## Electronic Supplementary Information

## Self-powered Fluorescence Display Devices Based on a Fast Self-charging/recharging Battery

## (Mg/Prussian Blue)

Hui Zhang<sup>a, b</sup>, You Yu<sup>a, b</sup>, Lingling Zhang<sup>a, b</sup>, Yiwen Zhai<sup>a, b</sup>, and Shaojun Dong<sup>a, b\*</sup>

<sup>a</sup> State Key Laboratory of Electroanalytical Chemistry, Changchun Institute of Applied Chemistry, Chinese Academy of Science, Changchun, Jilin, 130022 (PR China)

<sup>b</sup> University of Chinese Academy of Sciences, Beijing, 100049 (PR China)

\*Corresponding Author. Tel: +86 431 85692886. Fax: +86 431 85689711.

E-mail: dongsj@ciac.ac.cn.



Fig. S1. The SEM of the PB prepared film.



Fig. S2. Absorption (700 nm) response of the PB film versus time during consecutive switch cycles (0 V/0.4 V, 10 s) in 1 M KCl and 0.1 M phosphate buffer (pH 6) electrolyte solution.

When the PB film was subjected to repeated cyclic electroconversion, the absorbance change at 700 nm was monitored by absorption spectroscopy in situ. It is demonstrated that only 10 s was required to restore over 90% of the maximum absorbance variation. Moreover, the absorbance of PB was perfectly maintained with no sign of degradation after 50 test cycles, which indicated that this PB film was highly durable.



Fig. S3 In situ absorbance measurement of the Mg/PB battery connected at 50 s and then disconnected at 60 s. The red curve referred to the electrolyte containing 0.025 M  $(NH_4)_2S_2O_8$ , 1 M KCl and 0.1 M phosphate buffer (pH 6), and the black curve referred to the electrolyte containing 0.025 M  $H_2O_2$ , 1 M KCl and 0.1 M phosphate buffer (pH 6).



Fig. S4. The sheme of the Mg/PB battery.



Fig. S5. TEM of the  $Ru@SiO_2$  nanoparticle.



Fig. S6. Absorption spectra of  $Ru@SiO_2$  aqueous solution (0.05 mg / mL).



Fig. S7. Excitation (black dash curve) and emission (red solid curve) spectra of  $Ru@SiO_2$  aqueous solution (0.1 mg / mL).



Fig. S8. The SEM of the section of the  $PB/Ru@SiO_2/Nafion film$ .

			Anode	Cathode	Power
	Bleaching <sup>#</sup>	Colouration <sup>#</sup>	Potential /	Potential /	Density /
	_		V	V	mW cm <sup>-2</sup>
With	10 s, 95.1	6.2 min 00.0/	1.82	0.68	13.34
NaClO	%	0.2 mm, 90 %			
Without	10 s, 97.8	120h 0270/	1 79	0.27	2
NaClO	%	42.8 II, 82.7 %	1.70	0.37	3

Table S1 Difference of battery behavior between the electrolyte with and without NaClO.

<sup>#</sup> refer to the time required to achieve the given absorbance variation.

Cycles	Connecting Time / s	Disconnecting Time / s	Fluorescence Contrast / %
1	15	290	63
2	15	350	63
3	15	390	54
4	15	395	52
5	15	490	52
6	15	465	51
7	15	525	47
8	15	710	45
9	15	600	44
10	15	660	39
11	15	805	39

Table S2 Detailed information about Fig. 5.