

Isomorphous substitution induced ionothermal synthesis of magnesium aluminophosphate zeolite in fluoride-free media

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Supplementary Information

Characterizations

Chemical composition analysis was determined with a Philips Magix X-ray fluorescence spectrometer. Powder X-ray diffraction (XRD) analysis of the as-synthesized samples are recorded on a PANalytical X'Pert PRO diffractometer fitted with Cu K α radiation ($\lambda = 1.5418 \text{ \AA}$) operating at 40 mA and 40 kV. The Scanning electron microscopy (SEM) was performed on a Hitachi S4800 field-emission scanning electron microscopes. ³¹P and ²⁷Al magic angle spinning (MAS) and ¹³C cross polarization-magic angle spinning (CP-MAS) nuclear magnetic resonance (NMR) measurements were analyzed on a Varian Infinityplus-400 spectrometer. The acidity of the sample was characterized by using ammonia temperature programmed desorption (NH₃-TPD) with a Micromeritics Autochem 2910 automated catalyst characterization system. Thermal gravimetric/Differential scanning calorimetry (TG/DSC) was carried out under air atmosphere, using a NETZSCH STA 409PC thermoanalyser from room-temperature to 800 °C. The sample weight was 15 mg, and the heating rate was 10 °C/min, with an air flow rate of 30 ml/min.

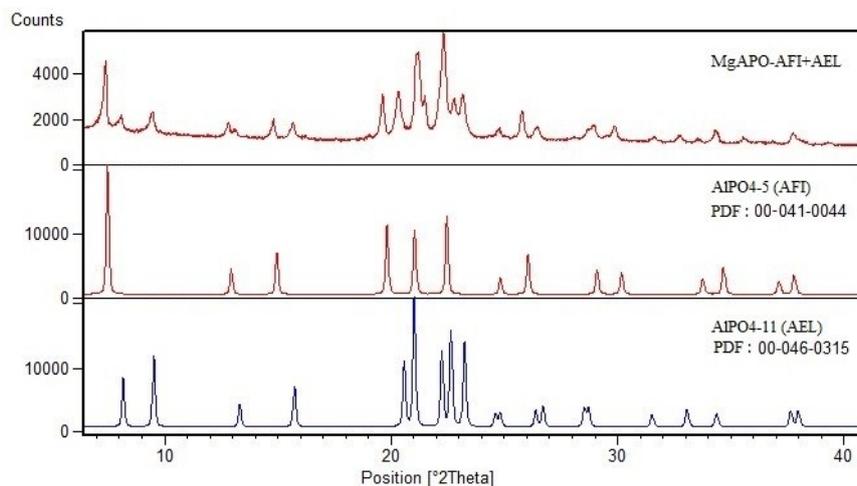


Figure S1. XRD pattern of MgAPOs with AFI and AEL structures and the reference XRD patterns of AlPO₄-11(AEL) and AlPO₄-5 (AFI).

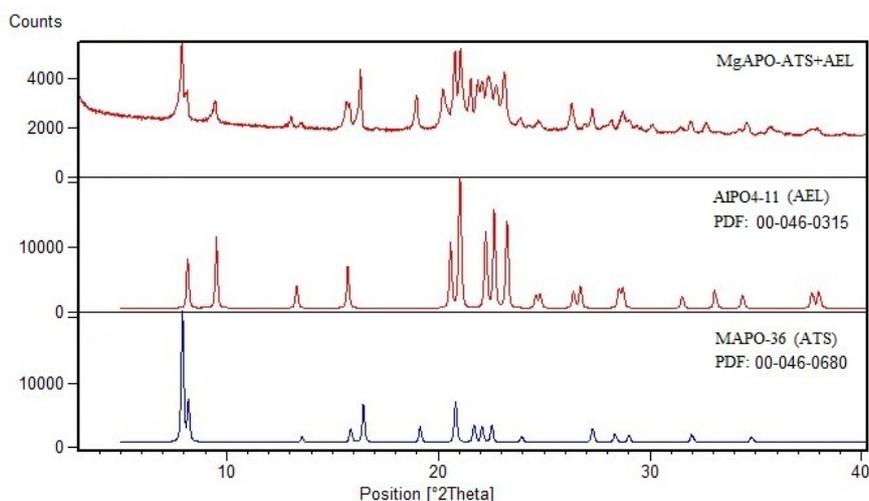


Figure S2. XRD pattern of MgAPOs with ATS and AEL structures and the reference XRD patterns of AlPO₄-11(AEL) and MAPO-36 (ATS).

