

## Supporting Information

### Highly Sensitive and Selective Fluorescent Probe for Hg<sup>2+</sup> in Ag(I)/Cu(II) 3D Supramolecular Architecture Based on Noncovalent Interactions

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**Table S1** Crystal data and structure refinement parameters of compounds **1**, **2** and **3**.

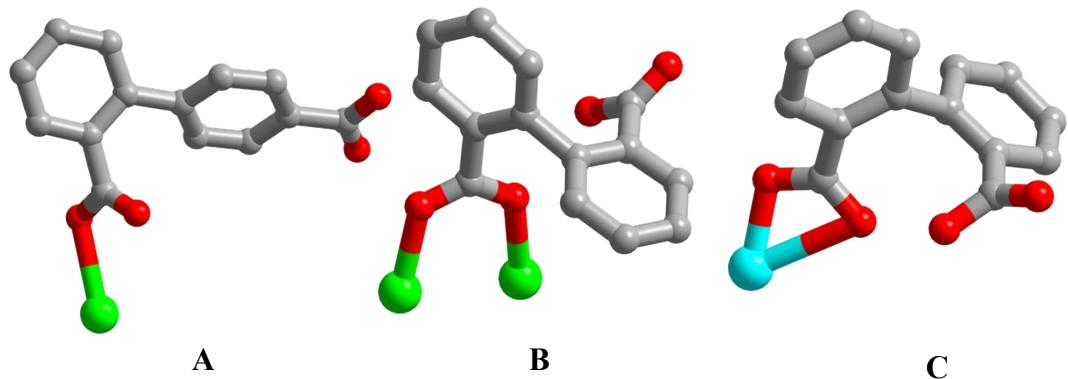
Identification code	<b>1</b>	<b>2</b>	<b>3</b>
Empirical formula	C <sub>24</sub> H <sub>17</sub> AgN <sub>2</sub> O <sub>4</sub>	C <sub>19</sub> H <sub>13</sub> AgNO <sub>4</sub>	C <sub>40</sub> H <sub>28</sub> CuN <sub>4</sub> O <sub>8</sub>
Formula mass	505.27	427.17	756.21
Crystal system	Monoclinic	Monoclinic	Triclinic
Space group	<i>C</i> 2/c	<i>C</i> 2/c	<i>P</i> Error!
<i>a</i> (Å)	21.460(6)	29.457(9)	8.0398(16)
<i>b</i> (Å)	11.416(3)	14.831(5)	9.1650(18)
<i>c</i> (Å)	17.625(5)	7.499(2)	12.290(3)
$\alpha$ (°)	90.00	90.00	111.73(3)
$\beta$ (°)	115.291(9)	99.616(5)	94.75(3)
$\gamma$ (°)	90.00	90.00	97.18(3)
<i>V</i> (Å <sup>3</sup> )	3904.0(19)	3230.1(17)	826.4(3)
<i>Z</i>	8	8	1
<i>D</i> <sub>v</sub> (g cm <sup>-3</sup> )	1.719	1.703	1.519
$\mu$ (Mo K $\alpha$ )/mm <sup>-1</sup>	1.069	1.270	0.726
<i>F</i> (000)	2032	1600	389
Crystal size	0.16 x 0.14 x 0.12 mm	0.30 x 0.28 x 0.24 mm	0.32 x 0.15 x 0.12
$\theta$ range (°)	3.36 – 25.00	1.40 – 26.88	3.23 – 27.56
Limiting indices	$-25 \leq h \leq 25$ $-11 \leq k \leq 13$ $-20 \leq l \leq 20$	$-36 \leq h \leq 34$ $-18 \leq k \leq 17$ $-9 \leq l \leq 9$	$-10 \leq h \leq 10$ $-11 \leq k \leq 11$ $-15 \leq l \leq 15$
Data/Restraints/Parameters	3392 / 7 / 291	3188 / 0 / 229	3745 / 0 / 241
GOF on <i>F</i> <sup>2</sup>	0.952	0.999	1.071
Final <i>R</i> indices [ <i>I</i> > 2σ( <i>I</i> )]			
<i>R</i> <sub>1</sub> <sup>a</sup>	0.0651	0.0566	0.0773
<i>wR</i> <sub>2</sub> <sup>b</sup>	0.1526	0.1428	0.2085
<i>R</i> indices (all data)			
<i>R</i> <sub>1</sub>	0.1079	0.0872	0.0885
<i>wR</i> <sub>2</sub>	0.1265	0.1660	0.2142
CCDC	1477431	1477429	1477430

<sup>a</sup>  $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ ; <sup>b</sup>  $wR_2 = [\sum [w(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}$ .

