

Electronic Supplementary Information

Direct conversion of coordination compounds into Ni<sub>2</sub>P  
nanoparticles entrapped in 3D mesoporous graphene for  
efficient hydrogen evolution reaction

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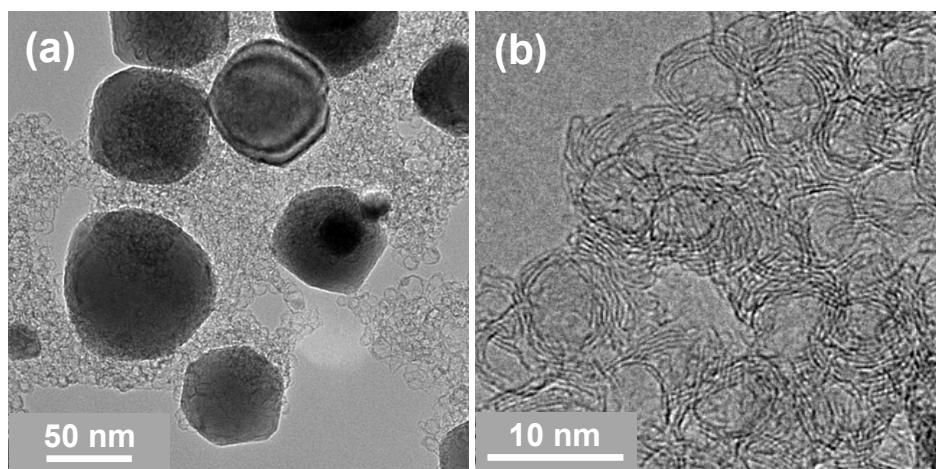
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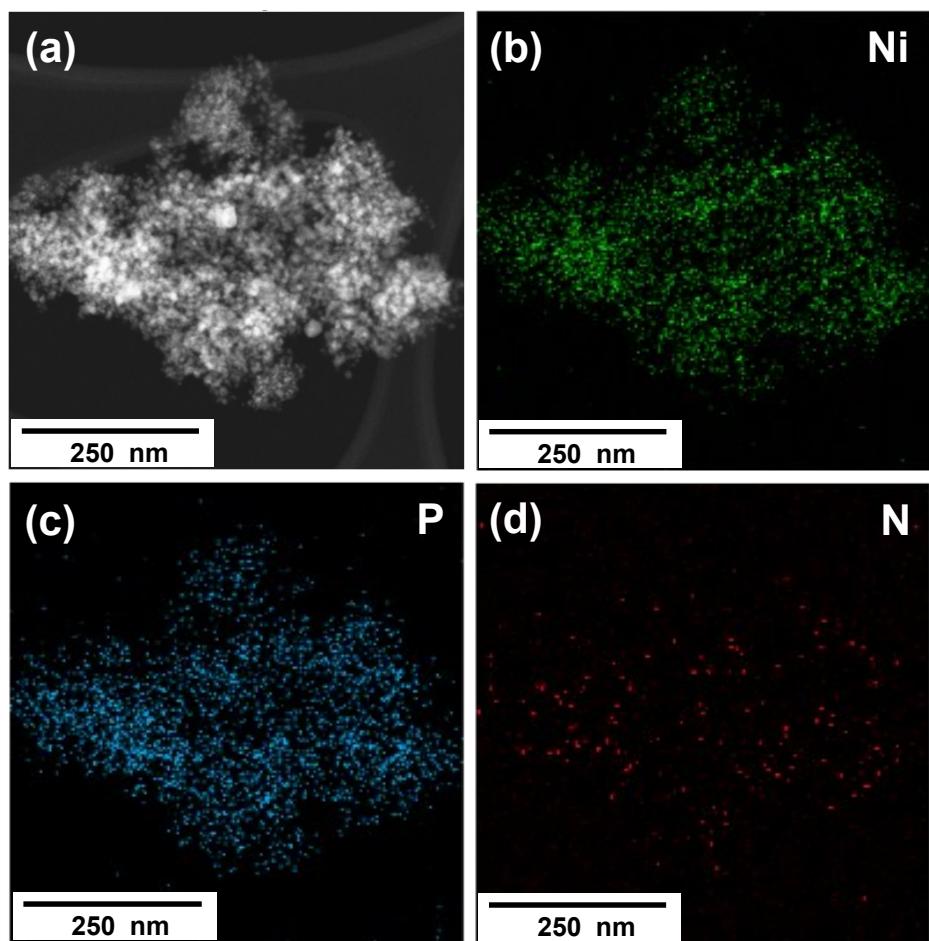
## 1. Supplementary Table S1 and Figures S1-S12

**Table S1.** Activity comparison table for HER in acidic media by nickel phosphide based catalysts.  $\eta_{10}$  indicates the overpotential required to drive a current density of -10 mA cm<sup>-2</sup>.

