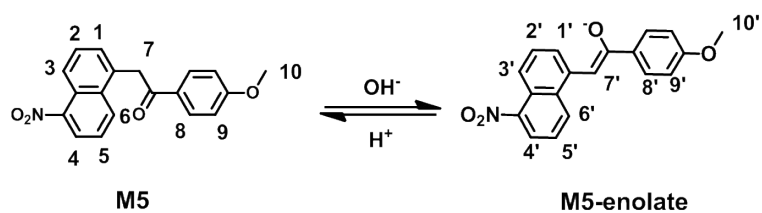


Figure S1: The chemical structures of **M5** and **M5-enolate**.

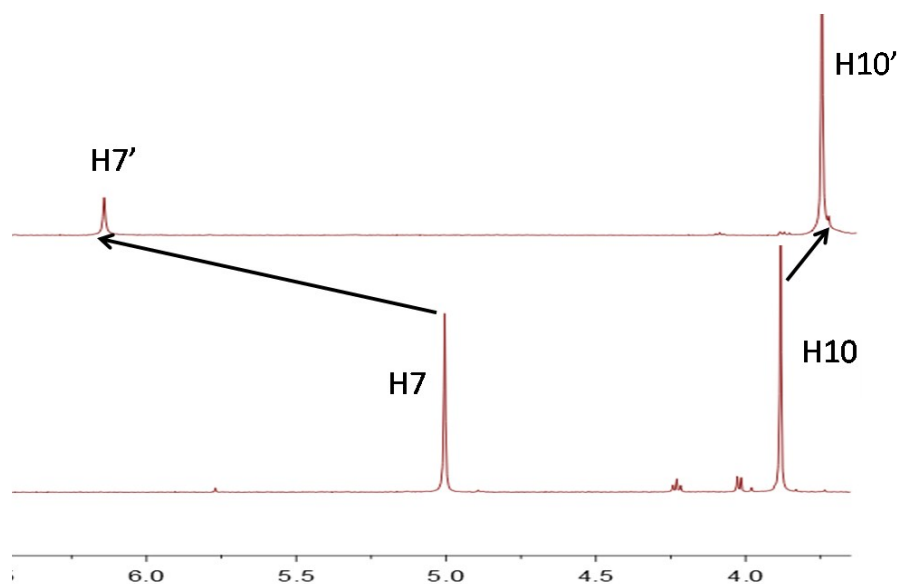


Figure S2: Partial ¹H NMR spectra of **M5** (bottom) and **M5-enolate** (top) obtained by adding potassium *tert*-butoxide to **M5** in DMSO-d⁶.

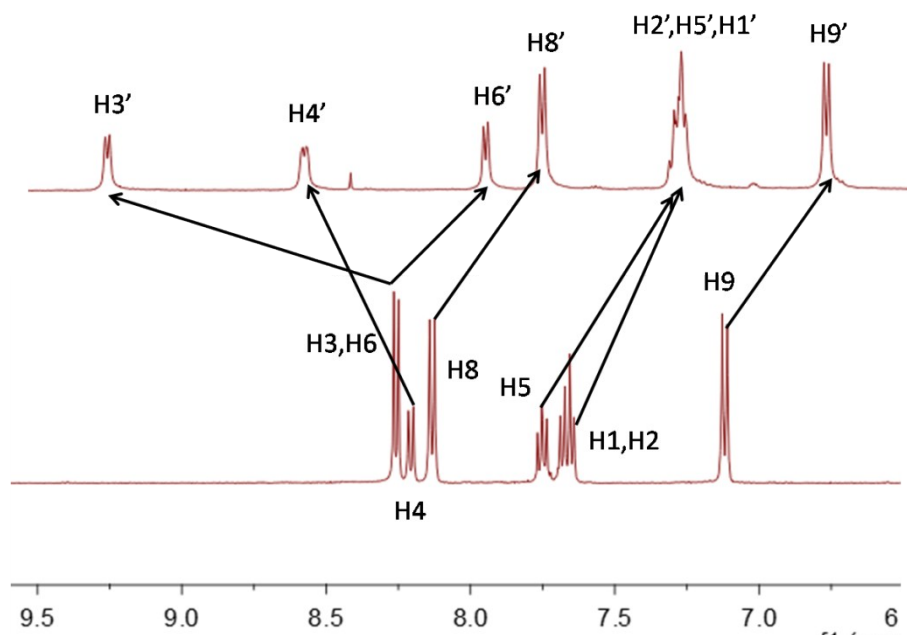


Figure S3: Partial ¹H NMR spectra of **M5** (bottom) and **M5-enolate** (top) obtained by adding potassium *tert*-butoxide to **M5** in DMSO-d⁶.

2. Basochromic and electrochromic properties of M5

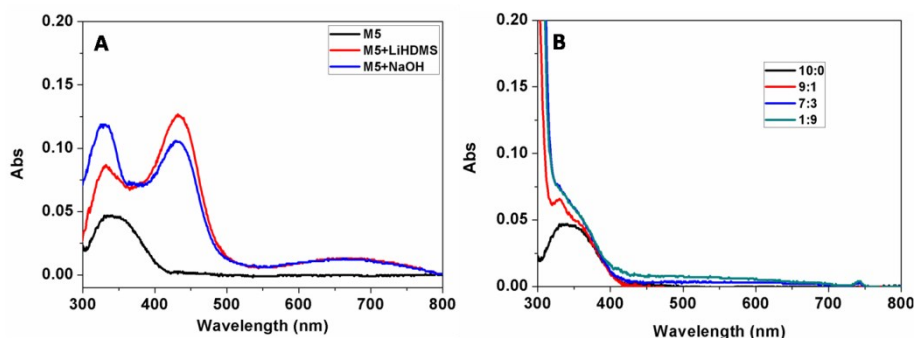


Figure S4: (a) Absorption spectra of **M5** (1.0×10^{-5} M, black curve), treated with 10eq Lithium bis(trimethylsilyl)amide (red curve) and 25 eq NaOH (blue curve) in DMF. (b) Absorption spectra of **M5** (1.0×10^{-5} M) in different DMF/ triethylamine.

