

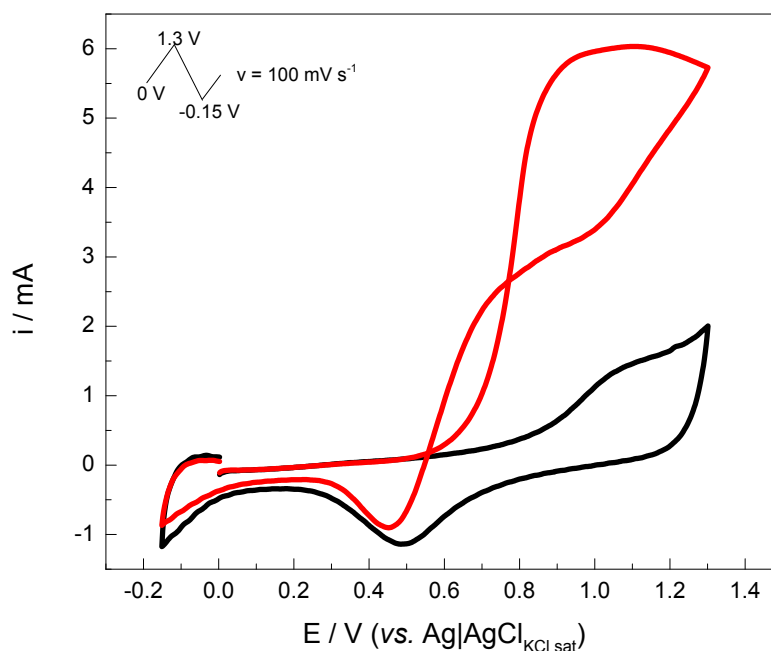
L-Cysteine Electrooxidation in Alkaline and Acidic Media: A Combined Spectroelectrochemical and Computational Study

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Surface Cleanness Process Efficiency

The process of cleaning the electrode, as shown in the experimental section, was verified by a cyclic voltammetry. For this, the electrode used for spectroelectrochemical measurements was cleaned in piranha solution and then placed into an electrochemical cell containing just the supporting electrolyte, H₂SO₄ 0.50 mol L⁻¹, and the scan was registered (S.F. 1). Just after this, some crystals of L-Cys were added to the system and the procedure was repeated.



S.F.1. Cyclic voltammetry to verify the cleanness efficiency. Black line in absence and red line in the presence of L-Cys, perturbation program inserted.

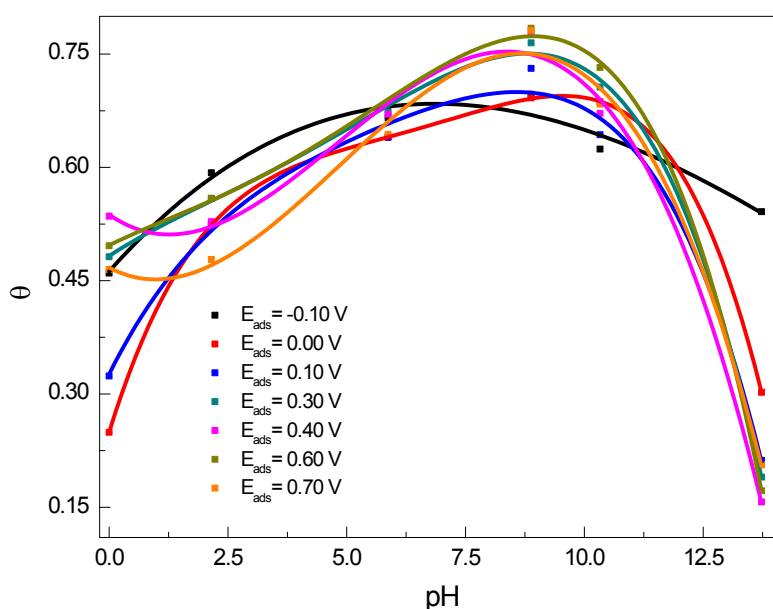
As could be seen, in absence of L-Cys a cyclic voltamogram with the predicted shape for Pt electrode in sulfuric acid media is observed, with a H_{UPD} region between -0.15 and 0.00 V, a oxide formation at potentials more positives than 0.90 V and reduction of it in the back scan around 0.50 V. In presence of L-Cys, even diluted, it

could be notice that the oxidation starts earlier, around 0.60 V, and the current increases at more positive potentials and almost did not reached a limit. A cross-over was observed at 0.80 V, but even with the positive current in a large region of the reverse scan, it was possible to observe the reduction of oxide in the same region as in the blank, just as the H_{UPD} , but with lower intensity, suggesting that some L-Cys is still adsorbed onto the electrode.

The cleanness could be verified, since the difference was very expressive between both situations and the shape of the blank experiment was in agreement with the expected for Pt electrodes.

Surface Coverage

To analyze the way in which the L-cys covers the electrode surface, some electrochemical measurements were performed. For this, the working electrode was inserted in solutions at the same pH conditions as for the IRRAS measurements. Then the electrode was submitted to a potential polarization that varied from -0.10 to 0.70 V (E_{ads}) and the coverage degree was estimated by the difference of charge of the H_{UPD} region before and after the modification with L-Cys, in a free amino acid solution.



S.F. 2. Coverage degree of Pt electrode in function of pH for different applied potentials. Lines does not present any physical meaning, just a guide for the tendency at each potential.

Spectroelectrochemical measurements

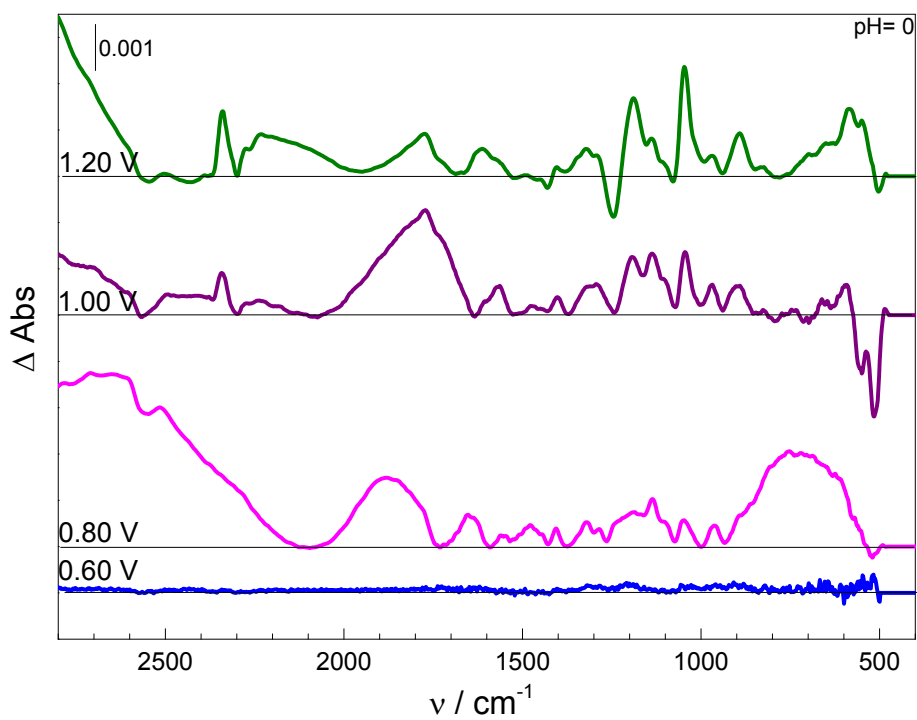
Because of the low intensity of the adsorption spectrum at pH 3 in Figure 3, it could be necessary a better comparison between spectra at pH 3 and 7. The type of signal observed (single/large band or multiplet) and the region observed at each condition are showed in S.T. 1. The small degree of variation at each pH condition corroborates the idea of similar spectra suggest in the text.

S.T. 1. Correspondence between the wavenumber observed in adsorption spectra at pH 3 and 7 (Figure 3).

