

Supporting Information

A novel histidine functionalized 1,8-naphthalimide-based fluorescent chemosensor for selective and sensitive detection of Hg²⁺ in water

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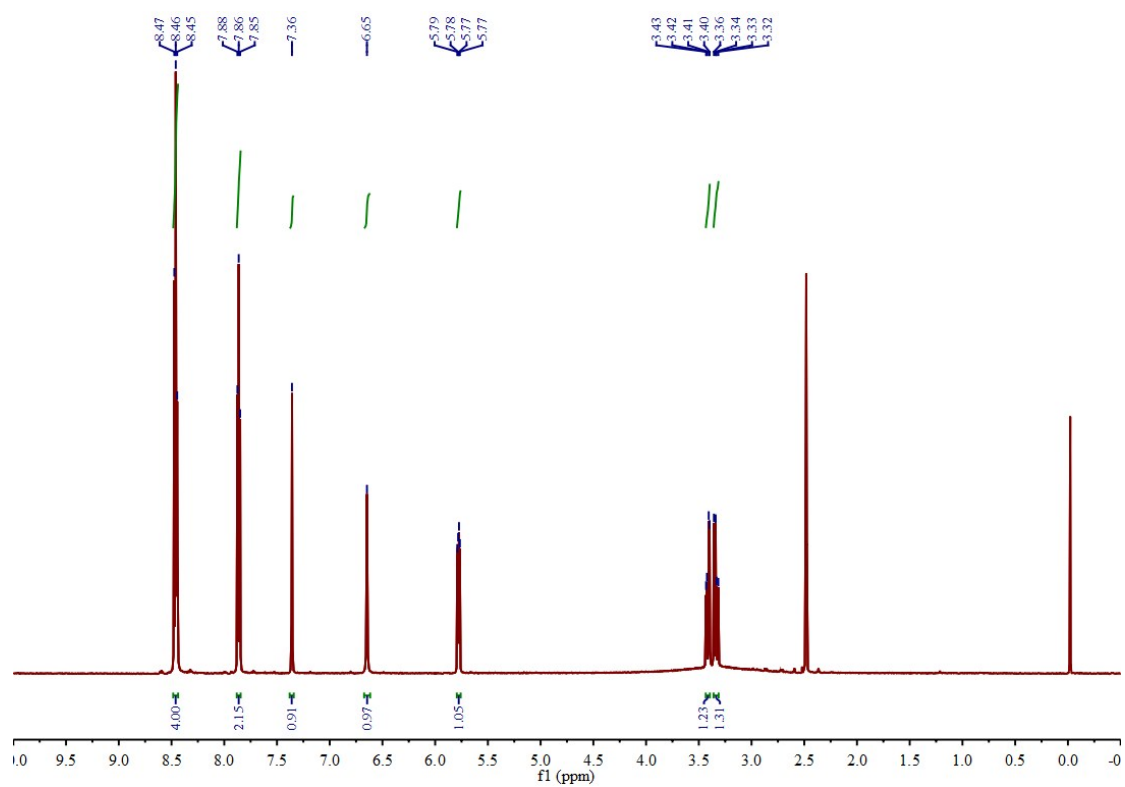


Fig. S1 ^1H NMR spectrum of Z_1 in DMSO-d_6 .

