The maiden report of a fluorescent-colorimetric sensor for expeditious detection of bifluoride ion in aqueous media

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Fig. S3. IR Spectra of L.



Fig. S4. Mass Spectra of L.



Fig. S6. ¹³C NMR Spectra of L_1 in CDCl₃.







Fig. S8. Mass Spectra of L₁.



Fig. S9. The experimental (black) and calculated (red, non-scaled) vibrational spectra and calculated vibrational transition (blue) of for L.



Fig. 10. The experimental (black), calculated (red) electronic absorption spectra and calculated electronic transition (blue) of L.



Fig. S11. Selected HOMOs and LUMOs of L.



Fig. S12. Changes in the absorption spectra of L (50 μ M) in the presence of 10 equiv. of different cations.



Fig. S13. Hill plot



Fig. S14. Detection limit



Fig. S15. The bar graph showing the relative absorption intensity of L upon treatment with various anions (1. L; 2. HF2-; 3. HSO_3^- ; 4.F⁻; 5.OAc⁻; 6.Cl⁻; 7. Br⁻; 8.I⁻; 9. $H_2PO_4^-$; 10. HCO_3^- ; 11. N_3^- ; 12. NO_3^- ; 13. SO_4^{2-} 14.CO₃²⁻; 15. CN- and 16.SO₃²⁻).



Fig. S16. The bar graph showing the relative absorption intensity of L upon treatment with various metal cations (1.L; $2.\text{HF}_2$; $3.\text{Mg}^{2+}$, $4.\text{Mn}^{2+}$, $5.\text{Fe}^{3+}$, $6.\text{Ba}^{2+}$, $7.\text{Co}^{2+}$, $8.\text{K}^+$, $9.\text{Al}^{3+}$, 10. Ni²⁺, 11. Ca²⁺, 12. Cu²⁺, 13. Cr³⁺, 14. Zn²⁺, 15. Cd²⁺, 16. Na⁺.)



Fig. S17. Change in absorbance spectra of L_1 in presence and absence of bifluoride ion.



Fig. S18. Job's plot.



Fig. S19. HOMO-LUMO of $L-2(HF_2)$ adduct















Fig. S20. ESI-MS of L at different pH, (a) pH = 4; (b) pH = 5; (c) pH = 6; (d) pH = 8 and (e) pH = 10



Fig. S21. Time evolution