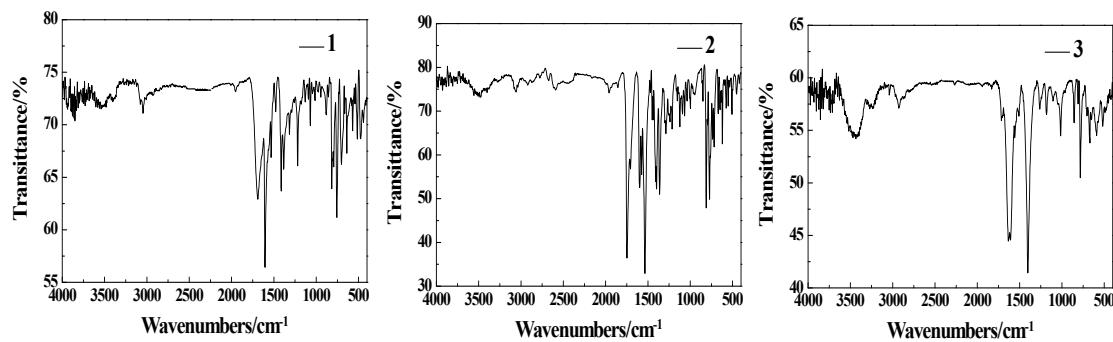
**Table S2** Selected bond lengths [ $\text{\AA}$ ] and angles [ $^\circ$ ] for **1**, **2** and **3**.

Bond length	( $\text{\AA}$ )	Bond angle	( $^\circ$ )
<b>1</b>			
Ag(1A)-N(1)	2.193(6)	N(1)-Ag(1A)-N(2)#1	159.7(3)
Ag(1A)-N(2)#1	2.214(6)	N(1)-Ag(1A)-O(2)	93.8(2)
Ag(1A)-O(2)	2.512(7)	N(2)#1-Ag(1A)-O(2)	105.9(2)
Ag(1B)-N(1)	2.201(6)	N(1)-Ag(1B)-N(2)#1	158.9(6)
Ag(1B)-N(2)#1	2.212(6)	N(1)-Ag(1B)-O(1)#2	106.6(4)
Ag(1B)-O(1)#2	2.572(11)	N(2)#1-Ag(1B)-O(1)#2	92.4(3)
Ag(1C)-N(1)	2.208(6)	N(1)-Ag(1B)-Ag(1B)#2	96.4(3)
Ag(1C)-N(2)#1	2.214(6)	N(2)#1-Ag(1B)-Ag(1B)#2	96.8(2)
Ag(1C)-O(2)	2.688(7)	O(1)#2-Ag(1B)-Ag(1B)#2	76.1(4)
Ag(1C)-Ag(1C)#2	2.807(9)	N(1)-Ag(1C)-N(2)#1	157.7(3)
Ag(1D)-N(1)	2.162(6)	N(1)-Ag(1C)-O(2)	88.8(2)
Ag(1D)-N(2)#1	2.187(6)	N(2)#1-Ag(1C)-O(2)	100.3(2)
N(2)-Ag(1B)#3	2.212(6)	N(1)-Ag(1C)-Ag(1C)#2	98.4(2)
N(2)-Ag(1A)#3	2.214(6)	N(2)#1-Ag(1C)-Ag(1C)#2	98.6(2)
O(1)-Ag(1B)#2	2.572(11)	O(2)-Ag(1C)-Ag(1C)#2	105.9(3)
		N(1)-Ag(1D)-N(2)#1	172.0(4)
<b>2</b>			
Ag(1)-O(1)	2.148(4)	O(1)-Ag(1)-O(1)#1	164.6(2)
Ag(1)-O(1)#1	2.149(4)	O(1)-Ag(1)-N(1)	97.71(10)
Ag(1)-N(1)	2.397(6)	O(1)-Ag(1)-Ag(2)	82.29(10)
Ag(1)-Ag(2)	2.9584(13)	N(1)-Ag(1)-Ag(2)	180.0
Ag(2)-O(2)	2.231(3)	O(2)-Ag(2)-O(2)#1	157.30(19)
Ag(2)-N(2)#2	2.392(6)	O(2)-Ag(2)-N(2)#2	101.35(9)
		O(2)-Ag(2)-Ag(1)	78.65(9)
<b>3</b>			
Cu(1)-N(2)	1.973(4)	N(2)-Cu(1)-N(2)#1	179.998(1)
Cu(1)-O(1)	1.978(4)	N(2)-Cu(1)-O(1)#1	88.07(17)
Cu(1)-O(2)	2.671(4)	N(2)-Cu(1)-O(1)	91.93(17)
		N(2)#1-Cu(1)-O(1)	88.07(17)
		O(1)#1-Cu(1)-O(1)	179.999(2)
		N(2)-Cu(1)-O(2)	96.09(16)
		N(2)#1-Cu(1)-O(2)	83.91(16)
		O(1)#1-Cu(1)-O(2)	125.72(14)
		O(1)-Cu(1)-O(2)	54.28(14)
		N(2)-Cu(1)-O(2)#1	83.91(16)
		N(2)#1-Cu(1)-O(2)#1	96.09(16)
		O(1)#1-Cu(1)-O(2)#1	54.28(14)
		O(1)-Cu(1)-O(2)#1	125.72(14)
		O(2)-Cu(1)-O(2)#1	180.00(18)

