

Supporting Information

Electrophilic Iodination: A Gateway to High Iodine Compounds and Energetic Materials

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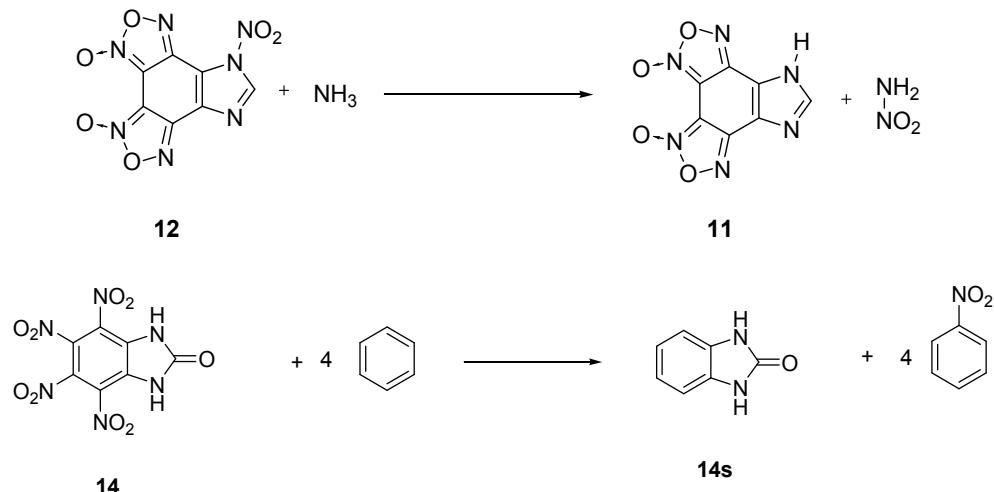
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Calculations – University of Idaho

Gaussian Calculations

Geometric optimization and frequency analyses of the compounds in the isodesmic reactions were accomplished by using the B3LYP functional with the 6-31+G** basis set from Gaussian 03.^[1] All of the optimized structures were characterized to be true local energy minima on the potential energy surface without imaginary frequencies. The enthalpy of reaction is obtained by combining the MP2/6-311++G** energy difference for the reactions, the scaled zero point energies (ZPE), values of thermal correction (H_T), and other thermal factors. The gas phase heat of formation at 298.15K for compound **11** was calculated from the atomization energy at the G2 level, gas phase enthalpies of **12** and **14** were calculated based on isodesmic reactions in Scheme S1 at MP2/6-311++G(d,p)//B3LYP(6-31+G(d,p)) level. The gas phase enthalpies of formation of **11**, **12** and **14** were converted to the solid state enthalpies of formation by subtraction of sublimation enthalpy calculated according to Trouton's rule ($\Delta H_{\text{sub}} = 188T_m$) (Table S1).^[2]



Scheme S1: Isodesmic reactions for compounds **12** and **14**.

Table S1. Calculated (MP2/6-311++G**//B3LYP/6-31+G**) total energy (E_0), corrected MP2 total energy (E_{corr}), zero-point energy (ZPE), thermal correction to enthalpy (H_T), heat of reaction for the isodesmic reaction (ΔH_R) and gas phase heats of formation (ΔH_m°) from Gaussian 03.^[1]

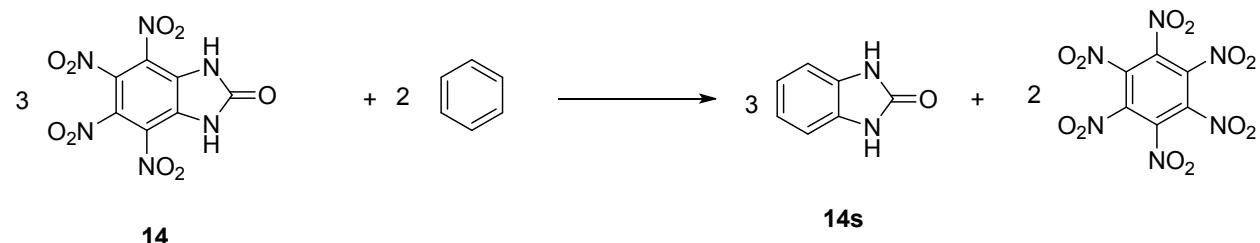
	ZPE [Hartree/Particle]	H _T [Hartree/Particle]	E ₀ [Hartree/Particle]	ΔH _R (Hartree/Particle)	ΔH _{Sub} ^a (kJ mol ⁻¹)	Δ _f H _m [°] (g) [kJ mol ⁻¹]	Δ _f H _m [°] (s) [kJ mol ⁻¹]
Benzene	0.100441	0.105785	-231.5842377	-	-	+82.9 ^b	-
Nitrobenzene	0.103043	0.110809	-435.6906555	-	-	+68.5 ^b	-
11	0.109795	0.1223	-895.031117	-	88.7	+604.2 ^c	+515.5
12	0.110929	0.126049	-1099.084817	0.05495	79.1	+604.3	+525.2
14s	0.122904	0.131114	-453.9743248	-	-	-63.9 ^d	-
14	0.131244	0.149867	-1270.361893	-0.03684	107.5	-24.8	-132.3

