

Supporting information

Development of the γ -polyglutamic acid binder for cathodes with high mass fraction of sulfur

Tingting Qiu^a, Hongyuan Shao^a, Weikun Wang^b, Hao Zhang^b, Anbang Wang^b, Zhenxing Feng^c, and Yaqin Huang^{*a}

^aState Key Laboratory of Chemical Resource Engineering, The Key Laboratory of Beijing City on Preparation and Processing of Novel Polymer Materials, Beijing University of Chemical Technology, 15 Beisanhuan East Road, Beijing, 100029, P.R. China. E-mail:

huangyq@mail.buct.edu.cn; Fax: +86 10 6443 8266; Tel: +86 10 6443 8266.

^bResearch Institute of Chemical Defense, 35 Huayuan North Road, Beijing, 100191, P.R. China.

^cSchool of Chemical, Biological, and Environmental Engineering, Oregon State University, Corvallis, OR, 97331, USA.

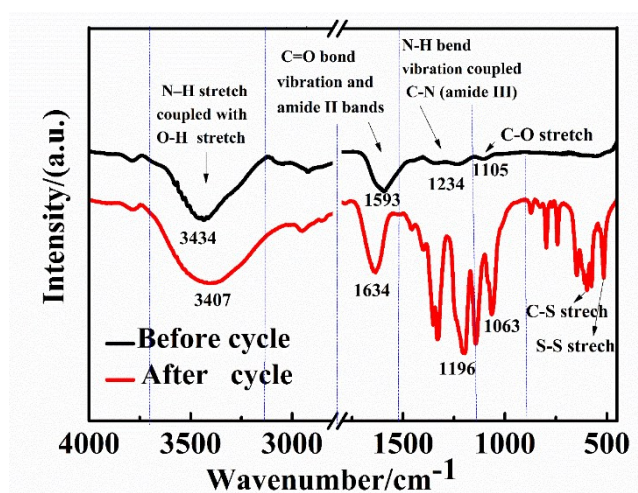


Fig.S1. FTIR of the PGA-70S cathode before cycle and after cycle.

The PGA-70S cathode after cycle was obtained through cycling the PGA-70S cathode for several times and stopping it at the upper discharge flat where high-order lithium polysulfides generated. For the PGA-70S cathode after cycle, the emergence of C–S bond at $\approx 600\text{ cm}^{-1}$ suggested that the PGA could be capable of forming chemical bonds with polysulfides.

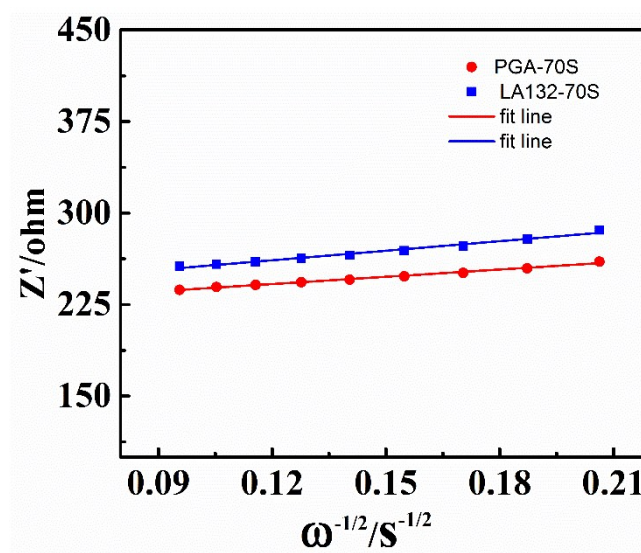


Fig. S2 The relationship between Z_{re} and $\omega^{-1/2}$ at low frequencies for PGA-70S cathode and LA132-70 cathode before cycle.

The lithium diffusion coefficient was calculated by using the following equations:

$$Z_{re} = R_{ct} + R_s + \sigma\omega^{-1/2} \quad (1)$$

where R_e is the resistance of the electrolyte, R_{ct} is the charge transfer resistance and ω is the angular frequency in the low frequency region and the σ is the Warburg factor, which can be obtained from the slopes of lines in Figure S 2.

$$D_{Li} = \frac{R^2 T^2}{2A^2 n^4 F^4 C_{Li}^2 \sigma^2} \quad (2)$$

where D_{Li} is the diffusion coefficient, R is the gas constant, T means the temperature, A is area of the electrode, n is the number of electrons involved, F is the Faraday constant, C_{Li} is the concentration of lithium ion in electrolyte.

Therefore, for the PGA-70S cathode and LA132 cathode

$$D_{\text{PGA-Li}} : D_{\text{LA132-Li}} = \sigma_{\text{LA}}^2 / \sigma_{\text{PGA}}^2 \quad (3)$$

where $D_{\text{PGA-Li}}$ and $D_{\text{LA132-Li}}$ is the diffusion coefficient of PGA-70 cathode and LA132-70 cathode, and is the Warburg factor of PGA-70 cathode and LA132-70 cathode.

Thus: $D_{\text{PGA-Li}} : D_{\text{LA132-Li}} = 1.7$

As confirmed by the diffusion coefficient, the PGA-70S cathode showed higher lithium ion conductivity. ^{1,2}

References

1. H. Wang, V. Sencadas, G. Gao, H. Gao, A. Du, H. Liu and Z. Guo, *Nano Energy*, 2016, **26**, 722–728.
2. Y. Cui, X. Zhao and R. Guo, *Electrochimica Acta*, 2010, **55**, 922-926.