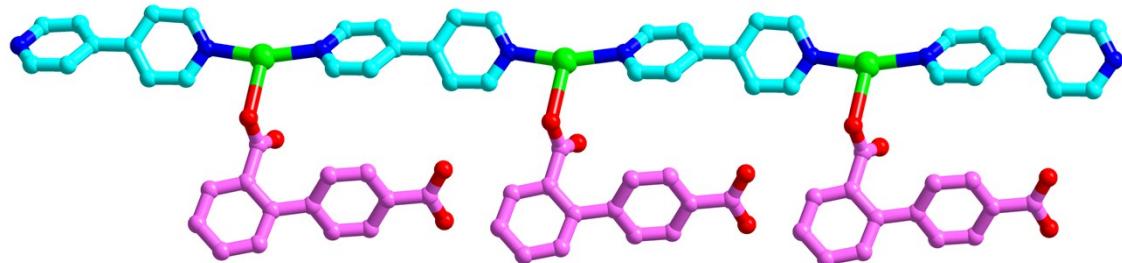
Symmetry transformations used to generate equivalent atoms: #1:  $-x, -y+2, -z$ ; #2:  $-x, -y+1, -z+1$ .



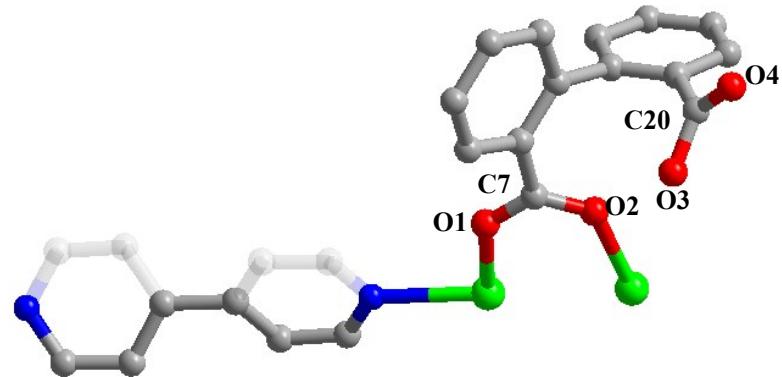
**Scheme S1** Coordination modes and configurations of  $\text{H}_2\text{pdc}$  ligands in compounds **1**, **2** and **3**.