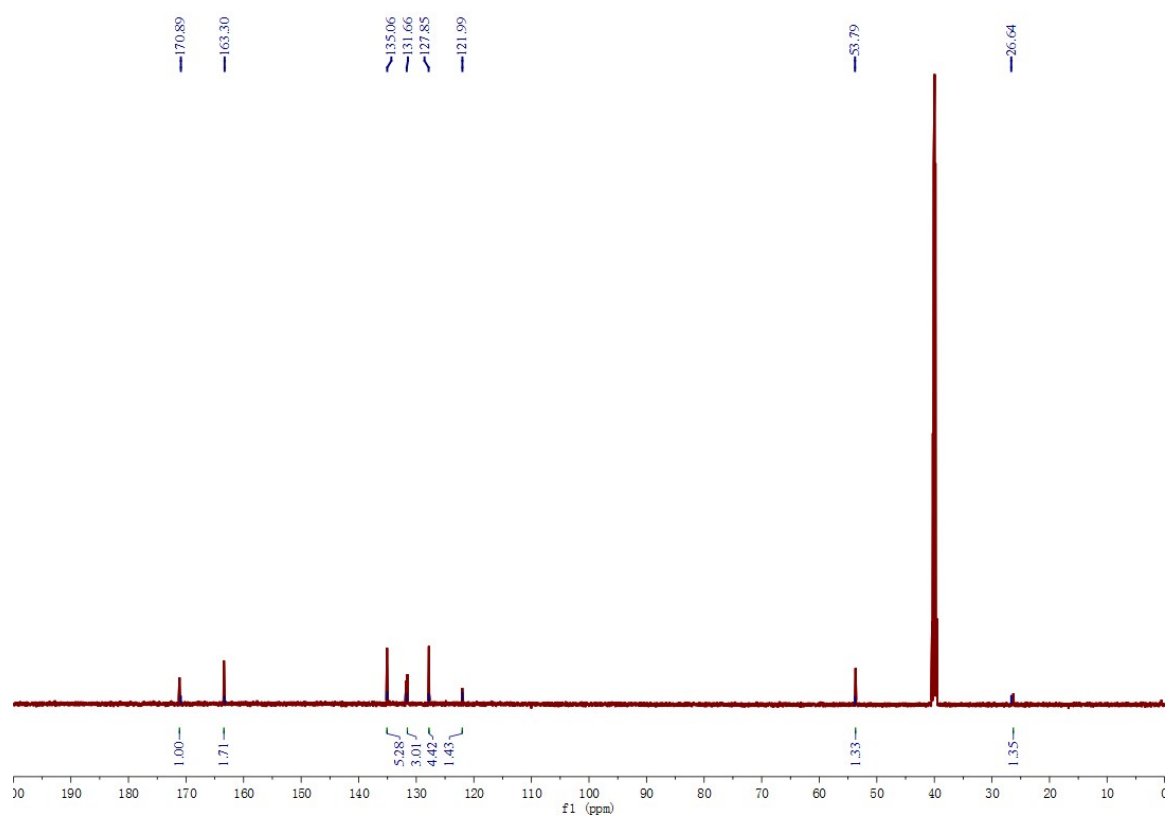


Fig. S2 ¹³C NMR spectrum of **Z**₁ in DMSO-d₆.

Generic Display Report

Analysis Info

Analysis Name D:\Data\YANGY\New\CHENSHENG PENG160331_4.d
Method POS_100-1200_For LC.m
Sample Name CHENSHENG PENG160331_F1
Comment

