

### Supporting Information

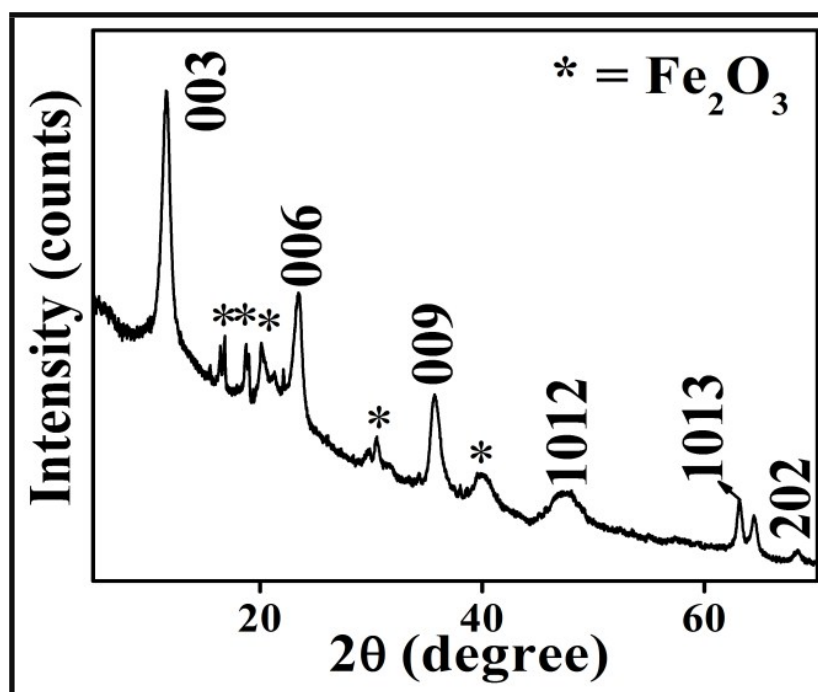
#### Ethylene glycol intercalated monometallic layered double hydroxide based on iron as an efficient bifunctional catalyst