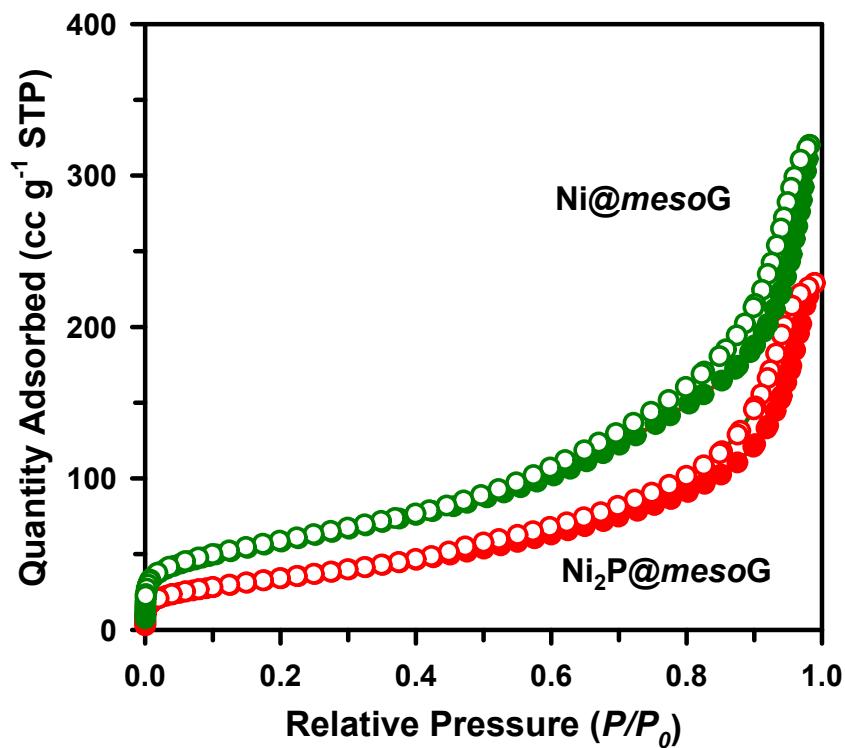
Reference	Catalyst	Loading (mg cm <sup>-2</sup> )	Electrolyte	$\eta_{10}$ (mV)	Tafel slope (mV dec <sup>-1</sup> )
This work	Ni <sub>2</sub> P@mesoG	1	0.5 M H <sub>2</sub> SO <sub>4</sub>	98	56
		3	0.5 M H <sub>2</sub> SO <sub>4</sub>	79	83
[S1]	Ni <sub>2</sub> P NPs/Ti	~1	0.5 M H <sub>2</sub> SO <sub>4</sub>	~118	~46
[S2]	Ni <sub>2</sub> P NPs	0.38	1 M H <sub>2</sub> SO <sub>4</sub>	122	87
[S3]	Ni <sub>12</sub> P <sub>5</sub> NPs/Ti	3	0.5 M H <sub>2</sub> SO <sub>4</sub>	107	63
[S4]	Ni <sub>2</sub> P/Ti	2	1 M H <sub>2</sub> SO <sub>4</sub>	122	60
[S5]	NiP <sub>2</sub> NS/CC	2.6	0.5 M H <sub>2</sub> SO <sub>4</sub>	75	~51
[S6]	Ni <sub>5</sub> P <sub>4</sub>	177	1 M H <sub>2</sub> SO <sub>4</sub>	23	33
[S7]	Ni <sub>5</sub> P <sub>4</sub> -Ni <sub>2</sub> P NS/Ni foam	-	0.5 M H <sub>2</sub> SO <sub>4</sub>	120	79.1
[S8]	Ni <sub>2</sub> P-G@Ni foam	-	0.5 M H <sub>2</sub> SO <sub>4</sub>	55	~30/107
[S9]	Ni <sub>5</sub> P <sub>4</sub>	1.99	0.5 M H <sub>2</sub> SO <sub>4</sub>	118	42



