# Electronic Supplementary Information 

## Are scaling relations truly universal?

Vladimir Tripkovic ${ }^{1}$<br>${ }^{1}$ Department of Energy Conversion and Storage, Technical University of Denmark, DK-2800 Kgs.<br>Lyngby, Denmark.

## Note 1: Differential adsorption energies for the $\mathbf{O H}$ and $\mathbf{O O H}$ intermediates

Only for $\operatorname{Pt}(111)$ and $\operatorname{Ir}(111)$, the lowest differential OH adsorption energy is at the $1 / 3 \mathrm{ML} \mathrm{OH}$ coverage, i.e. in a half dissociated water layer. ${ }^{30,32,33} \mathrm{On}$ all the other metals and alloys the OH is most stable in water bilayer at the $1 / 9 \mathrm{ML}$ OH coverage. As for OOH , the highest binding energy is obtained at $1 / 6 \mathrm{ML}$ coverage irrespective of the catalyst surface. For the strained $\operatorname{Pt}(111), \operatorname{Pd}(111)$ and $\mathrm{Ag}(111)$ we assume the same $\mathrm{OH} / \mathrm{OOH}$ coverages as those on pristine surfaces.


Figure S1 Differential adsorption energies of a) OH and b ) OOH on the (111) surfaces of late transition metals. Snapshots show structures at different $\mathbf{O H} / \mathbf{O O H}$ coverages in the water bilayer.


Figure S2 Differential adsorption energies of $\mathbf{O H}$ on the (111) surfaces of selected Pt-alloys. For the OOH the $\mathbf{1 / 6}$ ML coverage found on metals was assumed.

Note 2: Calculated values of lattice constants and binding energies of different intermediates
Table S1 The RPBE optimized lattice constants.

| Element | Lattice <br> Constant | Element/alloy | Lattice <br> Constant |
| :--- | :---: | :--- | :---: |
| Pt | 3.991 | Rh | 3.865 |
| Pd | 3.980 | Ru | 3.850 |
| Ag | 4.172 | $\mathrm{Pt}_{3} \mathrm{Ni}$ | 3.907 |
| Au | 4.186 | $\mathrm{Pt}_{3} \mathrm{Co}$ | 3.911 |
| Ir | 3.883 |  |  |

Table S2 Binding energies of O , 'dry' and 'hydrated' OH and OOH intermediates.

| Surface | OH | OH-wl | O | OOH | OOH-wl |
| :--- | :---: | :---: | :---: | :---: | :---: |
| $\operatorname{Pt}(111)$ | 0.961 | 0.427 | 1.457 | 3.948 | 3.792 |
| $+1 \%$ | 0.931 | 0.401 | 1.393 | 3.937 | 3.786 |
| $+2 \%$ | 0.871 | 0.201 | 1.308 | 3.891 | 3.718 |
| $\mathrm{Pt} / \mathrm{Cu} / \mathrm{Pt}(111)$ | 1.320 | 0.888 | 1.717 | 4.255 | 4.111 |
| $\mathrm{Pd} @ \operatorname{Pt-skin}(111)$ | 1.045 | 0.581 | 1.513 | 4.009 | 3.862 |
| $\mathrm{Pt}_{3} \mathrm{Ni}(111)$ | 1.080 | 0.660 | 1.719 | 4.075 | 3.987 |
| $\mathrm{Pt}_{3} \mathrm{Co}(111)$ | 1.128 | 0.725 | 1.811 | 4.128 | 4.010 |
| $\operatorname{Pd}(111)$ | 0.904 | 0.412 | 1.444 | 4.009 | 3.787 |
| $+1 \%$ | 0.853 | 0.308 | 1.385 | 4.008 | 3.741 |
| $+2 \%$ | 0.812 | 0.319 | 1.338 | 3.960 | 3.725 |
| $-1 \%$ | 0.936 | 0.462 | 1.499 | 4.040 | 3.816 |
| $-2 \%$ | 0.971 | 0.481 | 1.559 | 4.060 | 3.826 |
| $-3 \%$ | 1.005 | 0.492 | 1.631 | 4.088 | 3.893 |
| $-4 \%$ | 1.052 | 0.570 | 1.705 | 4.118 | 3.927 |
| $-5 \%$ | 1.064 | 0.700 | 1.756 | 4.128 | 3.987 |
| $\mathrm{Pt} @ \operatorname{Pd}-$-skin(111) | 0.740 | 0.148 | 1.360 | 3.945 | 3.668 |
| $\mathrm{Ag}(111)$ | 0.831 | 0.487 | 2.056 | 4.123 | 3.846 |
| $-1 \%$ | 0.849 | 0.510 | 2.069 | 4.111 | 3.804 |
| $-2 \%$ | 0.903 | 0.515 | 2.116 | 4.135 | 3.820 |
| $+1 \%$ | 0.799 | 0.468 | 2.032 | 4.107 | 3.830 |


| $+2 \%$ | 0.751 | 0.421 | 1.983 | 4.070 | 3.822 |
| :--- | :---: | :---: | :---: | :---: | :---: |
| $+3 \%$ | 0.696 | 0.403 | 1.926 | 4.032 | 3.817 |
| $+4 \%$ | 0.651 | 0.391 | 1.886 | 4.000 | 3.847 |
| $+5 \%$ | 0.607 | 0.387 | 1.839 | 3.965 | 3.809 |
| $\mathrm{Au}(111)$ | 1.482 | 1.057 | 2.610 | 4.677 | 4.35 |
| $\operatorname{Ir}(111)$ | 0.291 | -0.121 | 0.805 | 3.664 | - |
| $\operatorname{Rh}(111)$ | 0.463 | -0.049 | 0.697 | 3.716 | 3.575 |
| $\operatorname{Ru}(0001)$ | 0.027 | -0.497 | -0.150 | - | - |