Acquisition Date 3/31/2016 5:12:22 PM

Operator LZU
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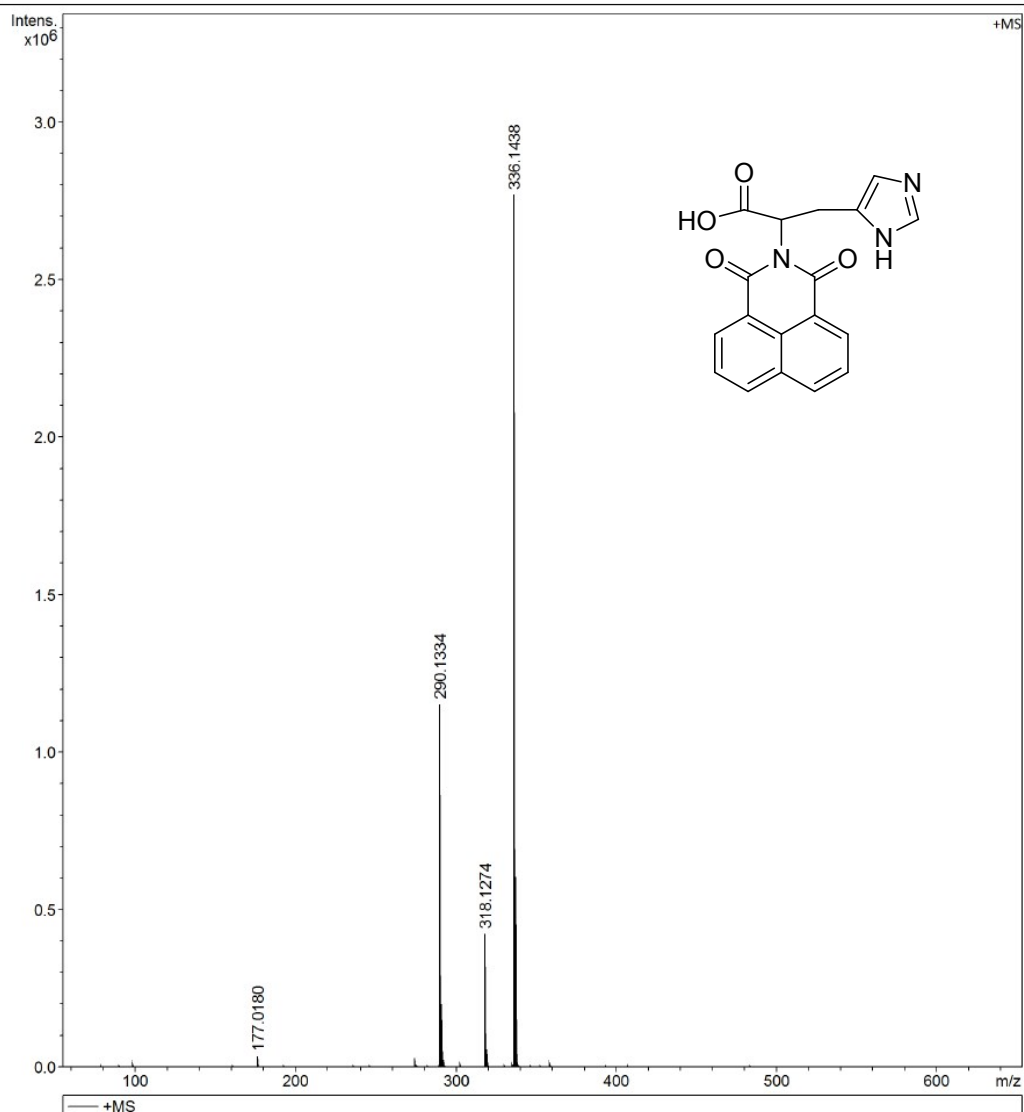


Figure S3 ESI-MS spectrum of **Z₁**.

