

**Copper- Coordination Polymers Controlled Cu@N-rGO or CuO@C Nanoparticles  
Formation: Reusable Green Catalyst for A-3 Coupling and Nitroarene Reduction**

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Page 2. PXRD pattern of Cu-COPs (Figure S1)

Page 3. FTIR spectra of Cu-COPs before and after calcination (Figure S2)

Page 4 -11. XPS spectra of CuNPs and CuONPs (Figure S3 to Figure S10).

Page 12. FE-SEM images of CuNPs-1 and CuONPs-4 (Figure S11).

Page 13. EDX spectra of CuNPs-1 and CuONPs-4 (Figure S12).

Page 14. Molecular packing of Cu-COP-1, Cu-COP-5, Cu-COP-6 and Cu-COP-7 (Figure S13).

Page 15 and 16. Atomic weight percentage of Cu/CuONPs (Table S1).

Page 17. Acid stability of CuNPs-1 and CuONPs-4 (Figure S14).

Page 18. The tentative mechanism of CuNPs-1 for A<sup>3</sup> coupling reaction (Scheme S1).

Page 19 to 26. <sup>1</sup>HNMR spectra A3 coupling reaction products (Figure S15).

Page 27 to 37. <sup>1</sup>HNMR spectra nitroarene reduction reaction products (Figure S16).

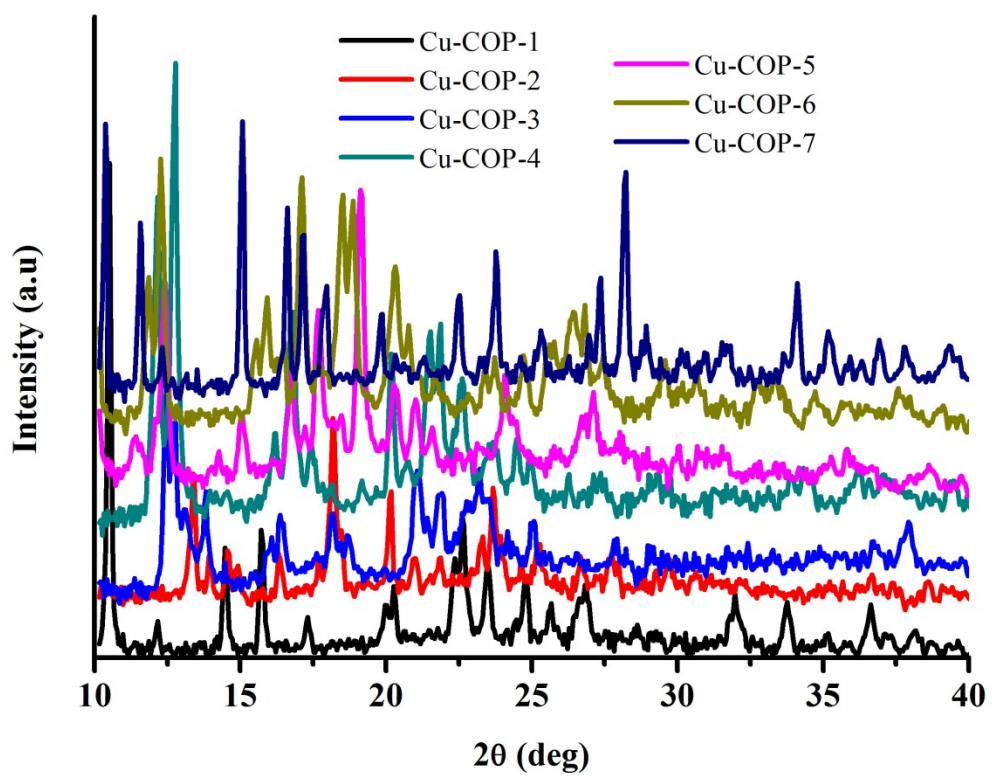


Figure S1. PXRD pattern of Cu-COPs.

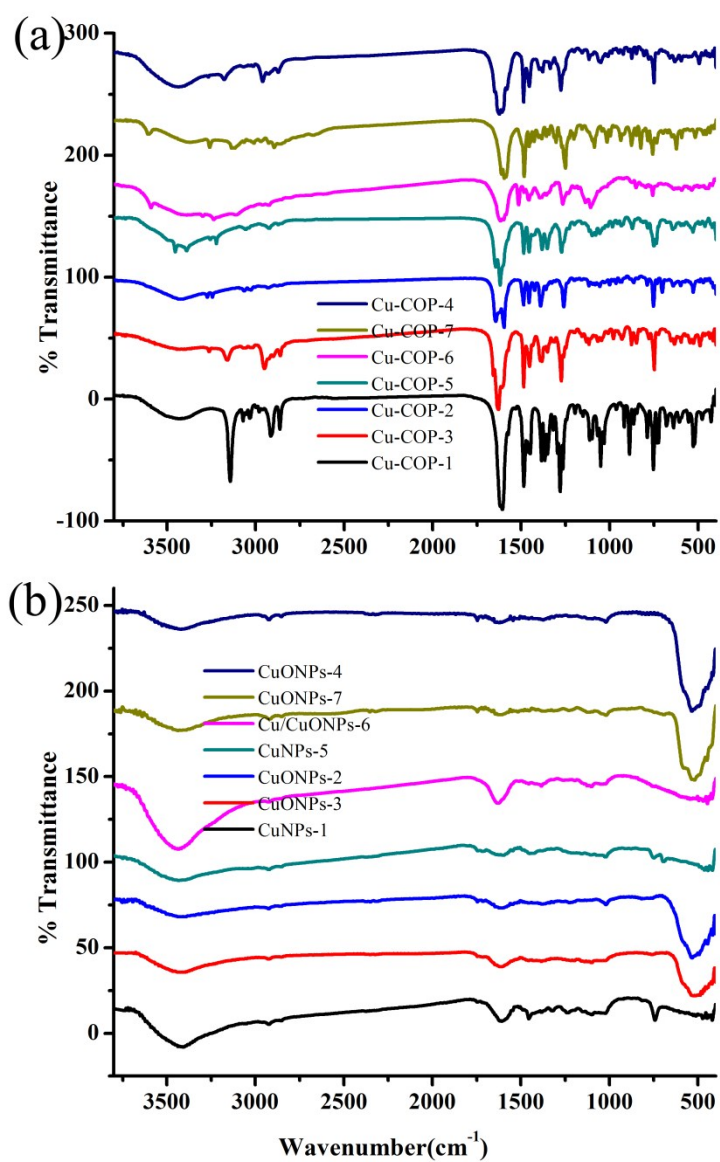


Figure S2. FTIR spectra of (a) Cu-COPs and (b) Cu-COPs after calcination.

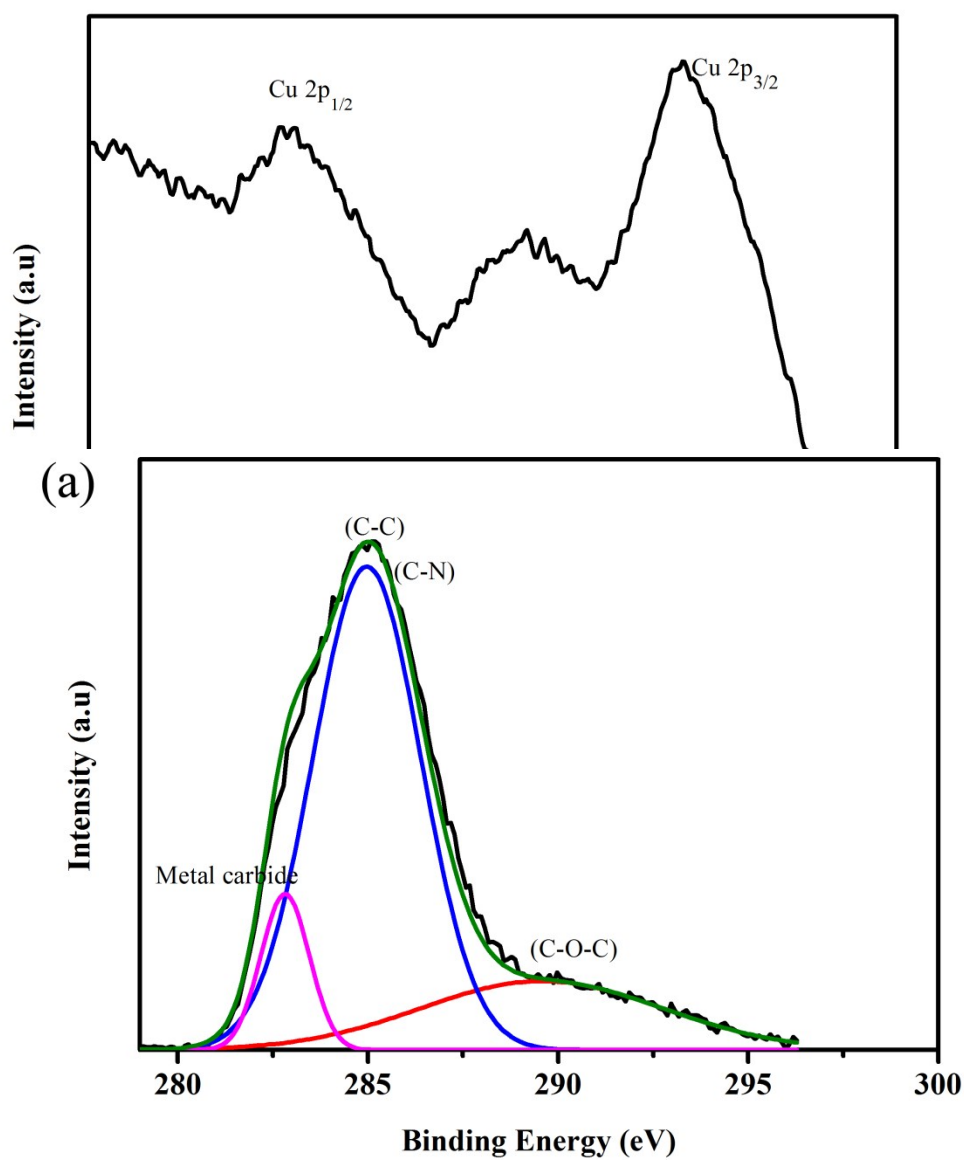


Figure S3.  
XPS  
spectra of  
CuNPs-1  
Cu.