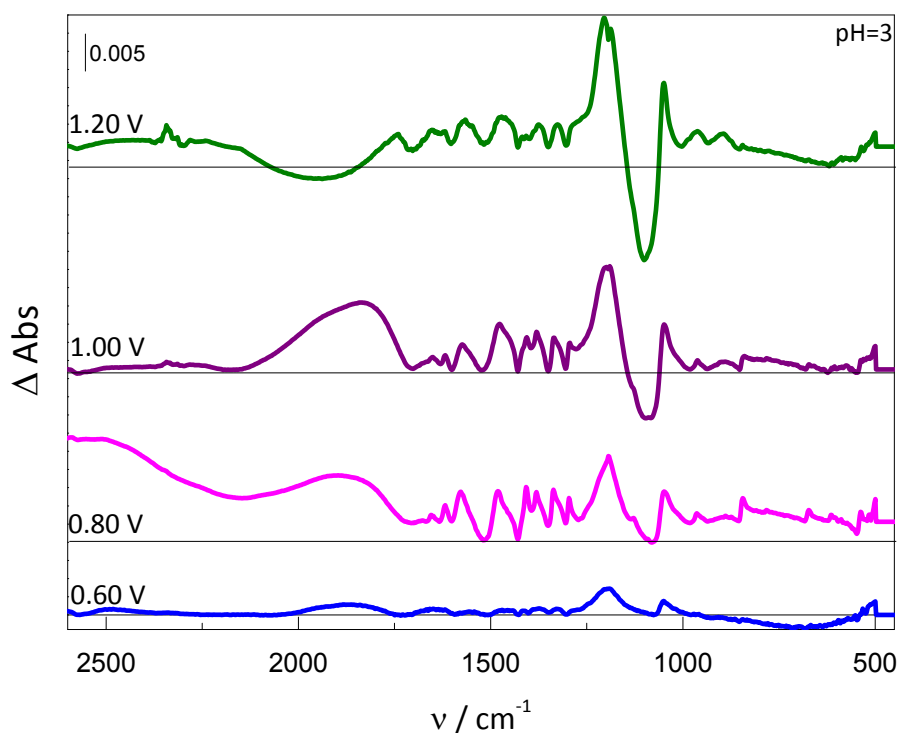
	large band	single band	single band	single band	single band	multiplet	single band	single band
pH 7	1898-1656	1609	1555	1551	1443	1401-1259	1172	1016
pH 3	2141-1692	1652	1616	1558	1486	1425-1299	1197	1049

Oxidation Spectra

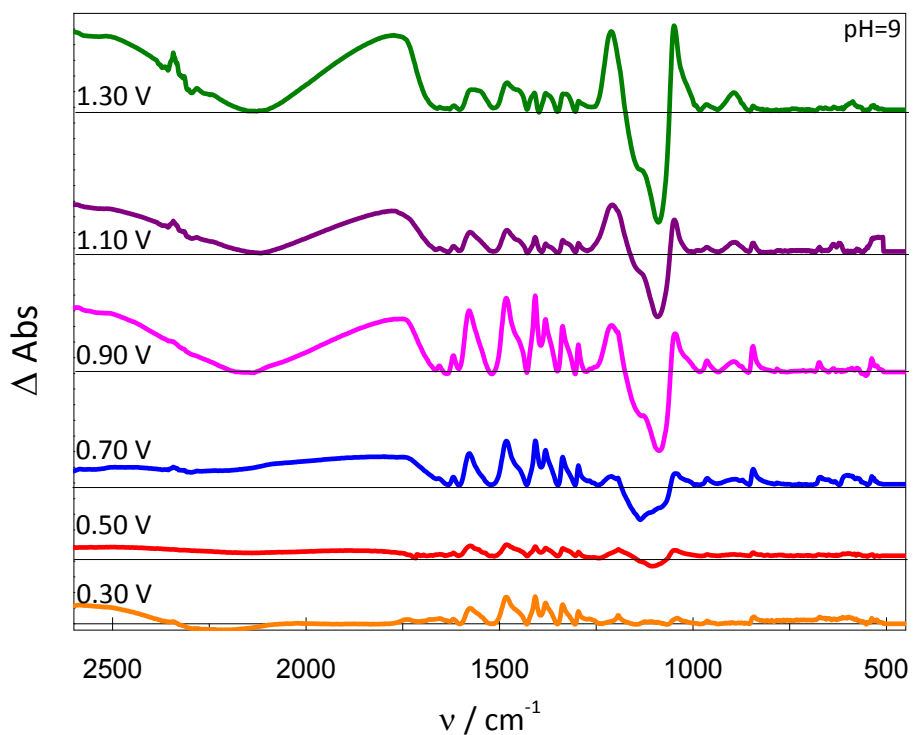
The following data presented in here are related to the oxidation spectra of L-Cys and L-Cys₂ for pH values different from 7.



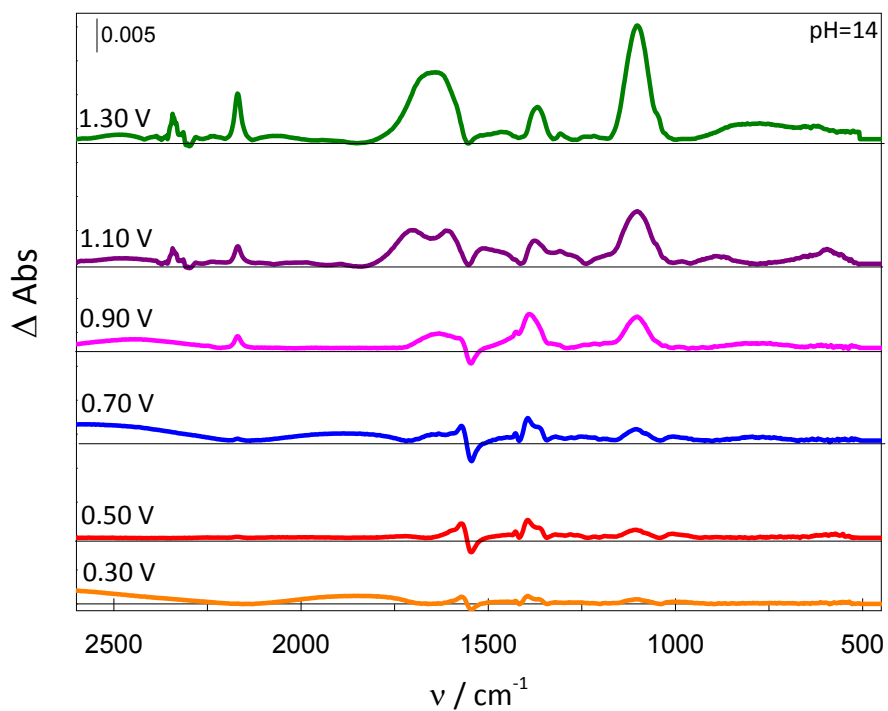
S.F.3. Oxidation spectra of L-Cys 0.10 mol L⁻¹ in H₂SO₄ 0.50 mol L⁻¹. Applied potentials showed.



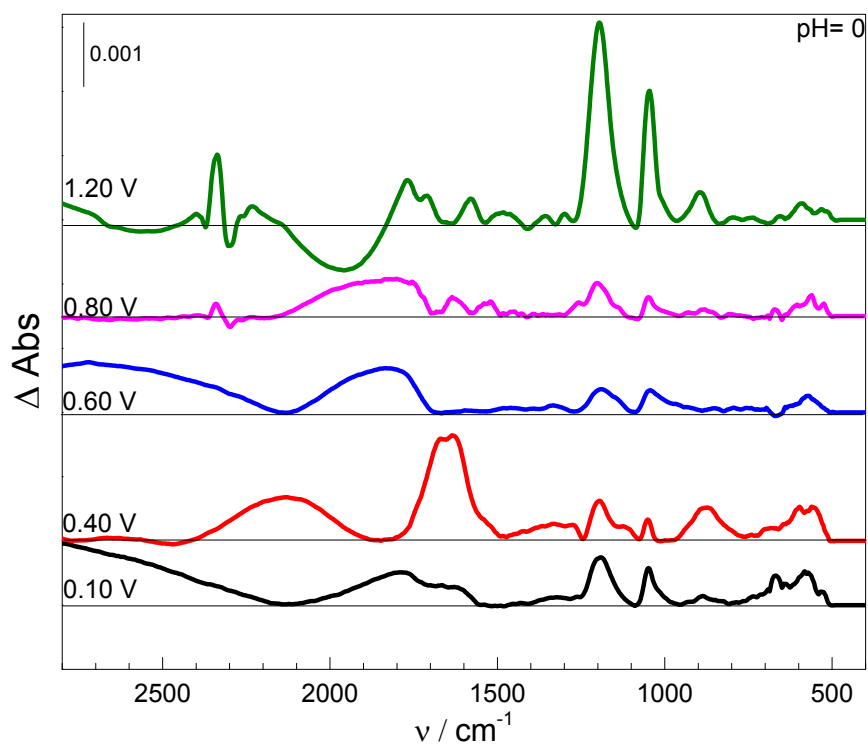
S.F.4. Oxidation spectra of L-Cys 0.10 mol L⁻¹ in Na₂SO₄ 0.50 mol L⁻¹, pH (in the figure) adjusted by H₂SO₄. Applied potentials showed.