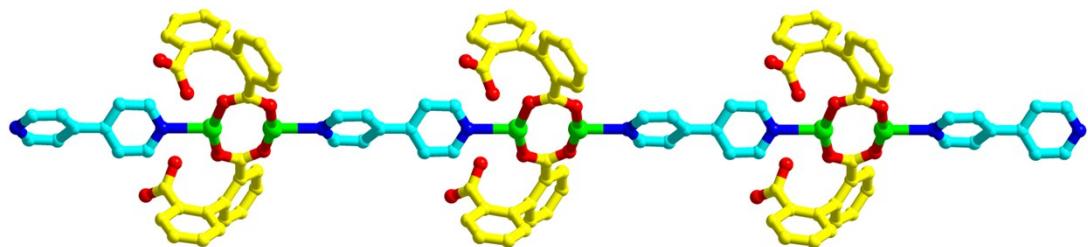
**Fig. S1** The IR spectra of **1**, **2** and **3**.



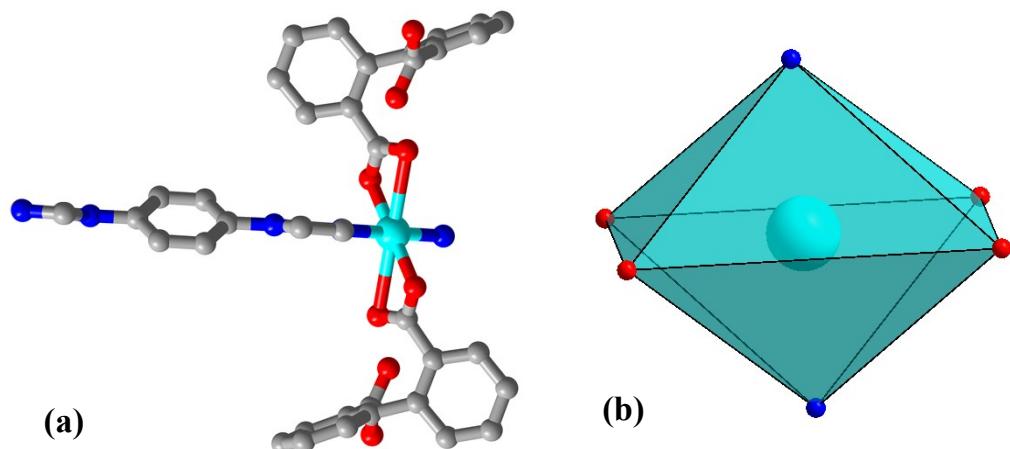
**Fig. S2** The 1D chain of **1**.



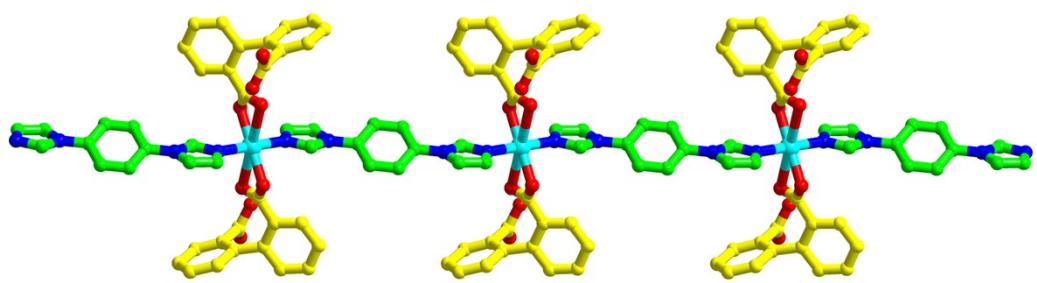
**Fig.S3** The structural unit of **2** with labeling scheme and 50% thermal ellipsoids (hydrogen atoms are omitted for clarity) symmetric components in 60% transparency.