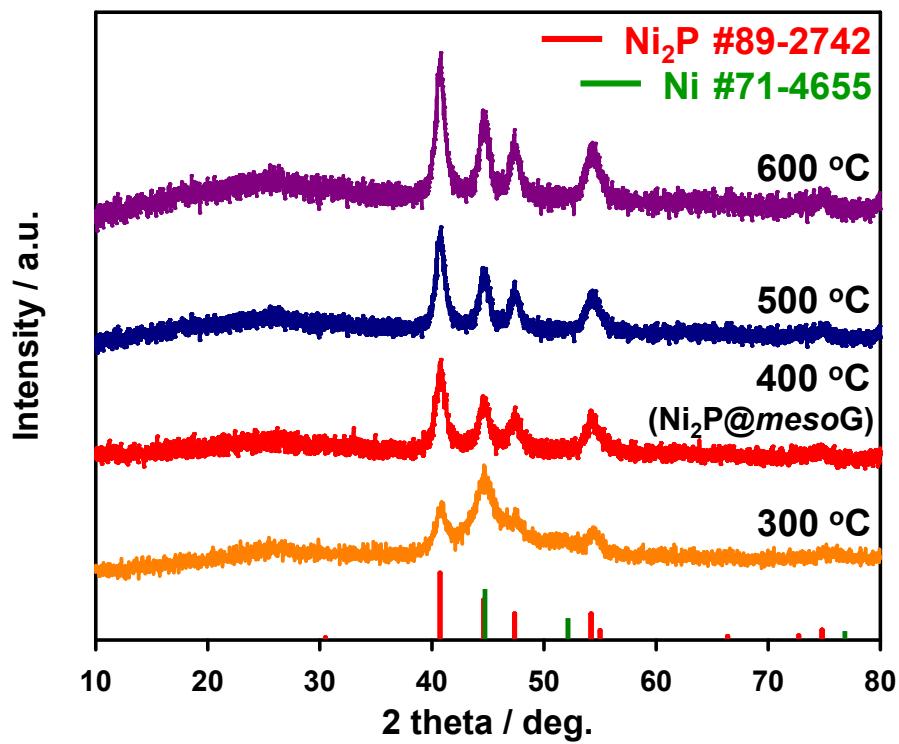
**Figure S1.** TEM images of mesoG before (a) and after (b) acid etching.



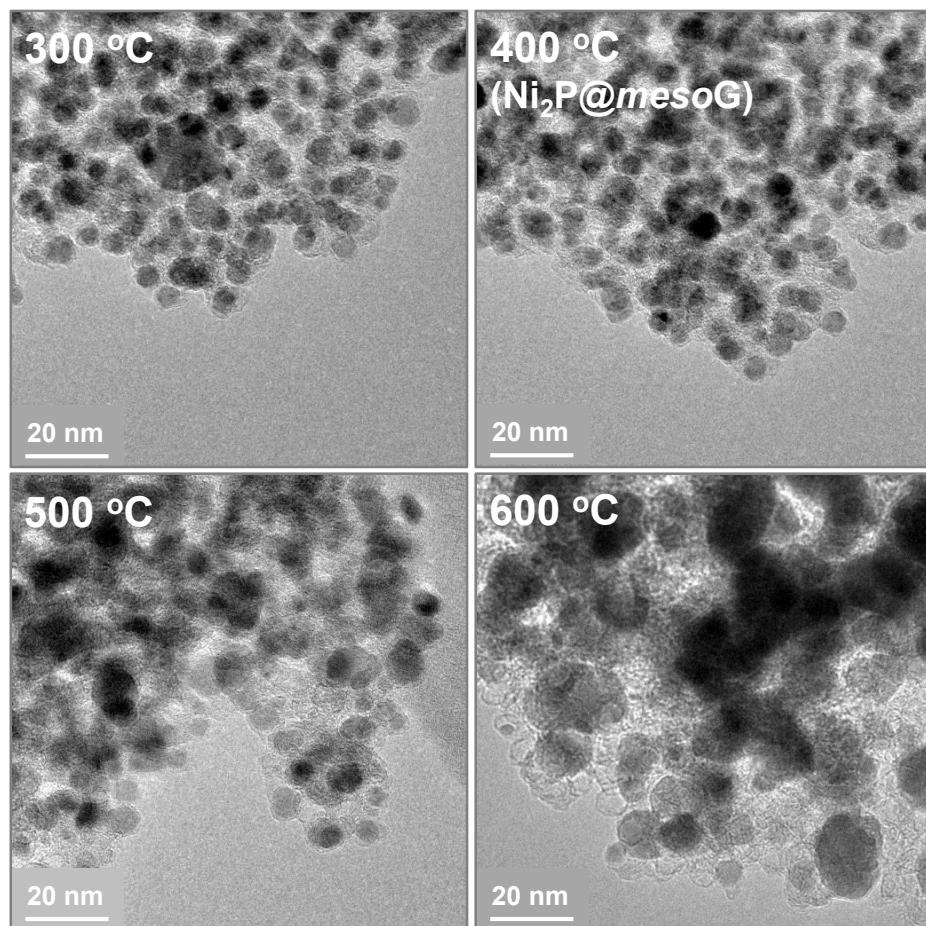
**Figure S2.** (a) STEM image of  $\text{Ni}_2\text{P}@\text{mesoG}$  and (b-d) corresponding EDS mapping images for (b) nickel, (c) phosphorus, and (d) nitrogen.