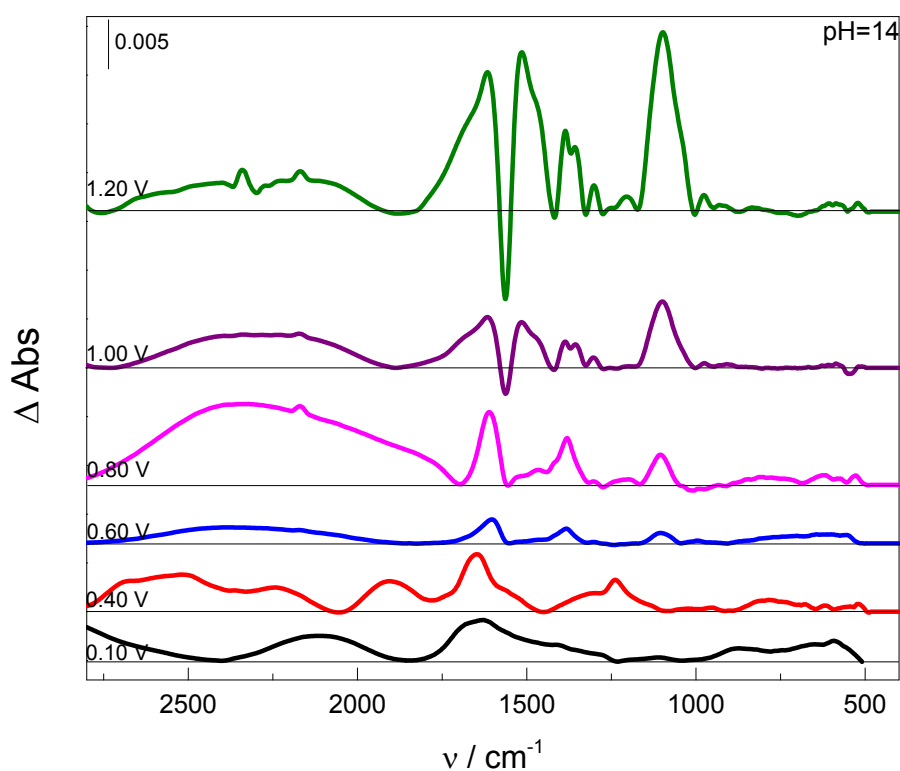
S.F.5. Oxidation spectra of cys 0.10 mol L^{-1} in Na_2SO_4 0.50 mol L^{-1} , pH showed in the figure (adjusted by NaOH addition). Applied potentials also inserted.



S.F.6. Oxidation spectra of cys 0.10 mol L^{-1} in NaOH 0.50 mol L^{-1} , pH showed in the figure. Applied potentials also inserted.

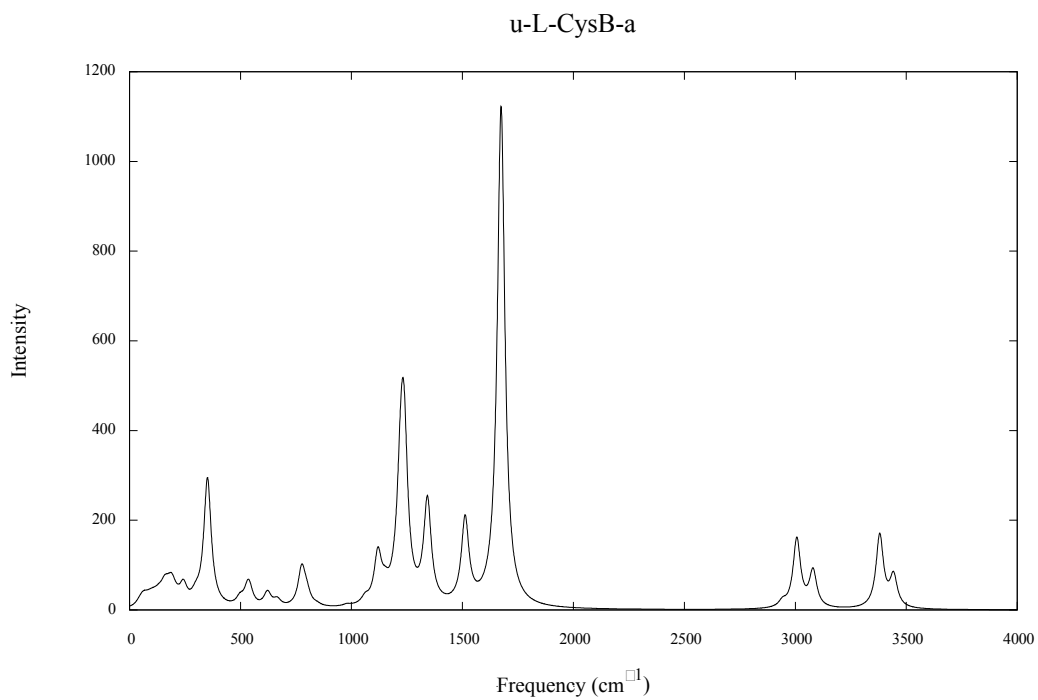


S.F.7. Oxidation spectra of L-Cys₂ sat. in H₂SO₄ 0.50 mol L⁻¹, pH showed in the figure. Applied potentials also inserted.

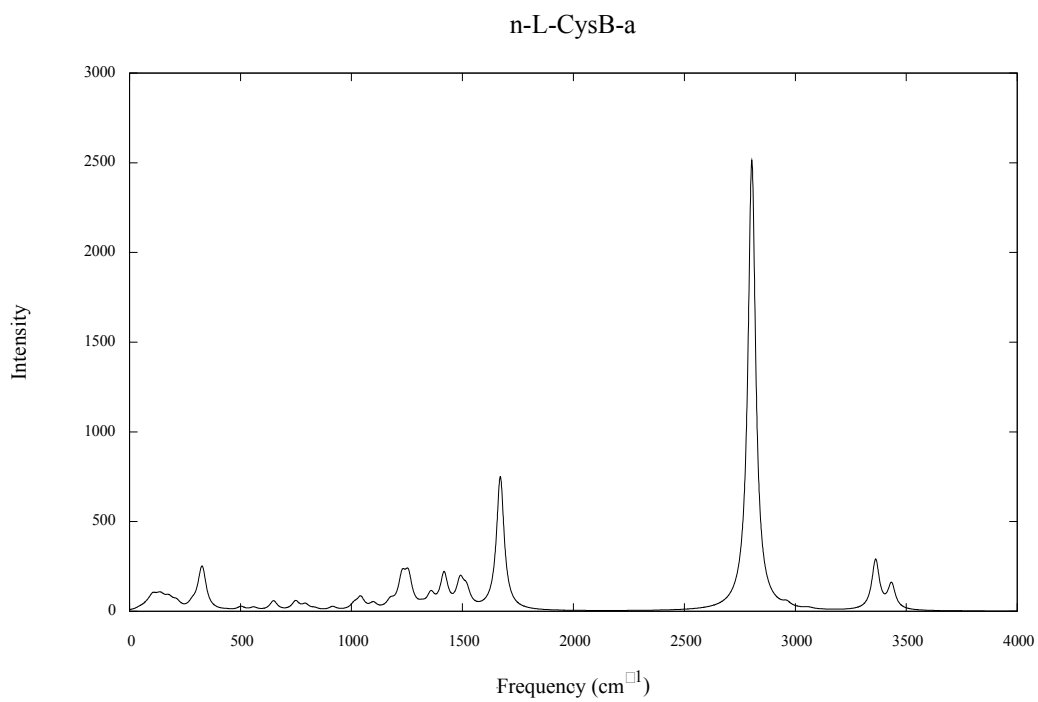


S.F.8. Oxidation spectra of cys₂ sat. in NaOH 0.50 mol L⁻¹, pH showed in the figure. Applied potentials also inserted.