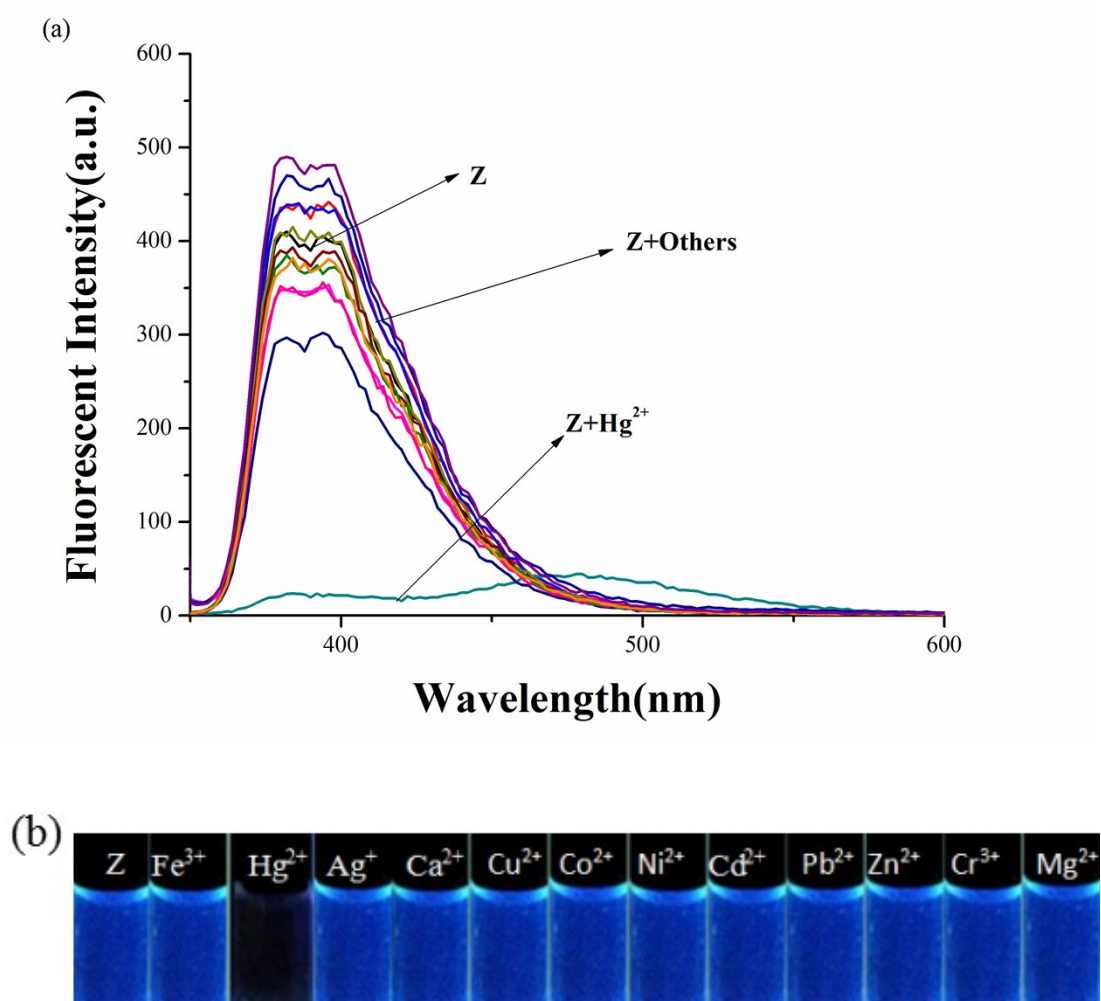


Figure S4 (a) Fluorescence emission data for a 1:10 mixture of **Z** (2.0×10^{-5} M) and different metal ions, as their perchlorate salts, in water. ($\lambda_{\text{ex}}=343$ nm). (b) Visual fluorescence emissions of sensor **Z** after the addition of various metal ions (10 equiv.) in water on excitation at 384 nm using UV lamp.

