The electronic supplementary information for

Metal-acid site synergistic catalysis in Ru–ZrO₂ toward selective hydrogenation of benzene to cyclohexene

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Computational details

Density functional theory (DFT),^[1] as implemented in the Vienna ab initio Simulation Package (VASP),^[2] was employed to perform the first-principles calculations. The DFT calculations were performed by using the electron projector-augmented wave methods^[3] with the PBE generalized gradient approximation (GGA) exchange-correlation functional,^[4] plus an on-site Zr d state U correction (DFT + U, or GGA + U).^[5] The value of U = 4.0 was applied for the Coulomb correction to the Zr 4d states, which was reported to well describe the electronic properties and defect states in crystalline Zirconium dioxide.^[6-9] The ZrO₂(-111) surface and B-doped ZrO₂(-111) surface with *p* (2×2) supercells are cleaved according to the XRD patterns (Fig. 1 in main manuscript). A plane-wave cut-off of 400 eV was used. The Brillouin zone was sampled using a $3\times3\times1$ Monkhorst-Pack *k*-point mesh.^[10] Spin polarization was taken into account in all calculations for comparison. The adsorption energy (*E*_{ad}) was defined as follows:

$$E_{\rm ad} = E_{\rm total} - (E_{\rm slab} + E_{\rm adsorbate}) \tag{1}$$

where E_{total} is the total energy of adsorbate-slab system; E_{slab} and $E_{\text{adsorbate}}$ are the energies of the slab and the gas phase adsorbate, respectively. A negative value of E_{ad} corresponds to an exothermic process.



Fig. S1 Ru 3p XPS spectra of (a) Ru/ZrO_2 , (b) $Ru/ZrO_2@ZrO_2-B(0\%)$, (c) $Ru/ZrO_2@ZrO_2-B(5\%)$, (d) $Ru/ZrO_2@ZrO_2-B(10\%)$ and (e) $Ru/ZrO_2@ZrO_2-B(15\%)$, respectively.



Fig. S2 Sideview of (A) $ZrO_2(-111)$ and (B) B- $ZrO_2(-111)$ facet. The red, cyan and pink ball represents O, Zr and B atom, respectively. The doped B atom and its original position of Zr atom are encircled.



Fig. S3 Nitrogen adsorption-desorption curves and the pore diameter distribution (inset) for the sample of $Ru/ZrO_2@ZrO_2-B(5\%)$.



Fig. S4 Catalytic performance of Ru/ZrO₂@ZrO₂(5%) in three consecutive recycles.

Reference

- (1) W. Kohn and L. J. Sham, Phys. Rev., 1965, 140, 1133-1138.
- (2) G. Kresse and J. Furthmuller, Phys. Rev. B: Condens. Matter Mater. Phys., 1996, 54, 11169-11186.
- (3) P. E. Blochl, Phys. Rev. B: Condens. Matter Mater. Phys., 1994, 50, 17953-17979.
- (4) J. P. Perdew, K. Burke and M. Ernzerhof, Phys. Rev. Lett., 1996, 77, 3865-3868.
- (5) S. L. Dudarev, G. A. Botton, S. Y. Savrasov, C. J. Humphreys and A. P. Sutton, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1998, **57**, 1505-1509.
- (6) R. P. Antonio and P. Gianfranco, Nanoscale, 2017, 9, 6866-6876.
- (7) S. Philomena, R. P. Antonio and P. Gianfranco, *Phys. Chem. Chem. Phys.*, 2015, **17**, 22342-22360.
- (8) T. C. Hsin-Yi, T. Sergio and P. Gianfranco, Surf. Sci., 2016, 652, 163-171.
- (9) T. C. Hsin-Yi, T. Sergio and P. Gianfranco, ACS Catal., 2015, 5, 5486-5495.
- (10) H. J. Monkhorst and J. D. Pack, Phys. Rev. B, 1976, 13, 5188-5192.