

**Dual role of Hydrogen Peroxide on the Oxidase-like Activity of Nanoceria and its  
Application for Colorimetric Hydrogen Peroxide and Glucose Sensing**

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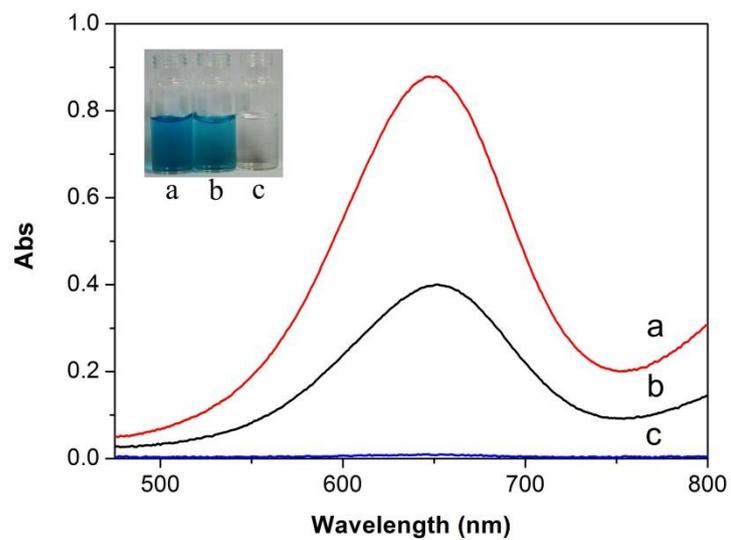
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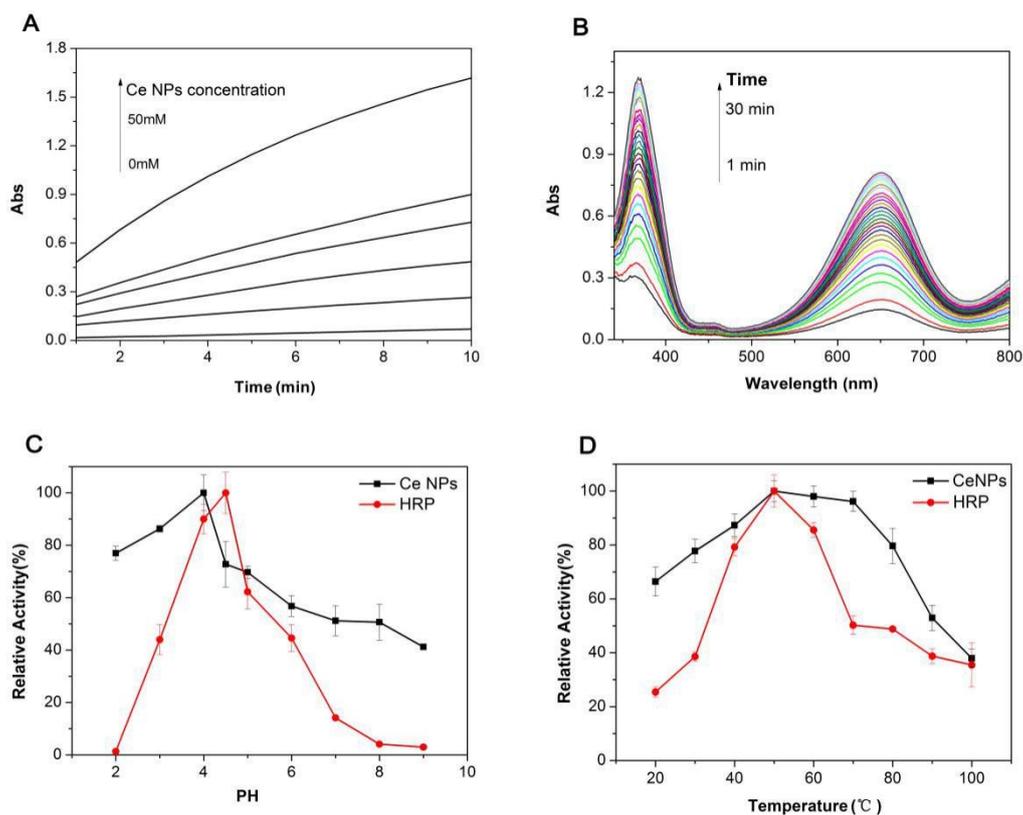
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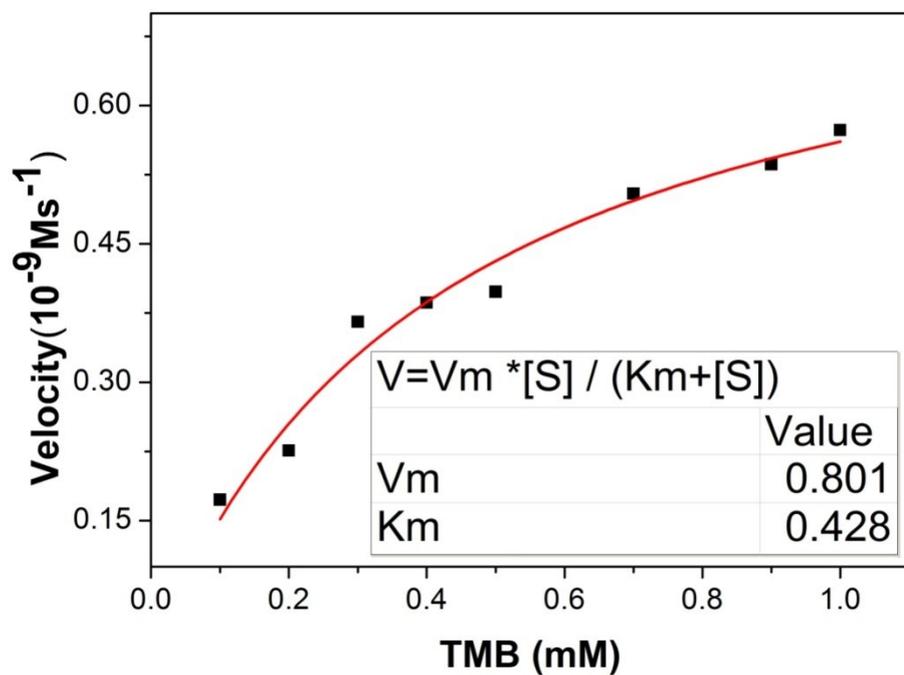
**Fig. S1** Photograph of nanoceria solution.