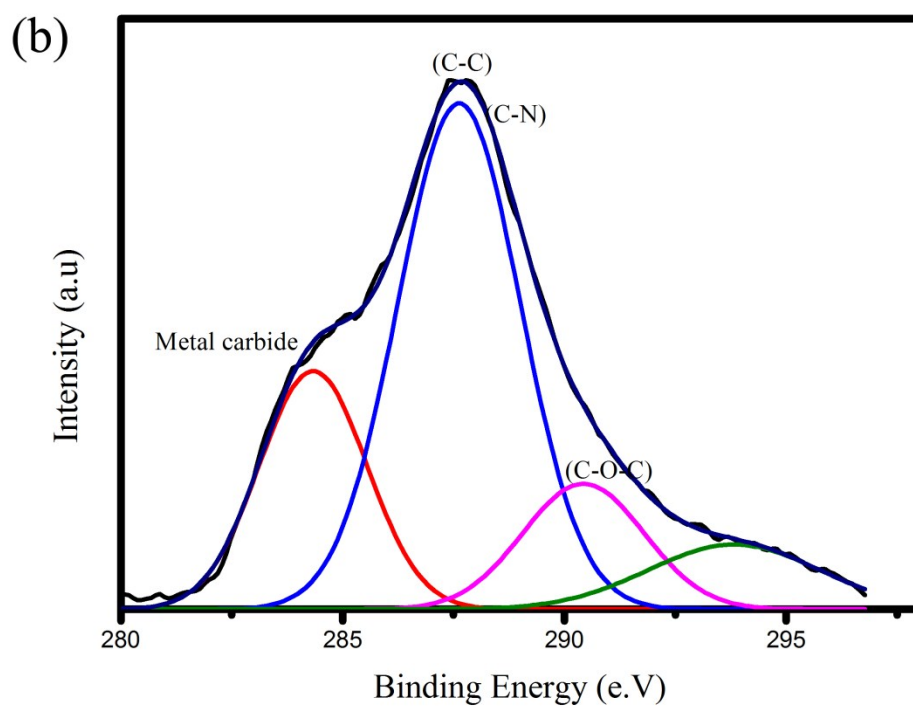


Figure S4. XPS spectra of (a) CuNPs-1 and (b) CuNPs-5 for C.

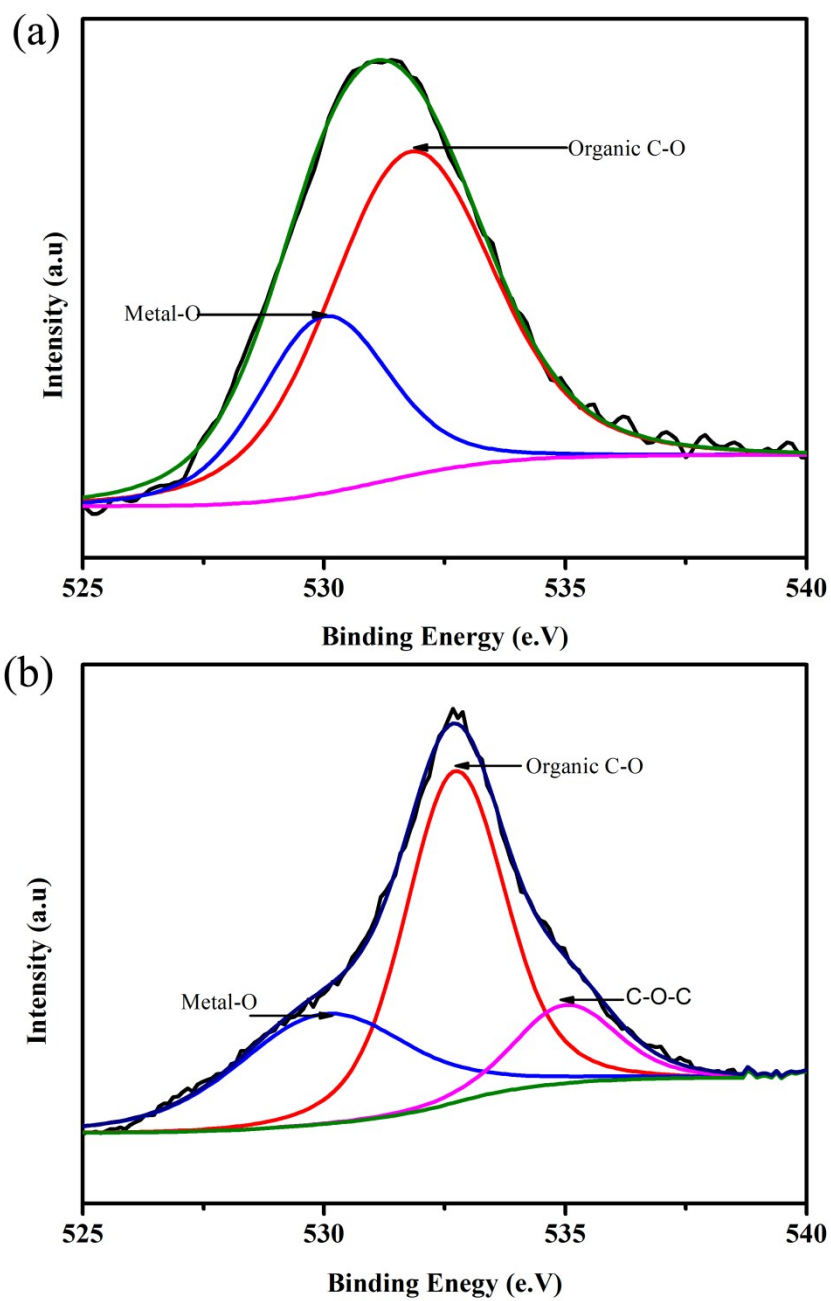


Figure S5. XPS spectra of (a) CuNPs-1 and (b) CuNPs-5 for O.

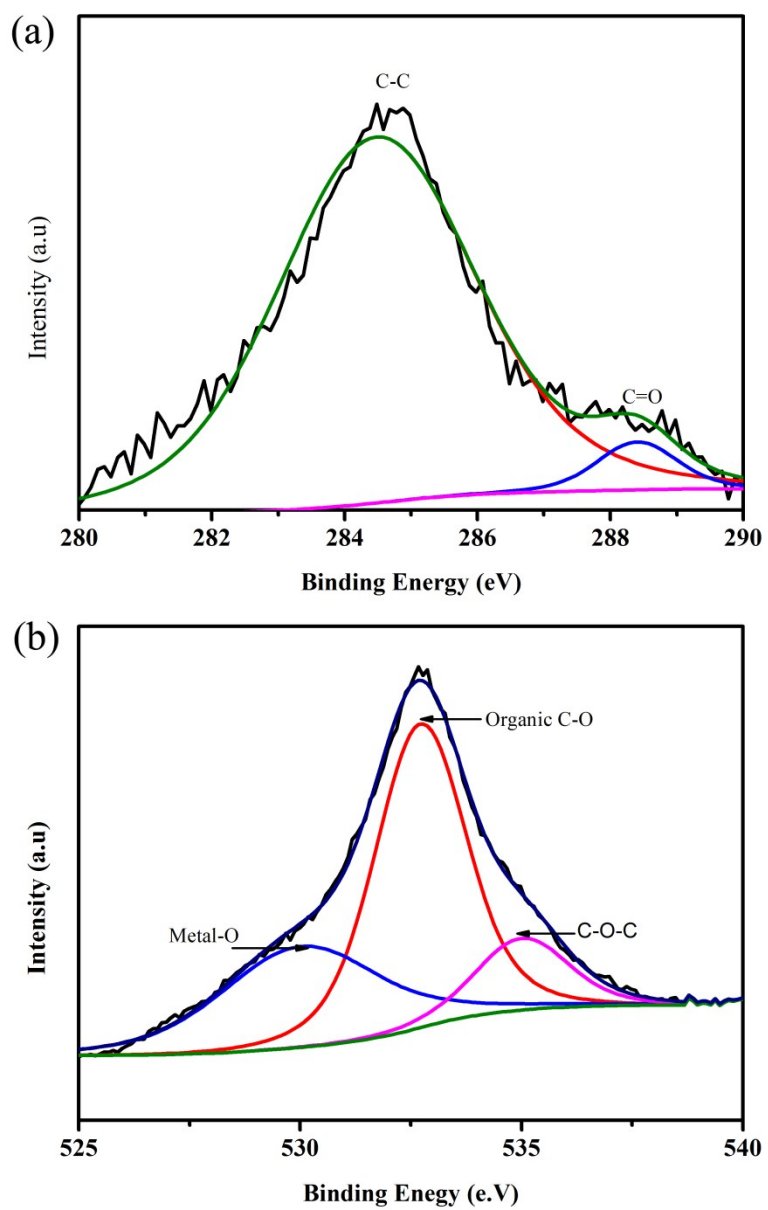


Figure S6. XPS spectra of CuONPs-2 (a) C and (b) O.

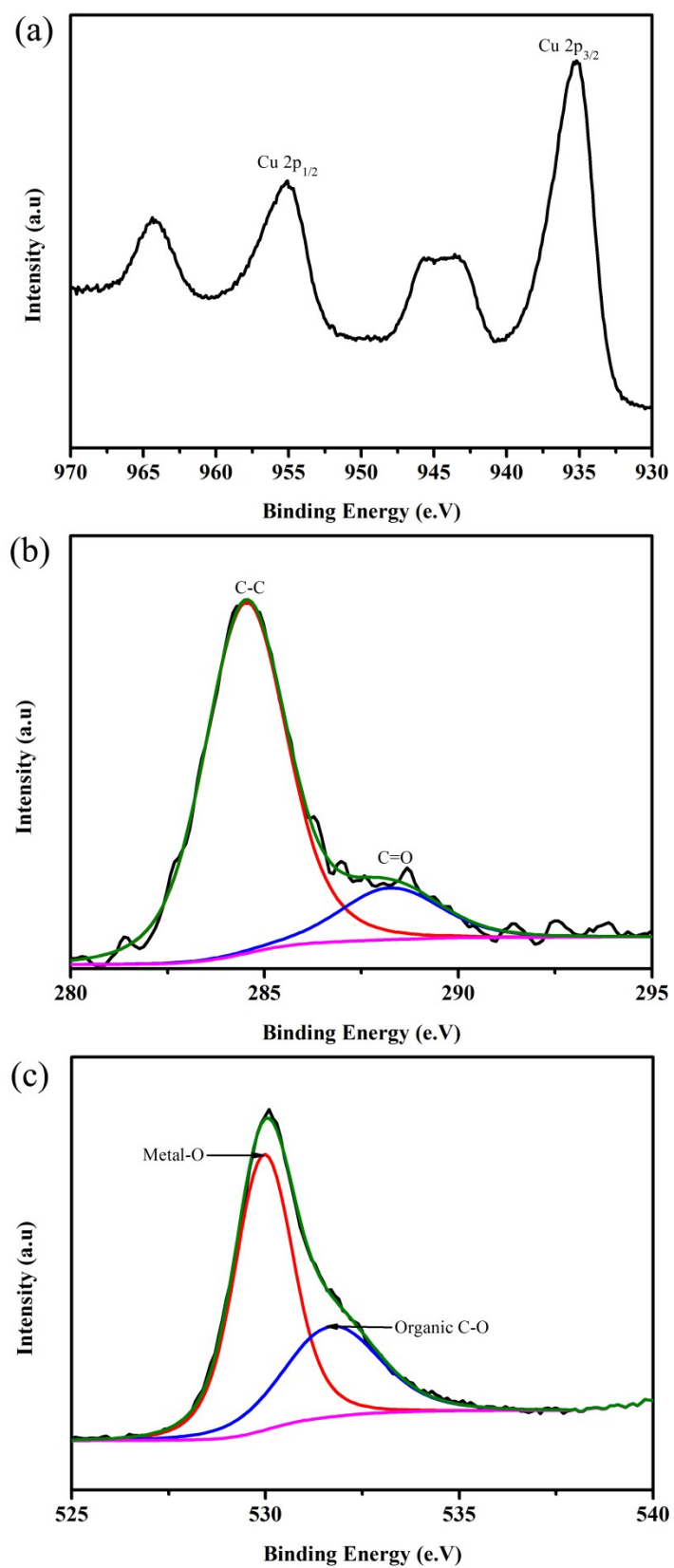


Figure S7. XPS spectra of CuONPs-3 (a) Cu, (b) C and (c) O.