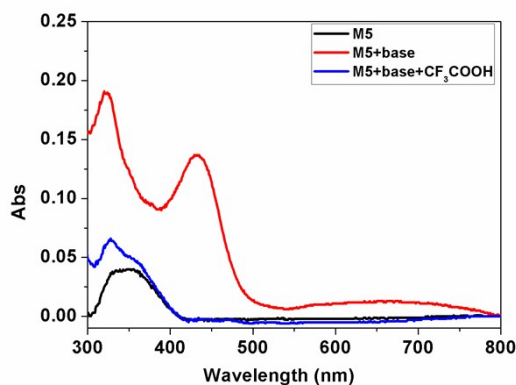


Figure S5: Absorption spectra of **M5** (1.0×10^{-5} M, black curve), treated with 50eq potassium *tert*-butoxide (red curve), then neutralized with CF₃COOH (blue curve) in DMF.

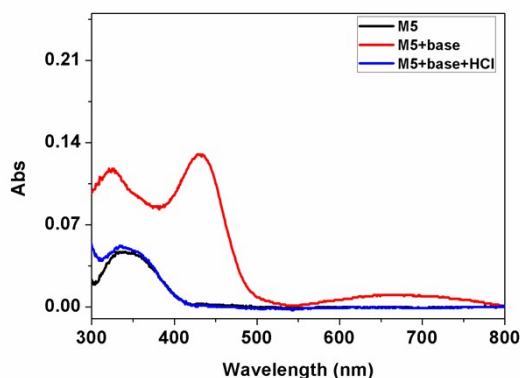


Figure S6: Absorption spectra of **M5** (1.0×10^{-5} M, black curve), treated with 50eq potassium *tert*-butoxide (red curve), then neutralized with HCl (blue curve) in DMF.

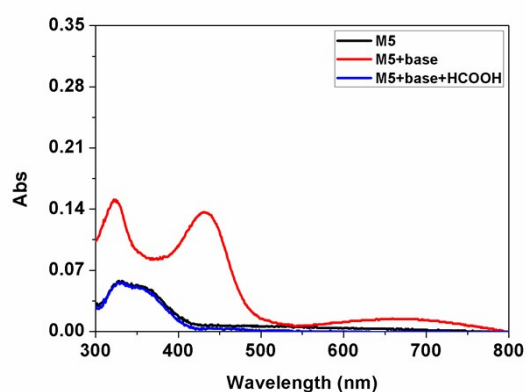


Figure S7: Absorption spectra of **M5** (1.0×10^{-5} M, black curve), treated with 50eq potassium *tert*-butoxide (red curve), then neutralized with HCOOH (blue curve) in DMF.

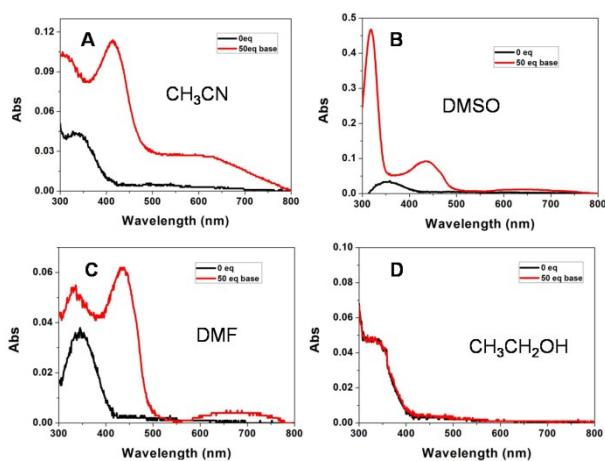


Figure S8: Absorption spectra of **M5** (1.0×10^{-5} M) in different solvents.

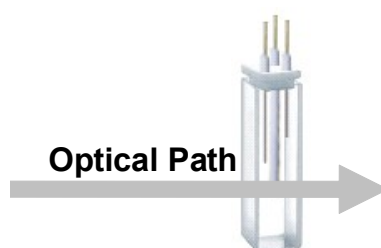


Figure S9: Illustration of *in-situ* Spectroelectrochemical Characterization.

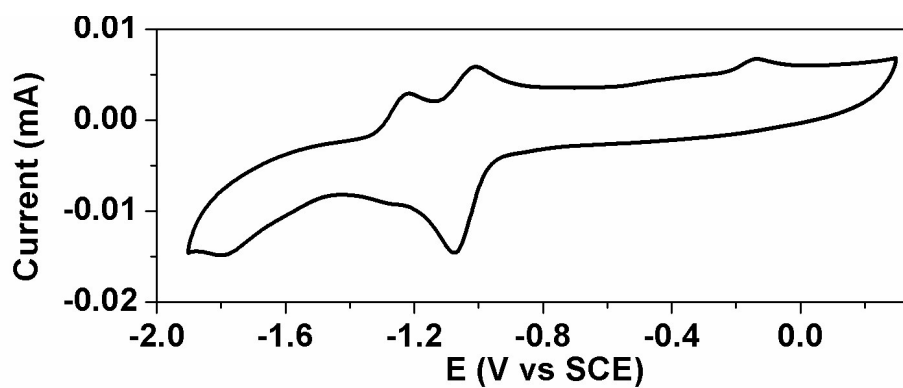


Figure S10: The cyclic voltammogram of **M5** (1.0×10^{-3} M) in DMF with 0.1 M TBAPF₆. Scan rate: 50 mV s⁻¹.