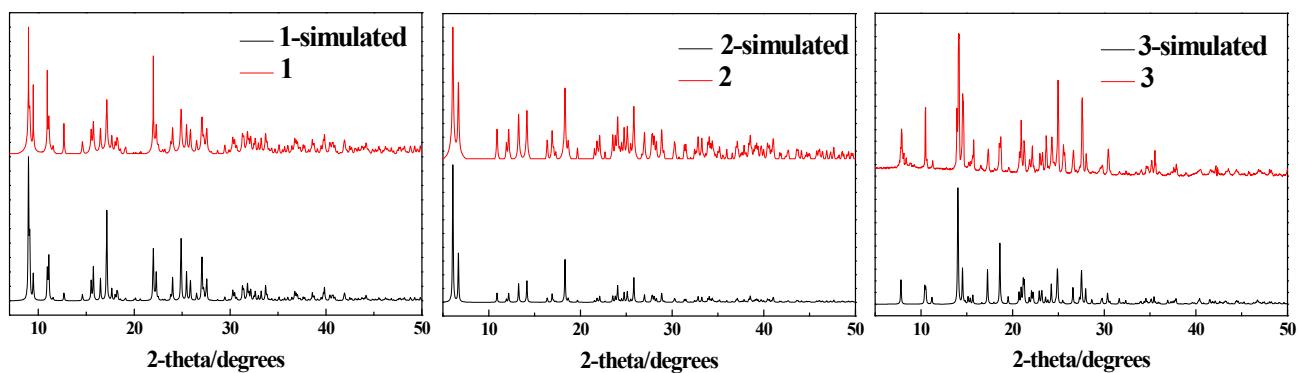
**Fig. S4** The 1D chain of **2**.



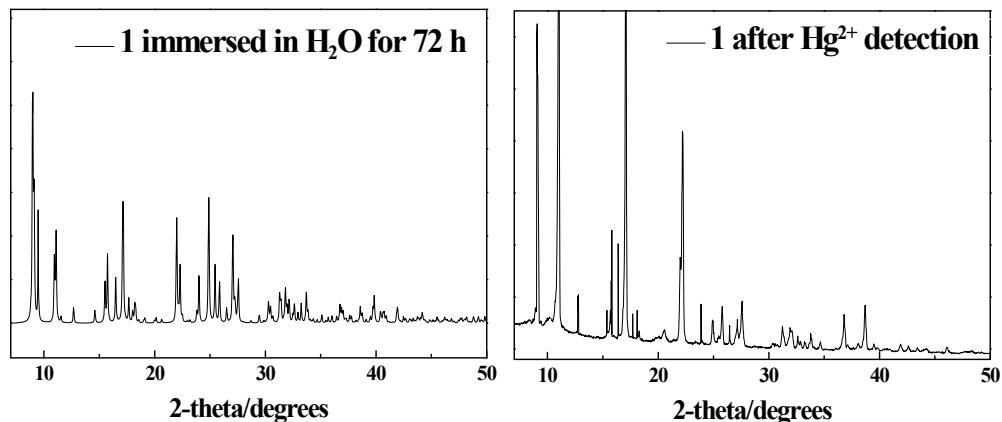
**Fig. S5** (a) The structural unit of **3** with labeling scheme and 50% thermal ellipsoids (hydrogen atoms are omitted for clarity). (b) Polyhedral representation of the coordination sphere of the  $\text{Cu}^{2+}$  centre.



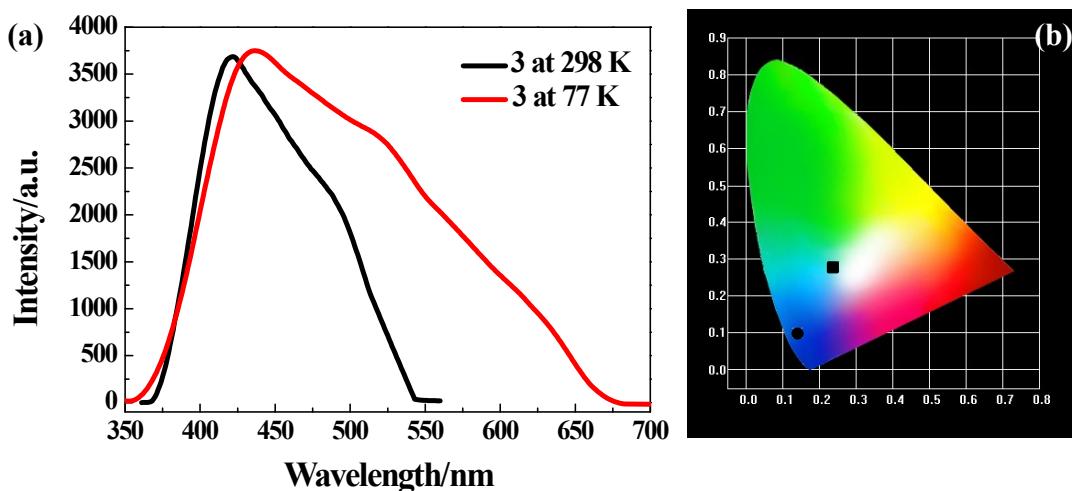
**Fig. S6** The 1D chain of **3**.