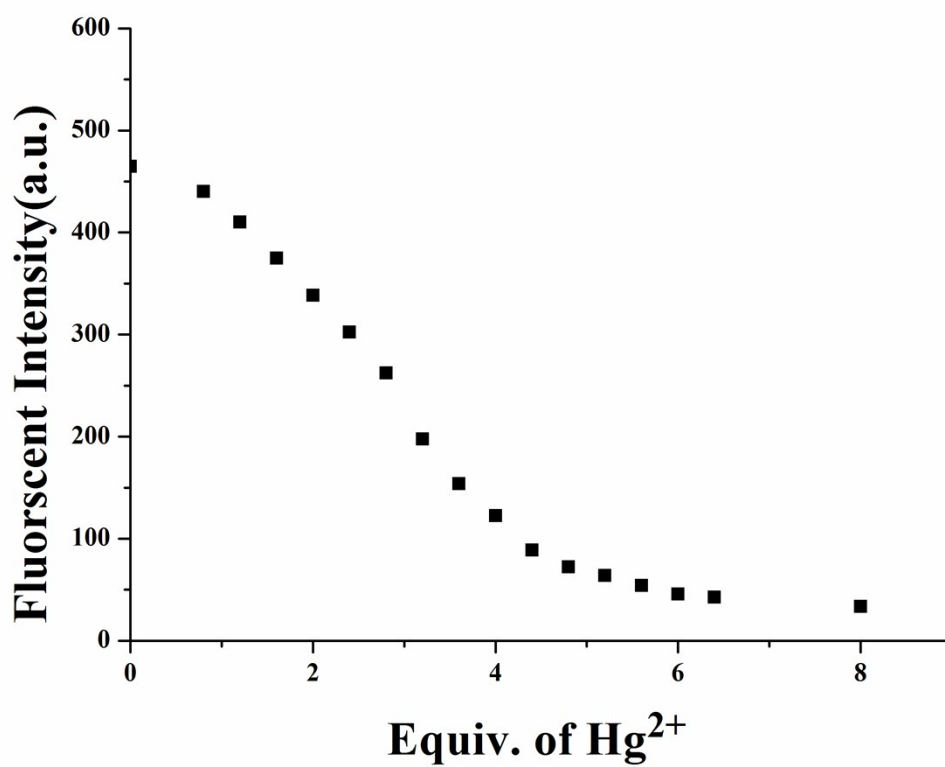


Fig. S5 A plot of fluorescence intensity of the sensor **Z** (2.0×10^{-5} M, water solution) depending on the concentration of Hg^{2+} in the range from 0 to 8 equivalents.

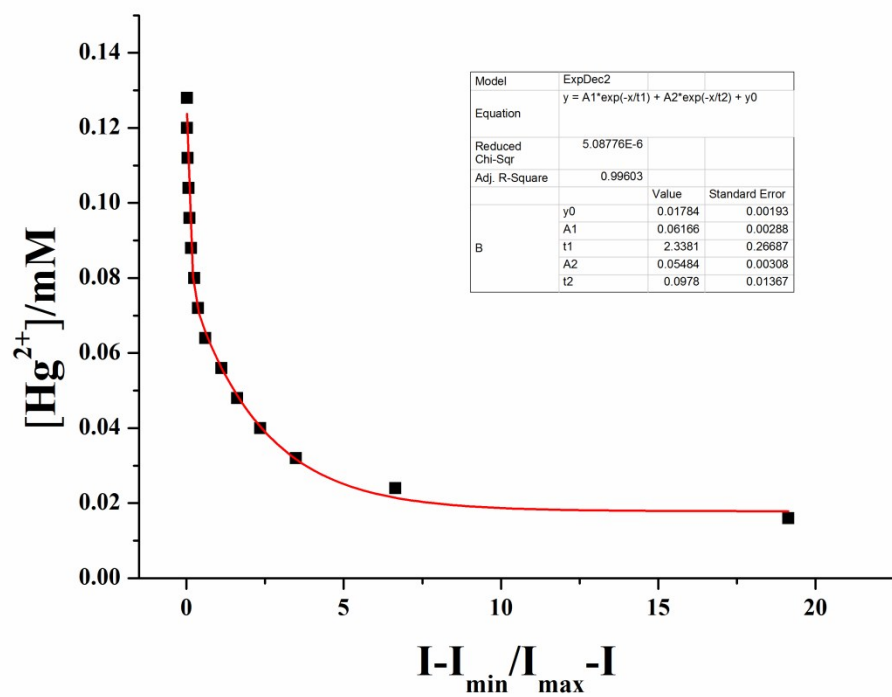


Figure S6 Non-linear least square fitting of intensity vs concentration of Hg^{2+} using 1:1 complex model.