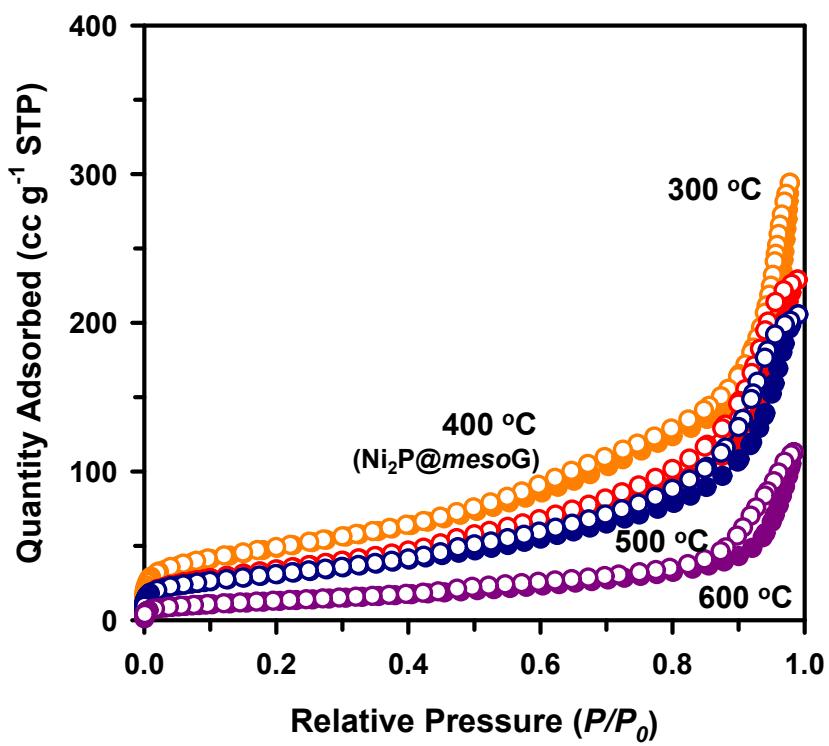
**Figure S3.** N<sub>2</sub> adsorption-desorption isotherms of Ni@mesoG and Ni<sub>2</sub>P@mesoG.



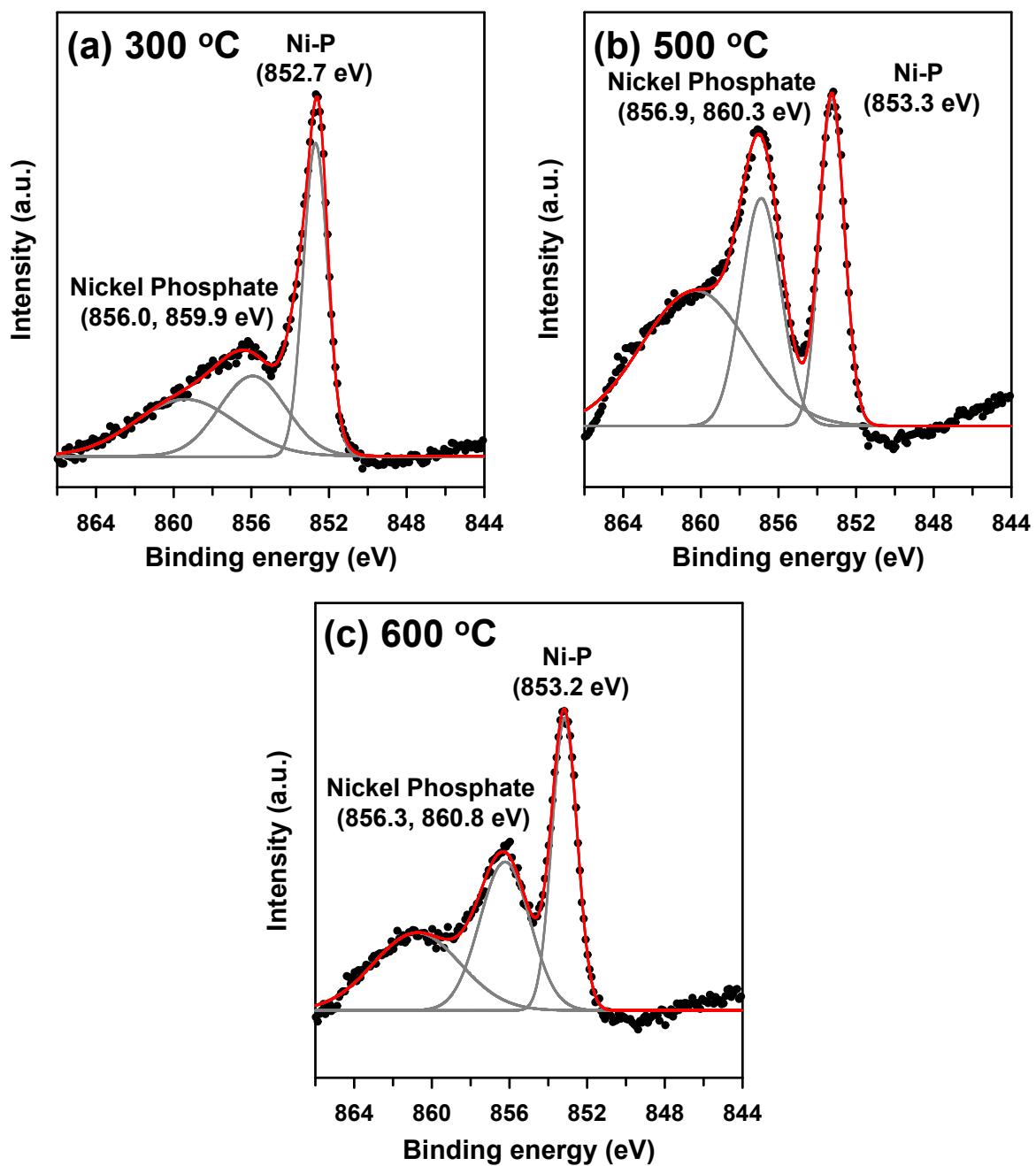
**Figure S4.** XRPD patterns of Ni@mesoG samples phosphidated at 300, 400 (Ni<sub>2</sub>P@mesoG), 500, and 600 °C.