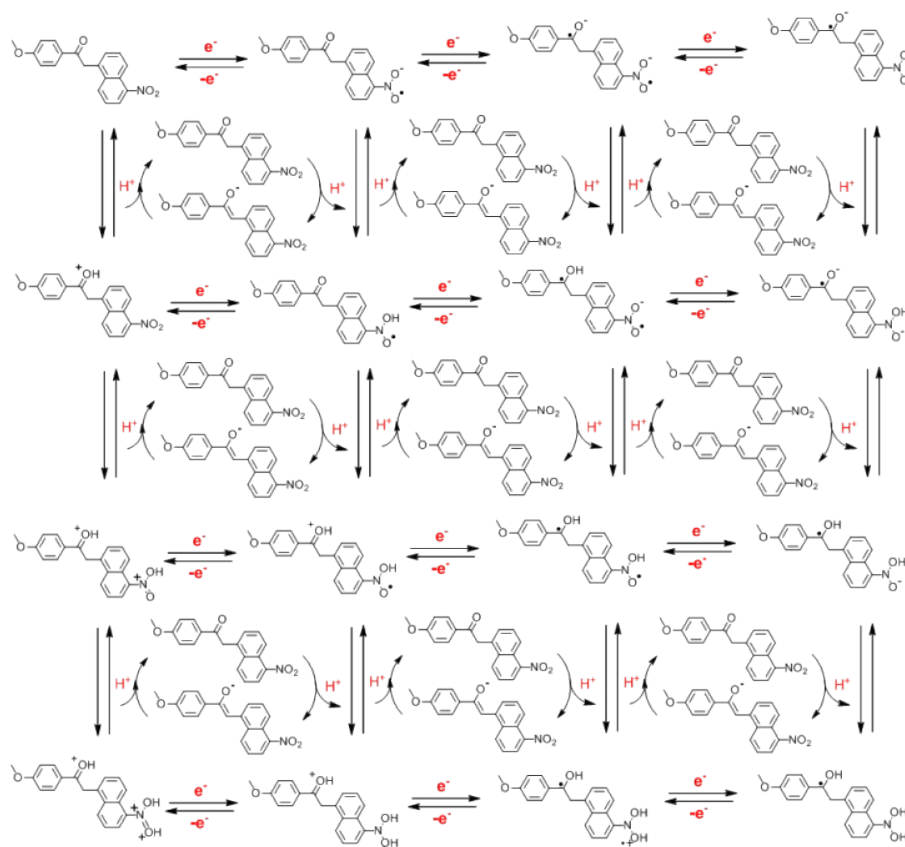


Figure S11. The electrochromic mechanism of **M5** with three-electron reduction mechanism.

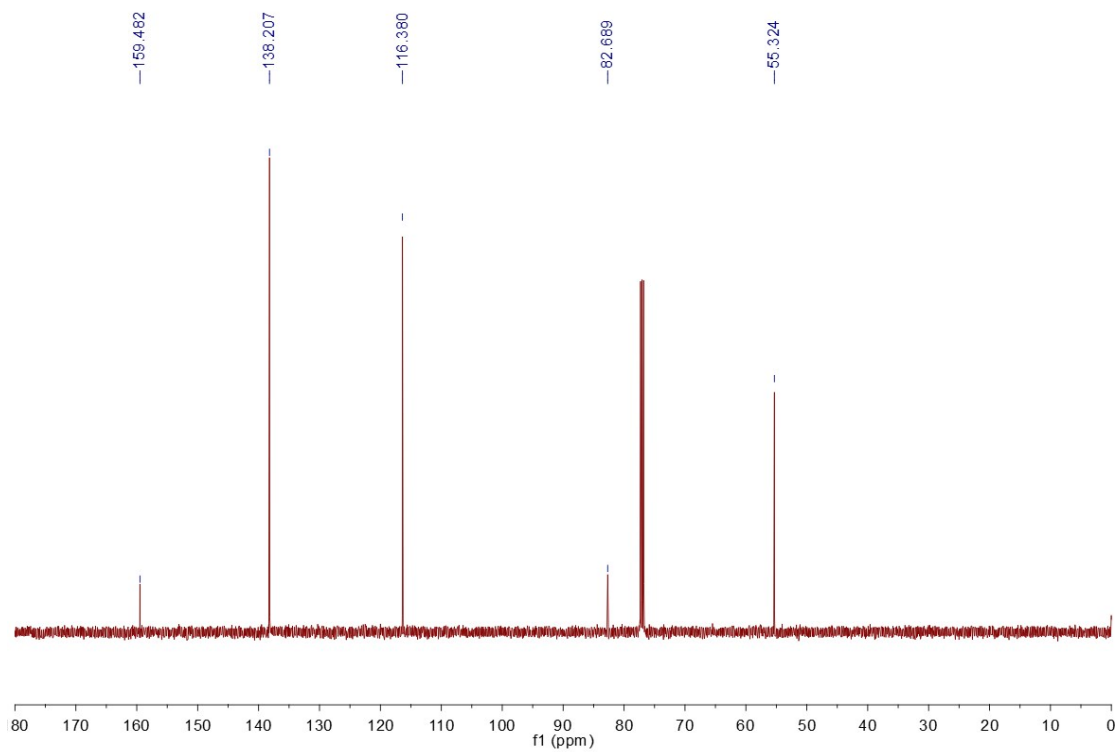
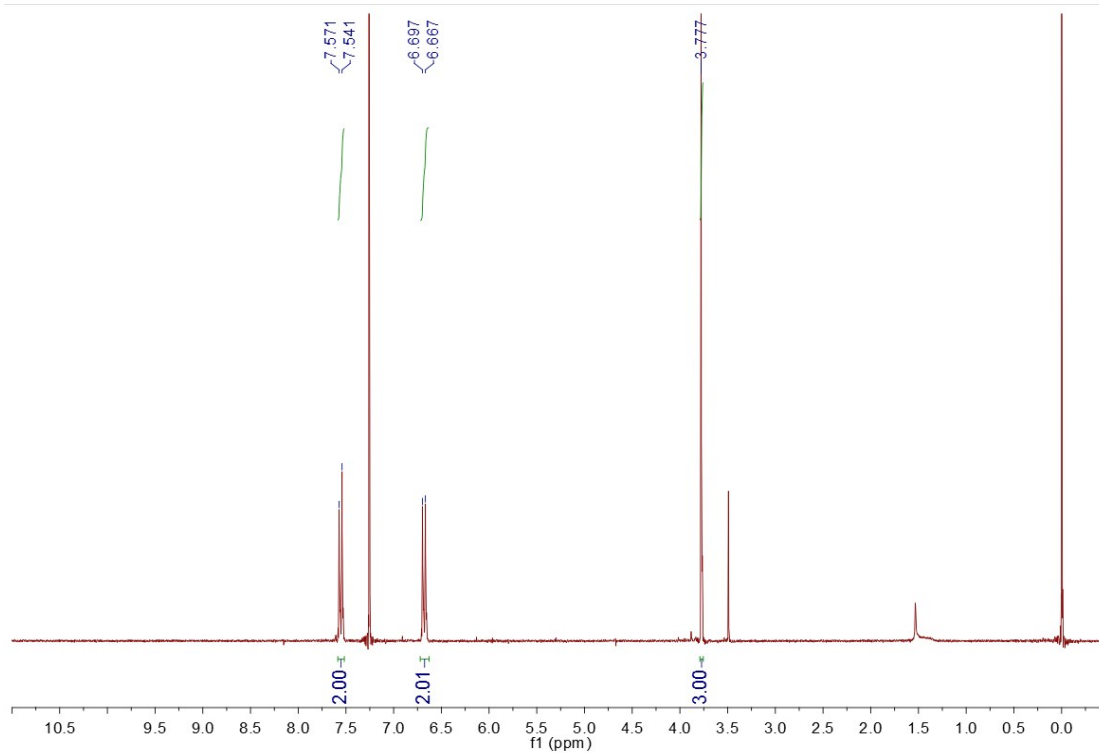
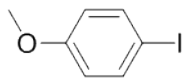
The calculation method of the coloration efficiency:

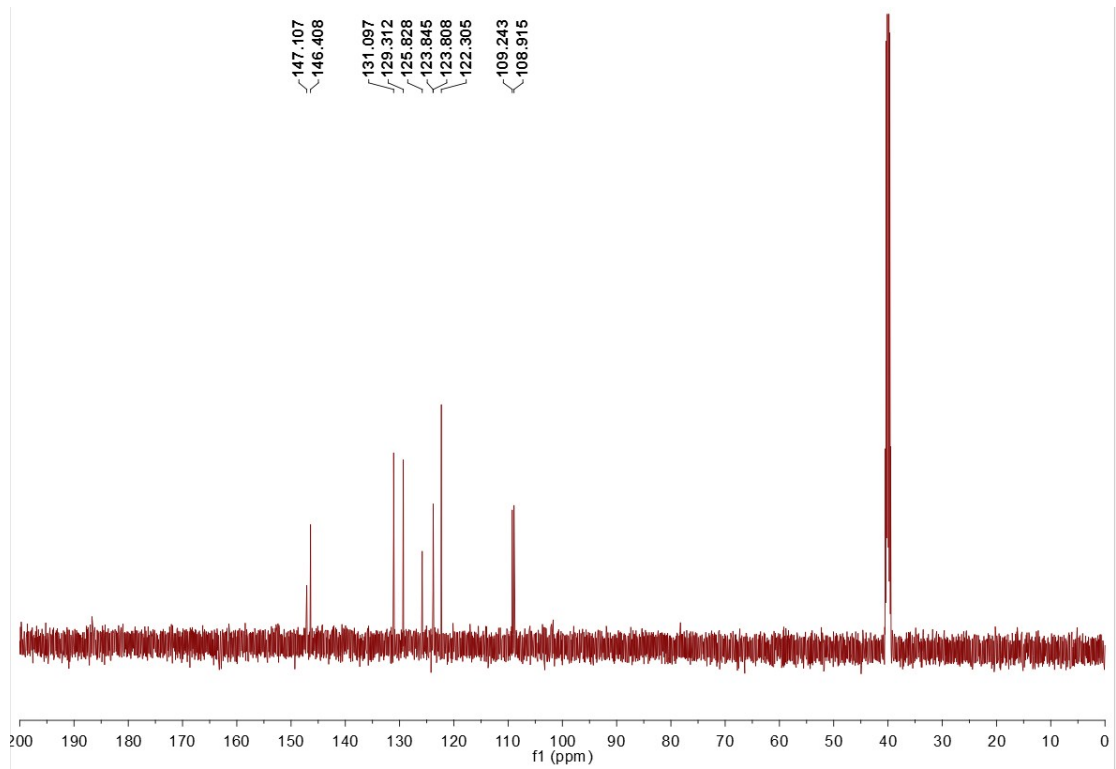
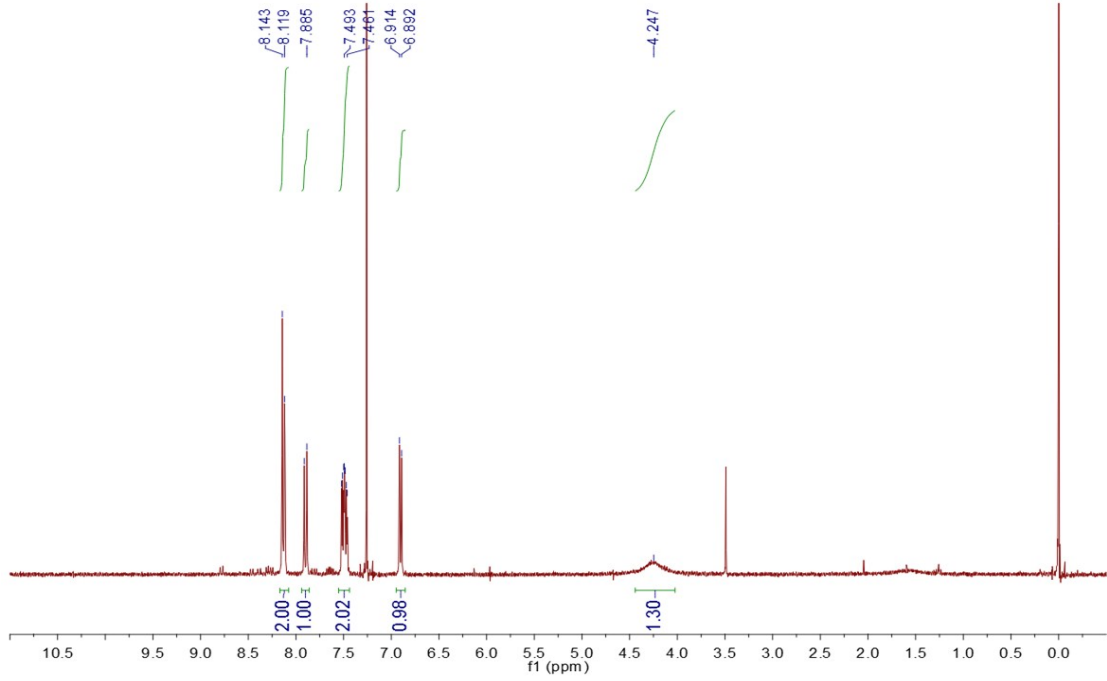
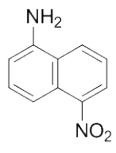
Although the coloration efficiency has been defined in several ways, all definitions are suitable only for an electrochromic polymer. We have modified the definition of the coloration efficiency of the device by eliminating the background interference of electrolytes for small organic electrochromic materials, in our earlier work (Light: Science & Applications, 2015, 4, e249).

