

Supporting Information

for

Transformation Kinetics of Silver Nanoparticles and Silver Ions in Aquatic Environments Revealed by Double Stable Isotope Labeling

Sujuan Yu, Yongguang Yin, Xiaoxia Zhou, Lijie Dong, Jingfu Liu*

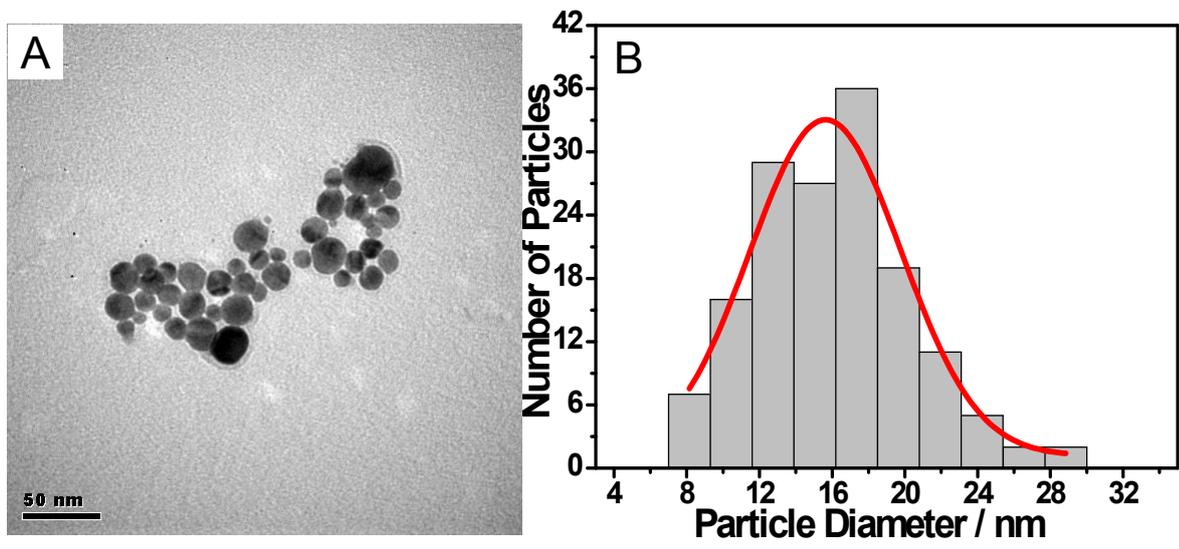
State Key Laboratory of Environmental Chemistry and Ecotoxicology, Research Center for Eco-
Environmental Sciences, Chinese Academy of Sciences, P.O. Box 2871, Beijing 100085, China

Total pages: 6

Total figures: 4

Total tables: 2

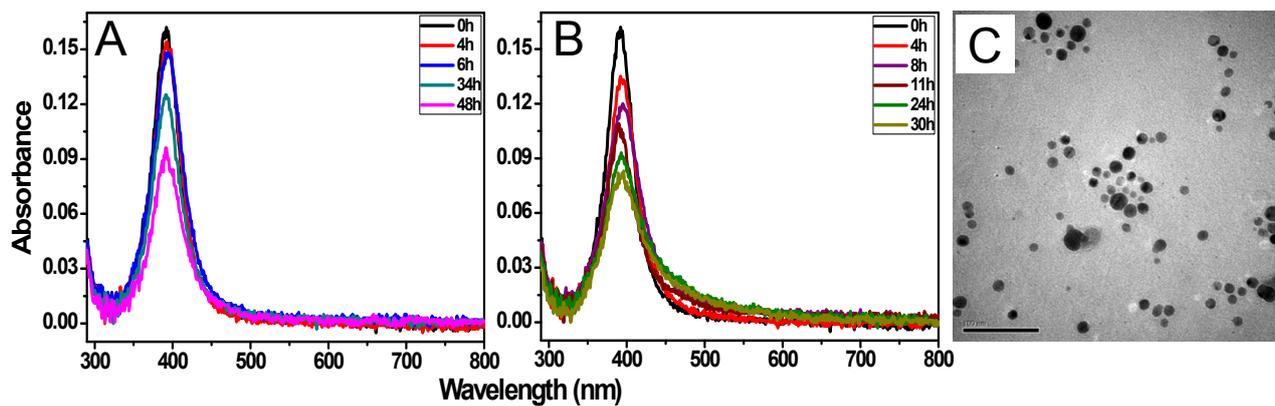
* Corresponding author: E-mail: jfliu@rcees.ac.cn