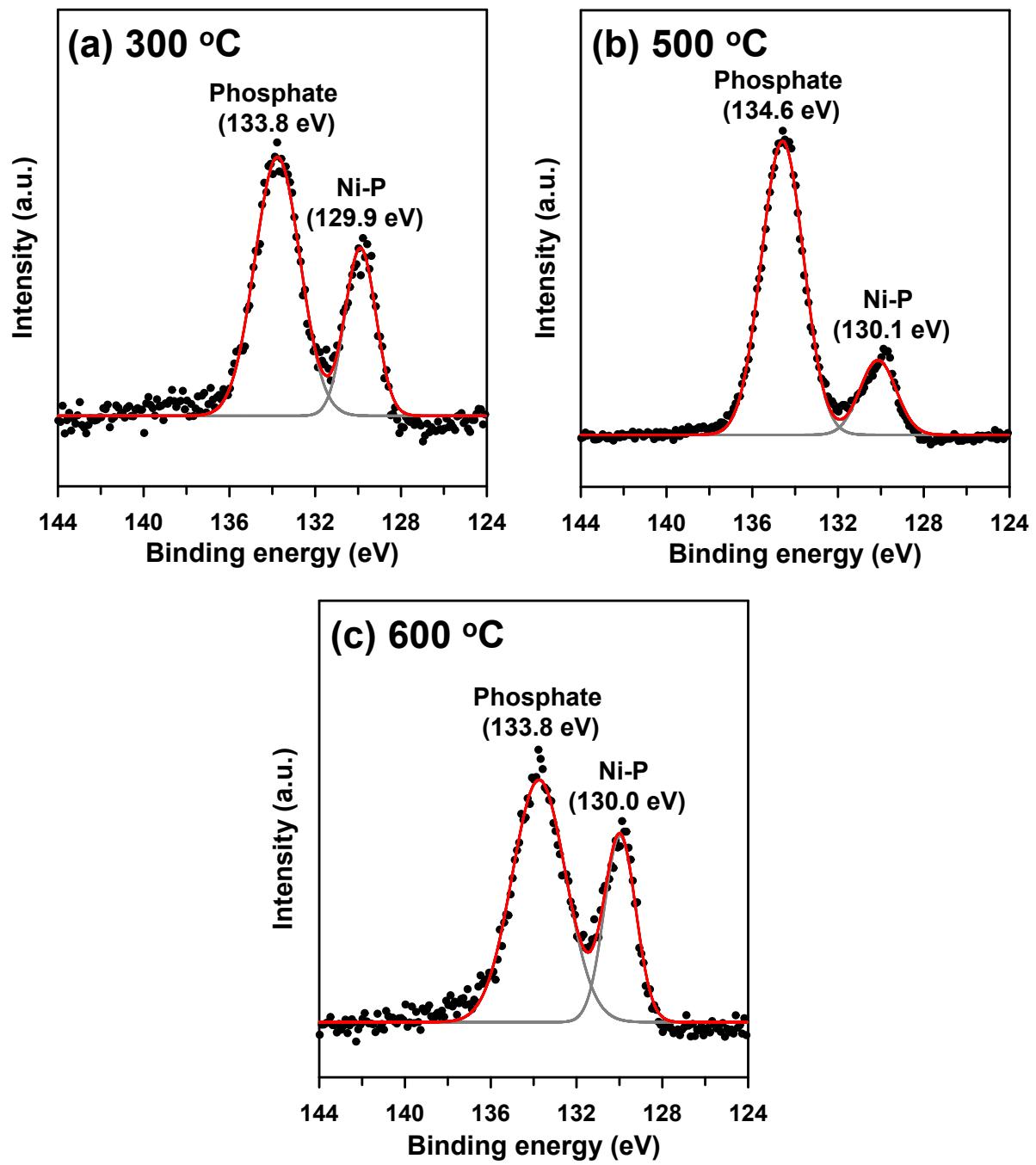
**Figure S5.** TEM images of Ni@mesoG samples phosphidated at 300, 400 (Ni<sub>2</sub>P@mesoG), 500, and 600 °C.