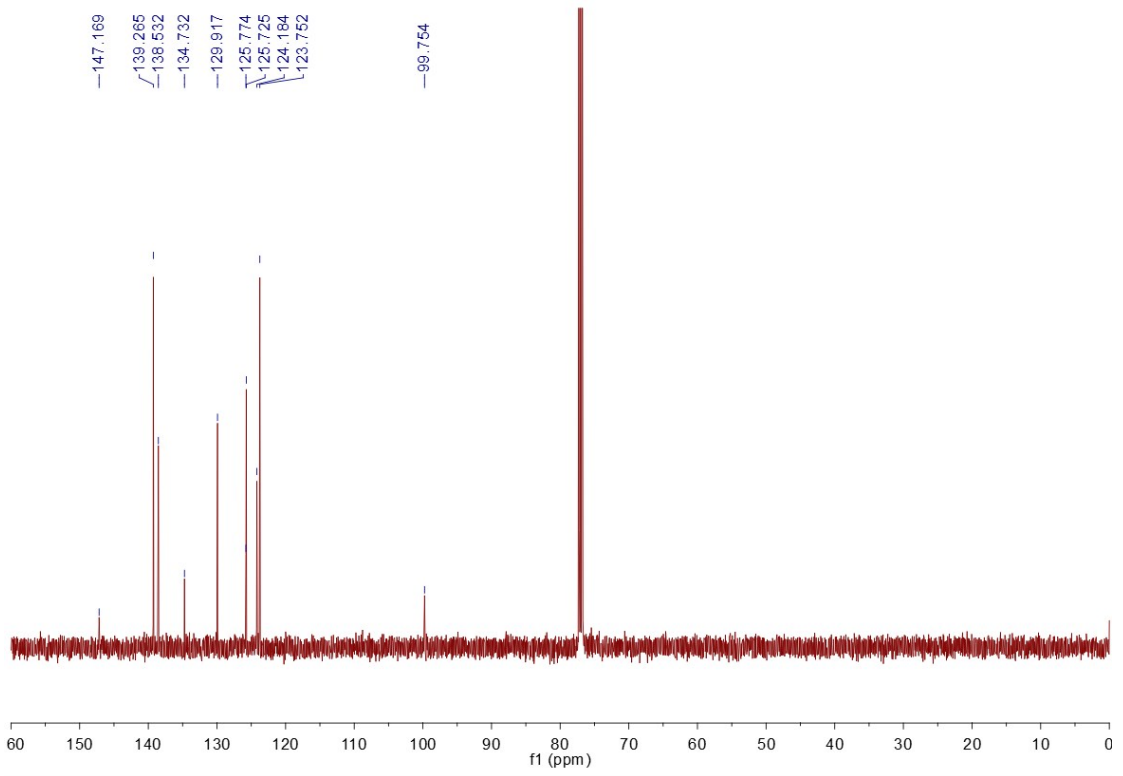
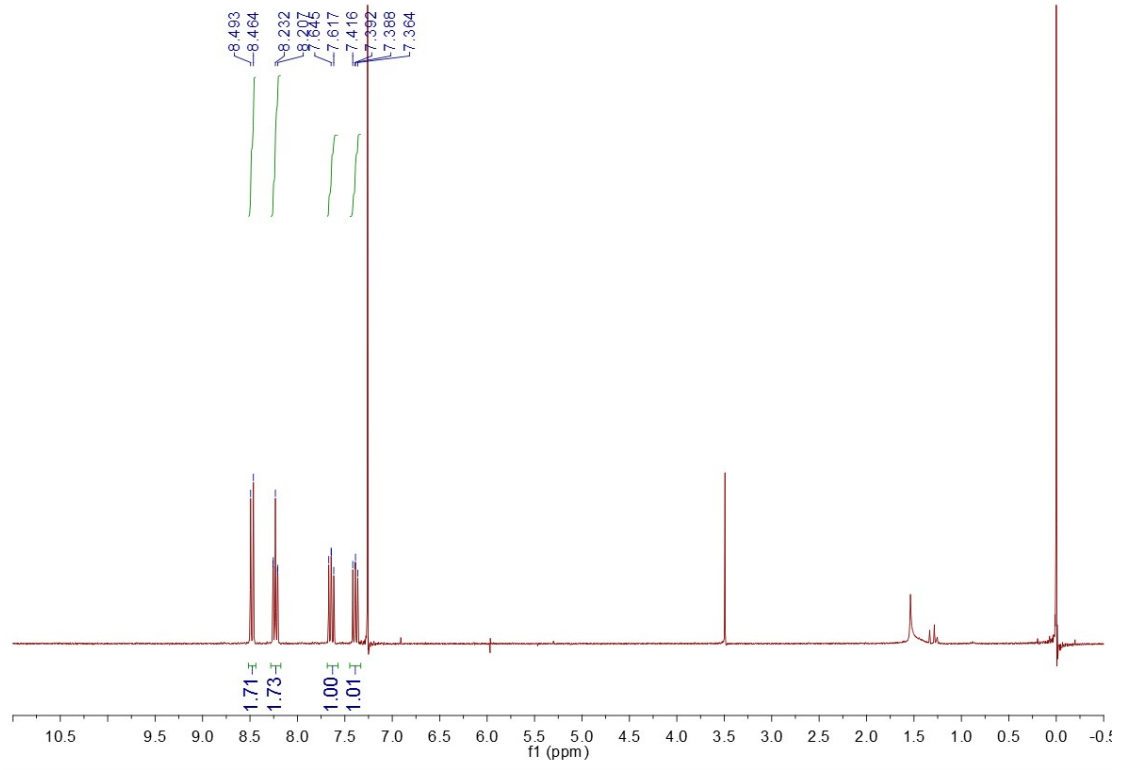
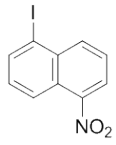
The newly defined coloration efficiency η is given by the optical absorbance change (ΔA) obtained for the change in injected charges per unit area (ΔQ) only, where (ΔQ) is calculated by subtracting the injected charge per unit area (Q_2) of a blank device with no switchable molecules from the injected charge per unit area (Q_1) of electrochromic device subtracts.