19

20 Figure S1. TEM image (A) and the size distribution (B) of the synthesized $^{107}\text{AgNPs}$.

21

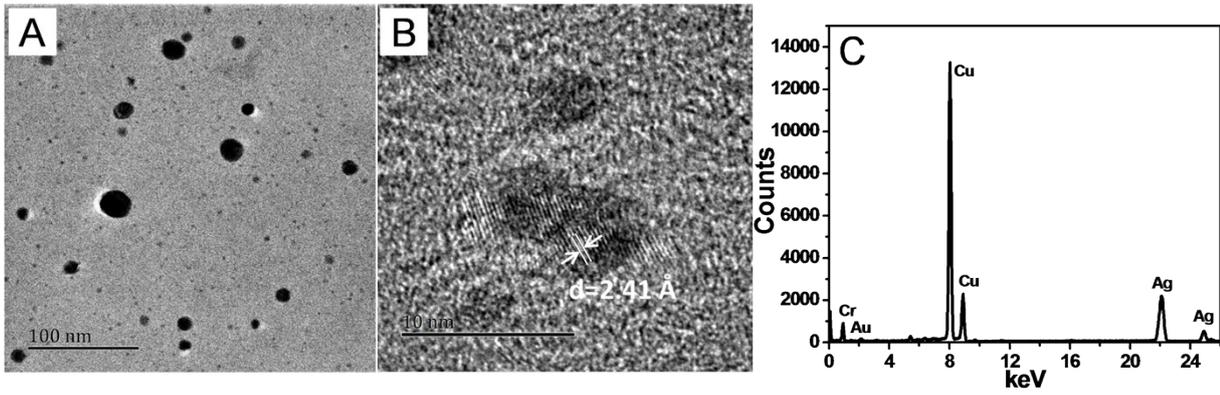


22

23 Figure S2. UV-vis absorption spectra of AgNPs in the dark (A) and light (B), and TEM image of

24 AgNPs (C) after solar irradiation for 4 h. A mixture of 1 mg/L $^{107}\text{AgNPs}$ and 1 mg/L $^{109}\text{Ag}^+$ in pure

25 water was treated at pH 7.4 and 30 °C.



