

Supporting Information

On the Design of Atropisomer-Separable Photochromic Diarylethene-based Metal-organic Framework Linkers

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Table of Contents

Table of Open-form Electronic Transitions	S2
Table of Closed-form Electronic Transitions	S3
Open-form Orbitals	S4
Closed-form Orbitals	S6
Open/Closed-form Energy Difference	S9
Reverse-Direction Rotational PES	S10

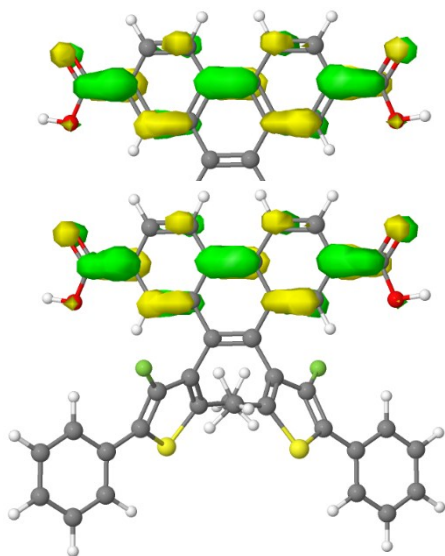
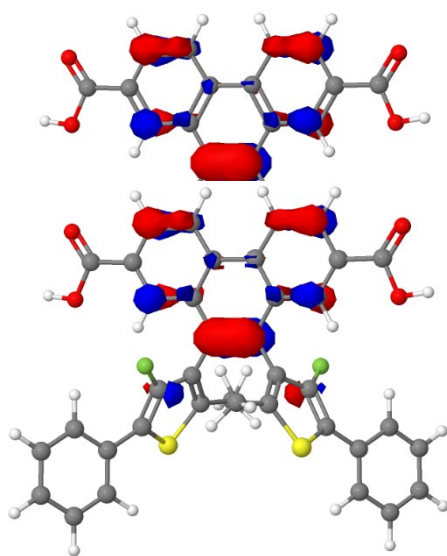
Table S1. Distances of potentially photoactive carbon atoms in each molecule, the excitation energy and oscillator strength of the lowest energy electronic transition, as well as the individual electronic transition components, and their associated CI expansion coefficients, which comprise the lowest energy transition in the open form of each molecule. Highlighted transitions are the major contributor to the excitation.

Molecule	Photoactive C-C Distance (Å)	Excitation Energy (nm)	Oscillator Strength	Electronic Transitions	CI Expansion Coefficients
1H	3.679	312.85	0.0565	HOMO-4 → LUMO	-0.12231
				HOMO-3 → LUMO+1	0.25866
				HOMO-2 → LUMO	0.55967
				HOMO → LUMO	-0.27451
1F	3.645	308.89	0.0469	HOMO-4 → LUMO	-0.13382
				HOMO-3 → LUMO+1	0.28594
				HOMO-2 → LUMO	0.56146
				HOMO → LUMO	-0.23242
1Cl	3.626	308.86	0.0440	HOMO-4 → LUMO	-0.11654
				HOMO-3 → LUMO+1	0.28683
				HOMO-1 → LUMO	0.46981
				HOMO → LUMO	0.38747
1Br	3.646	308.72	0.0432	HOMO-4 → LUMO	-0.11365
				HOMO-3 → LUMO+1	0.28667
				HOMO-2 → LUMO	0.47514
				HOMO → LUMO	-0.38064
1I	3.589	310.13	0.0423	HOMO-4 → LUMO	-0.11458
				HOMO-3 → LUMO+1	-0.28142
				HOMO-1 → LUMO	0.39880
				HOMO → LUMO	0.45890
1Me	3.581	314.64	0.0552	HOMO-4 → LUMO	-0.17467
				HOMO-3 → LUMO+1	0.24287
				HOMO-2 → LUMO	0.50810
				HOMO → LUMO	-0.34817
1OMe	3.675	318.30	0.0705	HOMO-5 → LUMO+1	-0.18896
				HOMO-4 → LUMO+1	0.15634
				HOMO-3 → LUMO	-0.25309
				HOMO-2 → LUMO	0.51021
				HOMO → LUMO	0.30074
2	3.711	307.95	0.0463	HOMO-3 → LUMO+1	0.25272
				HOMO-3 → LUMO+3	-0.12648
				HOMO-1 → LUMO	0.53361
				HOMO → LUMO	0.30852
3	3.685	310.42	0.0468	HOMO-5 → LUMO+1	0.27407
				HOMO-3 → LUMO	0.15461
				HOMO-2 → LUMO	0.52405
				HOMO → LUMO	-0.30915

Table S2. Excitation energies and oscillator strength of lowest energy transition in the closed-form of each system. Also shown are the individual electronic transition components and associated CI expansion coefficients which comprise the lowest energy transition. Highlighted transitions are the major contributor to the excitation.