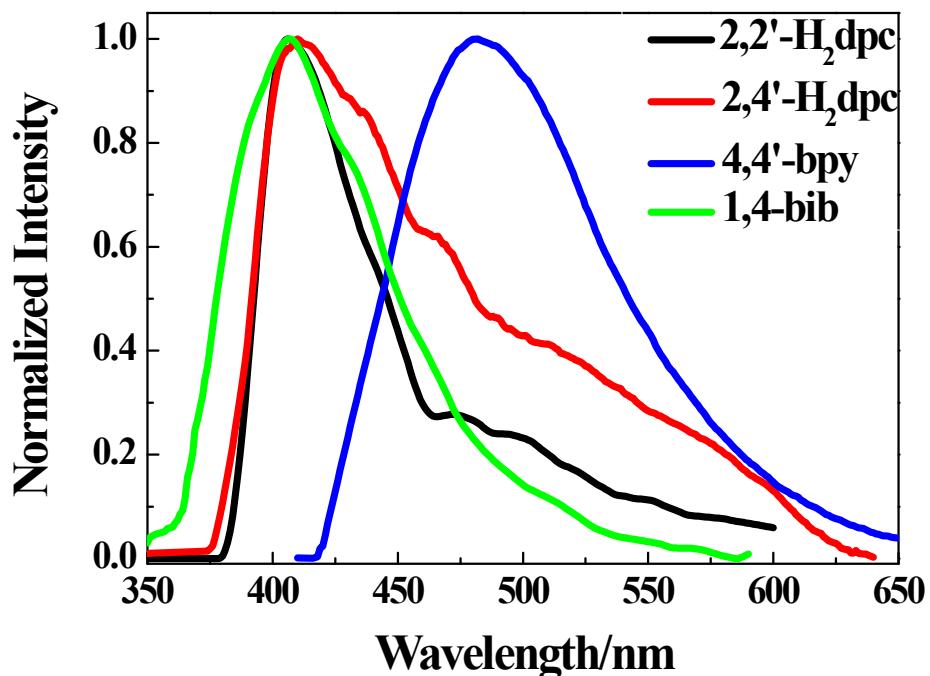
**Fig. S7** The PXRD patterns of **1**, **2** and **3** with the relevant simulated patterns.



**Fig. S8** The PXRD patterns of **1** immersed in H<sub>2</sub>O for 72 h and the sample **1** detected Hg<sup>2+</sup> ion.