**Fig. S2** Typical absorption curves of TMB reaction solutions mixing with (a) Ce NPs; (b) Ce NPs and H<sub>2</sub>O<sub>2</sub>; (c) H<sub>2</sub>O<sub>2</sub> at room temperature in acetate buffer after reaction for 10 min. Inset: photographs of different solutions each corresponding to (a) to (c).



**Fig.S3** Dependency of the Ce NPs oxidase-like activity on Ce NPs concentration (A), reaction time (B), pH (C) and temperature (D). Experiments were carried out using 0.8 mM TMB as a substrate and the reaction time was 10 min in the group of (A), 20 min in the group of (C) and (D). The H<sub>2</sub>O<sub>2</sub> concentration was 0 mM in all Ce NPs catalysis experiments and 10 mM in the HRP catalysis experiments. The error bars represent the standard deviation for three measurements.



**Fig. S4** Steady-state kinetic assay and catalytic mechanism of Ce NPs. The velocity ( $v$ ) of the reaction was measured using 50 mg/L Ce NPs in 1 mL acetate buffer at pH 4.0 and room temperature (about 15°C).

**Table S1.** Comparison of the proposed method with other colorimetric glucose sensors

Catalyst	Substrate	H <sub>2</sub> O <sub>2</sub>		Glucose		Reference
		LOD ( $\mu$ M)	Linear Range ( $\mu$ M)	LOD ( $\mu$ M)	Linear Range ( $\mu$ M)	
Nanoceria	TMB	2.5	4–40	2	4–40	This work
Gold clusters	TMB	–	–	–	2.0 $\times$ 10 <sup>3</sup> – 1.0 $\times$ 10 <sup>4</sup>	1
Fe <sub>3</sub> O <sub>4</sub>	ABTS	3.0	5–100	30	50–1000	2
NiO nitrogen- doped	TMB	8.0	0.02–0.10	20	50–500	3
graphene quantum dots	TMB	5.3	20–1170	16	25– 375	4

## References

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