Molecule	Excitation Energy (nm)	Oscillator Strength	Electronic Transitions	CI Expansion Coefficients
1H	625.5	0.2506	HOMO → LUMO	0.68818
1F	654.8	0.255	HOMO → LUMO	0.69216
1Cl	625.36	0.1939	HOMO → LUMO	0.69121
1Br	629.99	0.1965	HOMO → LUMO	0.69104
1I	620.42	0.183	HOMO → LUMO	0.69038
1Me	641.62	0.1903	HOMO → LUMO	0.68641
			HOMO → LUMO + 1	-0.11031
1OMe	638.45	0.2061	HOMO → LUMO	0.68808
			HOMO → LUMO + 1	-0.10008
2	490.01	0.2948	HOMO → LUMO	0.68485
3	565.88	0.1768	HOMO → LUMO	0.68756
			HOMO → LUMO + 1	-0.10849

1H



S3

1F

1Cl

1Br

1I

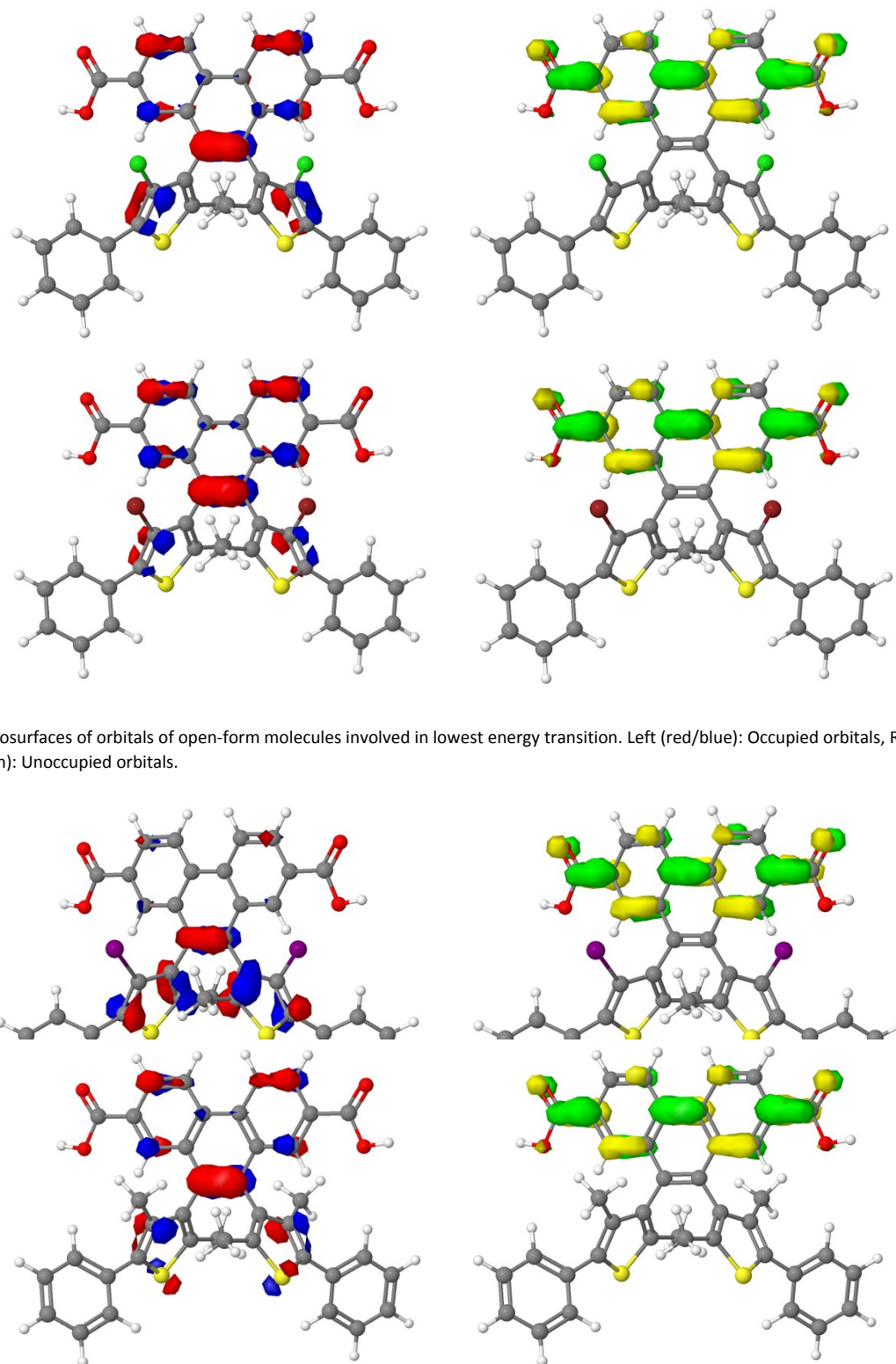


Figure S1a. Isosurfaces of orbitals of open-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

1Me

10Me

2

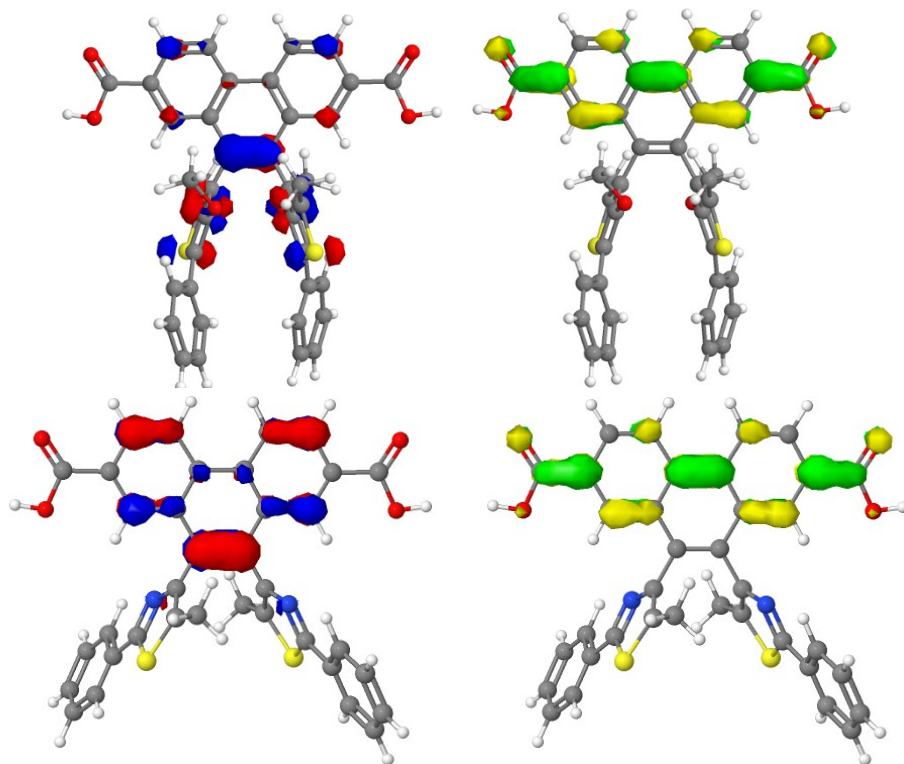


Figure S1b. Isosurfaces of orbitals of open-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