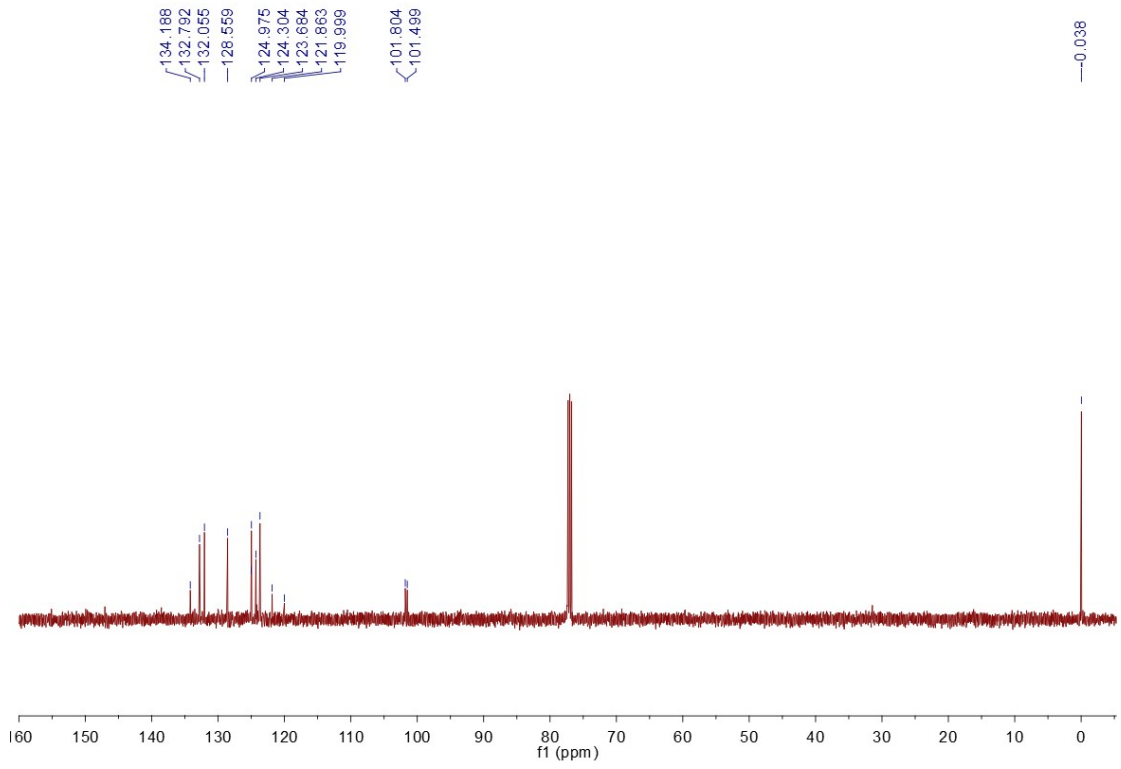
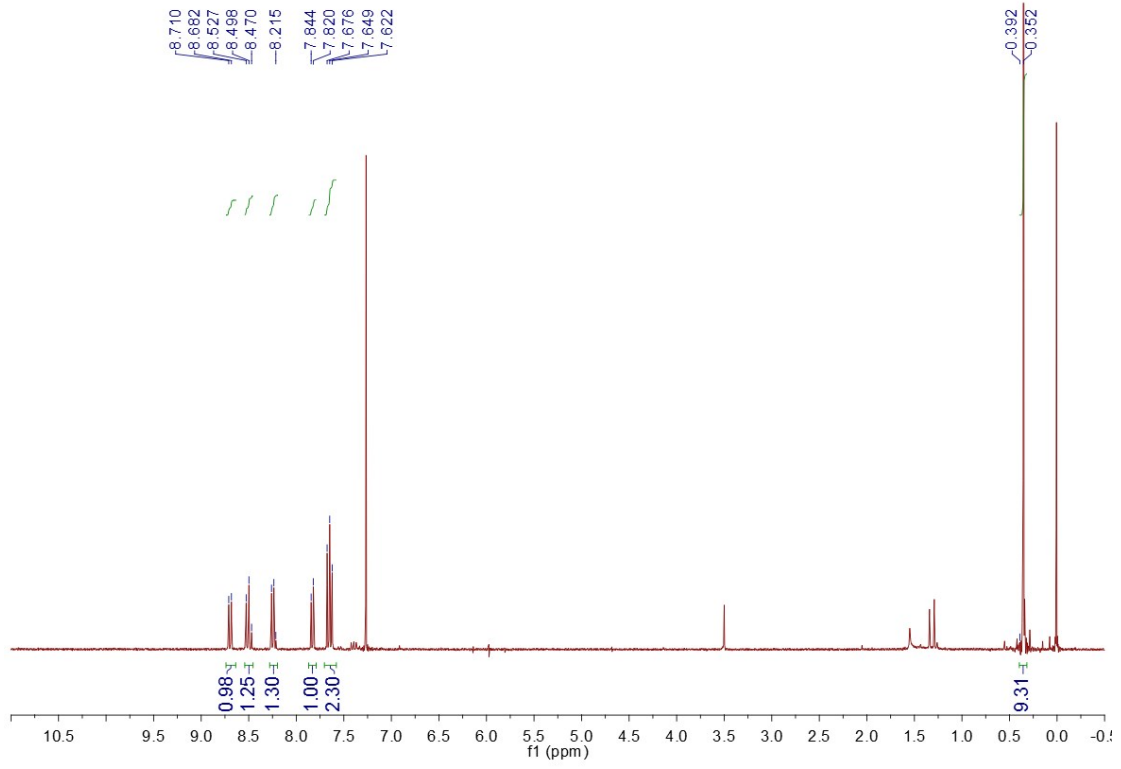
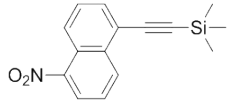
$$\eta = \Delta A / \Delta Q = \Delta A / (Q_1 - Q_2)$$