Rajamani Nagarajan\*, Pankaj Gupta, Poonam Singh and Pinki Chakraborty

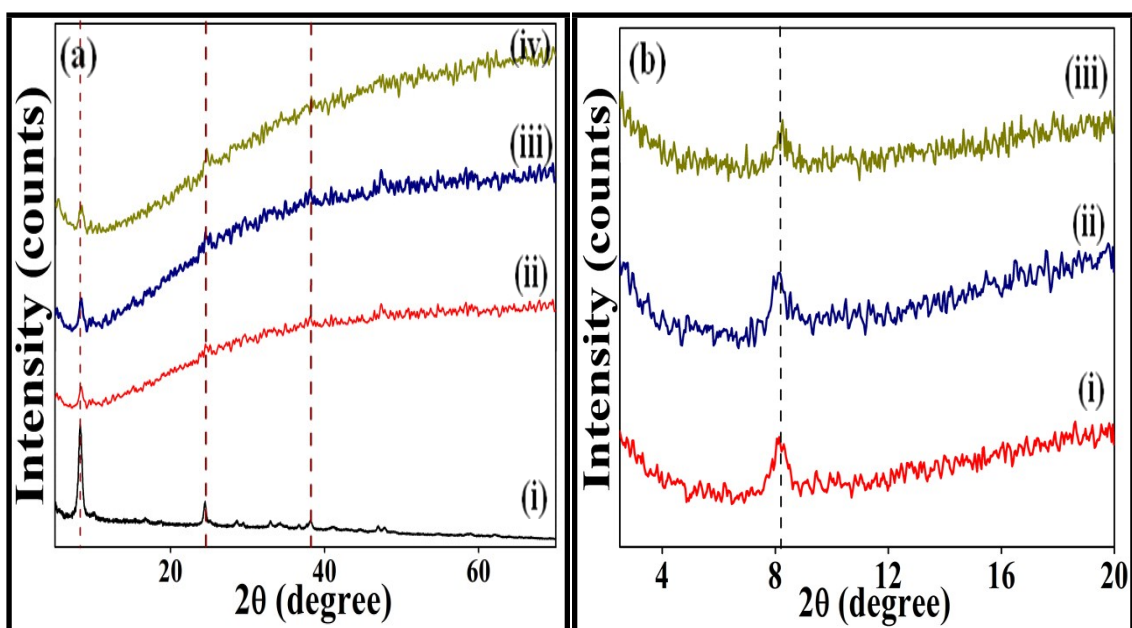
Materials Chemistry Group

Department of Chemistry, University of Delhi

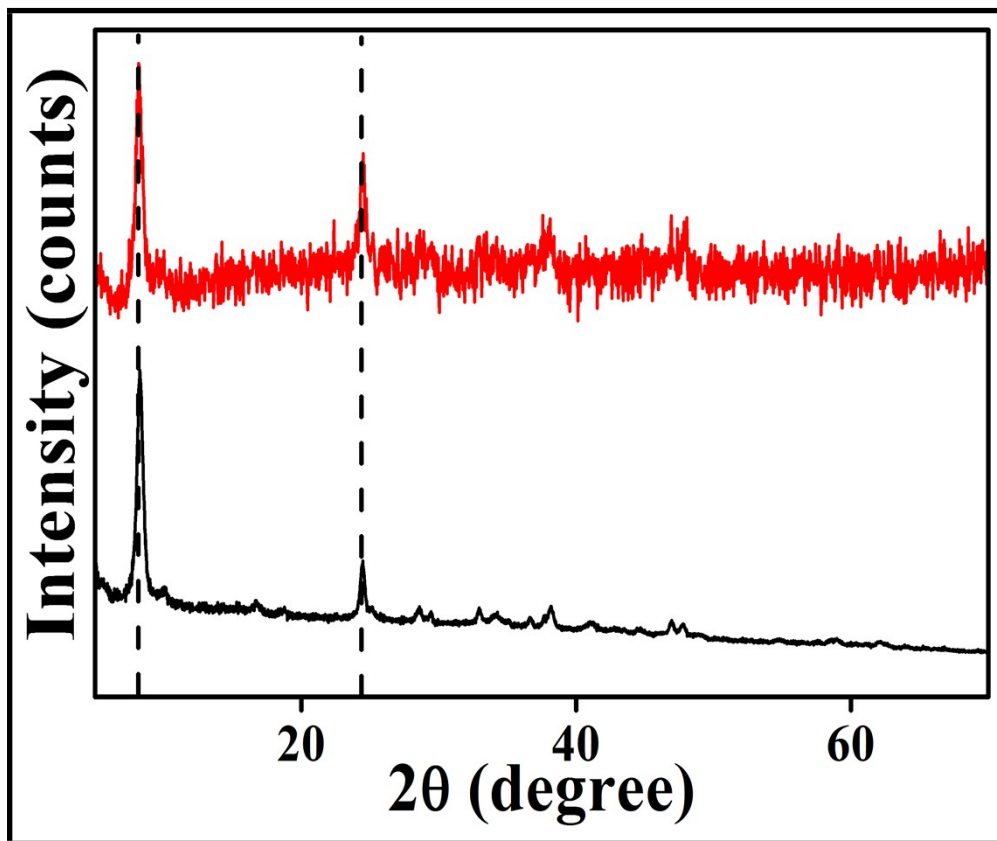
Delhi 110007 INDIA



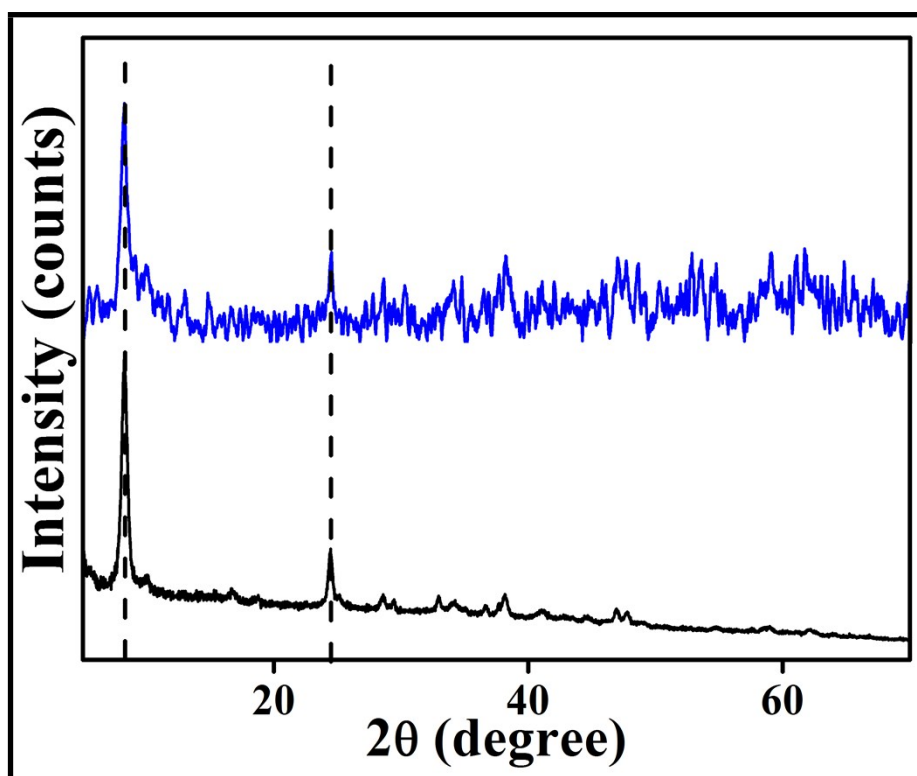
**Fig. S1** PXRD pattern of the product from the reaction of FeCl<sub>3</sub> and urea employing water as the solvent under hydrothermal conditions.



**Fig. S2** (A) PXRD pattern of EG intercalated Fe<sup>II</sup>- Fe<sup>III</sup>- LDH (i), after soaking 50 mg of the catalyst in XO (ii), MO (iii) and MB (iv) dye solutions ( $1 \times 10^{-4}$  M and for 24 h) at room temperature. (B) PXRD patterns of these samples in low angle region (2-20°) of the product after soaking the soaking experiments in XO (i), (ii) in MO (ii) and MB (iii).



**Fig. S3** (a) PXRD pattern of the EG intercalated Fe<sup>II</sup> - Fe<sup>III</sup> - LDH synthesized and (b) after its use as a catalyst (first cycle) for the oxidative degradation of XO.