3

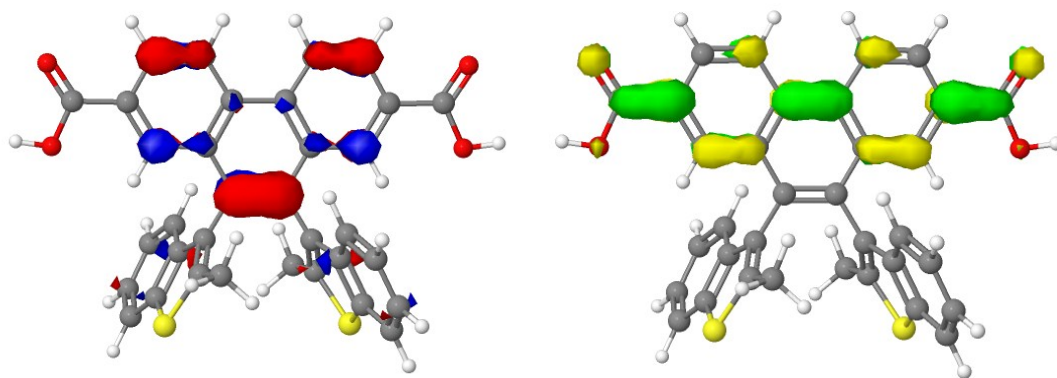
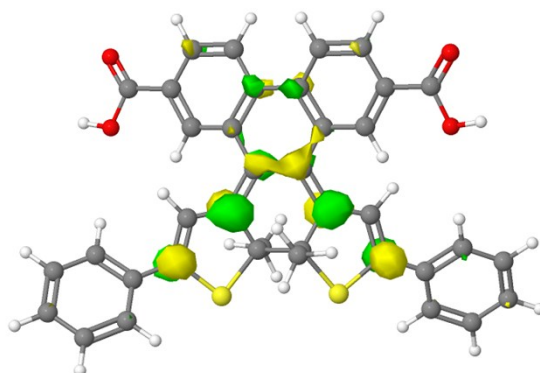
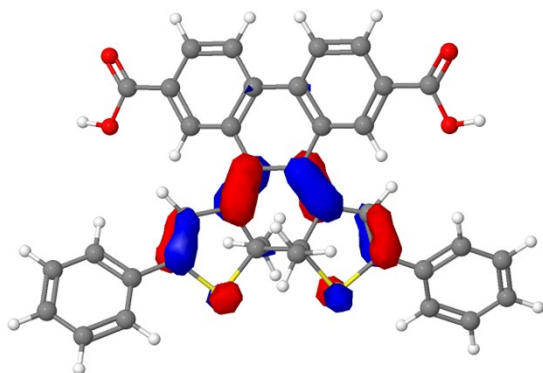


Figure S1c. Isosurfaces of orbitals of open-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

1H



1F

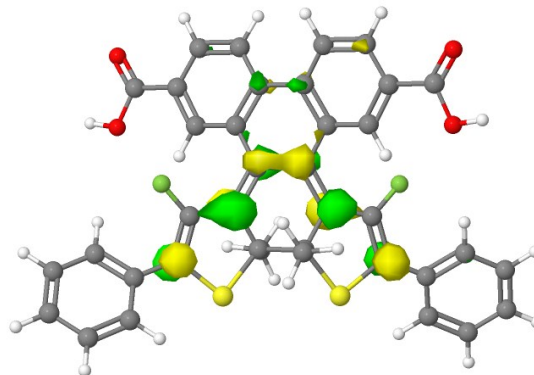
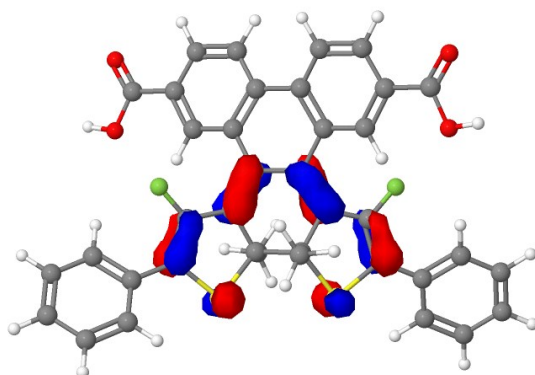


Figure S2a. Isosurfaces of orbitals of closed-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