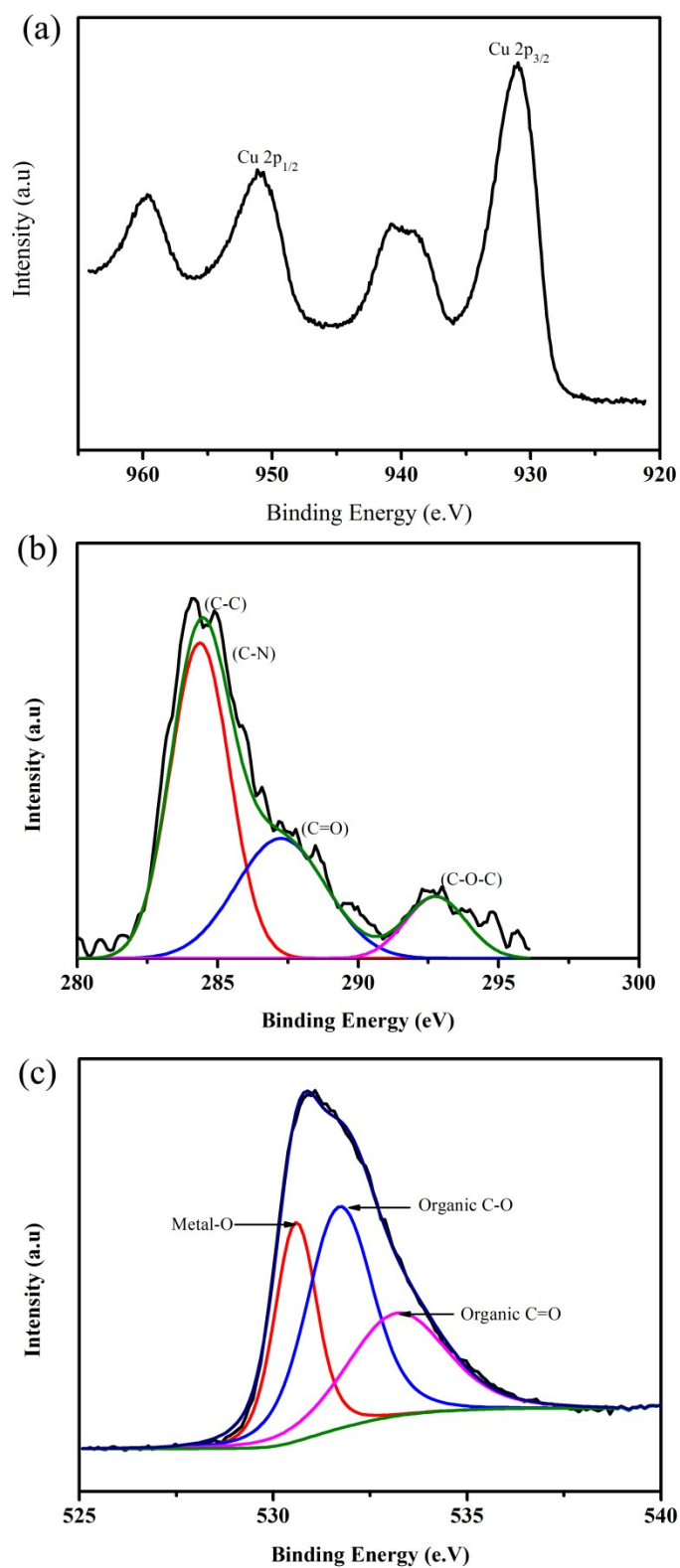


Figure S8. XPS spectra of CuONPs-4 (a) Cu, (b) C and (c) O.

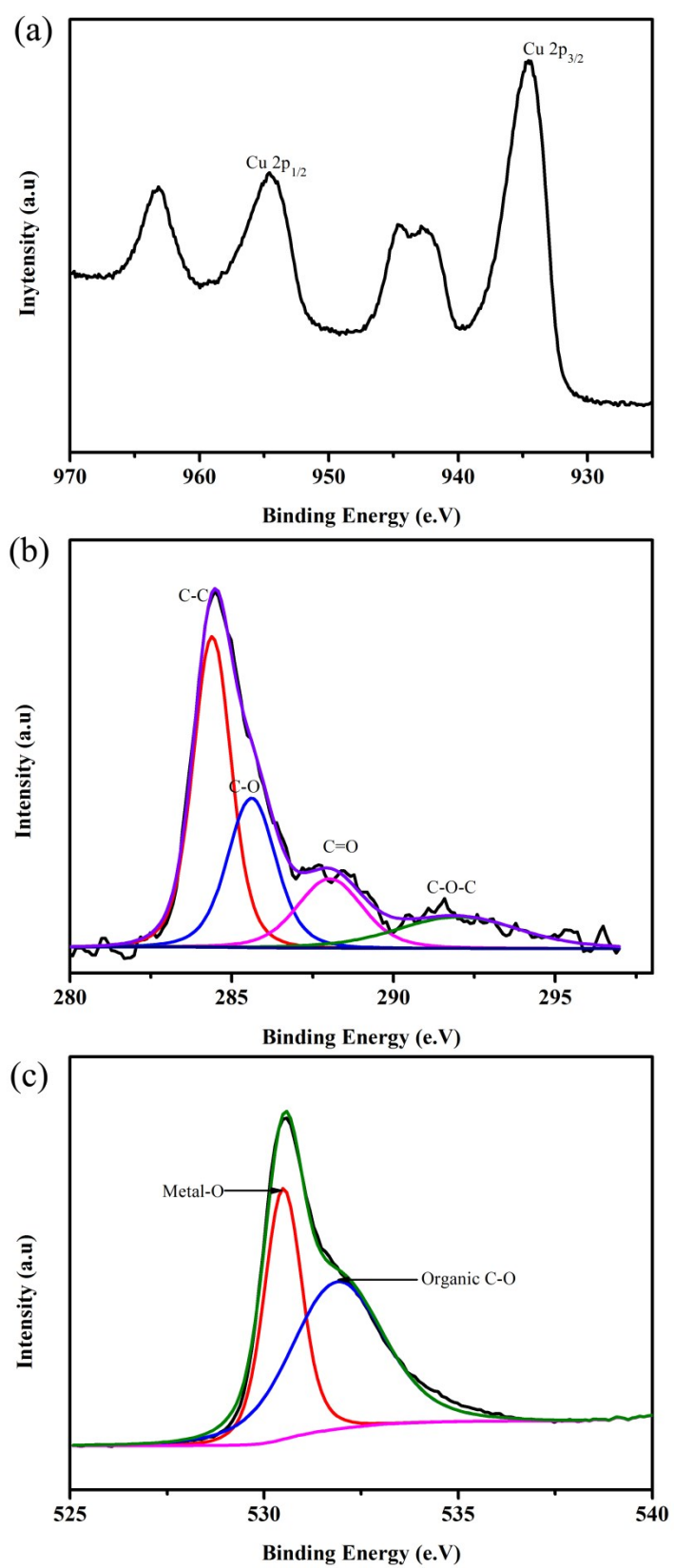


Figure S9. XPS spectra of CuONPs-7 (a) Cu, (b) C and (c) O.

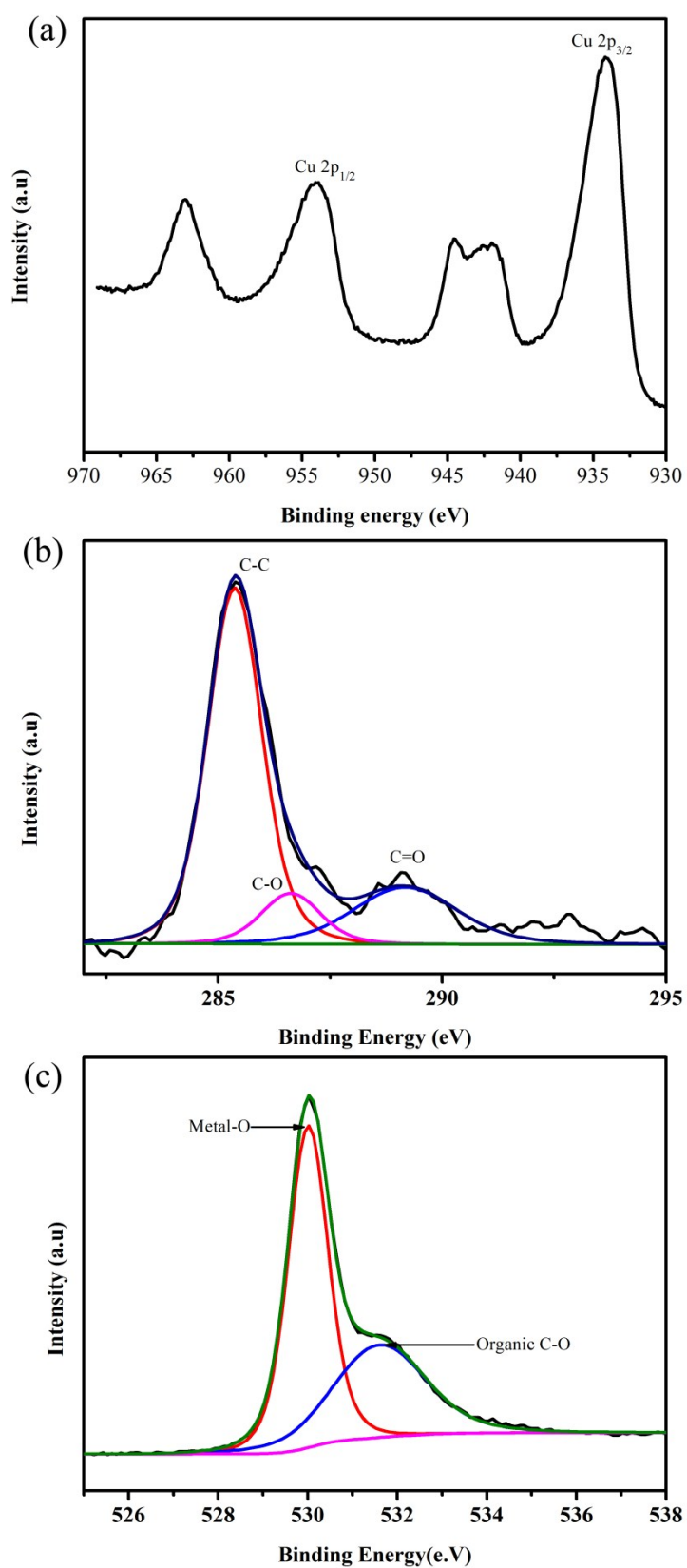


Figure S10. XPS spectra of CuONPs-6 (a) Cu, (b) C and (c) O.

