## **Electronic Supporting Information**

## **Dissolution Induced Self-Selective Zn- and Ru-doped TiO<sub>2</sub>** Structure for Electrochemical Generation of KClO<sub>3</sub>

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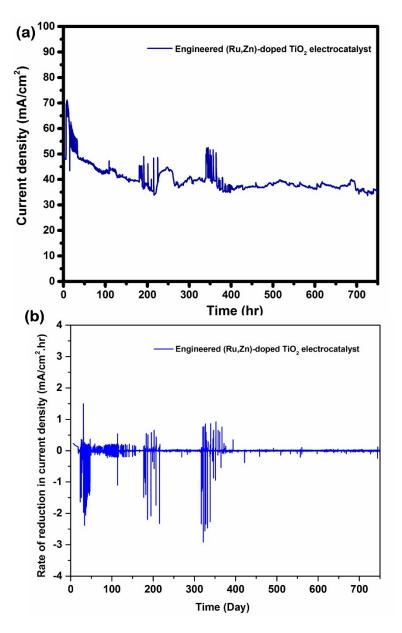
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*<u>#equally contributed to the work</u>* 

Section S1: (a) The current density and (b) the rate of reduction in current density of engineered (Zn, Ru)-doped TiO<sub>2</sub> (Ru<sub>0.26</sub>Ti<sub>0.73</sub>Zn<sub>0.01</sub>O<sub>x</sub>) electrocatalyst at an electrochemical potential of 1.35 V vs Ag/AgCl in neutral 0.5 M KCl solution. The electrolyte is change after 7 hours. (Page 3)

Section S2: Individual fitted XPS peaks of Ru-3d zone (278 eV-292 eV) of engineered (Zn, Ru)-doped TiO<sub>2</sub> (Ru<sub>0.26</sub>Ti<sub>0.73</sub>Zn<sub>0.01</sub>O<sub>x</sub>), untreated (Zn, Ru)-doped TiO<sub>2</sub> (Ru<sub>0.31</sub>Ti<sub>0.67</sub>Zn<sub>0.02</sub>O<sub>x</sub>), Ru-doped TiO<sub>2</sub> having a composition of Ru<sub>0.3</sub>Ti<sub>0.7</sub>O<sub>2</sub> and RuO<sub>2</sub>. (Page 4-7)



**Fig. S1:** (a) The current density and (b) the rate of reduction in current density of engineered (Zn, Ru)-doped TiO<sub>2</sub> (Ru<sub>0.26</sub>Ti<sub>0.73</sub>Zn<sub>0.01</sub>O<sub>x</sub>) electrocatalyst at an electrochemical potential of 1.35 V vs Ag/AgCl in neutral 0.5 M KCl solution. The electrolyte was changed after 7 hours. The positive values and the negative values in Figure (b) represent reduction in current density and increase in current density due to (1) increase in ECSA or (2) change in electrolyte.

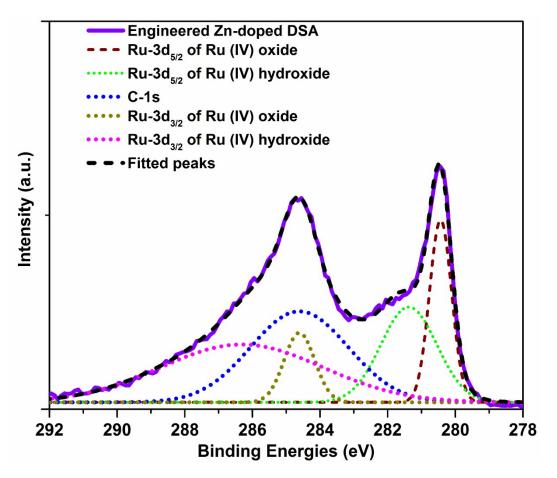
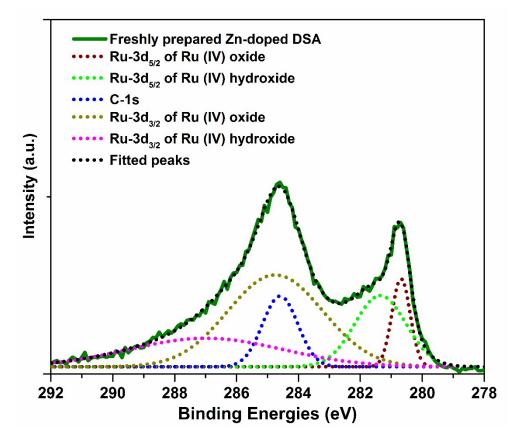


Fig. S2: Fitted XPS peaks of Ru-3d zone (278 eV-292 eV) of engineered Zn-doped DSA electrocatalyst.



**Fig. S3:** Fitted XPS peaks of Ru-3d zone (278 eV-292 eV) of freshly prepared Zn-doped DSA electrocatalyst.

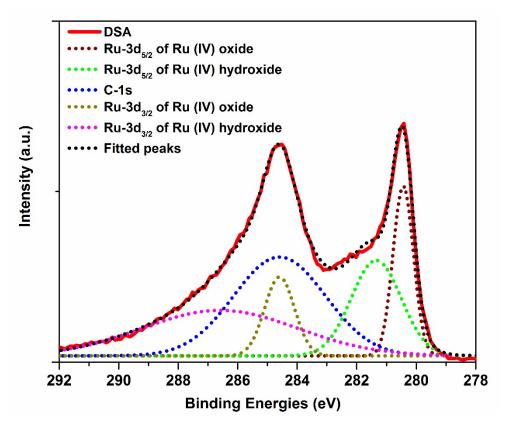


Fig. S4: Fitted XPS peaks of Ru-3d zone (278 eV-292 eV) of freshly prepared DSA electrocatalyst

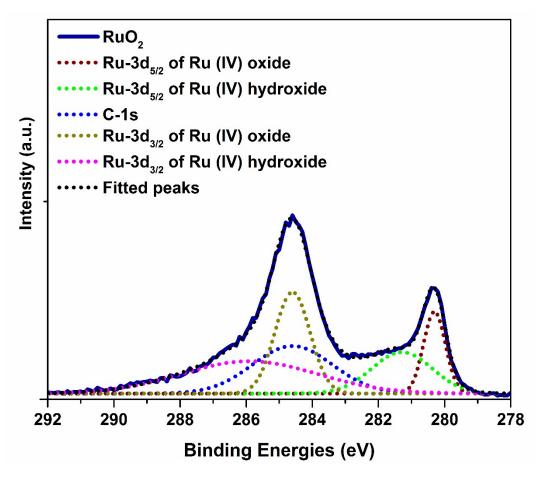


Fig. S5: Fitted XPS peaks of Ru-3d zone (278 eV-292 eV) of RuO<sub>2</sub> electrocatalyst.