Electronic Supplementary Information (ESI)

Multi-responsive turn-on flurogenic probe to sense Zn²⁺, Cd²⁺ and Pb²⁺: Left- Right-Center emission signal swing

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Figure S2: Expanded (aromatic region) ¹H-NMR spectra of L



Figure S3: Expanded (aliphatic region) ¹H-NMR spectra of L



Figure S4: ¹³C-NMR spectra of L



Figure S5: Mass spectrum of L



Figure S6: UV-Vis spectra of L (10 μ M) in presence of excess (20 equivalents) of various metal ions in CH₃OH/aqueous HEPES buffer (5mM, pH~7.3; 4:1, v/v) mixed solvent



Figure S7: UV-Vis titration spectra of L ($20\mu M$) with incremental addition of Cu²⁺ in mixed solvent. INSET: Changes in the emission intensity at 573 nm with different concentration of Cu²⁺.



Figure S8: Job's plot for Cu²⁺ from the titration spectra.



Figure S9: Changes in the emission intensity of L at 582 nm with time upon interaction with Pb^{2+} ; $\lambda_{ex} = 450 \text{ nm}$



Figure S10: Fluorescence intensity vs. concentration of Zn^{2+} plot for determination of detection limit



Figure S11: Fluorescence intensity vs. concentration of Cd²⁺ plot for determination of detection limit



Figure S12: Fluorescence intensity vs. concentration of Pb²⁺ plot for determination of detection limit



Figure S13: Job's plot for Zn^{2+} from the fluorescence titration spectra



Figure S14: Job's plot for Cd²⁺ from the fluorescence titration spectra



Figure S15. Mass spectrum of L in presence of Zn²⁺



Figure S16. Mass spectrum of L in presence of Cd²⁺



Figure S17: Mass spectrum of L in presence of Pb²⁺



Figure S18: Optimized structures of L and its Cd^{2+} and Zn^{2+} complexes. For L and it's Zn^{2+} complex the calculations were performed using B3LYP/6-31 G (d,p) as implemented on Gaussian 09. For L-Cd²⁺ complex calculation was performed using B3LYP/6-31 G (d) basis set for all the atoms except for Cd^{2+} , where LANL2DZ effective core potential (ECP) was employed.

References:

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