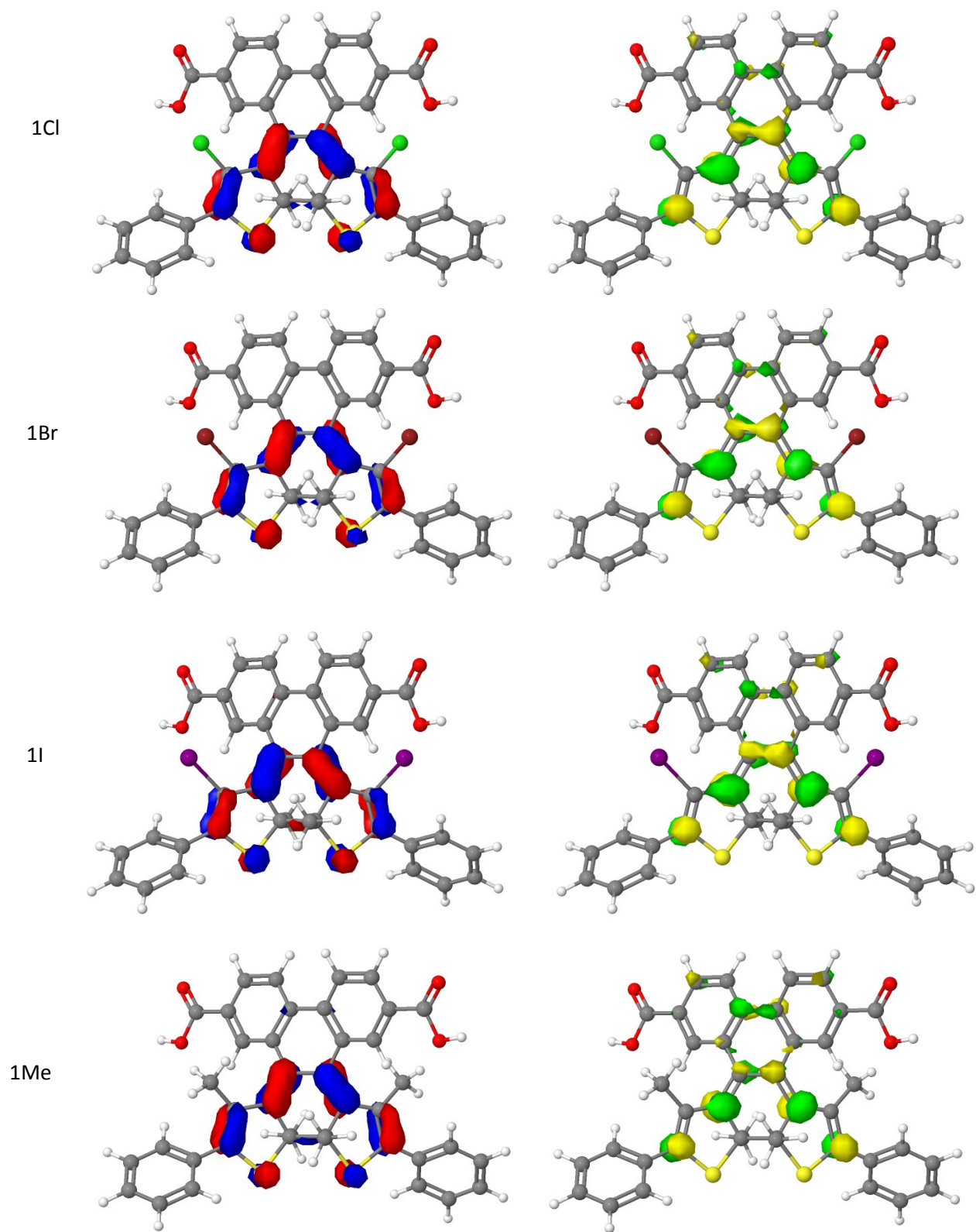


Figure S2b. Isosurfaces of orbitals of closed-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