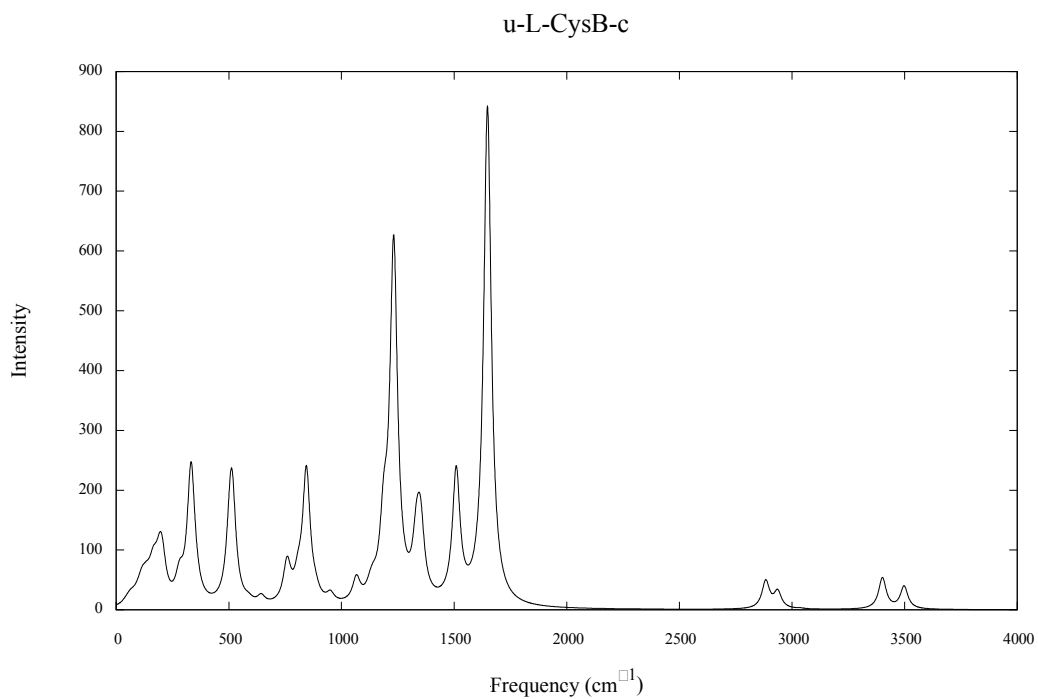
Computational Data Supporting Information



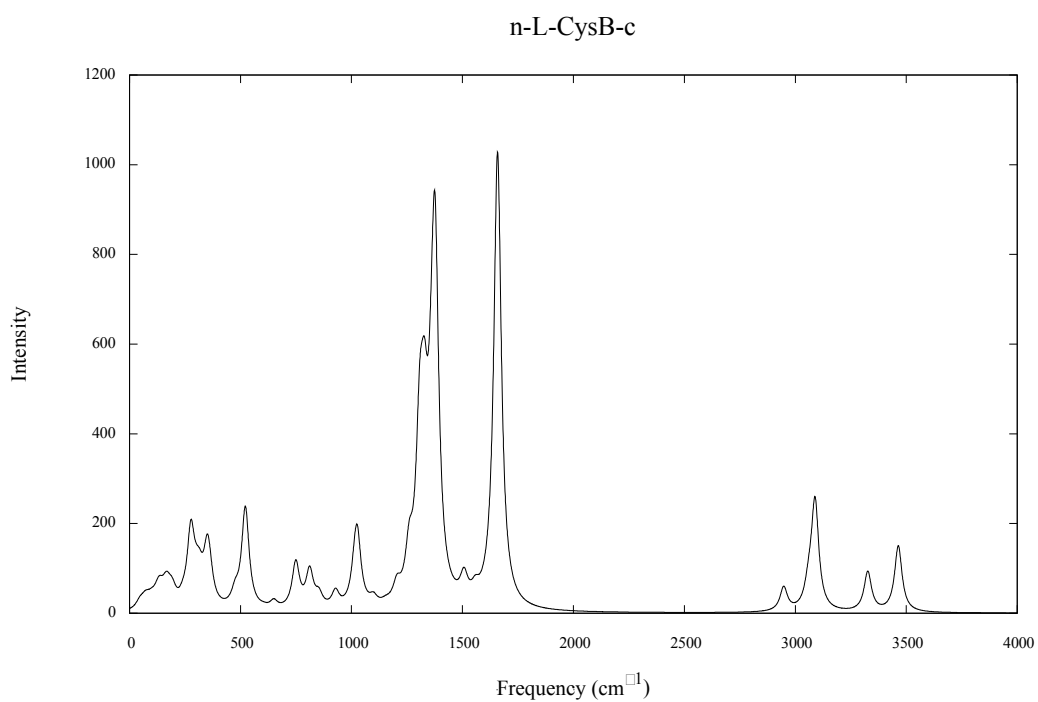
S.F.9. Computed infrared spectrum for u-L-CysB-a.



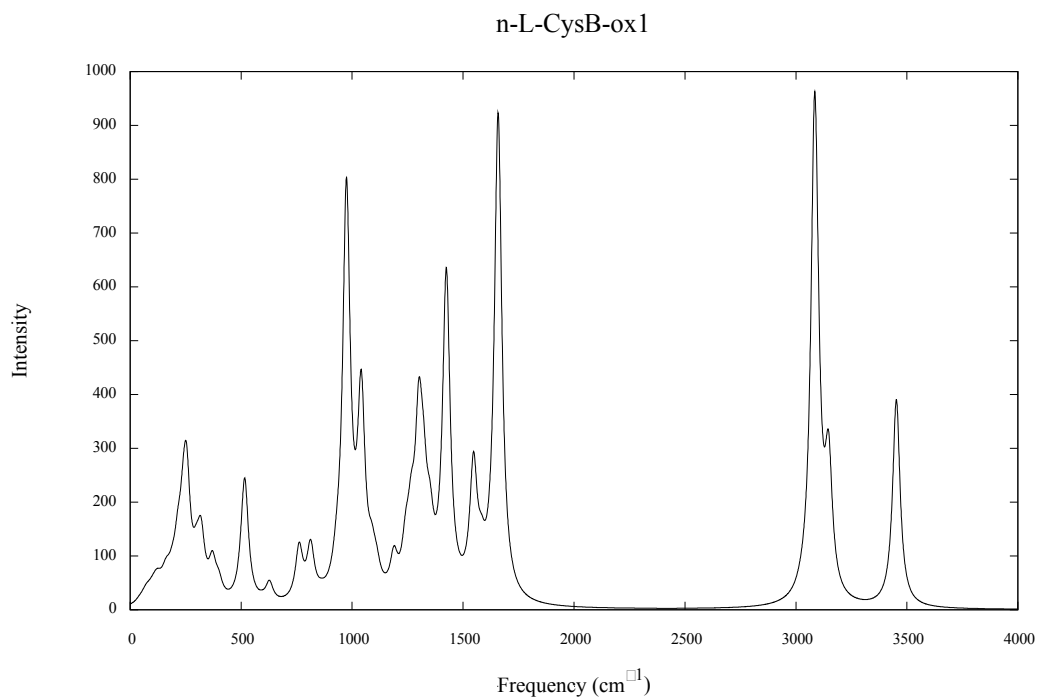
S.F.10. Computed infrared spectrum for n-L-CysB-a.



S.F.11. Computed infrared spectrum for u-L-CysB-c.