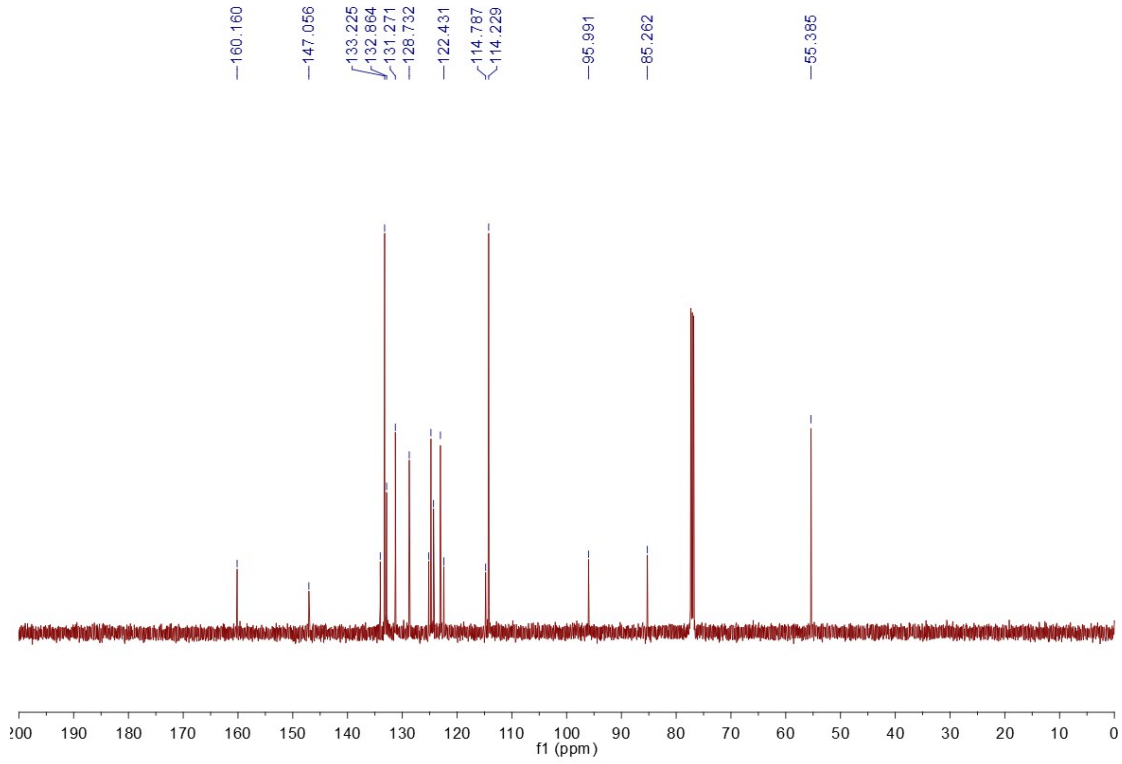
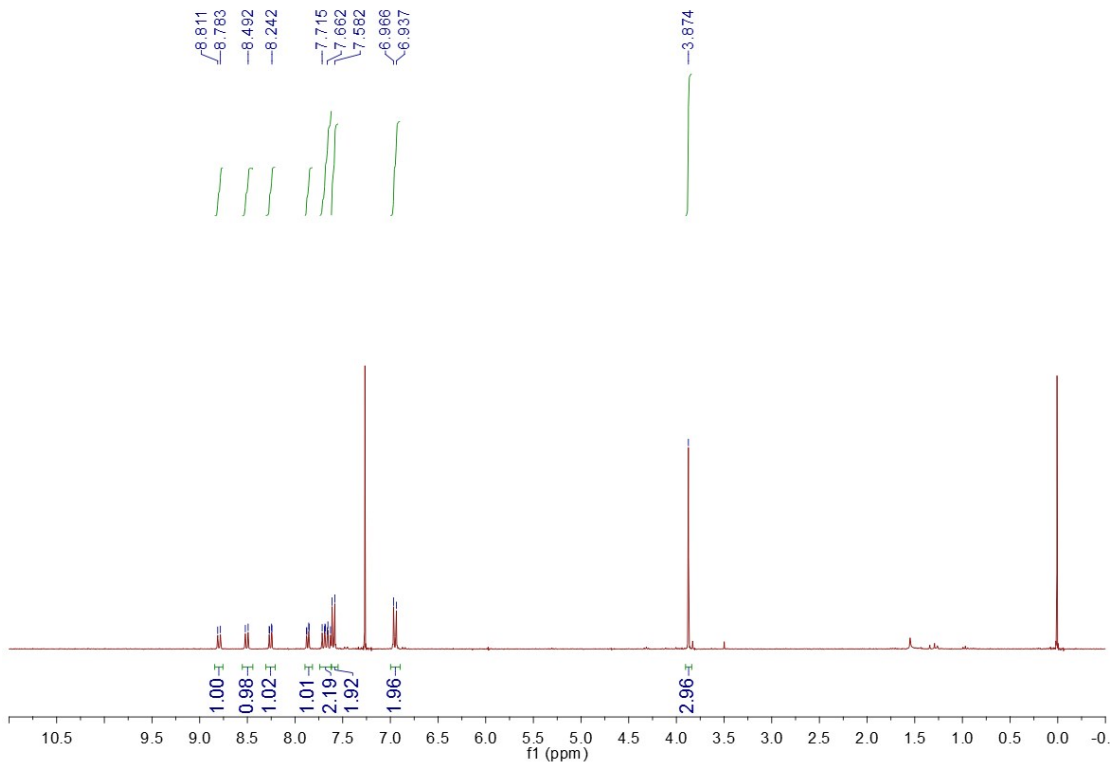
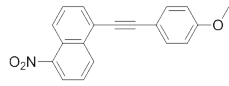
3. ^1H NMR and ^{13}C NMR