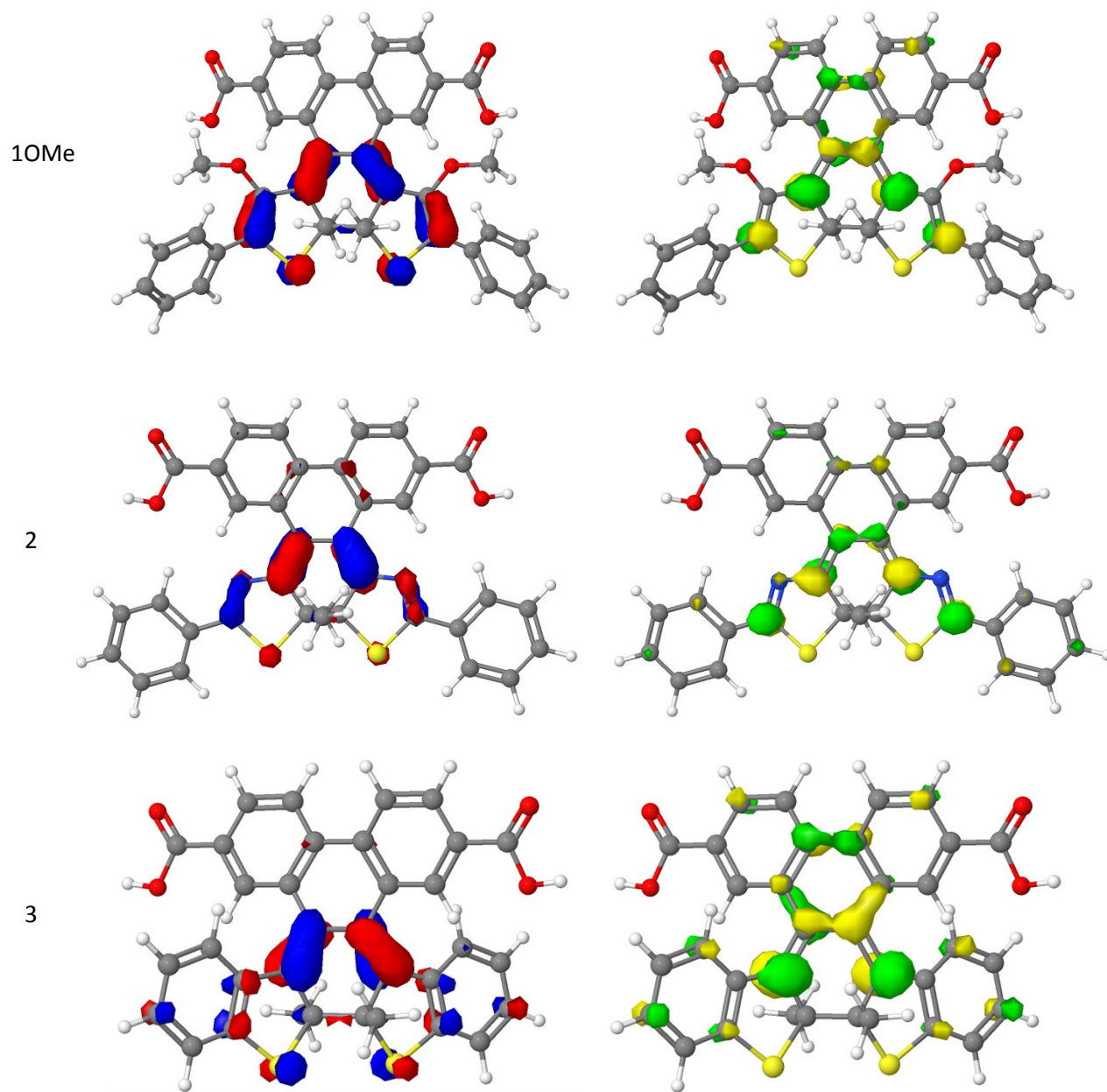


Figure S2c. Isosurfaces of orbitals of closed-form molecules involved in lowest energy transition. Left (red/blue): Occupied orbitals, Right (yellow/green): Unoccupied orbitals.

Table S3. The calculated energies of open and closed forms of all systems, and the energy difference between the two forms.

Molecule	Open Energy (Hartree)	Closed Energy (Hartree)	ΔE (Hartree)	ΔE (kcal/mol)
1H	-2560.45440742	-2560.40847009	0.04593733	28.83
1F	-2758.86443943	-2758.80664197	0.05779746	36.27
1Cl	-3479.61113779	-3479.54496100	0.06617679	41.53
1Br	-7702.72440358	-7702.65340612	0.07099746	44.55
1I	-2581.99449640	-2581.92637963	0.06811677	42.74
1Me	-2639.07408409	-2639.01367516	0.06040893	37.91
1OMe	-2789.42339159	-2789.36339165	0.05999994	37.65
2	-2592.54155117	-2592.51719905	0.02435212	15.28
3	-2405.69613621	-2405.63964215	0.05649406	35.45