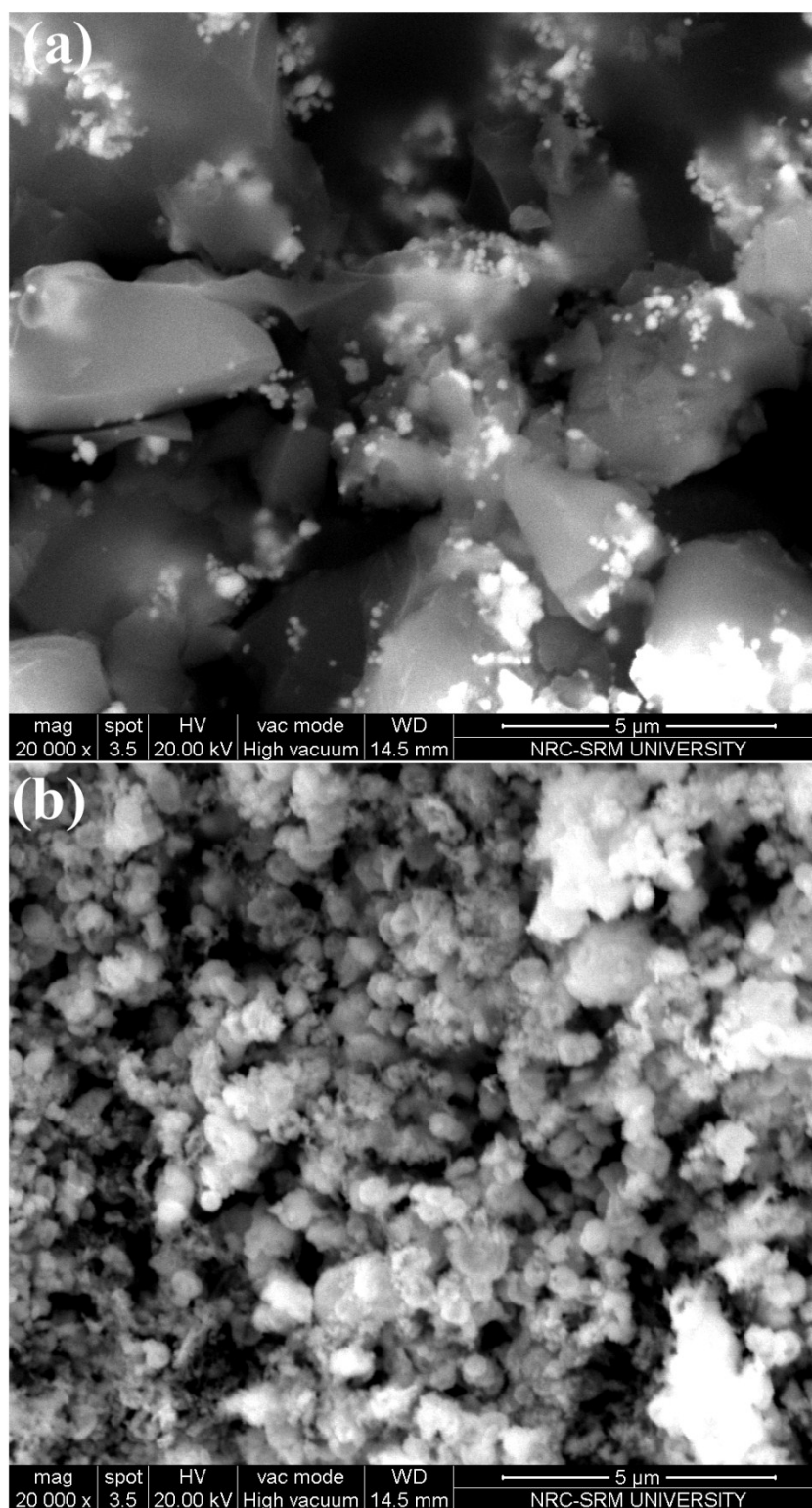


Figure S11. FE-SEM images of (a) CuNPs-1 and (b) CuONPs-4.

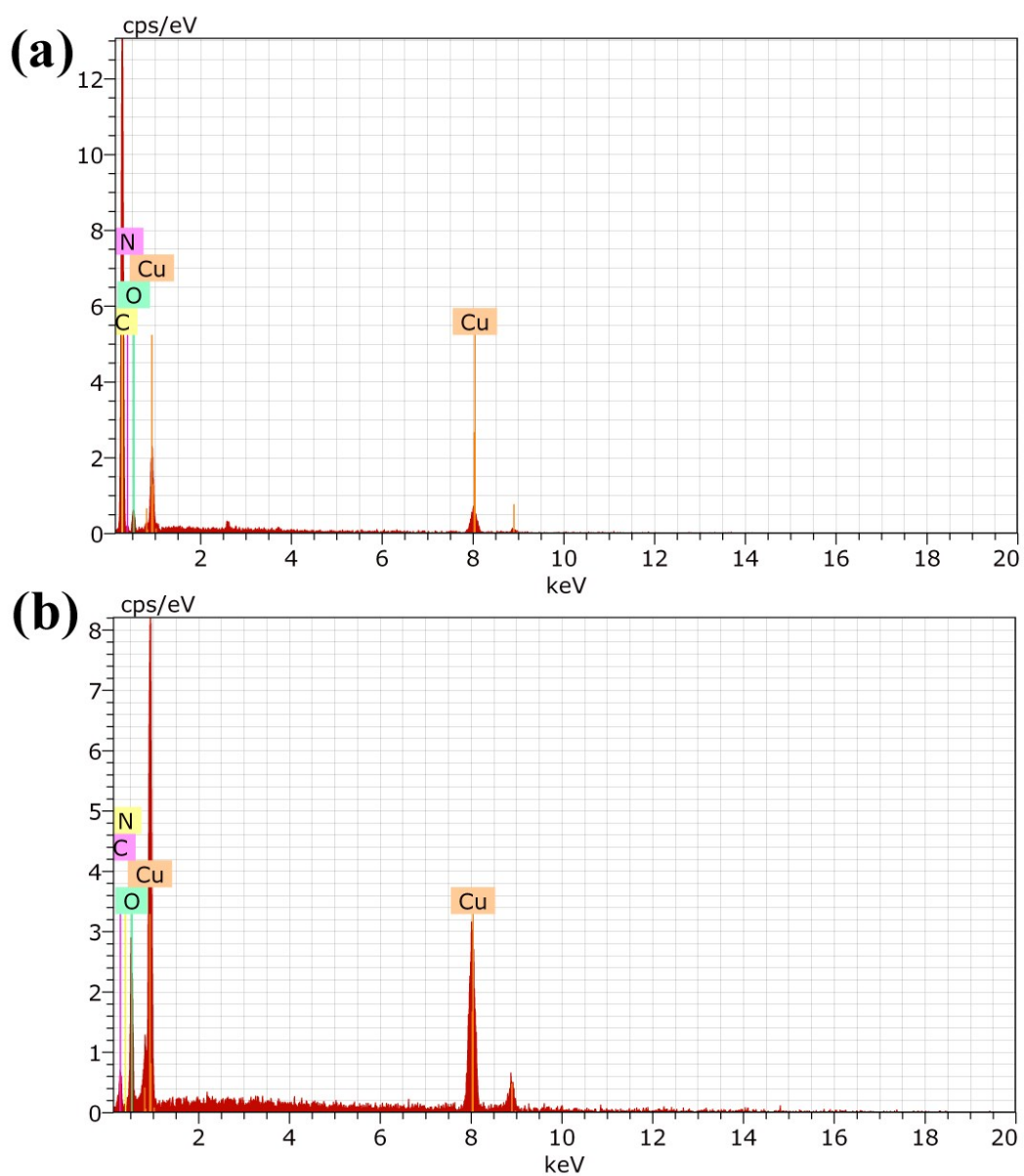


Figure S12. EDAX spectra of (a) CuNPs-1 and (b) CuONPs-4.

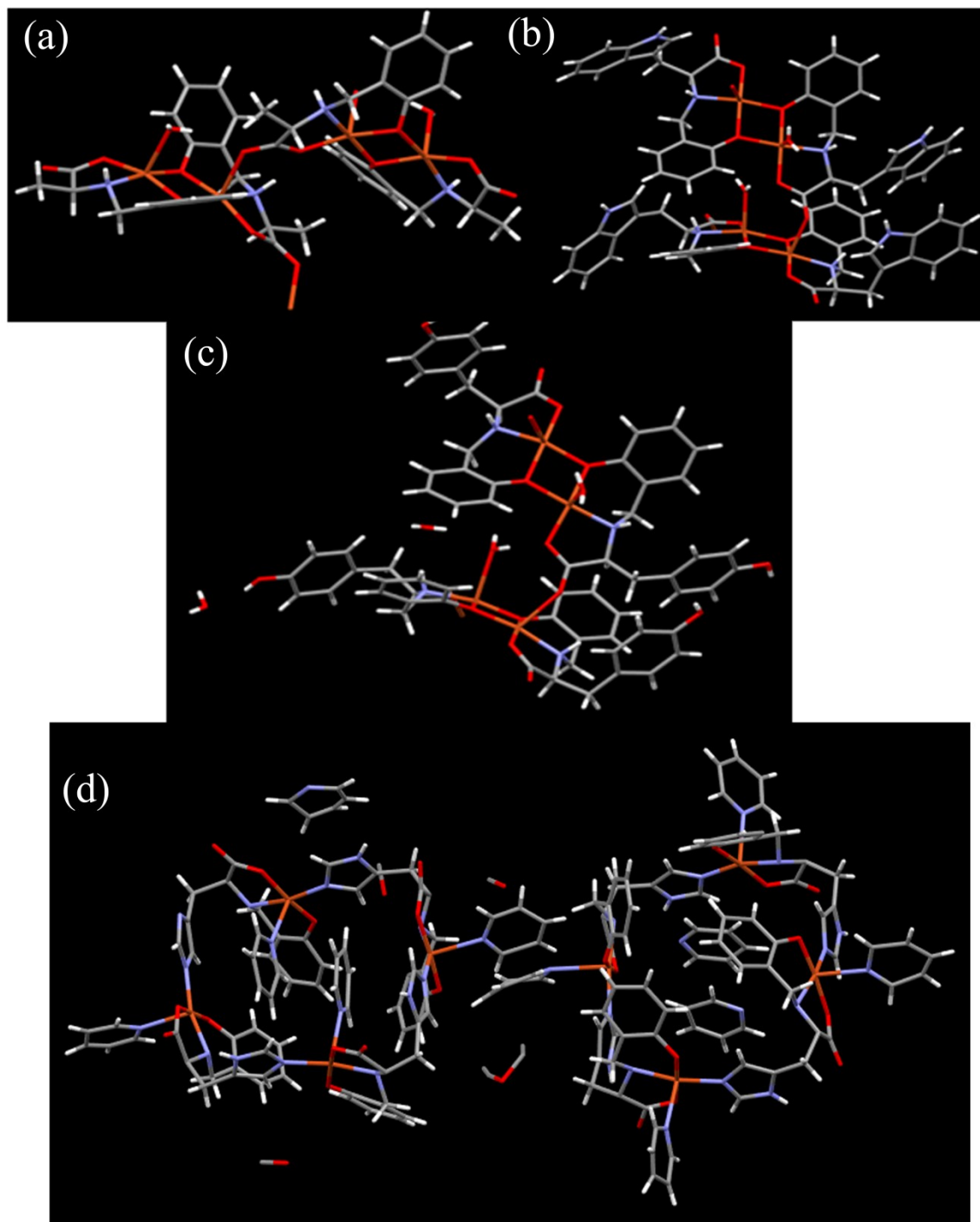


Figure S13. Molecular packing of (a) Cu-COP-1 (CCDC No. 133087), (b) Cu-COP-5 (CCDC No. 209715), (c) Cu-COP-6 (CCDC No. 209714) and (d) Cu-COP-7 (CCDC No. 197602) in solid state.

Table S1. Atomic weight percentage of Cu/CuONPs obtained from XPS analysis.