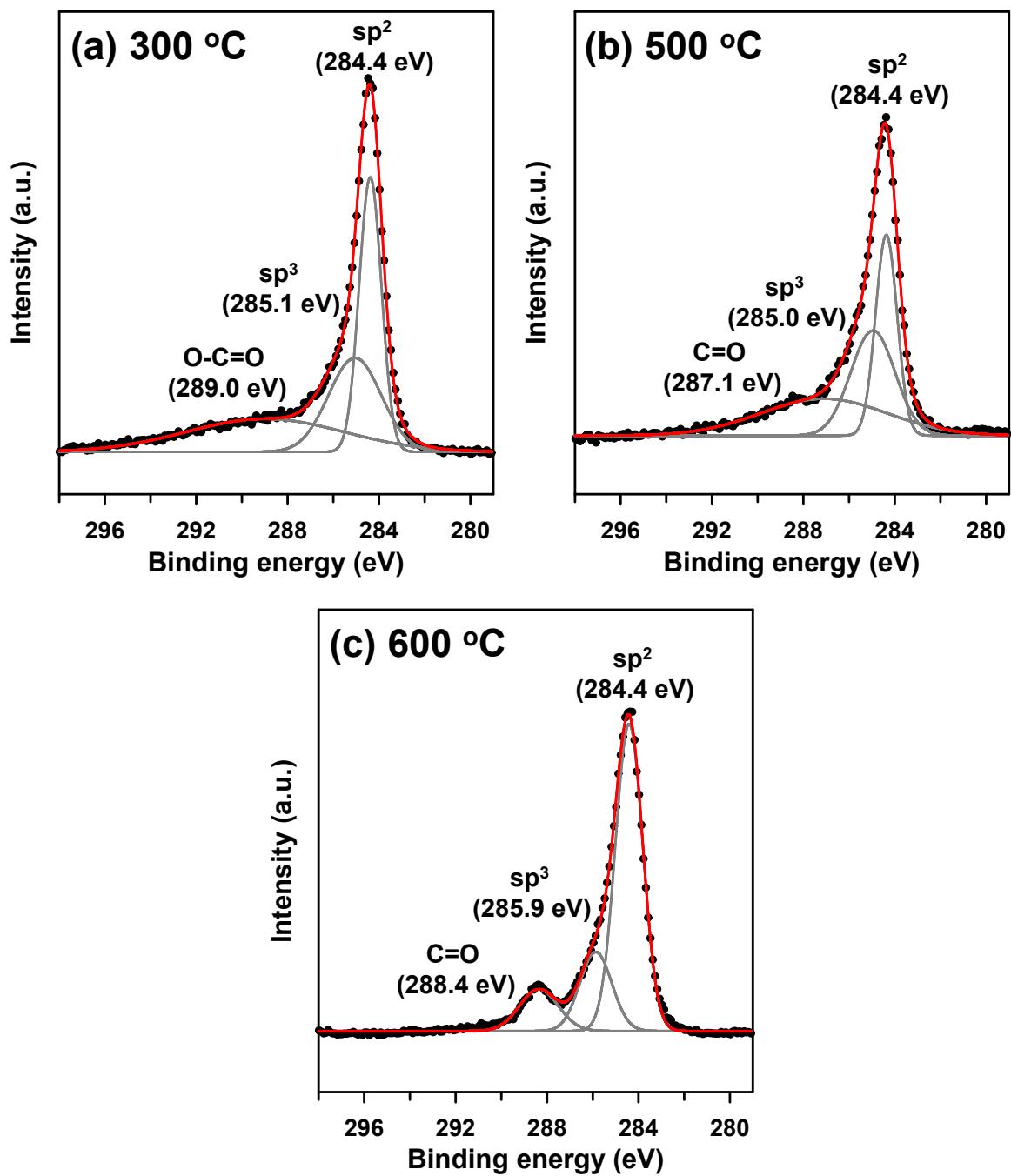
**Figure S6.** N<sub>2</sub> adsorption-desorption isotherms of Ni@mesoG samples phosphidated at 300, 400 (Ni<sub>2</sub>P@mesoG), 500, 600 °C.



