

Electronic Supplementary Information

Understanding the role of hydrogen bonding in Brønsted acidic ionic liquid-catalyzed transesterification: A combined theoretical and experimental investigation

Supplementary Figures

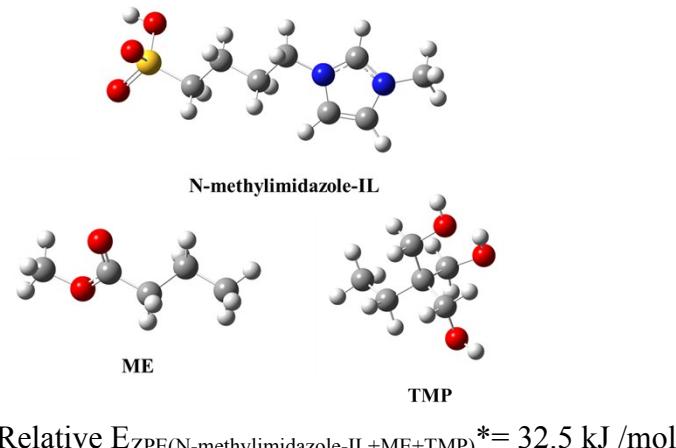
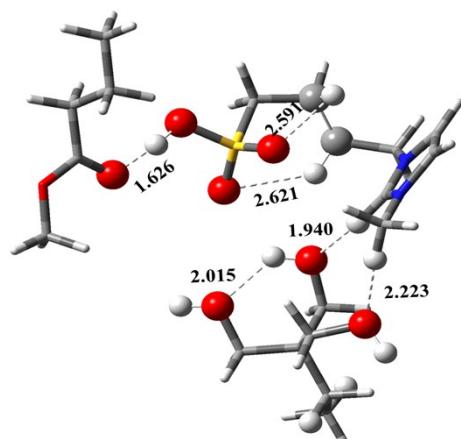


Fig. S1 The most stable geometries of the N-methylimidazole-IL/TMP/ME mixture with intra- and inter-molecular hydrogen bonds (a1) and without hydrogen bonds (a2)

It can be seen that the relative energy E_{ZPE}^* of geometry a2 was much higher, implying the significant contribution of HBs for stabilizing this system.