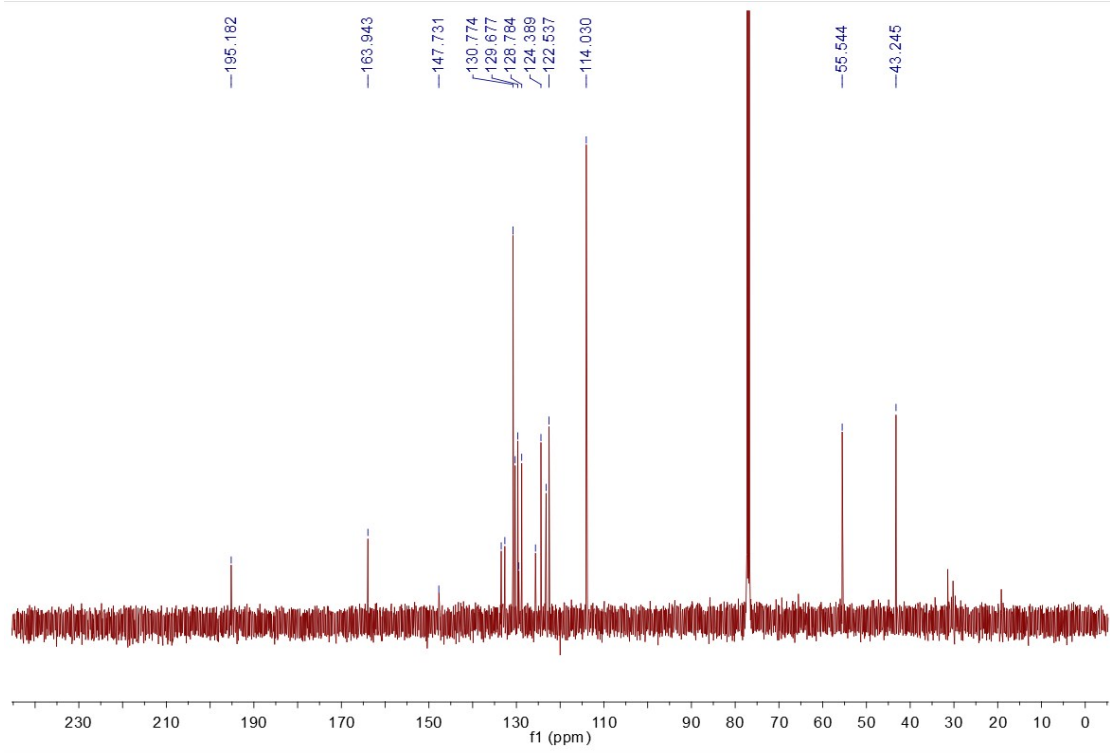
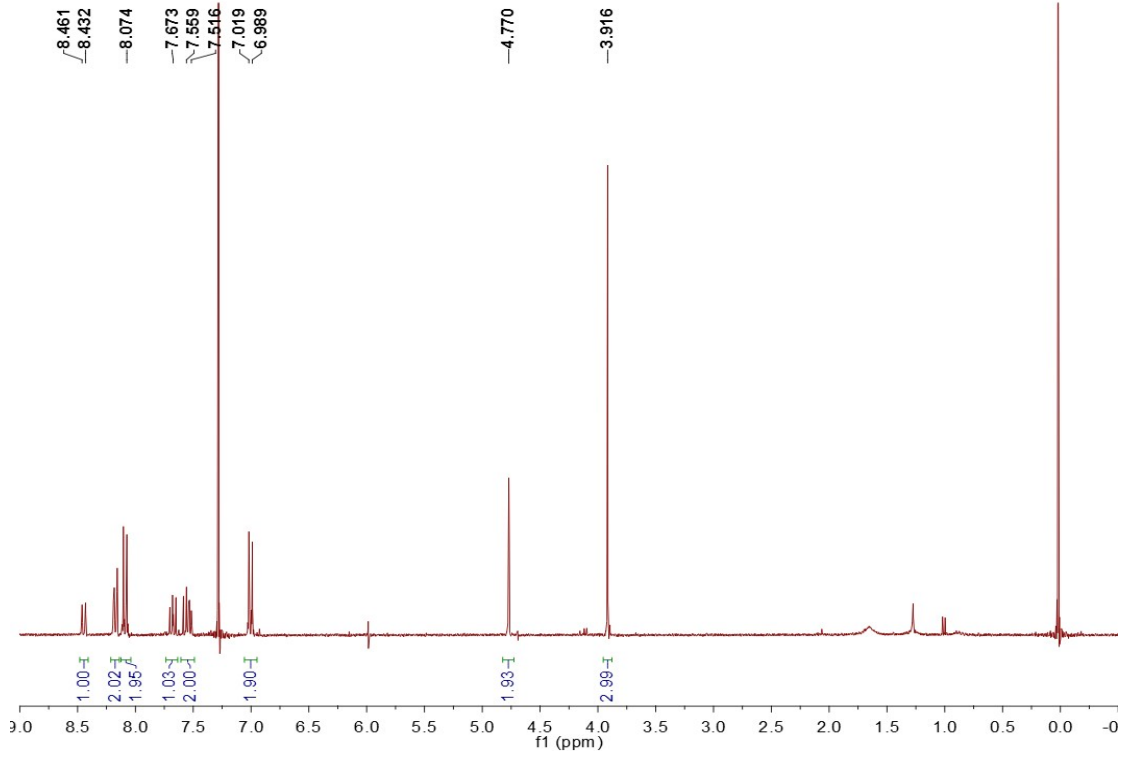
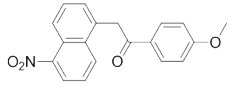




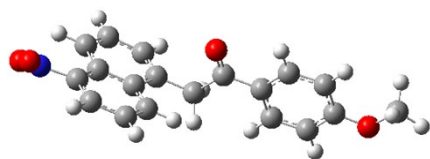




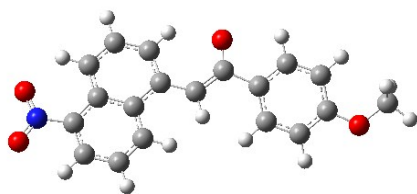




4. The Cartesian Coordinate



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