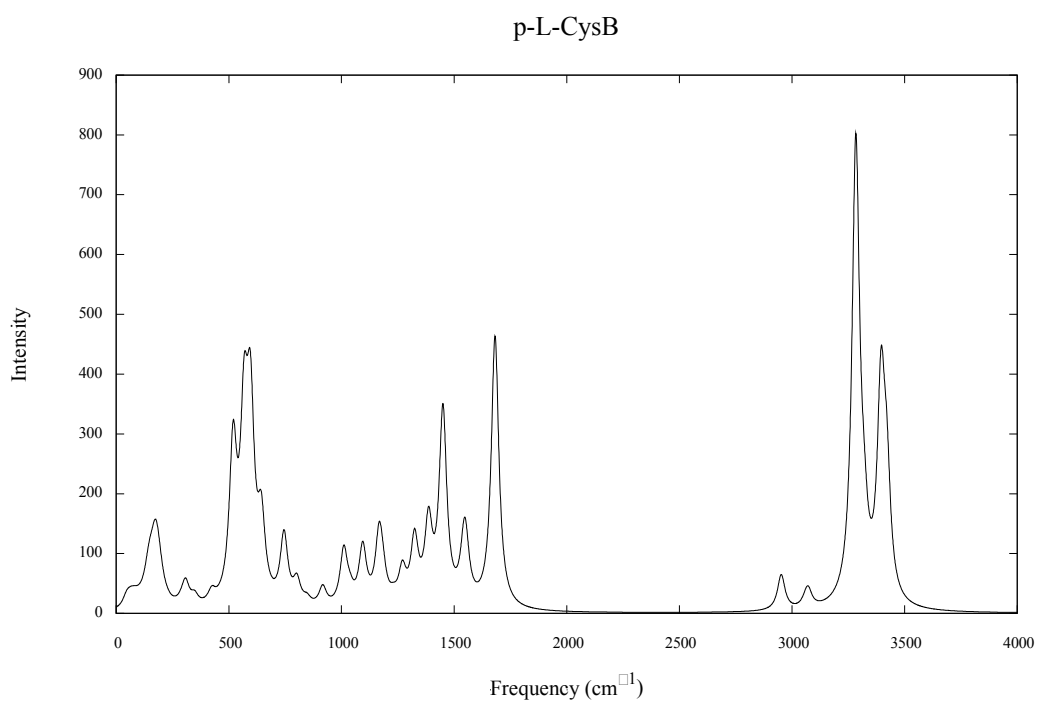


S.F.12. Computed infrared spectrum for n-L-CysB-c.

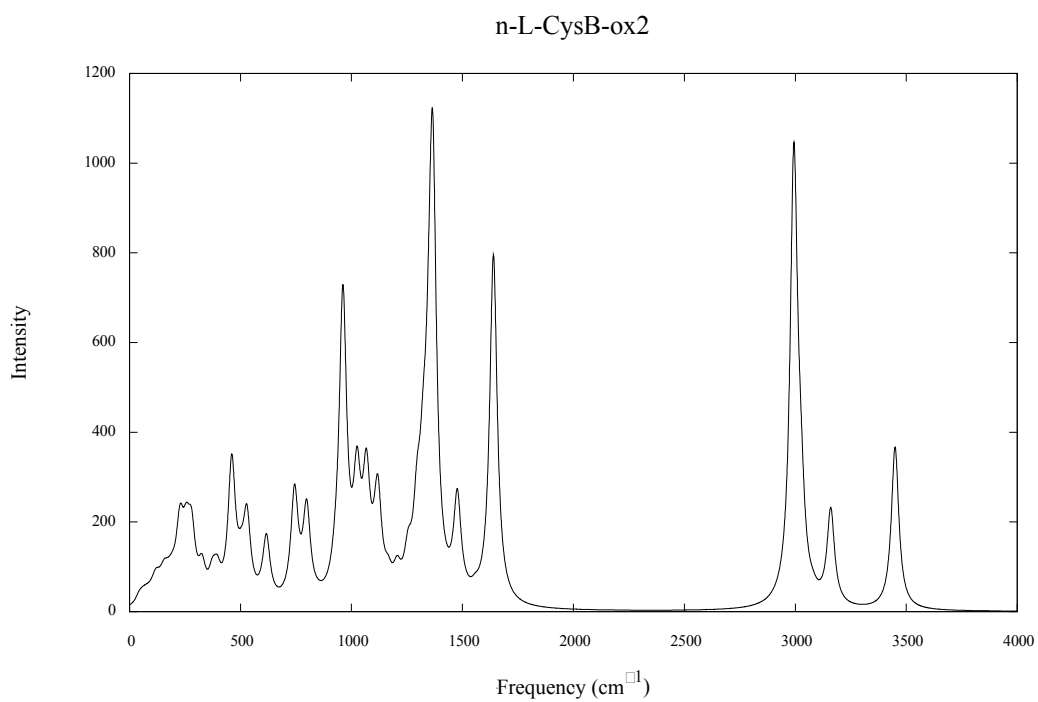


S.F.13

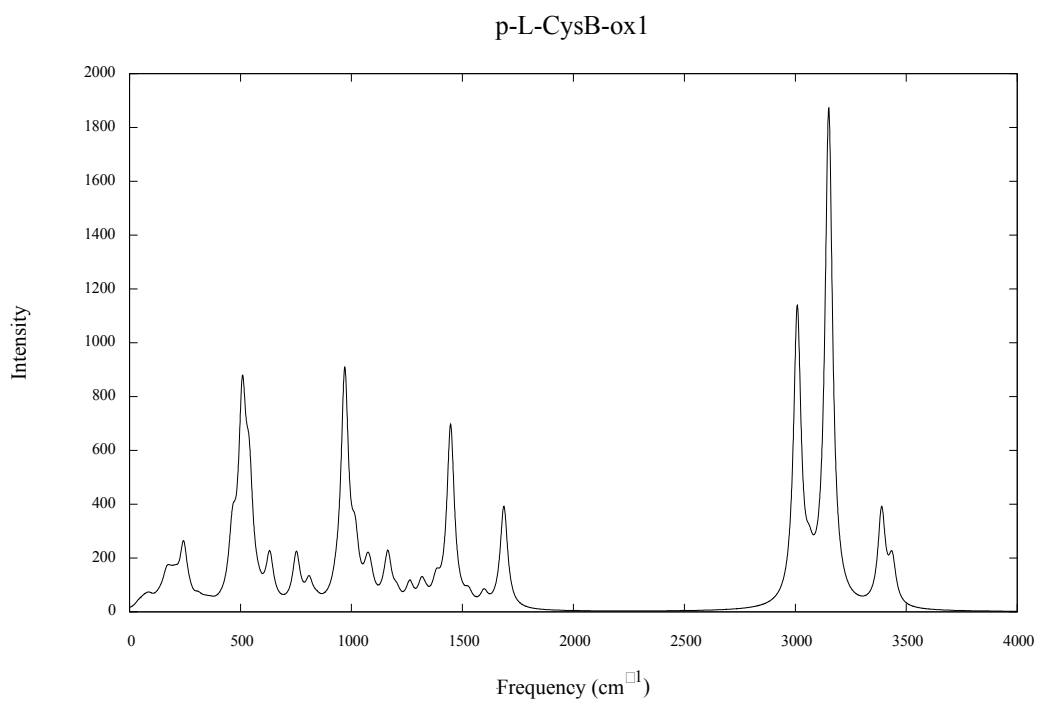
. Computed infrared spectrum for n-L-CysB-ox1.