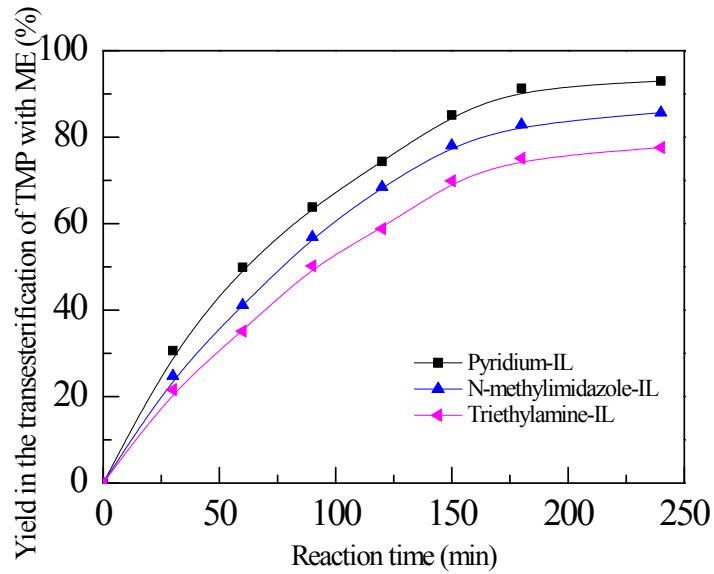


Fig. S2 Transesterification of TMP with ME catalyzed by these three ILs

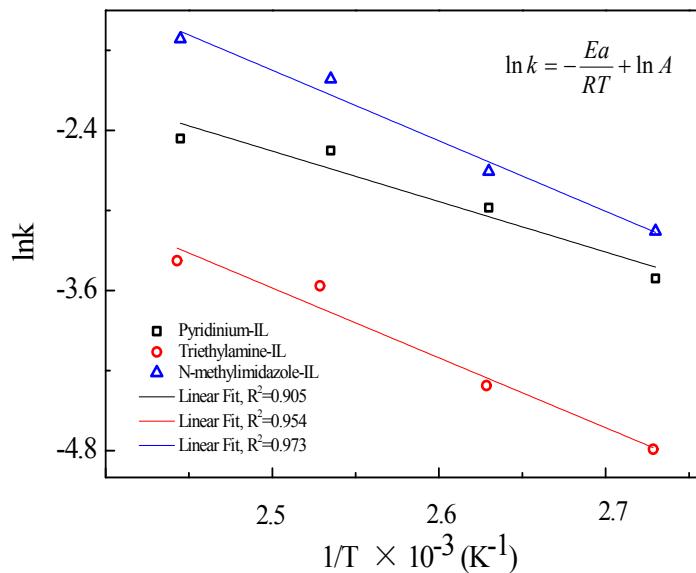


Fig. S3 Kinetics plot of ln k vs. 1/T for the triethylamine-IL catalyzed transesterification, the slope represents [-Ea/R] and the intercept represents lnA respectively.

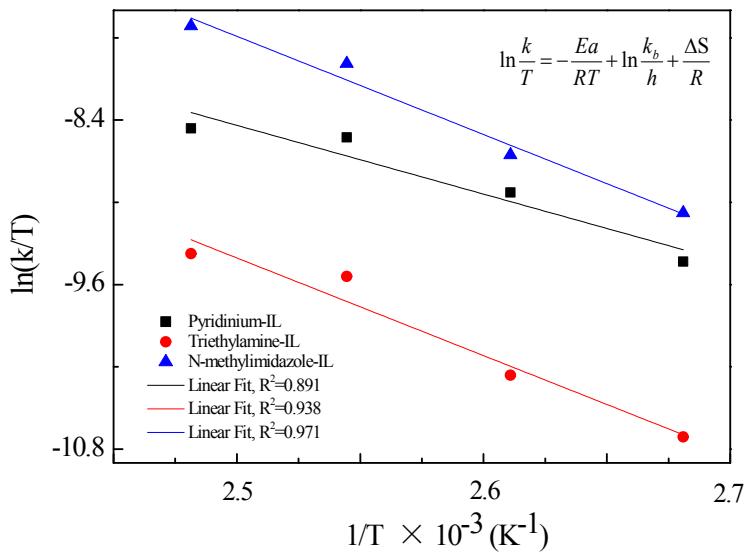


Fig. S4 Kinetics plot of $\ln(k/T)$ vs. $1/T$ for the triethylamine-IL catalyzed transesterification, the slope represents $[-\Delta H/R]$ and the intercept represents $[\ln(kB/h) + \Delta S/R]$ in the Eyring equation respectively.

Supplementary Table

Table S1 Binding energies (ΔE_{bin}) and normalized delocalization energies (DE_{na}) calculated for the three BAIL-containing systems at the B3LYP/6-311++G level of theory ^a

IL	ΔE_{bin} (kJ/mol)	DE_{na} (kJ/mol)
pyridinium-IL	132.56	80.2
triethylamine-IL	130.45	0
N-methylimidazole-IL	152.16	55.2

^aThe binding energies included BSSE and ZPVE corrections based on the results. The normalized delocalization energies were calculated according to the methods proposed by Sabine et al^[54]

Table S2 Comparison for the activation energy in the transesterification between DFT computed value and experiment result.

IL	Activation Energy ΔE (kJ/mol)			
	DFT computed result	DFT with HSO_4^-	DFT with HSO_4^- under implicit solvation	Experiment result
pyridinium-IL	199.6	157.3	116.8	108.3
triethylamine-IL	212.7	172.1	126.5	113.3
N-methylimidazole-IL	201.2	162.5	120.1	109.0