Supporting Information

Mesoporous TiO₂@N-doped Carbon Composite Nanospheres Synthesized by Direct Carbonization of Surfactants after Sol-gel Process for Superior Lithium Storage

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Figure S1. SEM image of the amorphous TiO₂ nanospheres synthesized via double surfactant assisted assembly sol-gel process.



Figure S2. TEM image of the amorphous TiO₂ nanosphere synthesized via double surfactant assisted assembly sol-gel process.



Figure S3. FTIR spectrum of the amorphous TiO₂ nanospheres synthesized via double surfactant assisted assembly sol-gel process.



Figure S4.Monodisperse amorphous TiO₂ nanospheres with different diameter size synthesized via double surfactant assisted assembly sol-gel process(a,150 nm; b,250 nm).



Figure S5. The energy dispersed X-ray spectrum (EDS) of the TiO₂@N-doped carbon composite nanospheres synthesized via double surfactant assisted assembly sol-gel process followed by sequential carbonization of surfactants in N₂ atmosphere.



Figure S6. The high-resolution XPS spectra of N1s in $TiO_2@N$ -doped carbon composite nanospheres synthesized via double surfactant assisted assembly sol-gel process followed by sequential carbonization of surfactants in N₂ atmosphere.



Figure S7. Raman spectrum of the mesoporous TiO_2 nanospheres obtained through solvothermal treatment followed by calcination in air at 700 °C.



Figure S8. TG curve of the $TiO_2@N$ -doped carbon composite nanospheres synthesized via double surfactant assisted assembly sol-gel process followed by sequential carbonization of surfactants in N₂ atmosphere.



Figure S9. Nitrogen sorption isotherms of the mesoprous TiO_2 nanospheres obtained obtained through solvothermal treatment followed by calcination in air at 700 °C.



Figure S10. (a) Nitrogen adsorption/desorption isotherms and (b) pore size distribution curves of the mesoporous $TiO_2@N$ -doped carbon composite nanospheres after 200 cycles.