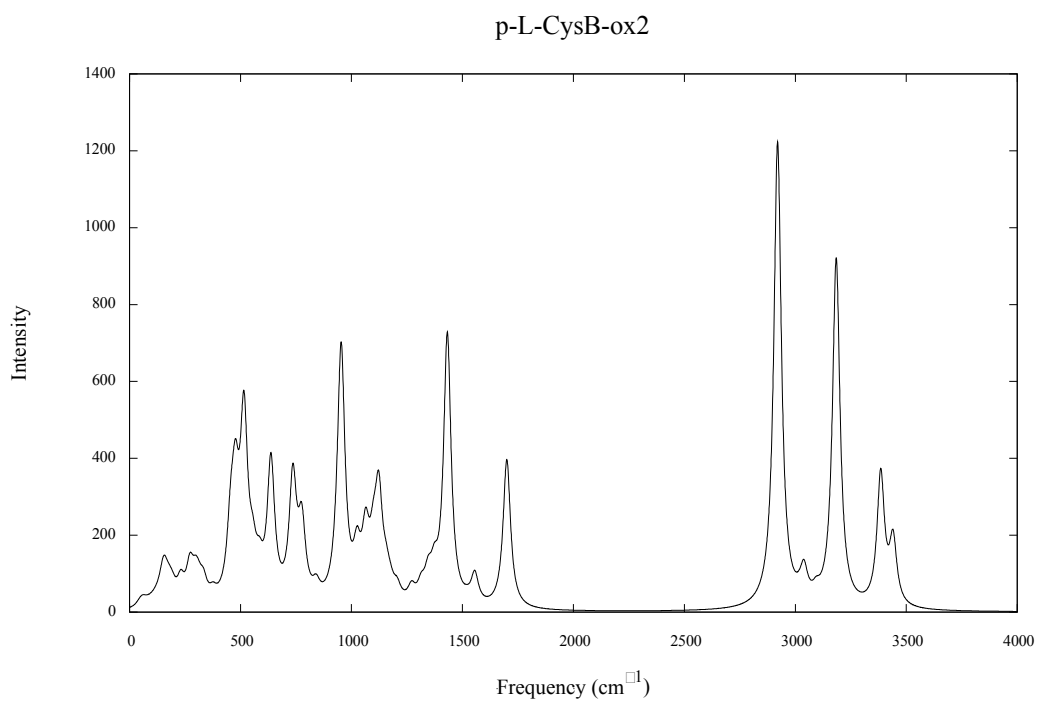
S.F.14. Computed infrared spectrum for p-L-CysB.



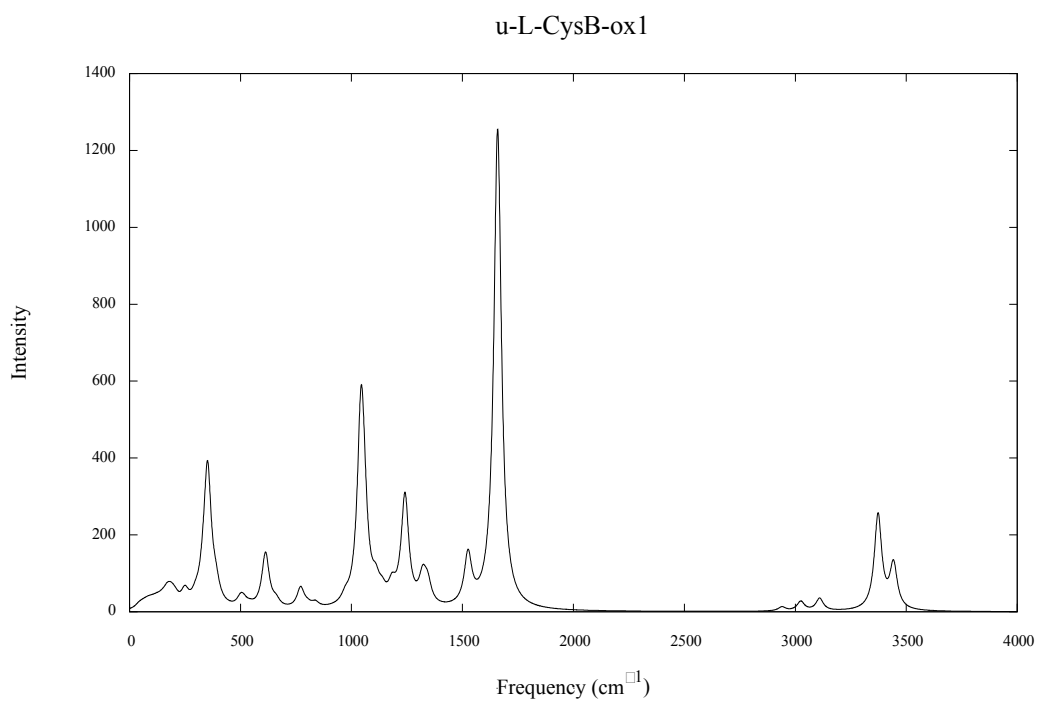
S.F.15. Computed infrared spectrum for n-L-CysB-ox2.



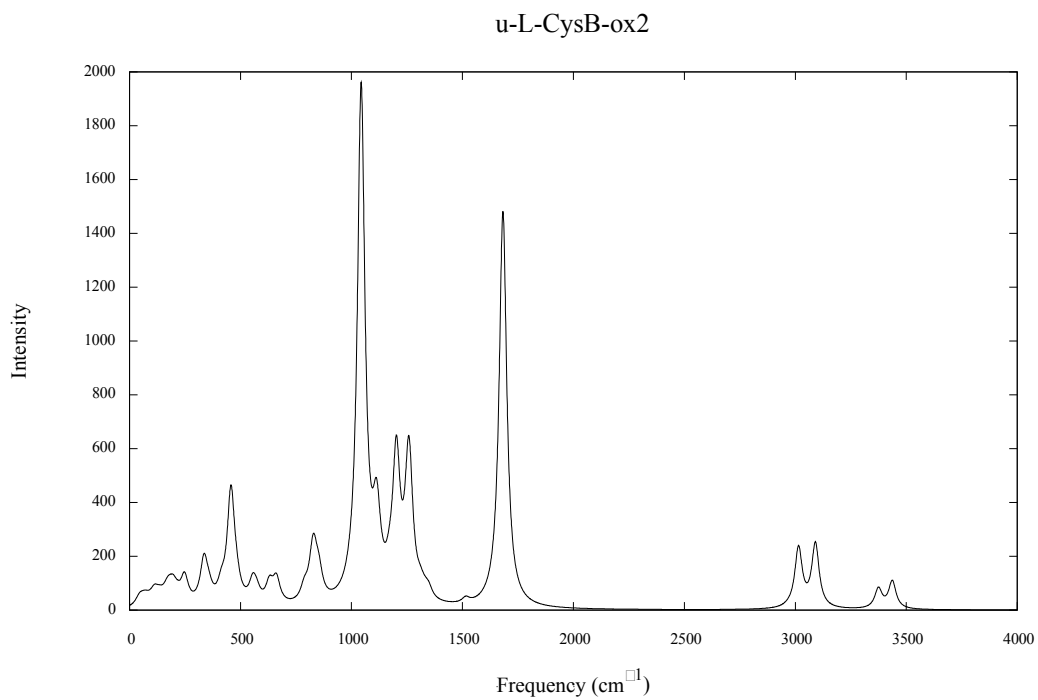
S.F.16. Computed infrared spectrum for p-L-CysB-ox1.



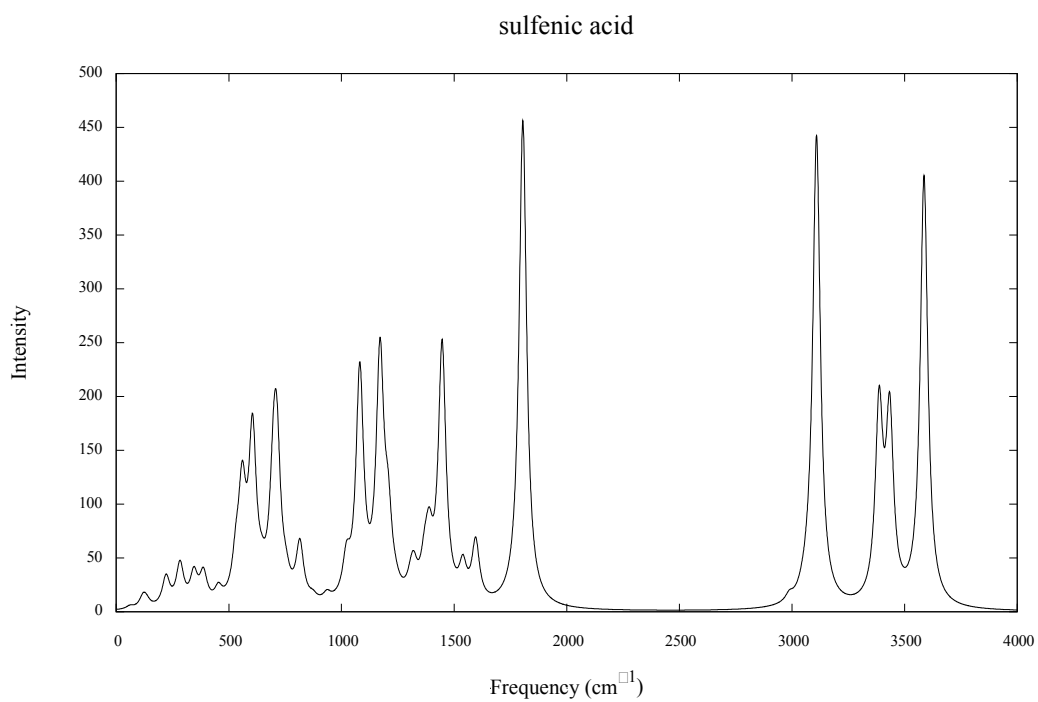
S.F.17. Computed infrared spectrum for p-L-CysB-ox2.