<b>CuNPs-1</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
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C1s	286.62	4.20	78104.71	81.57
N1s	401.30	4.11	13099.65	8.07
O1s	533.98	4.54	22406.13	9.17
Cu2p	935.73	6.09	17999.53	1.18

<b>CuNPs-5</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	289.08	7.27	556227.95	73.63
N1s	402.08	7.72	37517.10	2.76
O1s	537.08	8.01	486249.09	21.97
Cu2p	938.08	9.33	314433.41	1.64

<b>CuONPs-2</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	939.76	5.87	2823127.13	23.51
N1s	401.16	0.54	1121.04	0.72
O1s	536.81	4.69	1048264.88	54.32
Cu2p	939.76	5.87	2823127.13	23.51

<b>CuONPs-3</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	287.23	2.51	15598.50	15.86
N1s	401.67	0.76	1829.15	1.10
O1s	531.96	2.13	108172.39	43.05
Cu2p	935.26	3.28	626315.17	39.99



<b>CuONPs-4</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	286.40	3.15	12199.88	11.73
N1s	400.99	0.30	687.03	0.39
O1s	531.18	3.08	121812.44	45.85
Cu2p	934.90	3.78	695901.47	42.02

<b>CuONPs-7</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	286.22	3.33	109236.79	13.34
N1s	400.86	0.46	986.68	0.82
O1s	531.35	3.33	989035.13	47.28
Cu2p	935.01	4.05	3383743.20	39.38

<b>Cu/CuONPs-6</b>	Peak BE	FWHM eV	Area (P) CPS.eV	Atomic %
C1s	285.98	2.77	149377.78	14.66
N1s	401.42	0.51	1012.24	0.91
O1s	531.09	2.99	1222401.21	46.96
Cu2p	934.46	3.94	4105417.78	38.38

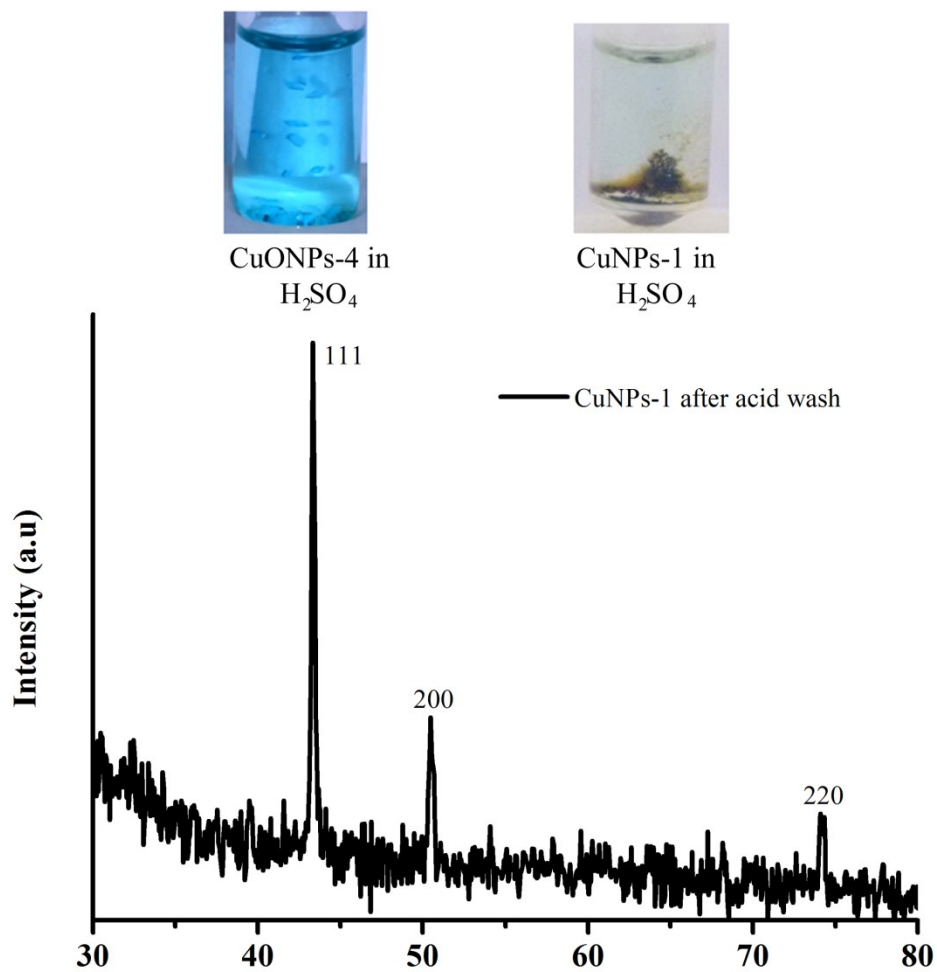
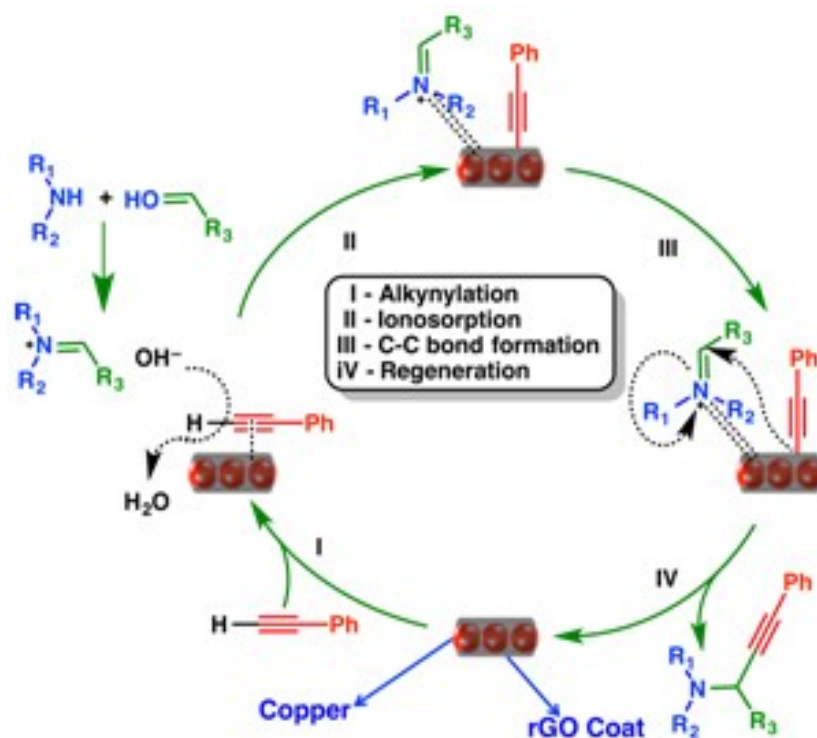
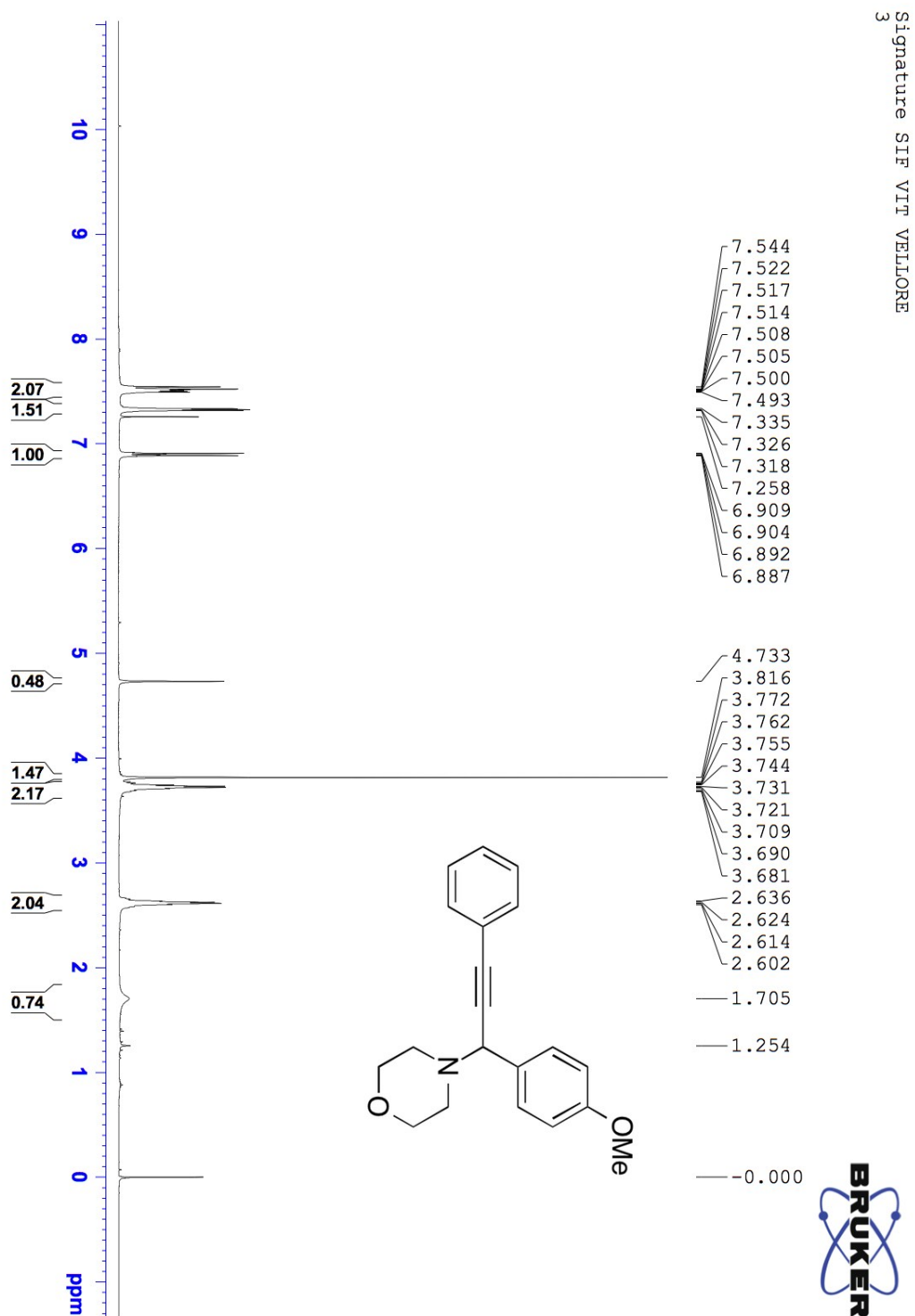


