Electronic Supporting Information

Boron doped reduced graphene oxide based bimetallic Ni/Fe nanohybrids for rapid dechlorination of trichloroethylene

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Materials	Initial weight (%)	Residual weight (%)	Residue	Weight ratio
B-rGO	100	38.3	Boron and carbon	100%
B-rGO/Fe/Ni	100	76.9	Boron and carbon Fe/Ni bimetal	37.4% 62.6%

Table S1 Determination of composition of B-rGO/Fe/Ni from thermogravimetry.

Table S2. The comparison of pseudo-first order rate constant (k_{obs}) for TCE dechlorination by bimetallic Fe/Ni nanoparticles immobilized onto different supports under anoxic conditions.

Bimetallic nanomaterials	pН	$k_{obs} (h^{-1})$	References
Free Fe/Ni		8.2	
B-rGO/Fe/Ni (5%)	7.0	10.9	This study
B-rGO/Fe/Ni (5%) (mass normalized)		18.1	
Zeolite/Fe/Ni	N.A	2.2	38
Fe/Ni (10%)	6.5	1.2	
Fe/Ni (22%)	7.0	3.0	39
Fe/Ni (50%)	7.5	0.36	
Polystyrene resin/Fe/Ni (10%)	7.0	5.84	25
Polystyrene resin/Fe/Ni (20%)		6.95	
Polystyrene resin/Fe/Ni (25%)		6.40	
Fe/Ni/nylon-66		6.47	
Fe/Ni/PVDF	7.0	1.77	8
Fe/Ni/MCEM		1.04	
Fe/Ni/Millex GS membranes		0.56	
Free Fe/Ni (5%)	9.2 ± 0.1	0.0984	40
Bentonite/Fe/Ni (5%)	8.2 ± 0.1	0.181	

N.A.: not available.



Figure S1 Adsorption isotherm of ferric ions by B-rGO under anoxic conditions .



Figure S2. The calibration curve of TCE prepared by using the external standard method with the known concentrations of TCE in aqueous solutions.



Figure S3. The energy-dispersive X-ray spectra (EDS) of B-rGO/Fe/Ni nanocomposites. Inset is the analytical area of SEM image for EDS.



Figure S4. The Nyquist plot of B-rGO, Fe/Ni and B-rGO/Fe/Ni nanocomposites by using 1.0 M Na₂SO₄ as the electrolyte.



Figure S5. FTIR spectra of GO and B-rGO/Fe/Ni nanocomposites.



Figure S6. Dechlorination of 20 μ M TCE by 1.0 g/L graphene oxide immobilized Fe/Ni nanocomposite at pH 7 under anoxic conditions.



Figure S7. The (a) XRD patterns and (b) SEM images of B-rGO/Fe/Ni nanocomposite before and after the dechlorination of TCE at pH 7 under anoxic conditions.