Electronic Supplementary Information

Selective oxidation of organic sulfides to sulfoxides using sugar derived *cis*-dioxo molybdenum(VI) complexes: it's kinetic and mechanistic studies

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1. Characterization data of sulfoxides

1.1 Diphenyl sulfoxide (**SO2**)

White crystal; m.p. 68-70 °C; IR (KBr; cm⁻¹): 1041 ($\nu_{S=0}$); ¹H NMR (CDCl₃, 400 MHz, ppm): δ 7.65 (4H, m, ArH), 7.46 (6H, m, ArH) ¹³C NMR (100 MHz, CDCl₃, ppm): δ 145.8, 131.2, 129.5, 125.0; ESI-MS: *m/z* calcd for (M+H)⁺ C₁₂H₁₁OS 203.05; found 203.11.

1.2 Benzyl phenyl sulfoxide (SO3)

White crystal; m.p. 124-126 °C; IR (KBr; cm⁻¹): 1034 ($v_{S=O}$). ¹H NMR (CDCl₃, 400 MHz, ppm): δ 7.41 (5H, m, ArH), 7.24 (3H, m, ArH) 6.99 (2H, m, ArH), 4.0 (2H, m, SOCH₂) ¹³C NMR (100 MHz, CDCl₃, ppm): δ 142.8, 131.1, 130.3, 129.1, 128.8, 128.4, 128.2, 124.4, 63.5; ESI-MS: *m/z* calcd for (M+H)⁺ C₁₃H₁₃OS 217.07; found 216.93.

1.3 2-(Phenylsulfinyl)aniline (**SO4**)

Brown liquid; IR (KBr; cm⁻¹): 1018 ($\nu_{S=0}$). ¹H NMR (CDCl₃, 400 MHz, ppm): δ 7.58 (2H, m, ArH), 7.44 (4H, m, ArH) 7.25 (1H, m, ArH) 6.78 (1H, m, ArH), 6.59 (1H, m, ArH), 4.94 (2H, br, NH₂) ¹³C NMR (100 MHz, CDCl₃, ppm): δ 147.6, 143.2, 133.0, 130.3, 128.9, 128.4, 124.7, 123.5, 117.6, 117.2; ESI-MS: *m/z* calcd for (M+H)⁺ C₁₂H₁₂NOS 218.06; found 218.05.

2. Spectroscopic data



Fig. S1: ¹H NMR of complex 5 recorded in DMSO-d₆ (400 MHz)



Fig. S2: ¹³C NMR of complex 5 recorded in DMSO-d₆ (100 MHz)



Fig. S3: HRMS of complex 5



Fig. S4: ¹H NMR of complex 6 recorded in DMSO-d₆ (400 MHz)



Fig. S5: ¹³C NMR of complex 6 recorded in DMSO-d₆ (100 MHz)



Fig. S6: HRMS of complex 6



Fig. S7: Representative HPLC of diphenyl sulfoxide (**S02**), formed from catalyst **5**; (peaks in the range of 2.5-3.5 RT are from the blank)



Fig. S8: Representative HPLC of benzyl phenyl sulfoxide (SO3), formed from catalyst 1; (peaks in the range of 2.5-3.5 RT are from the blank)



Fig. S9: Representative HPLC of 2-(phenylsulfinyl)aniline (**SO4**), formed from catalyst **3**; (peaks in the range of 2.5-3.5 RT are from the blank)



Fig. S10: ¹H NMR of SO2 recorded in CDCI₃ (400 MHz)



Fig. S11: ¹³C NMR of SO2 recorded in CDCI₃ (100 MHz)







Fig. S13: ¹H NMR of SO3 recorded in CDCl₃ (400 MHz)



Fig. S14: ¹³C NMR of **SO3** recorded in CDCl₃ (100 MHz)



Fig. S15: ESI-MS of SO3



Fig. S16: ¹H NMR of SO4 recorded in CDCl₃ (400 MHz)



Fig. S17: ¹³C NMR of SO4 recorded in CDCI₃ (100 MHz)







Fig. S19: ¹H NMR of S5 recorded in CDCl₃ (400 MHz)



Fig. S20: ¹³C NMR of S5 recorded in CDCl₃ (100 MHz)



Fig. S21: ¹H NMR of SO5 recorded in CDCI₃ (400 MHz)



Fig. S22: ¹³C NMR of SO5 recorded in CDCl₃ (100 MHz)



Fig. S23: Comparative UV-visible absorption spectra of reactants (S5) (black), (SO5) (red), complex 1(blue), UHP (pink)



Fig. S24: UV-Vis spectrum of diphenyl sulphide **(S2)** [10⁻⁵ M], Complex **1** [10⁻⁵ M] and mixture of **S2** and complex **1**



Fig. S25: UV-Vis spectrum of Diphenylsulfide (S2) [10⁻⁵ M], UHP [10⁻⁵ M] and mixture of S2 and UHP



Fig. S26: Absorption spectra of complex **1** [10⁻⁵ M], UHP [10⁻⁵ M] and mixture of UHP and complex