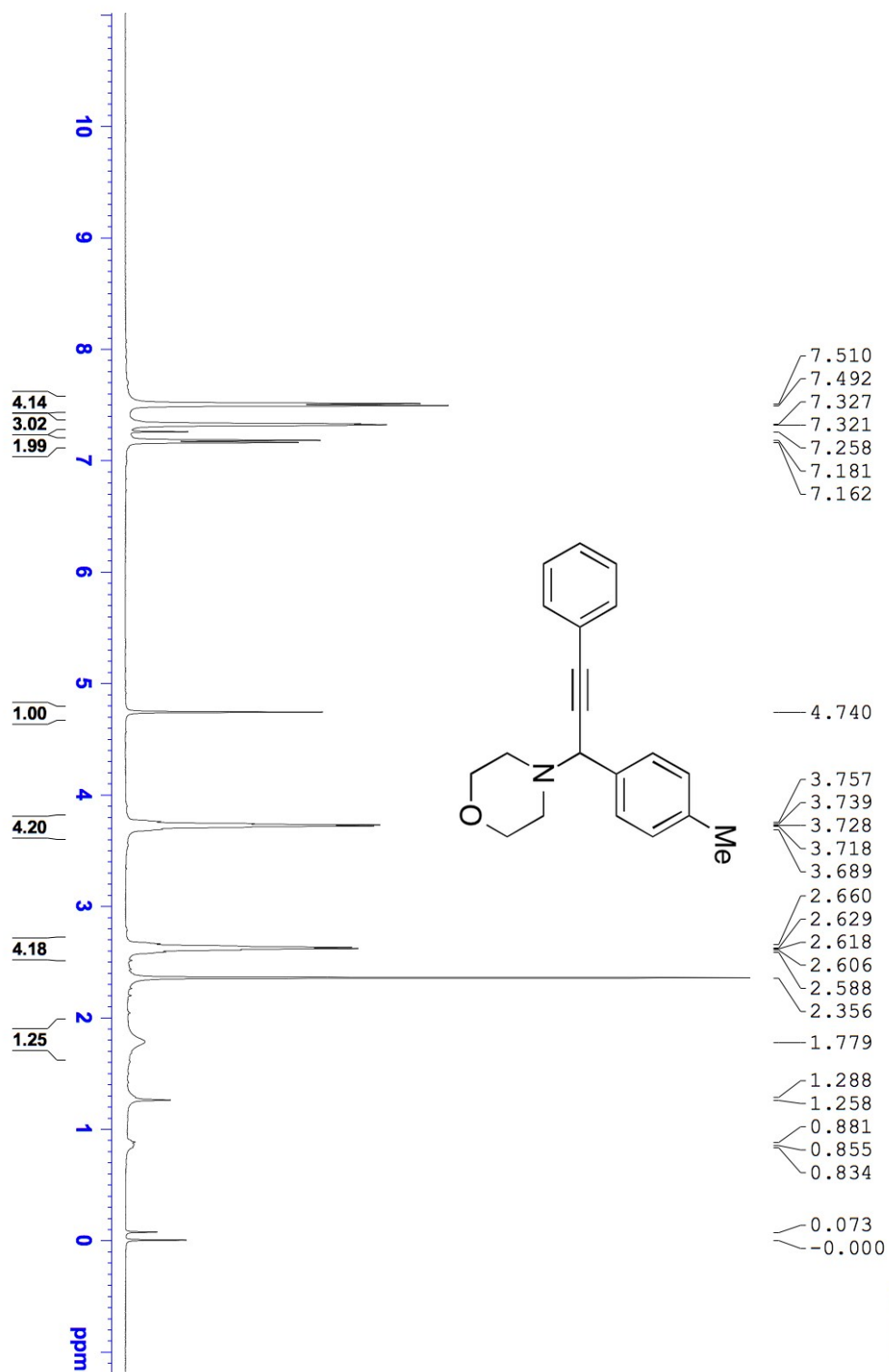
Figure S14. Digital images of CuONPs-4 and CuNPs-1 and PXRD pattern of CuNPs-1 after immersing in concentrated H<sub>2</sub>SO<sub>4</sub> (10 M).



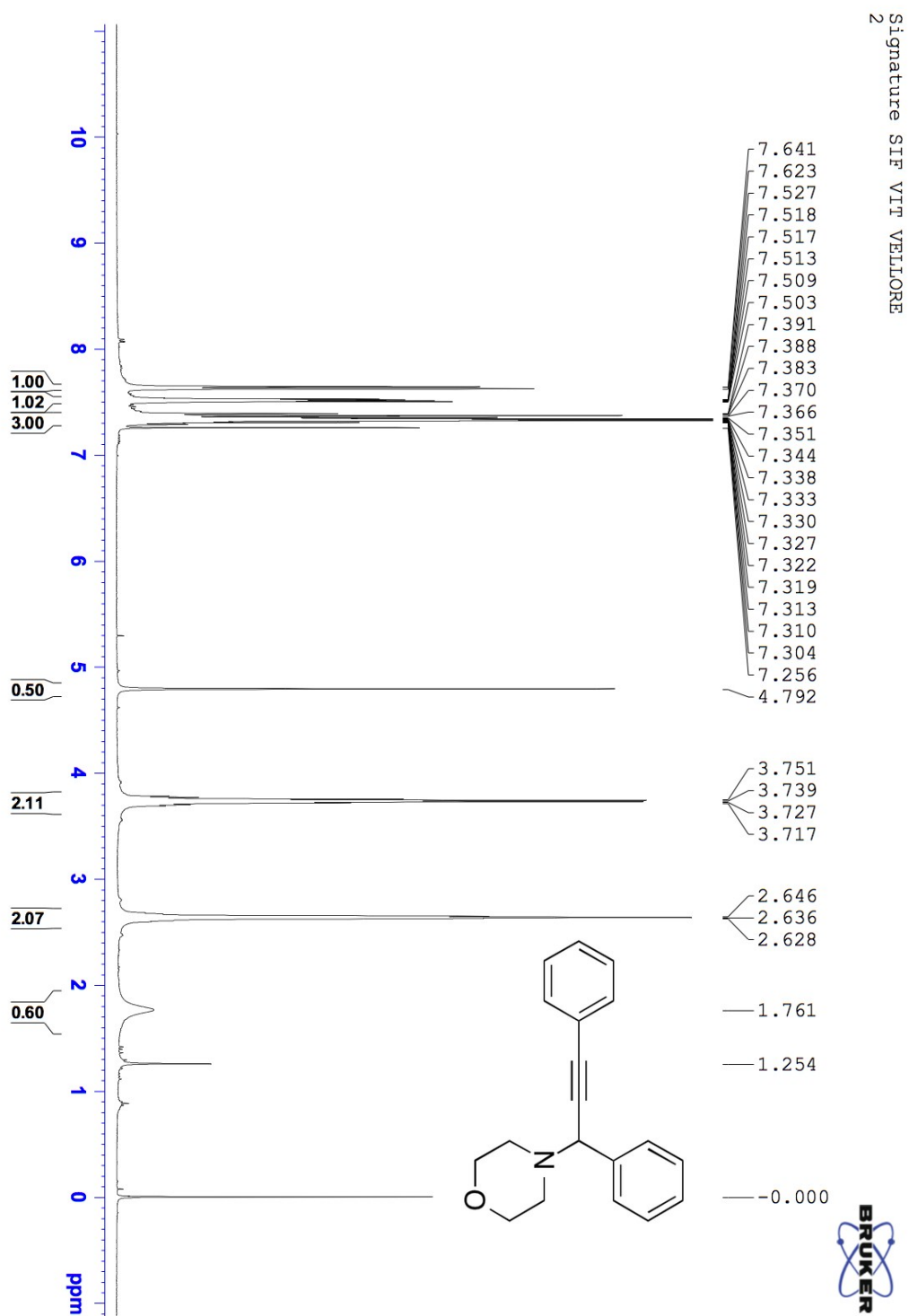
Scheme S1. The first step in A3 coupling is deprotonation of acetylene. Chelation of acetylene with N-rGO encapsulated CuNPs-1 might increase the labile nature of C-H bond in acetylene, thus weak base in the reaction medium could deprotonate the proton. The resulting iminium ion might get ionosorped on the surface of N-rGO encapsulated CuNPs-1. Further, electron transfer and C-C bond formation results in the formation of product and the regeneration of catalyst.

Figure S15. <sup>1</sup>H-NMR spectra of propargylamines obtained using CuNPs-1 catalyst.

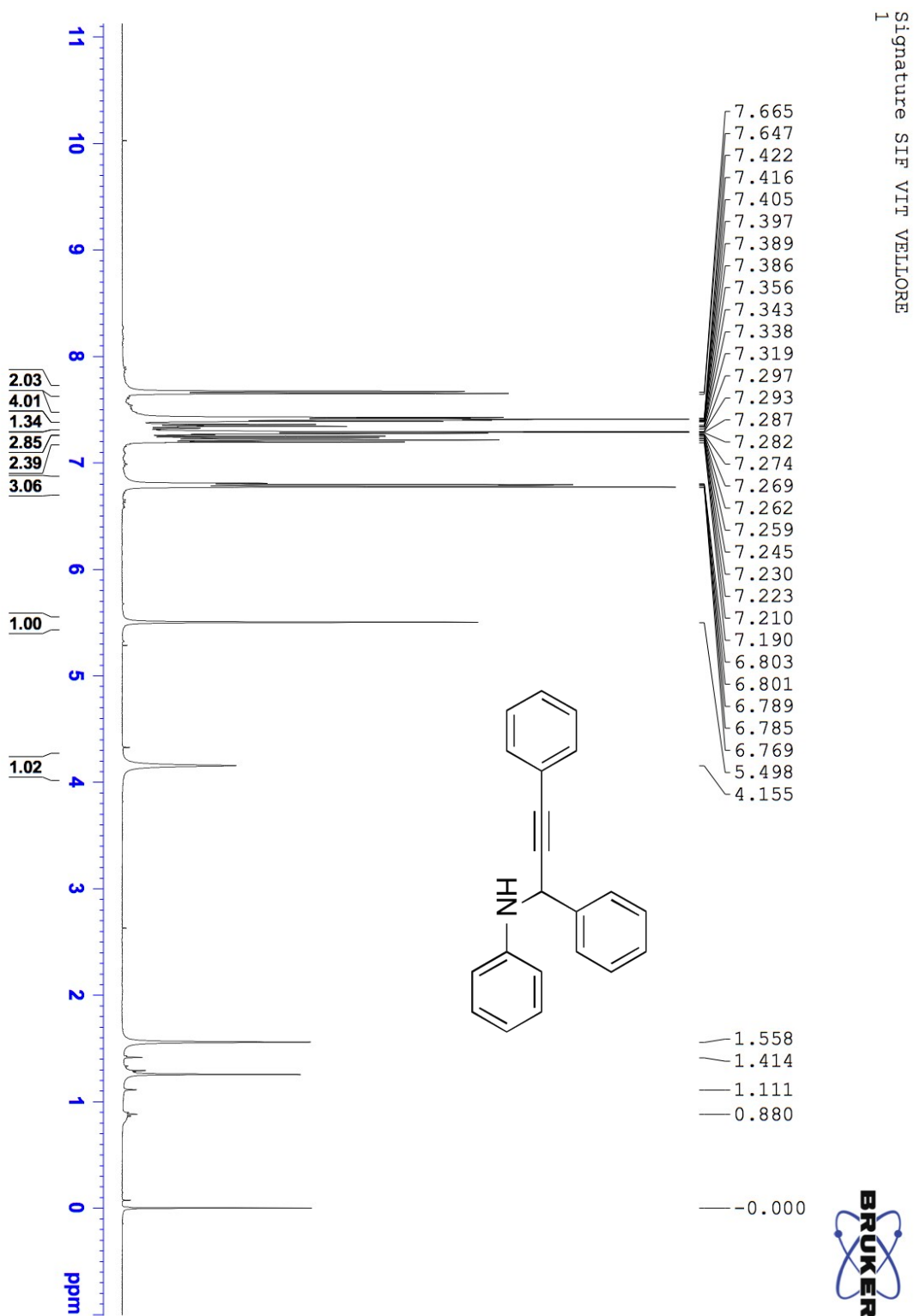
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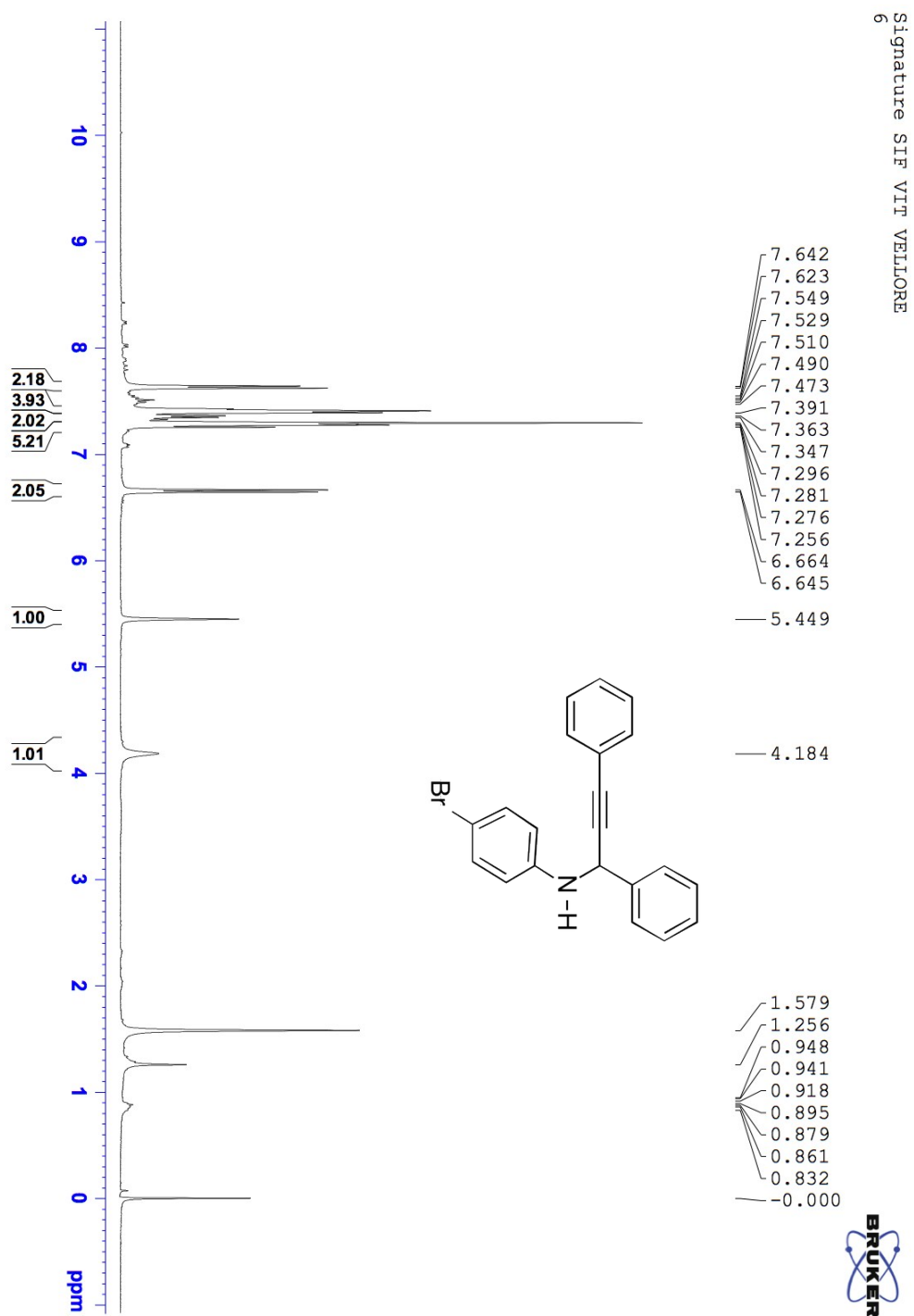
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Entry-4