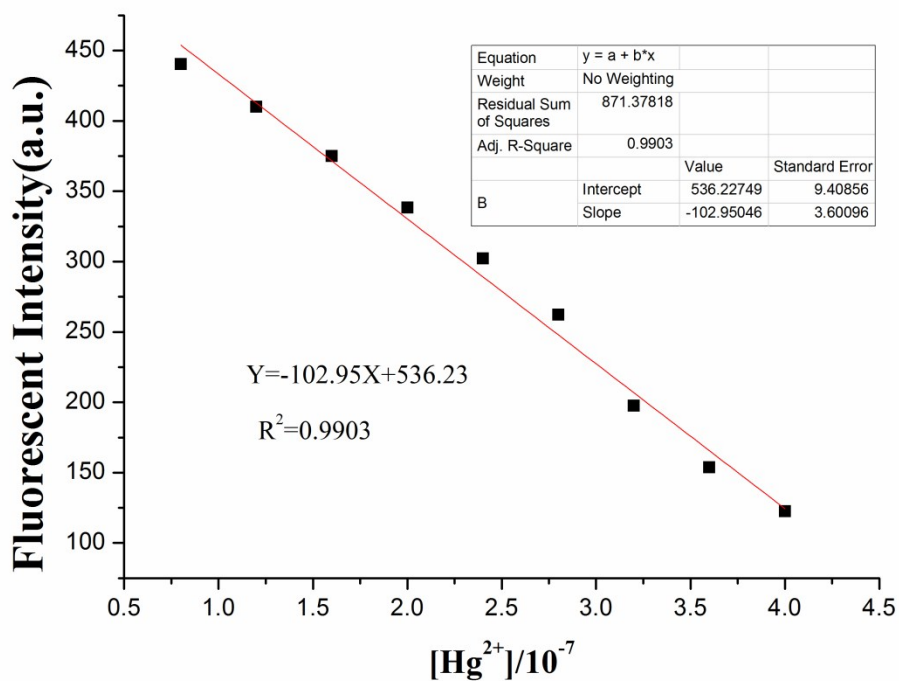


Figure S7 Plot of the intensity at 384 nm for a mixture of **Z** and Hg^{2+} in water in the range 1.0×10^{-7} – 2.0×10^{-6} M ($\lambda_{\text{ex}} = 343$ nm). Linear Equation: $Y = -102.950X + 536.227$, $R^2 = 0.9903$

$$\delta = \sqrt{\frac{\sum (F_0 - F_1)^2}{N - 1}} = 6.069, K=3$$

$$S = 1.03 \times 10^8$$

$$\text{LOD} = K \times \delta / S$$

$$\text{LOD} = 1.785 \times 10^{-7} \text{ M}$$

F_0 is the fluorescence intensity of **Z**; F_1 is the average of the F_0 .

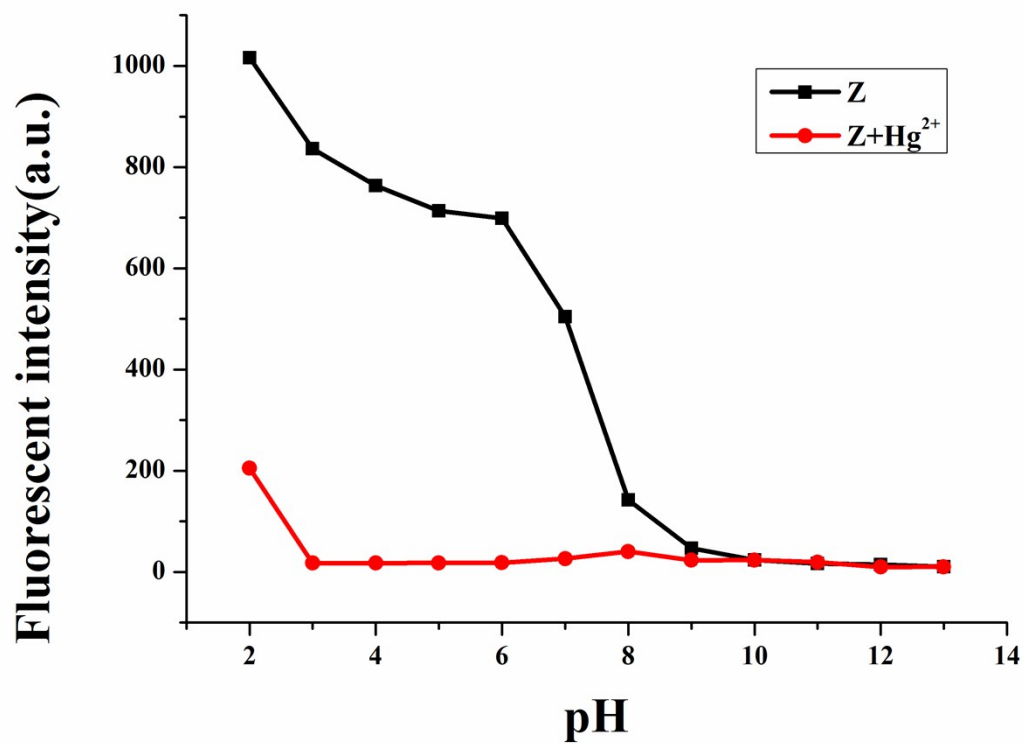


Figure S8 Influence of pH on the fluorescence of **Z** and **Z+Hg²⁺** in HEPES buffered solution in water.