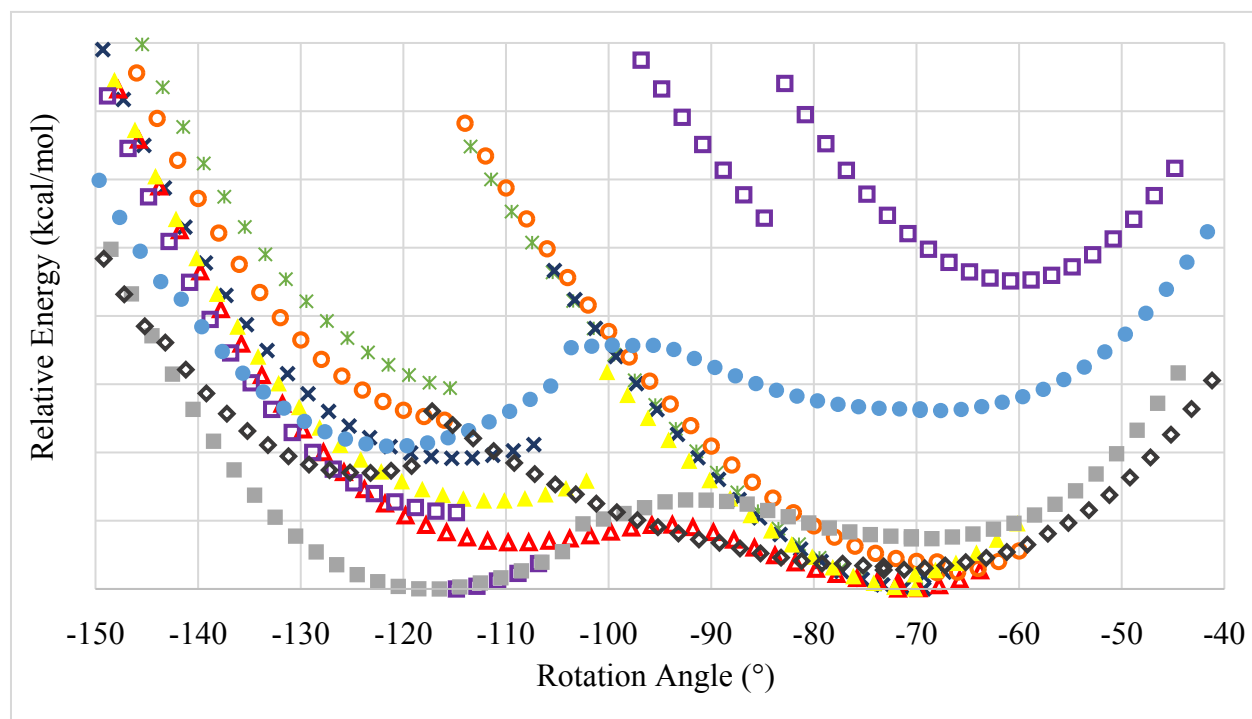
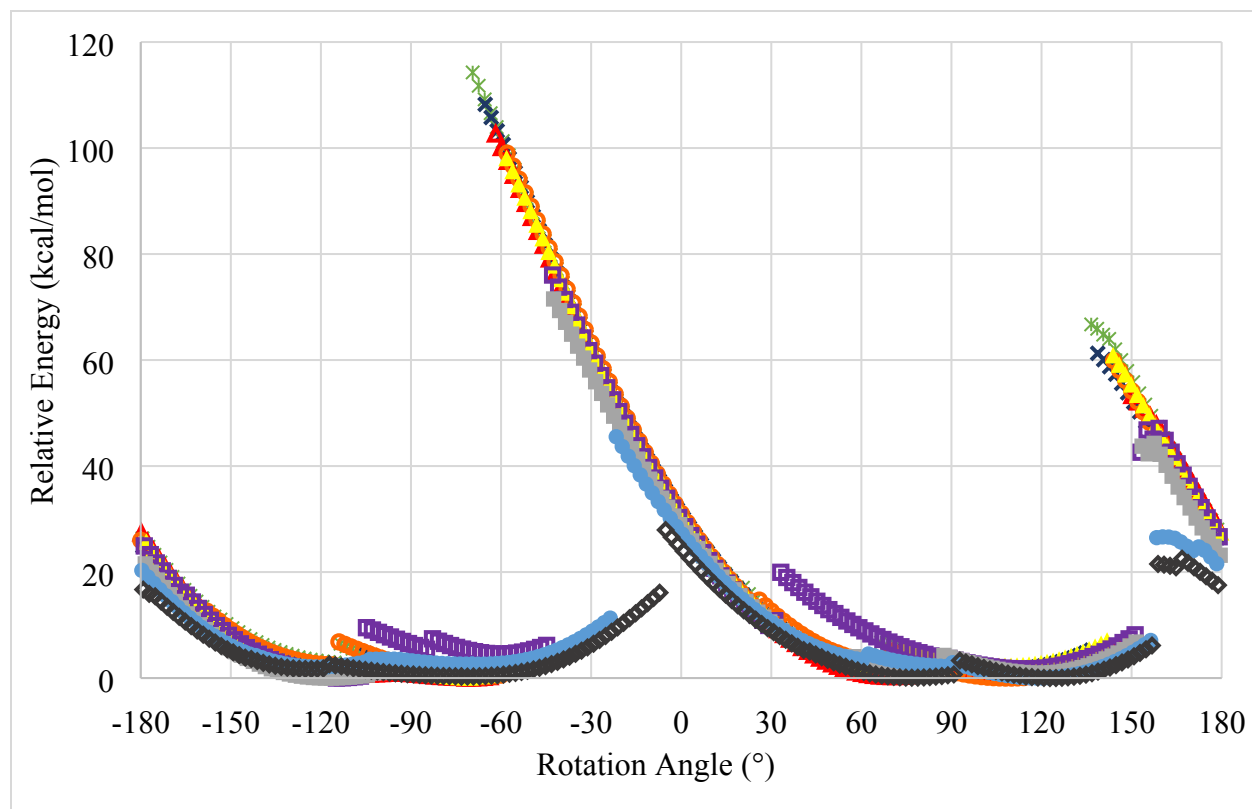