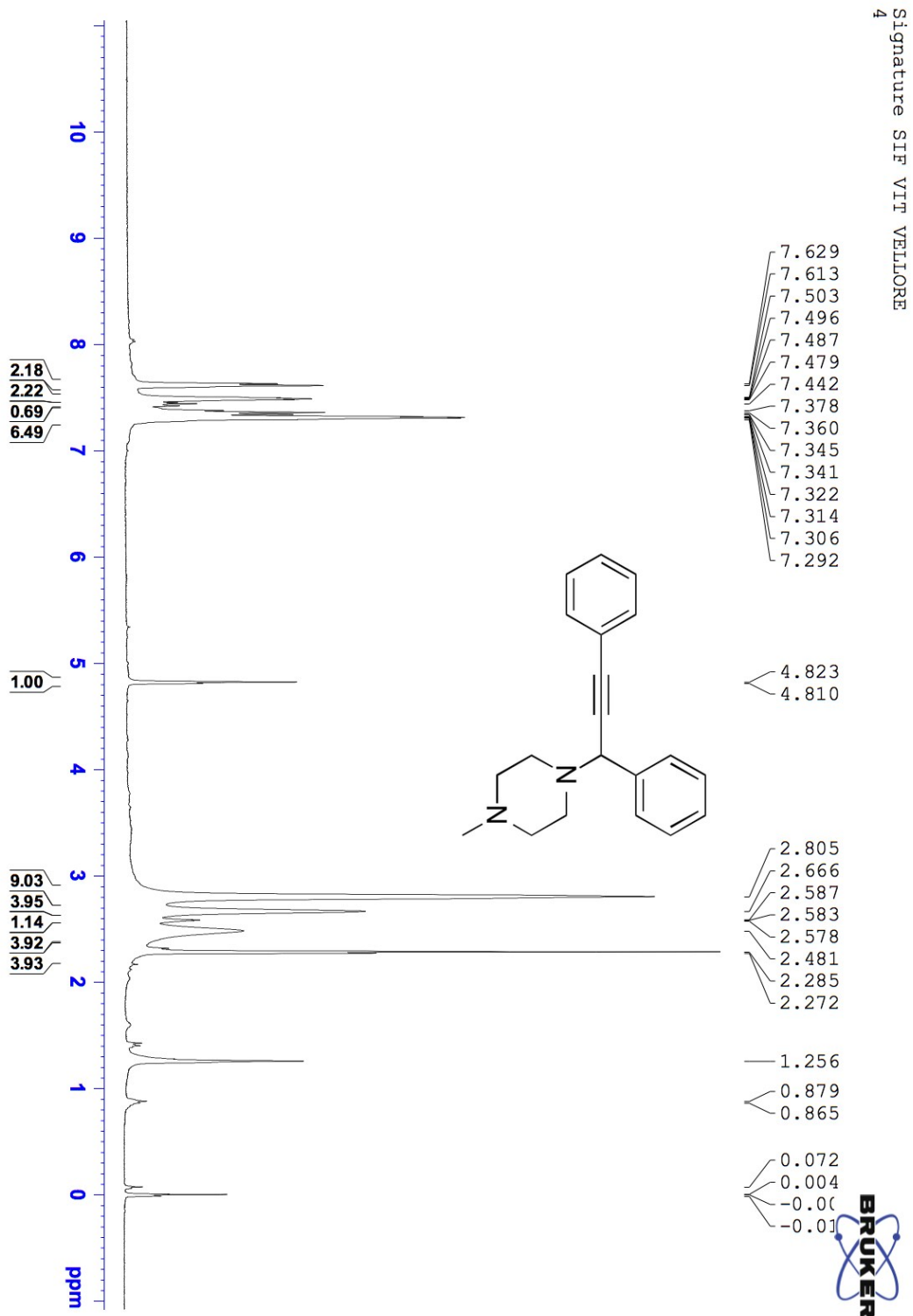


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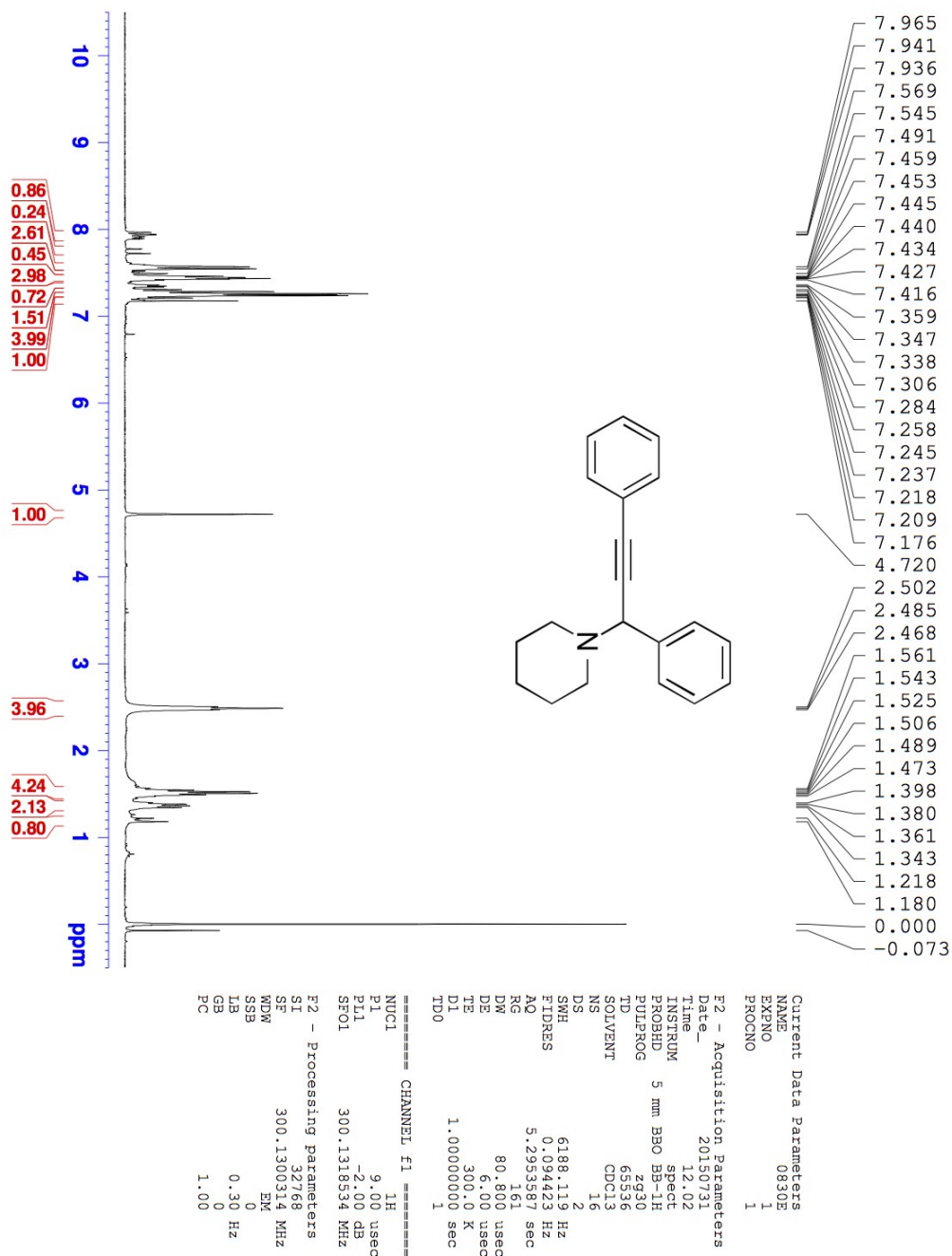