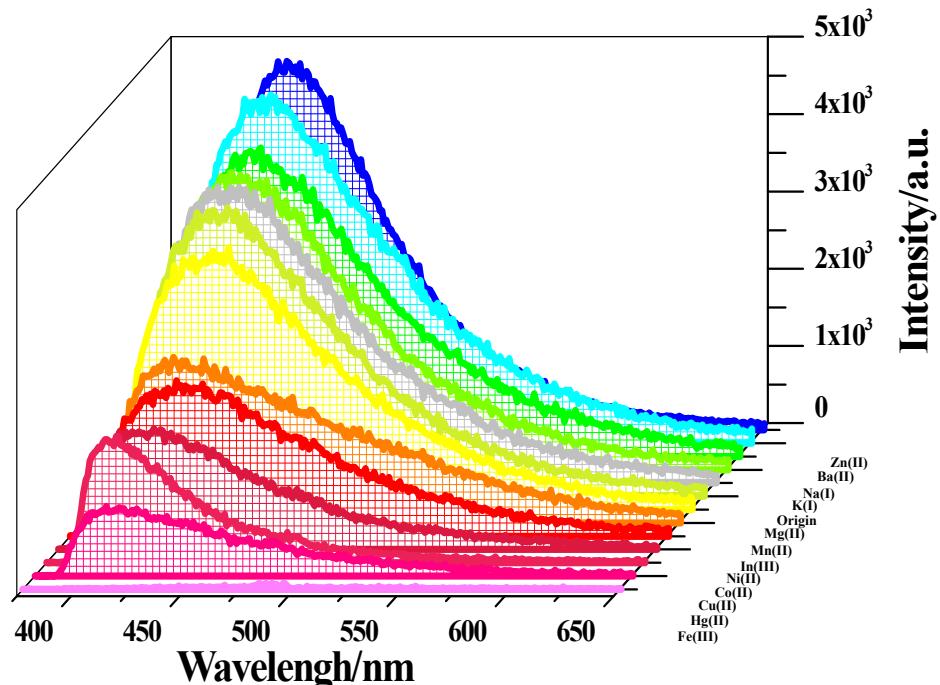
**Fig. S9** (a) Emission spectra of compound **3** in the solid state at 298 K and 77 K; (b) The corresponding color coordinate diagram of emission for **3** (circle symbol for 298 K and square symbol for 77 K).



**Fig. S10** Emission spectra of ligands in the solid state at 298 K.

**Table S3** Luminescent data for **1**, **2**, **3** in the solid state at 298 K and 77K, and corresponding ligands at 298 K.

Compound	Temperature (K)	Excitation ( $\lambda_{\text{ex}}$ , nm)	Emission ( $\lambda_{\text{em}}$ , nm)	CIE (x, y)
<b>1</b>	298	350	414	(0.14, 0.03)
	77	350	516	(0.29,0.34)
<b>2</b>	298	350	519	(0.29, 0.38)
	77	350	537	(0.30,0.43)
<b>3</b>	298	350	420	(0.14, 0.10)
	77	350	436	(0.24, 0.28)
2,2'-H <sub>2</sub> dpc	298	300	410	(0.18,0.15)
2,4'-H <sub>2</sub> dpc	298	300	412	(0.20,0.21)
4,4'-bpy	298	300	481	(0.21,0.32)
1,4-bib	298	300	403	(0.16, 0.08)



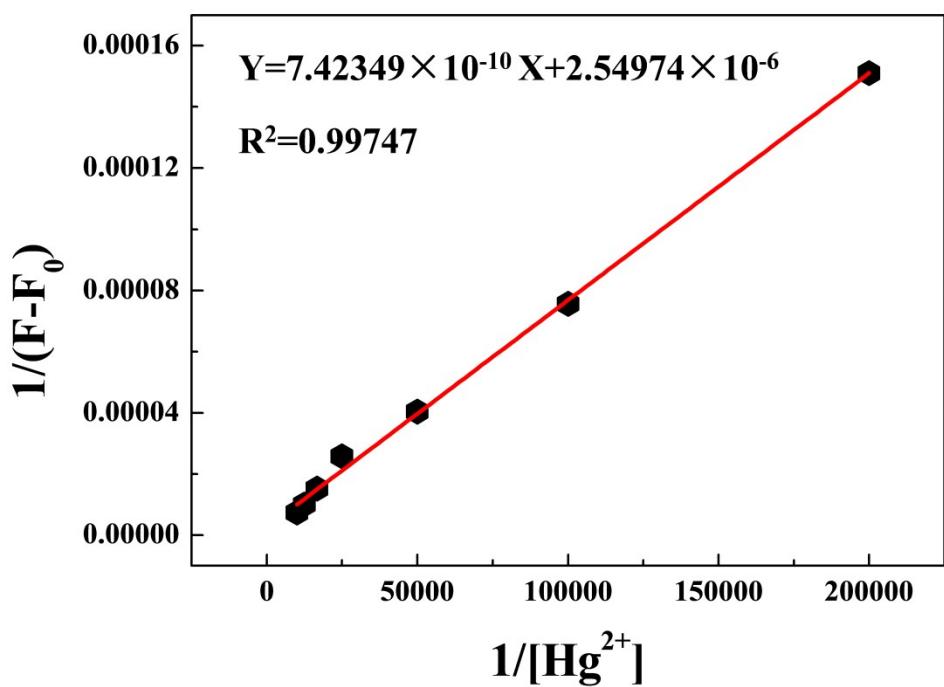
**Fig. S11** Suspension-state PL spectra of **2** at 426 nm dispersed in aqueous solutions containing different metal ions (50 $\mu$ M) when excited at 320 nm.