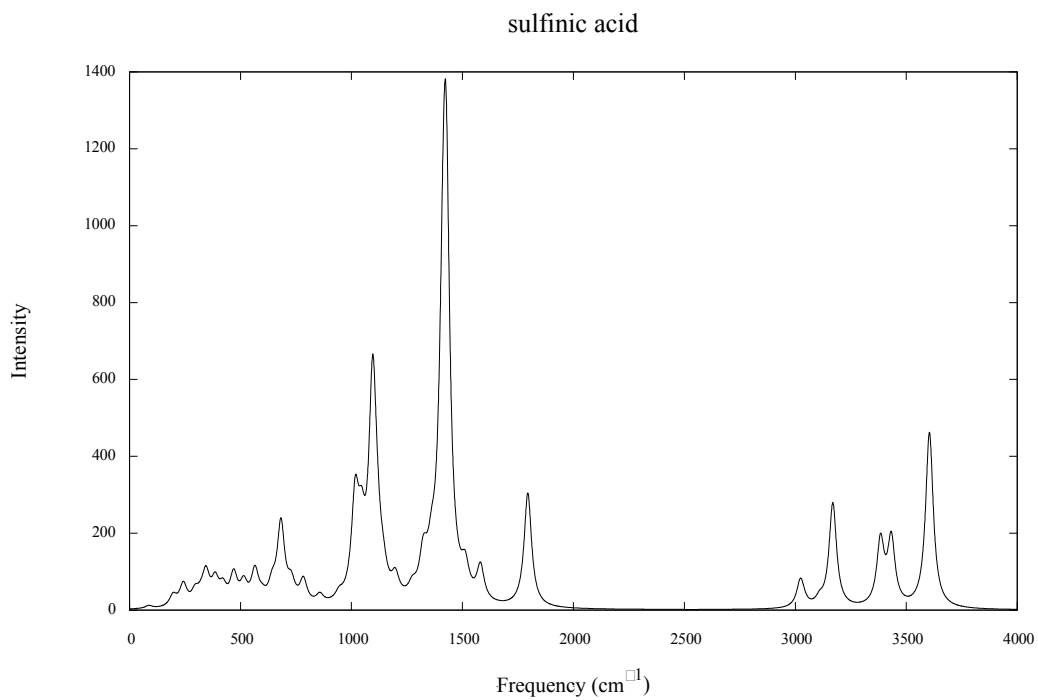
S.F.18. Computed infrared spectrum for u-L-CysB-ox1.



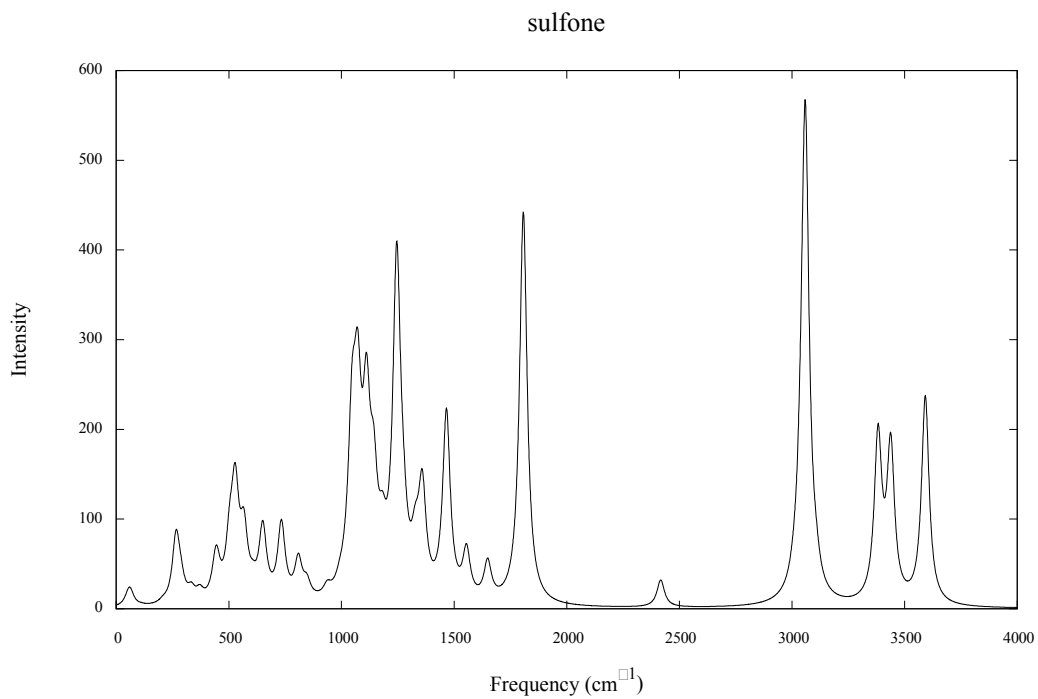
S.F.19. Computed infrared spectrum for u-L-CysB-ox2.



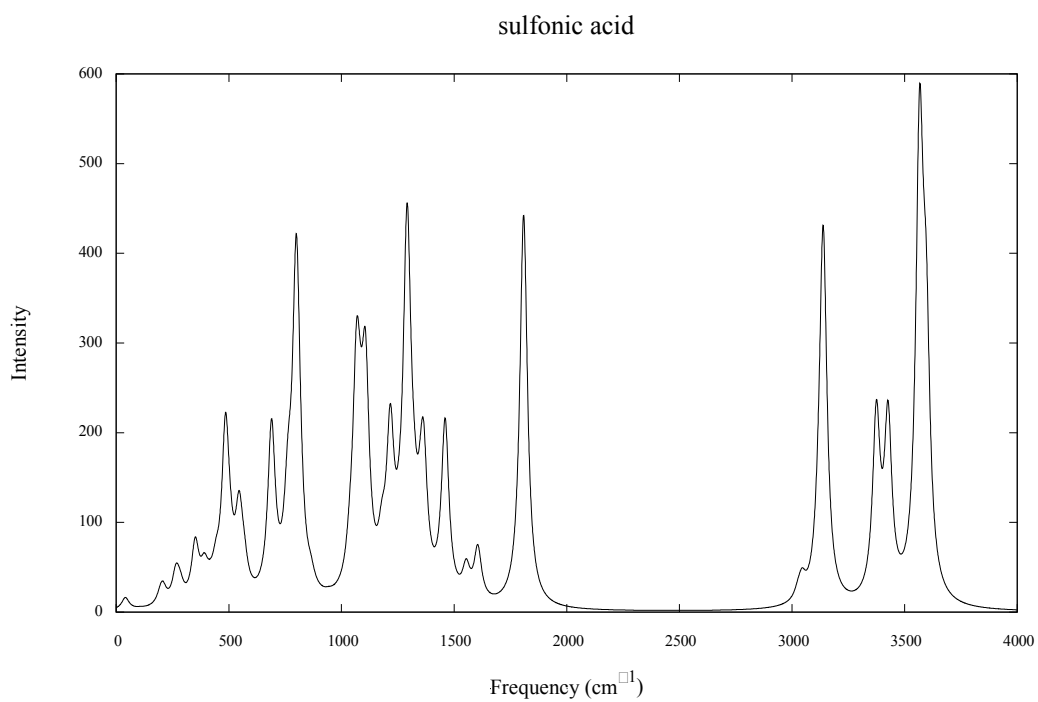
S.F.20. Computed infrared spectrum for sulfenic acid.