Table S1 Product structure and the MgO/Al₂O₃ ratio for the samples synthesized with different magnesium content in the initial mixture. Initial mixture composition (molar ratio): 1 Al₂O₃ : 3 P₂O₅ : x MgO : 80BMIMBr_s, crystallization temperature : 200°C, crystallization time: 1 day.

Sample	x	Structure	MgO/Al ₂ O ₃ in the products*
A	0.1	AEL	0.042
B	0.2	AFI+AEL	0.071
C	0.3	AFI+AEL	0.094
D	0.4	AFI+AEL	0.108
E	0.5	AFI+AEL	0.141

* Determined by X-ray fluorescence spectrometer element analysis.

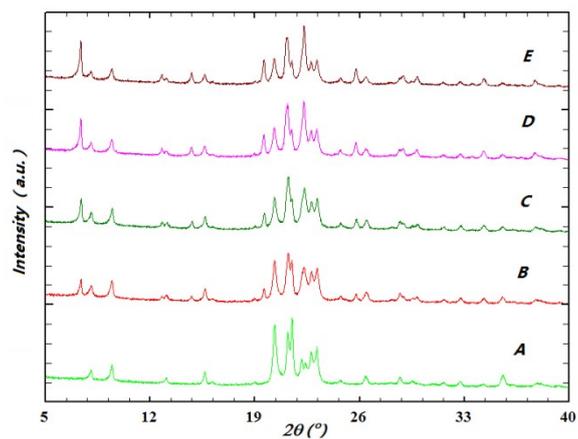


Figure S3. XRD patterns of the samples A-E synthesized with different magnesium concentration in the initial mixture.

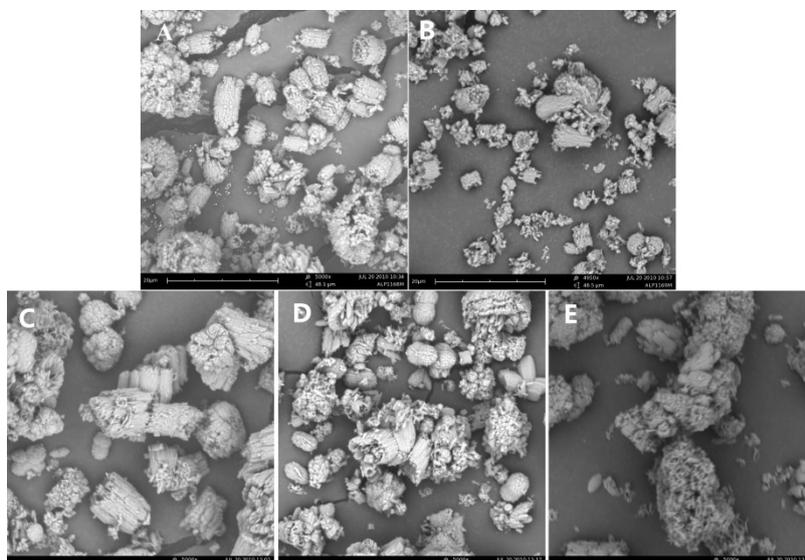


Figure S4. SEM images of the samples A-E synthesized with different magnesium concentration in the initial mixture.

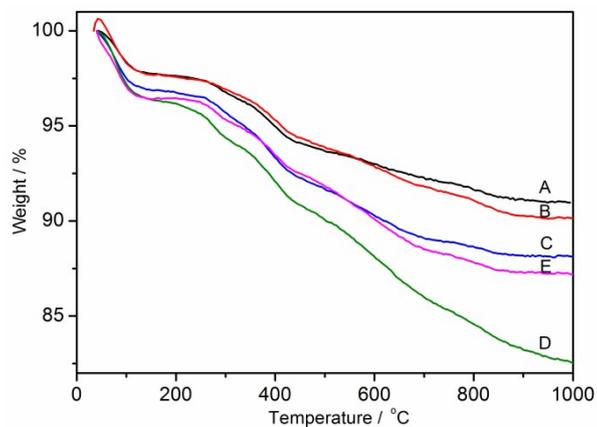


Figure S5. TG profiles of the samples A-E synthesized with different magnesium concentration in the initial mixture. Note: The weight loss increase with the magnesium content in the framework, for sample E possibly some extra-framework magnesium exists.