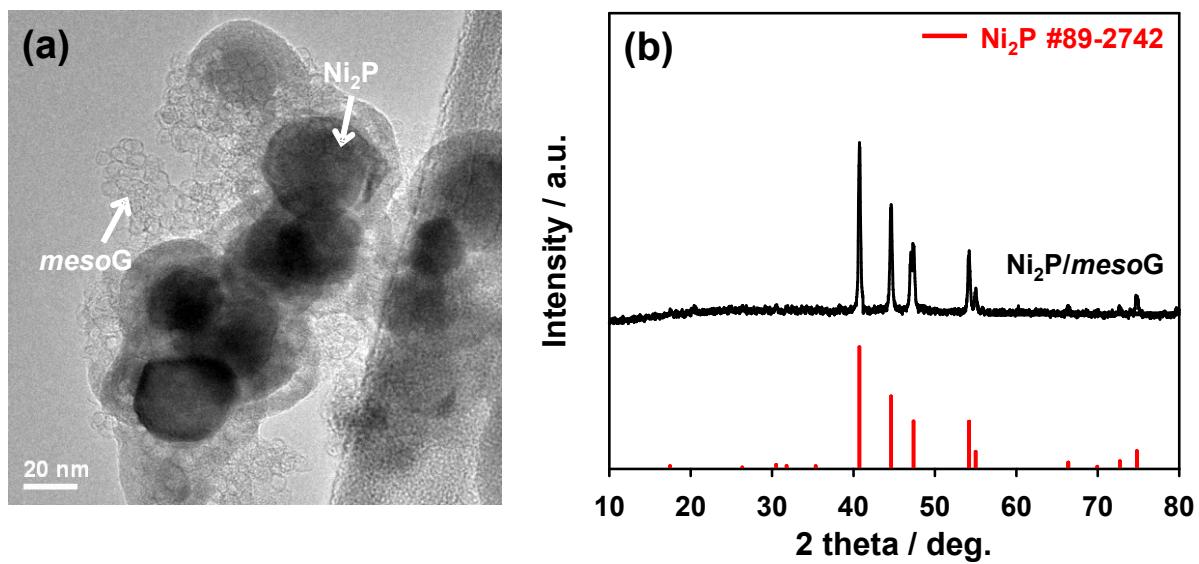
**Figure S7.** Ni 2p<sub>3/2</sub> XPS spectra of Ni@mesoG samples phosphidated at (a) 300, (b) 500, and (c) 600 °C.



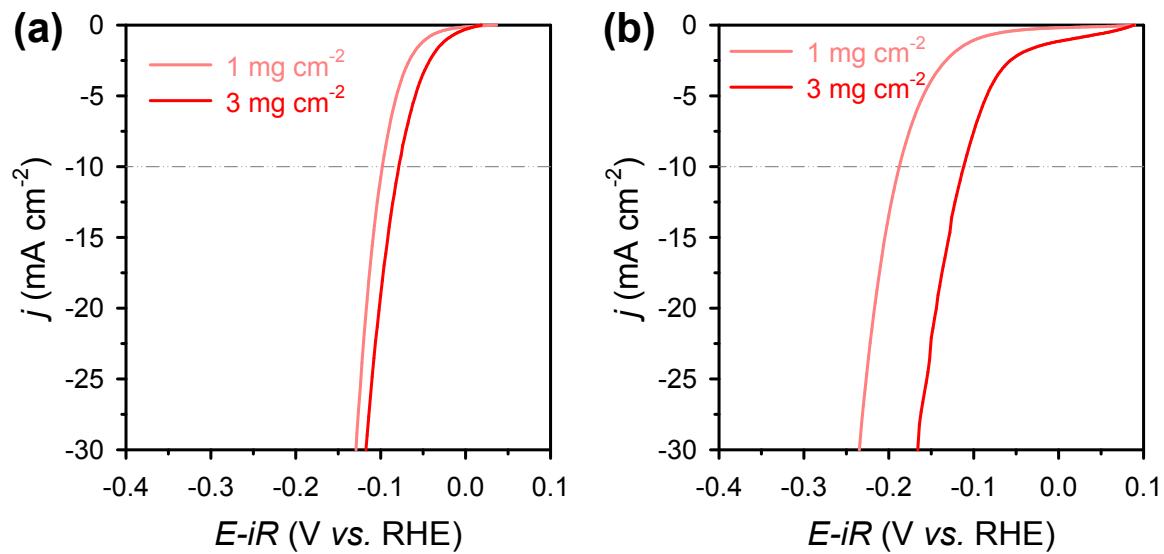
**Figure S8.**  $P\ 2p_{3/2}$  XPS spectra of Ni@mesoG samples phosphidated at (a) 300, (b) 500, and (c) 600 °C.