^a Heat of sublimation calculated according to Trouton's rule ($\Delta H_{\text{Sub}} = 0.188T_d$) ref: [2]. ^b Gas phase heat of formation NIST ref [3]; ^c Calculated from G2. ^d Experimental gas phase heat of formation ref [4]

Calculations – Dr. Jerry Boatz, Air Force Research Laboratory, Edwards Air Force Base, CA

GAMESS Calculations

The calculations for compound **14** were also performed by using GAMESS quantum chemistry program^[5,6] with a different isodesmic reaction shown in Scheme S2. The heat of reaction is obtained by combining corrected



Scheme S2: Isodesmic reaction for **14**

B3LYP/6-311++(d,p) energy difference of the products and reactants calculated at B3LYP /6-311++g(d,p)//B3LYP6-311++g(d,p)) level. The gas phase heats of fomation for the related species were calculated from the atomization energy at the G3(MP2) level. The heat of sublimation was obtained using an approach developed by Rice and Byrd (Table S2).^[7,8]

Table S2: Calculation of heat of formation of **14**

	B3LYP/6-311++(d,p)	B3LYP ZPE	B3LYP scaled ZPE	b3lyp elect+scaled ZPE	b3lyp thermal correction	ΔH_R (Hartree/Particle)	ΔH_f kcal/mol	ΔH_f kJ/mol	ΔH_{Sub} (kJ mol ⁻¹)	$\Delta_f H_m^\circ(s)$ [kJ mol ⁻¹]
14	-1272.842640	0.130986	0.128759	-1272.713881	0.018791	0.07378	1.5	6.3	120.9	-114.6
14s	-454.987584	0.122195	0.120118	-454.867467	0.008385				-62.9 ^a	
Hexanitro benzene	-1458.901699	0.111421	0.109527	-1458.792172	0.021663				279.1 ^a	
Benzene	-232.157071	0.099939	0.098240	-232.058831	0.005412				78.2 ^a	

^a Calculated based on G3(MP2) using GAMESS.

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Table S3. Crystallographic data for **5**, **10** and **14**

Compound	5	10•C₂H₆OS	14•C₄H₈O₂
Formula	C ₈ H ₅ I ₂ N ₇ O ₄	C ₉ H ₈ I ₂ N ₄ O ₅ S	C ₁₁ H ₁₀ N ₆ O ₁₁
CCDC number	1424001	1424002	1424003
M _w	516.99	538.05	402.25
Crystal size [mm ³]	0.209 x 0.034 x 0.005	0.123 x 0.057 x 0.032	0.310 x 0.175 x 0.058
Crystal system	Monoclinic	Monoclinic	Triclinic
Space group	P2 ₁ /c	P2 ₁ /c	P-1
a [Å]	16.6797(6)	5.34160(10)	7.8304(7)
b [Å]	98.956(2)	25.7684(6)	9.3182(8)
c [Å]	9.1213(3)	11.3694(3)	12.2176(11)
α [°]	90	90	102.459(3)
β [°]	107.606(19)	97.2720(10)	104.030(3)
γ [°]	90	90	100.776(3)
V [Å ³]	1339.59(8)	1552.35(6)	817.12(13)
Z	4	4	2
T [K]	150(2)	273(2)	296(2)
ρ _{calcd} [Mg m ⁻³]	2.498	2.302	1.635
μ [mm ⁻¹]	37.190	4.211	0.149
F(000)	960	1008	412
θ [°]	2.682 to 68.705	3.613 to 26.381	2.506 to 26.422
Index ranges	-17<=h<=20 -10<=k<=10 -10<=l<=10	-6<=h<=6 -32<=k<=32 -14<=l<=14	-8<=h<=9 -11<=k<=11 -15<=l<=15
Reflections collected	7543	14724	7885
Independent reflections (R _{int})	2389 [R _(int) = 0.0453]	3168 [R _(int) = 0.0346]	3321 [R _(int) = 0.0157]
Data/restraints/parameters	2389 / 1 / 193	3168 / 0 / 196	3321 / 0 / 275
GOF on F ²	1.024	0.827	1.069
R ₁ (I > 2δ(I)) ^a	0.0322	0.0275	0.0381
wR ₂ (I > 2δ (I)) ^b	0.0718	0.0758	0.1067
R ₁ (all data)	R ₁ = 0.0536	0.0410	0.0536
wR ₂ (all data)	0.0797	0.0853	0.1197
Largest diff. peak and hole [e. Å ⁻³]	0.830 and -0.874	0.855 and -0.449	0.257 and -0.203

$$^a R_1 = \sum |F_0| - |F_c| / \sum |F_0| \quad ^b R_2 = [\sum w(F_0^2 - F_c^2)^2 / \sum w(F_0^2)^2]^{1/2}$$