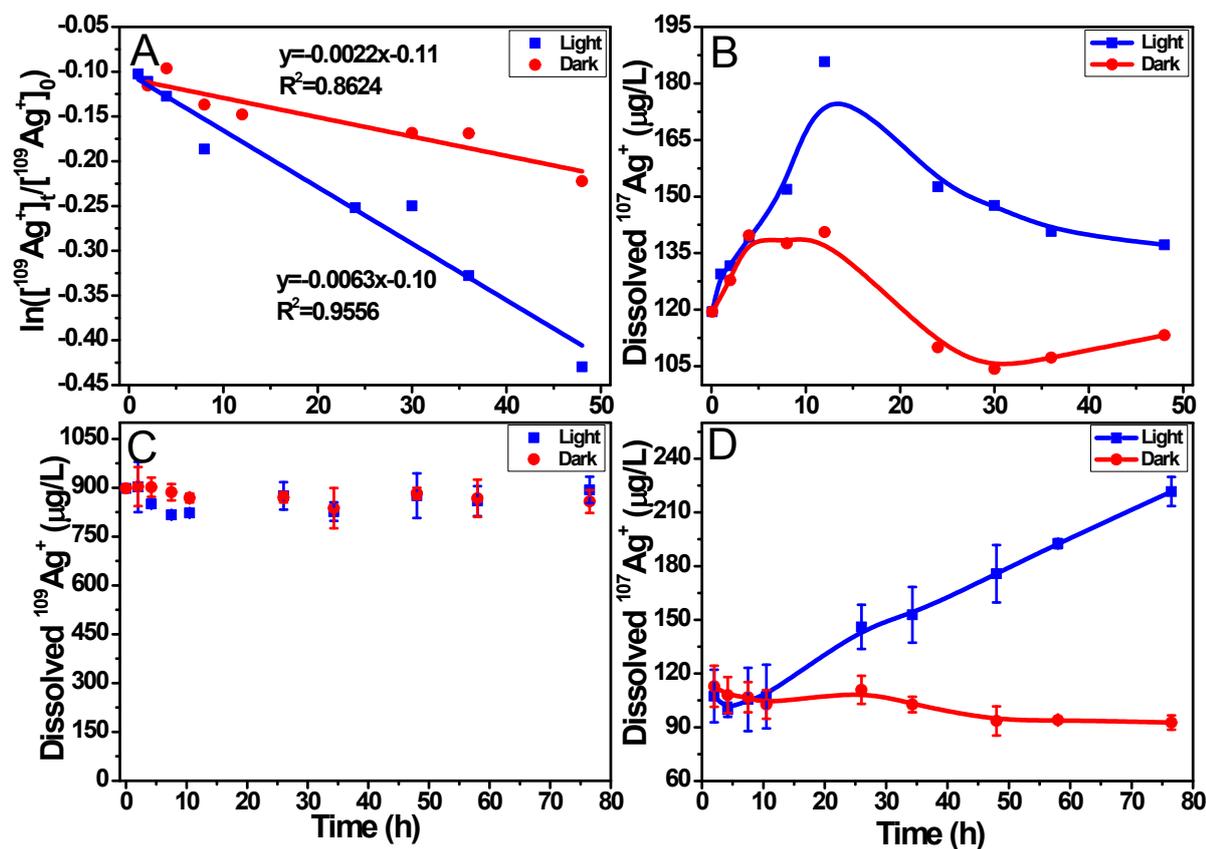
26

27 Figure S3. TEM image of AgNPs in the dark after 24 h (A), and HRTEM image (B) and EDS

28 analysis (C) of the newly formed AgNPs in the presence of SRHA. A mixture of 1 mg/L $^{107}\text{AgNPs}$,

29 1 mg/L $^{109}\text{Ag}^+$ and 5 mg C/L SRHA was treated at pH 7.4 and 30 °C.

30



31

32 Figure S4. Fraction of dissolved $^{109}\text{Ag}^+$ and ion release kinetics of $^{107}\text{AgNPs}$ over time at different
 33 temperatures. Fraction of dissolved $^{109}\text{Ag}^+$ *verse* time at 50 °C (A), ion release kinetics of $^{107}\text{AgNPs}$
 34 at 50 °C (B), concentration change of dissolved $^{109}\text{Ag}^+$ at 6 °C (C), ion release kinetics of $^{107}\text{AgNPs}$
 35 at 6 °C (D). A mixture of 1 mg/L $^{107}\text{AgNPs}$, 1 mg/L $^{109}\text{Ag}^+$ and 5 mg C/L SRHA was treated at pH
 36 7.4. Error bars in Figure C and D represent the standard deviation for three independent
 37 measurements.

38 Table S1 Comparison of the reduction rate constant k of $^{109}\text{Ag}^+$ under different conditions

Reaction condition	k/h^{-1} (Light)	k/h^{-1} (Dark)
without SRHA, pH 7.4, 30 °C	-	-
5 mg C/L SRHA, pH 5.6, 30 °C	0.0025	0.0015
5 mg C/L SRHA, pH 7.4, 30 °C	0.0051	0.0017
5 mg C/L SRHA, pH 8.5, 30 °C	0.0098	0.0051
5 mg C/L SRHA, pH 7.4, 50 °C	0.0063	0.0022
5 mg C/L SRHA, pH 7.4, 6 °C	-	-
5 mg C/L SRHA, pH 7.4, 30 °C, with Ca^{2+} and Mg^{2+}	0.0048	0.0017

39 -: The reduction of $^{109}\text{Ag}^+$ was not obvious.

40

41 Table S2 Comparison of the oxidation rate constant k of $^{107}\text{AgNPs}$

Reaction condition	k/h^{-1} (Light)	k/h^{-1} (Dark)
without SRHA, pH 7.4, 30 °C	0.0366 (the first 12 h) 0.0081(after 12 h)	0.0078

42