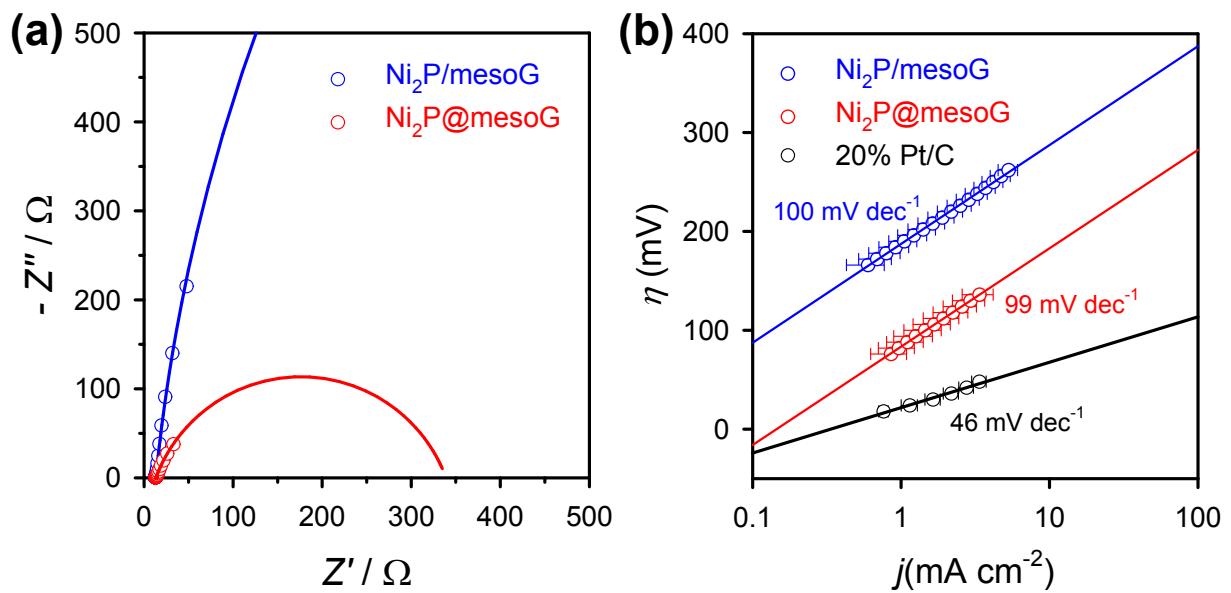
**Figure S9.** C 1s XPS spectra of Ni@mesoG samples phosphidated at (a) 300, (b) 500, and (c) 600 °C.



**Figure S10.** (a) TEM image and (b) XRPD pattern of Ni<sub>2</sub>P/mesoG.



**Figure S11.** Polarization curves for  $\text{Ni}_2\text{P}@\text{mesoG}$  with different catalyst loadings in (a) 0.5 M  $\text{H}_2\text{SO}_4$ , and (b) 1 M  $\text{KOH}$ .



**Figure S12.** (a) Nyquist plots for the impedance spectra obtained at  $\eta = 100$  mV (vs RHE) in 1 M KOH. The empty triangles and the solid lines represent numerical raw data and fitting results, respectively. (b) Tafel plots in 1 M KOH. The Tafel plots were depicted with error bar and without correction for ohmic drop.

## 2. References for Electronic Supplementary Information

- S1. E. J. Popczun, J. R. McKone, C. G. Read, A. J. Biacchi, A. M. Wiltrot, N. S. Lewis and R. E. Schaak, *J. Am. Chem. Soc.*, 2013, **135**, 9267.
- S2. L. Feng, H. Vrubel, M. Bensimon and X. Hu, *Phys. Chem. Chem. Phys.*, 2014, **16**, 5917.
- S3. Z. Huang, Z. Chen, Z. Chen, C. Lv, H. Meng and C. Zhang, *ACS Nano*, 2014, **8**, 8121.
- S4. Z. Pu, Z. Liu, C. Tang, A. M. Asiri and X. Sun, *Nanoscale*, 2014, **6**, 11031.
- S5. P. Jiang, Q. Liu and X. Sun, *Nanoscale*, 2014, **6**, 13440.
- S6. A. B. Laursen, K. R. Patraju, M. J. Whitaker, M. Retuerto, T. Sarkar, N. Yao, K. V. Ramanujachary, M. Greenblatt and G. C. Dismukes, *Energy Environ. Sci.*, 2015, **8**, 1027.
- S7. X. Wang, Y. V. Kolen'ko, X.-Q. Bao, K. Kovnir and L. Liu, *Angew. Chem. Int. Ed.*, 2015, **54**, 8188.
- S8. A. Han, S. Jin, H. Chen, H. Ji, Z. Sun and P. Du, *J. Mater. Chem. A*, 2015, **3**, 1941.
- S9. Y. Pan, Y. Liu, J. Zhao, K. Yang, J. Liang, D. Liu, W. Hu, D. Liu, Y. Liu and C. Liu, *J. Mater. Chem. A*, 2015, **3**, 1656.