### **Standard deviation and detection limit calculation**

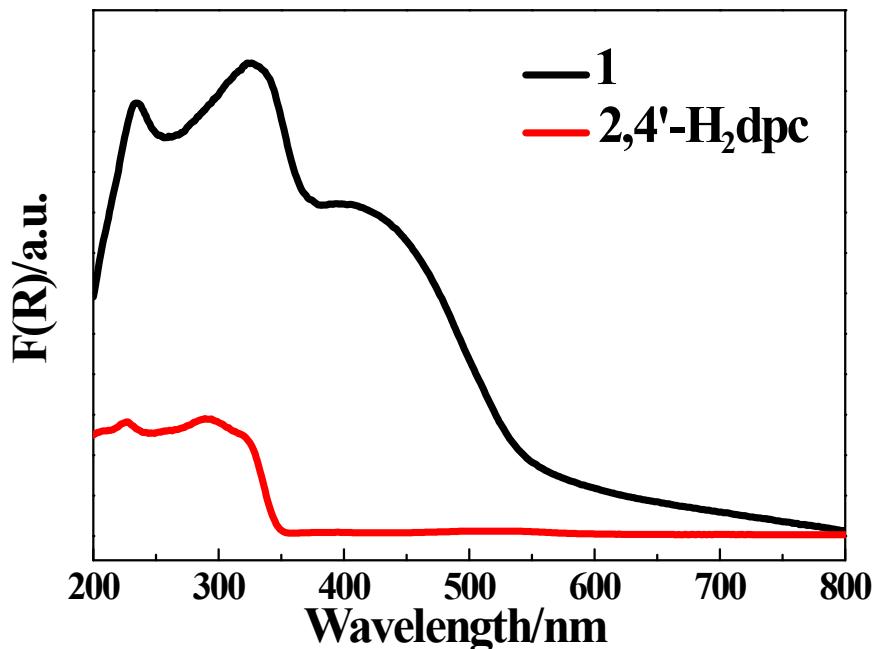
To calculate the standard deviation and detection limit of this detection method, compound **1** with fine particles was made into a suspension. Then, Hg<sup>2+</sup> solution (5-100 μL, 100 μM) was added into the suspension and the fluorescent intensities were recorded. Standard deviation ( $\sigma$ ) was calculated from five blank tests of **1** suspension and the detection limit was calculated via the formula:  $3\sigma/m$  (m: the slope of the linear region).

**Table S4** Standard deviation calculation and detection limit calculation.

Text 1 Fluorescence intensity	22206.56 a.u.
Text 2 Fluorescence intensity	22214.65 a.u.
Text 3 Fluorescence intensity	22208.56 a.u.
Text 4 Fluorescence intensity	22216.65 a.u.
Text 5 Fluorescence intensity	22212.35 a.u.
Standard Deviation ( $\sigma$ )	4.18
Slope (m)	1302.7022 $\mu\text{M}^{-1}$
Detection limit ( $3\sigma/m$ )	0.00963 $\mu\text{M}$



**Fig. S12** Benesi–Hildebrand plot of **1** assuming 1:1 stoichiometry between **1** and  $Hg^{2+}$  in water. The association constant of **1**- $Hg^{2+}$  is  $3.44 \times 10^3 \text{ M}^{-1}$ .



**Fig. S13** Plots of UV-vis absorption spectra of **1** and 2,4'-H<sub>2</sub>pdc ligand.