**Fig. S4** (a) PXRD pattern of the EG intercalated  $\text{Fe}^{\text{II}}$  -  $\text{Fe}^{\text{III}}$  - LDH synthesized and (b) after its use as a catalyst (first cycle) for the reduction of p-nitrophenol.

**Table S1** Summary of the rate constants ( $k$ ) obtained using our sample (EG intercalated Fe (II) - Fe (III) - LDH)

Substrate	1 <sup>st</sup> cycle	2 <sup>nd</sup> cycle	3 <sup>rd</sup> cycle	4 <sup>th</sup> cycle
Xylenol orange	$348 \times 10^{-3} \text{ min}^{-1}$ ( $5.8 \times 10^{-3} \text{ s}^{-1}$ )	$330 \times 10^{-3} \text{ min}^{-1}$ ( $5.5 \times 10^{-3} \text{ s}^{-1}$ )	$204 \times 10^{-3} \text{ min}^{-1}$ ( $3.4 \times 10^{-3} \text{ s}^{-1}$ )	$108 \times 10^{-3} \text{ min}^{-1}$ ( $1.8 \times 10^{-3} \text{ s}^{-1}$ )
<i>p</i> -nitrophenol	$258.6 \times 10^{-3} \text{ min}^{-1}$ <sup>1</sup> ( $4.31 \times 10^{-3} \text{ s}^{-1}$ )	$251.4 \times 10^{-3} \text{ min}^{-1}$ ( $4.19 \times 10^{-3} \text{ s}^{-1}$ )	$198 \times 10^{-3} \text{ min}^{-1}$ ( $3.3 \times 10^{-3} \text{ s}^{-1}$ )	$114 \times 10^{-3} \text{ min}^{-1}$ ( $1.9 \times 10^{-3} \text{ s}^{-1}$ )