Electronic Supporting Information

New Promising Lithium Malonatoborate Salts for Lithium Ion Battery Application

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Fig. S1. Cyclic voltammogramms (CVs) of $LiNi_{0.5}Mn_{1.5}O_4$ in (a) 1.0M LiDFMFMB, (b) 1.0M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte at a scan rate of 0.05 mV/s. Lithium is used as both counter and reference electrode.



Fig. S2. Electrochemical impedance spectra of $LiNi_{0.5}Mn_{1.5}O_4$ in (a) 1.0 M LiDFMFMB, (b) 1.0 M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte at different CV cycles at a scan rate of 0.05 mV/s.



Fig. S3. Cyclic voltammogramms (CVs) of graphite in (a) 1.0M LiDFMFMB, (b) 1.0M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte at a scan rate of 0.05 mV/s. Lithium is used as both counter and reference electrode.



Fig. S4. Electrochemical impedance spectra of graphite in (a) 1.0M LiDFMFMB, (b) 1.0M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte at different CV cycles at a scan rate of 0.05 mV/s.



Fig. S5. Electrochemical impedance spectra of the $LiNi_{0.5}Mn_{1.5}O_4||Li$ half-cell before and after cycling at a current rate of C/2 in (a) 1.0 M LiDFMFMB, (b) 1.0 M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte.



Fig. S6. Electrochemical impedance spectra of the NG||Li half-cell before and after cycling at a current rate of C/5 in (a) 1.0 M LiDFMFMB, (b) 1.0 M LiDFEFMB, and (c) 1.0 M LiDFPFMB electrolyte.



(a) NG-Ref



(b) LNMO-Ref



(c) NG-LiDFMFMB



(d) LNMO-LiDFMFMB

(f) LNMO-LiDFEFMB



(e) NG-LiDFEFMB



(h) LNMO-LiDFPFMB



(g) NG-LiDFPFMB



Fig. S7. SEM images with EDX of pristine graphite (a) and LNMO (b) electrode and cycled graphite electrodes (c, e and g) and LNMO (d, f and h) from the LNMO||NG full cells in 1.0 M LiDFMFMB (c and d), LiDFEFMB (e and f), and LiDFPFMB (g and h), respectively.