Figure S3a. Reverse-direction rotational potential energy surfaces of all molecules with expanded views of the energy wells. Light-blue circles: **1H**, Light-grey squares: **1F**, Yellow triangles: **1Cl**, Dark-blue X's: **1Br**, Green stars: **1I**, Orange open circles: **1Me**, Violet open squares: **1OMe**, Dark-grey open diamonds: **2**, Red open triangles: **3**.

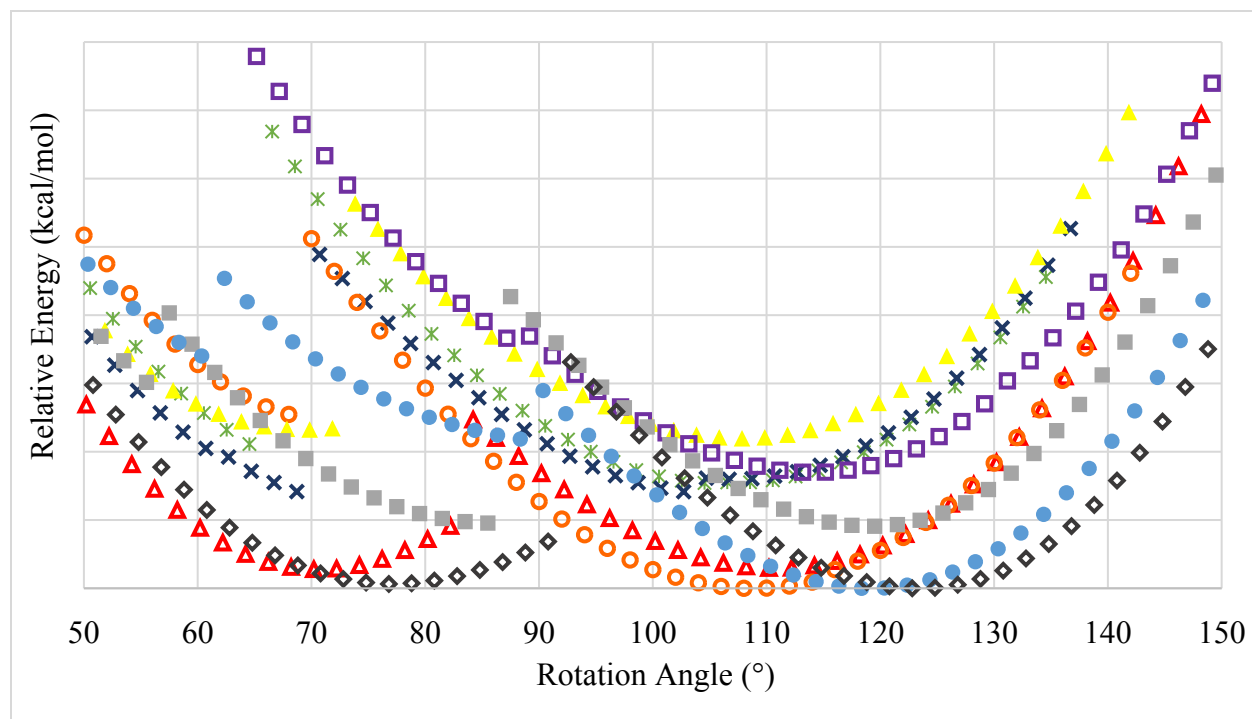


Figure S3b. Reverse-direction rotational potential energy surfaces of all molecules with expanded views of the energy wells. Light-blue circles: **1H**, Light-grey squares: **1F**, Yellow triangles: **1Cl**, Dark-blue X's: **1Br**, Green stars: **1I**, Orange open circles: **1Me**, Violet open squares: **1OMe**, Dark-grey open diamonds: **2**, Red open triangles: **3**.