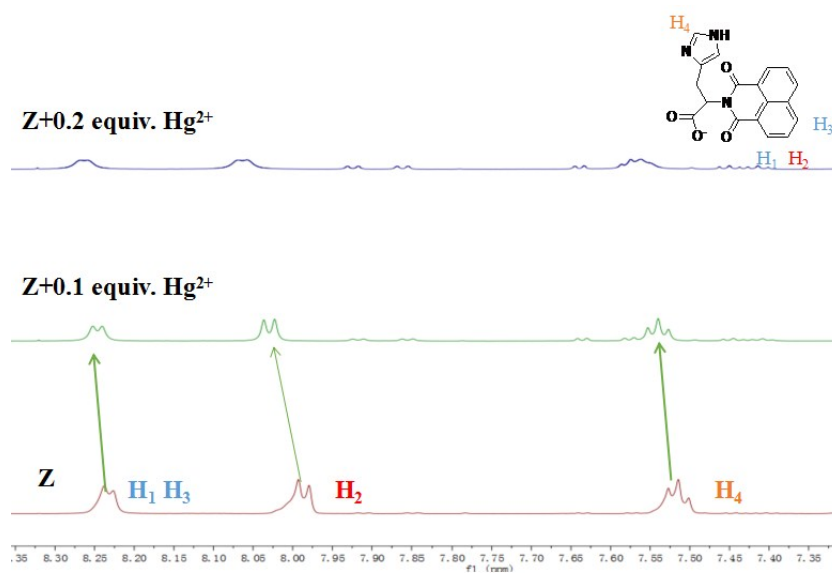


Fig. S9 Partial ¹H NMR spectra of **Z** (0.05 M, D₂O), and **Z** in the presence of varying amounts of Hg²⁺.

Measurement of fluorescence quantum yields

Fluorescence quantum yields were determined by the following equation.

$$\Phi = \Phi_R \times \frac{I A_R}{I_R A}$$

Where Φ is fluorescence quantum yield, I is the integrated fluorescence intensity and A is the optical density (absorption). The subscript R refers to the reference of Quinine hemesulfate salt.”

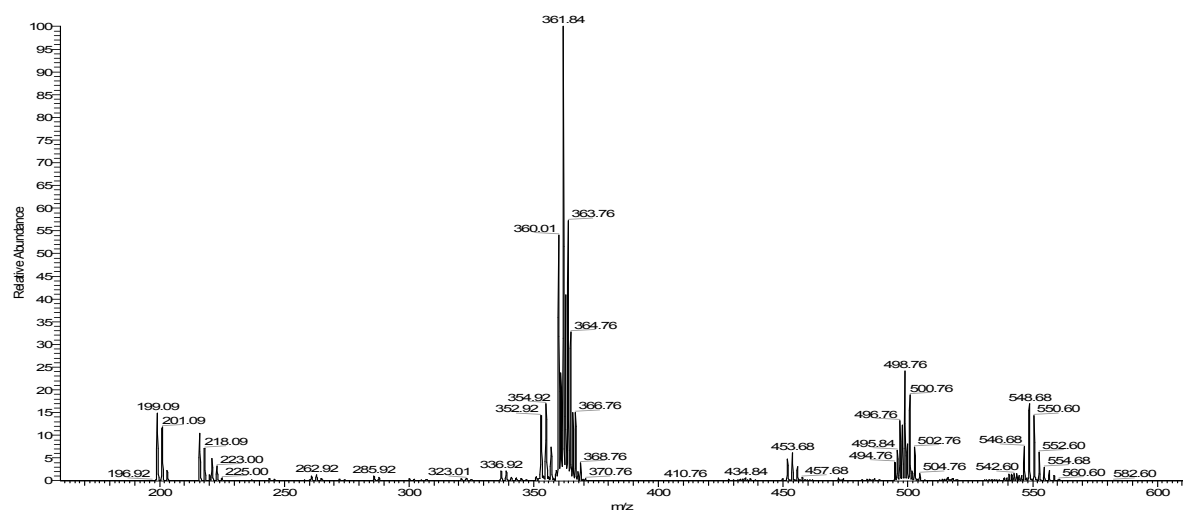


Figure S10 ESI-MS spectrum of $Z+Hg^{2+}$ complex.