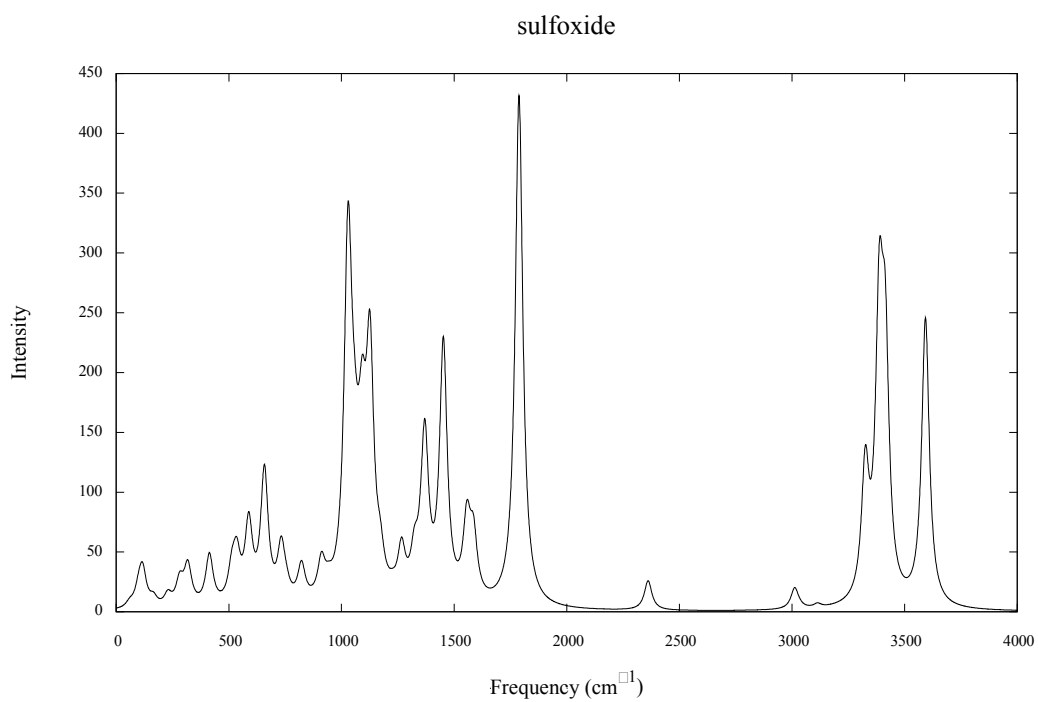
S.F.21. Computed infrared spectrum for sulfinic acid.



S.F.22. Computed infrared spectrum for sulfone.



S.F.23. Computed infrared spectrum for sulfonic acid.



S.F.24. Computed infrared spectrum for sulfoxide.

Pt	2.79472	0.61787	0.77919
Pt	-1.32483	0.53980	-1.67800
Pt	1.41037	0.50157	-1.66080
Pt	-1.34420	3.00193	1.41643
Pt	1.43365	2.96193	1.42104
Pt	0.06694	2.84258	-1.00990
Pt	2.84705	2.75190	-0.99206
Pt	-2.68216	2.88543	-1.00752
O	-2.65558	-3.14694	3.43711
H	-0.17249	-3.60096	2.96224
H	-1.59105	-5.63568	3.44498
H	0.52896	-5.13345	2.28967
C	-0.32153	-4.42571	2.23714
C	-2.85751	-4.10568	2.59445
C	-1.66651	-5.10278	2.47601
O	-3.86553	-4.32576	1.90618
S	-0.13667	-3.66189	0.56263
N	-2.01839	-6.09578	1.41584
H	-3.02786	-5.90620	1.19423
H	-1.45322	-5.93284	0.54432
H	-1.90476	-7.07872	1.70707
O	-0.25198	-4.87040	-0.38263

n-L-CysB-ox2

>>>>>>>>>>>>>>>>>>>> n-L-CysB-ox2.xyz <<<<<<<<<<<<<<<<<<<<

35

temp.out Energy: -3259.5359326

Pt	-2.55713	-1.27900	2.40527
Pt	0.03681	-1.20568	2.52900
Pt	2.64043	-1.20328	2.61841
Pt	-3.98589	-1.43444	0.01022
Pt	-1.32405	-1.80614	0.15517

Pt	1.31369	-1.90558	0.29925
Pt	4.02498	-1.61328	0.26510
Pt	-2.64408	-1.65301	-2.22367
Pt	0.00081	-1.76571	-2.15744
Pt	2.69396	-1.76235	-2.00936
Pt	-2.74891	0.72451	0.76140
Pt	0.02051	0.66800	0.75138
Pt	2.79500	0.62140	0.76330
Pt	-1.36770	0.52960	-1.63899
Pt	1.40900	0.47750	-1.63220
Pt	-1.32220	3.01660	1.39500
Pt	1.45060	2.96430	1.39741
Pt	0.06250	2.81860	-1.00430
Pt	2.83890	2.76730	-0.99750
Pt	-2.71540	2.86850	-1.00070
O	-1.81953	-3.24594	3.25476
H	0.83592	-3.79812	2.86244
H	-0.77064	-5.71635	3.32470
H	1.41399	-5.45306	2.44809
C	0.67284	-4.67004	2.19771
C	-1.94637	-4.20030	2.39735
C	-0.75264	-5.20426	2.34041
O	-2.93229	-4.45526	1.68284
S	1.28725	-4.15171	0.50536
N	-1.09556	-6.21160	1.29664
H	-2.12497	-6.09242	1.13365
H	-0.60859	-5.95348	0.39533
H	-0.87461	-7.18490	1.55441
O	2.69931	-4.66352	0.46751
O	0.33626	-4.80075	-0.49365

p-L-CysB-ox1

S	-0.06440	-3.67591	0.59925
N	-1.58563	-6.33955	1.58846
H	-2.57778	-6.62279	1.50839
H	-1.24094	-6.04389	0.63445
H	-3.88496	-3.51699	1.41711
H	-1.05224	-7.18255	1.86121
O	-0.24913	-4.87746	-0.34499

p-L-CysB-ox2

>>>>>>>>>>>>>>> p-L-CysB-ox2.xyz <<<<<<<<<<<<<<<<<<<<<<

36

temp.out Energy: -3259.9660322

Pt	-2.57701	-1.23650	2.43106
Pt	0.02315	-1.22022	2.52369
Pt	2.63320	-1.25442	2.58329
Pt	-3.99105	-1.46494	0.02635
Pt	-1.30363	-1.78289	0.18048
Pt	1.35016	-1.90066	0.28337
Pt	4.06184	-1.56102	0.19460
Pt	-2.62494	-1.66641	-2.19388
Pt	0.02501	-1.75796	-2.14417
Pt	2.71794	-1.72335	-2.05785
Pt	-2.75490	0.72500	0.74950
Pt	0.01770	0.66740	0.74730
Pt	2.79230	0.61580	0.76580
Pt	-1.36440	0.53680	-1.64250
Pt	1.41210	0.48400	-1.63160
Pt	-1.32550	3.01170	1.40150
Pt	1.44820	2.95760	1.41050
Pt	0.06610	2.82480	-0.99570
Pt	2.84140	2.77080	-0.98320
Pt	-2.71000	2.87930	-1.00040

Pt	0.05839	-1.53941	-2.65424
Pt	2.78426	-1.56942	-2.41403
Pt	-2.75490	0.72500	0.74960
Pt	0.01760	0.66730	0.74720
Pt	2.79230	0.61580	0.76580
Pt	-1.36430	0.53680	-1.64250
Pt	1.41210	0.48410	-1.63160
Pt	-1.32540	3.01180	1.40150
Pt	1.44820	2.95760	1.41050
Pt	0.06600	2.82480	-0.99570
Pt	2.84150	2.77080	-0.98320
Pt	-2.71000	2.87930	-1.00040
O	-2.48474	-5.82009	-1.74371
H	0.17138	-6.03782	-1.25353
H	-0.60053	-5.52391	-3.37057
H	1.41604	-4.96955	-2.00597
C	0.38059	-5.01169	-1.61281
C	-2.12726	-4.73147	-2.19446
C	-0.64728	-4.65081	-2.68114
O	-2.95041	-3.73440	-2.39760
S	0.46196	-4.09759	-0.00103
N	-0.28630	-3.47074	-3.50739
H	0.58294	-3.68676	-4.02114
H	-1.00664	-3.35243	-4.23643
O	-1.03865	-4.04333	0.45395
O	1.30307	-5.00977	0.85066