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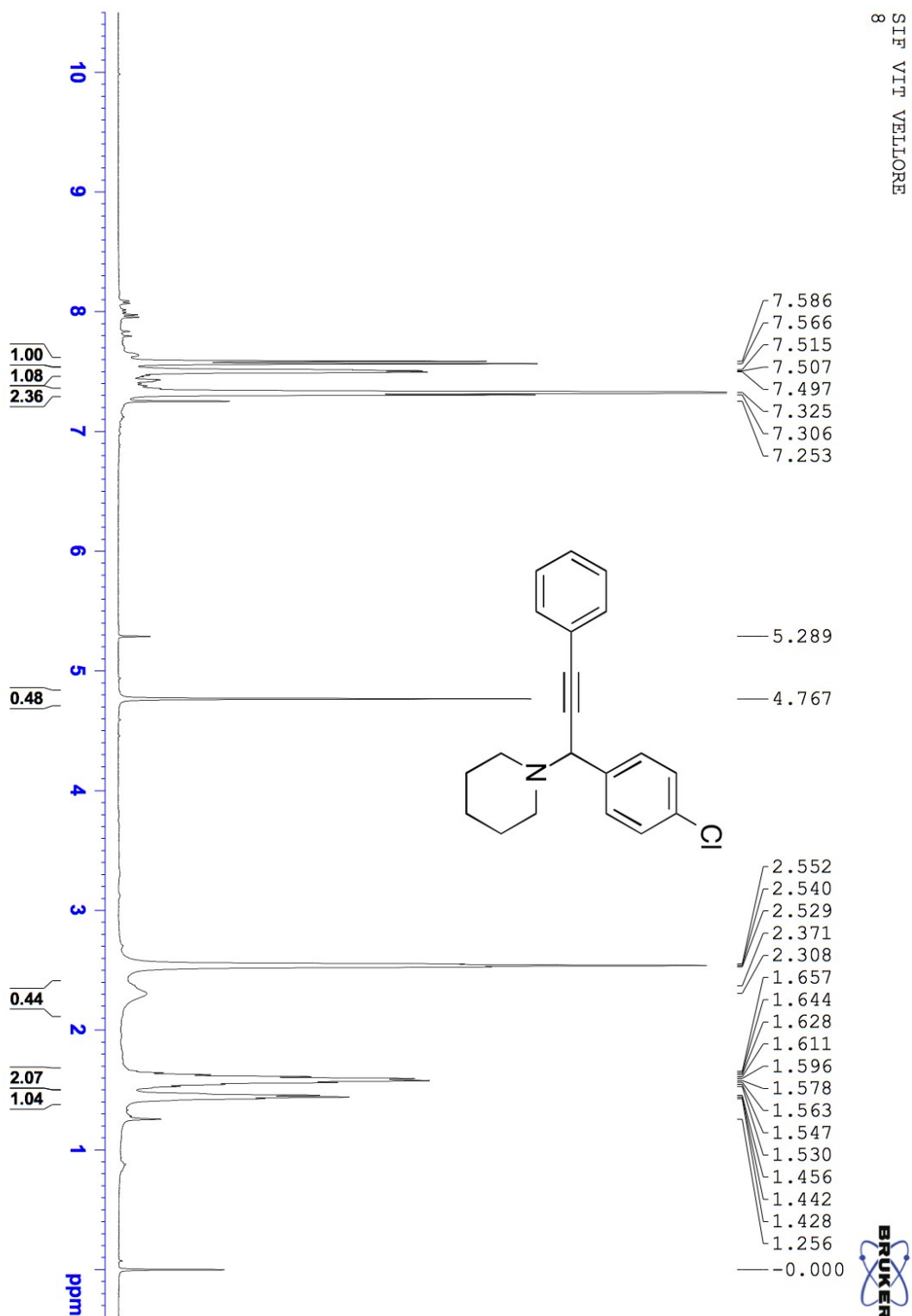




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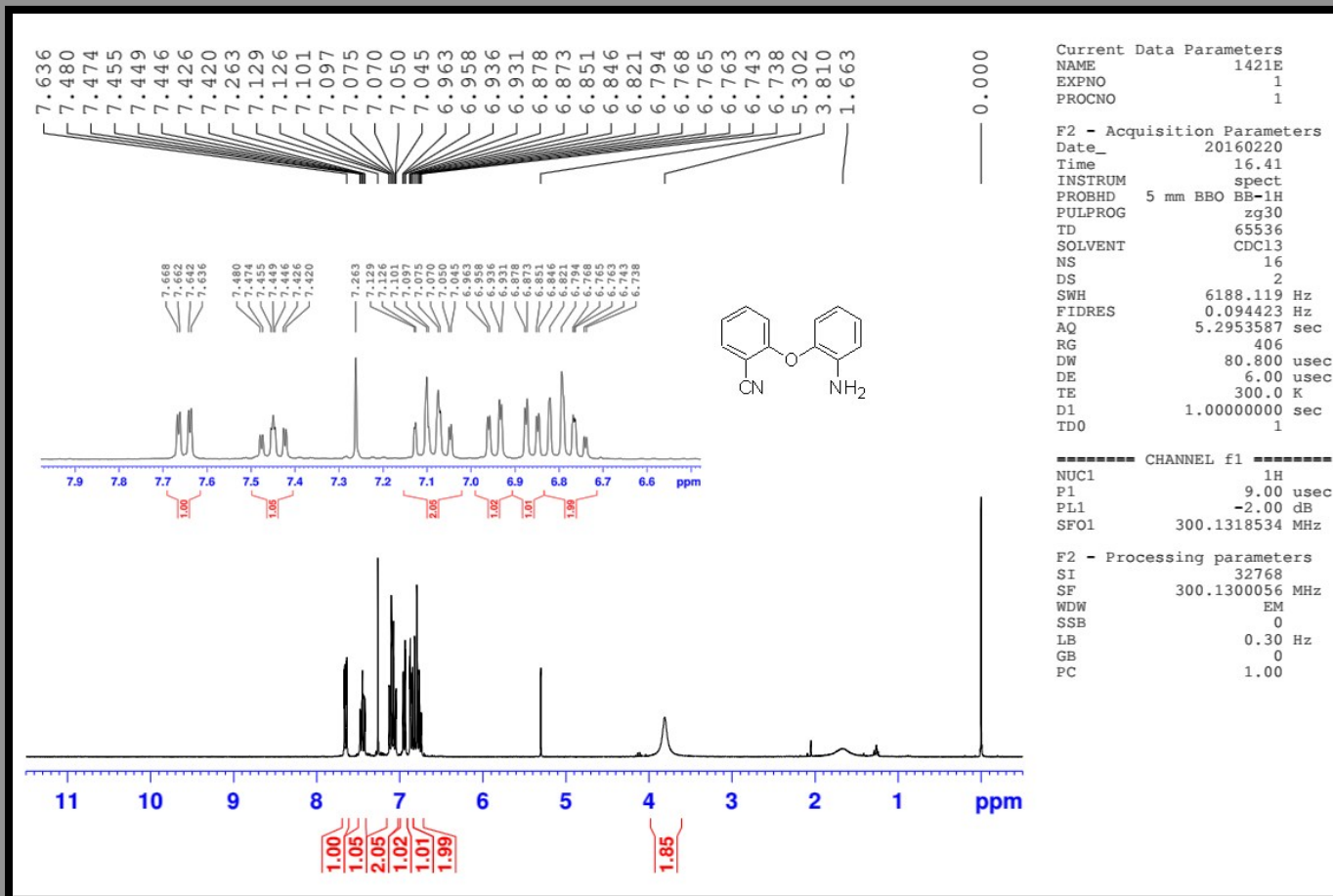


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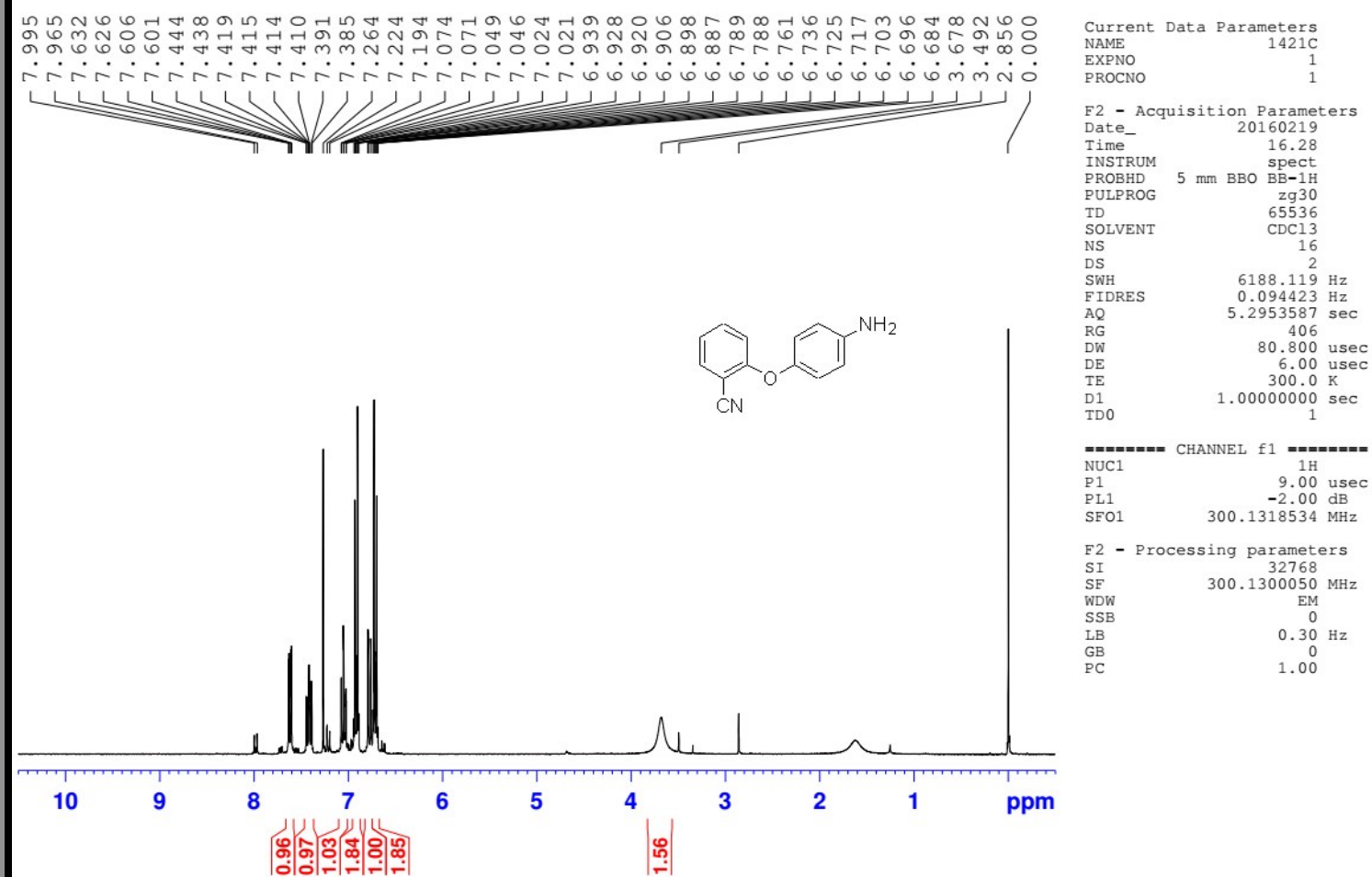


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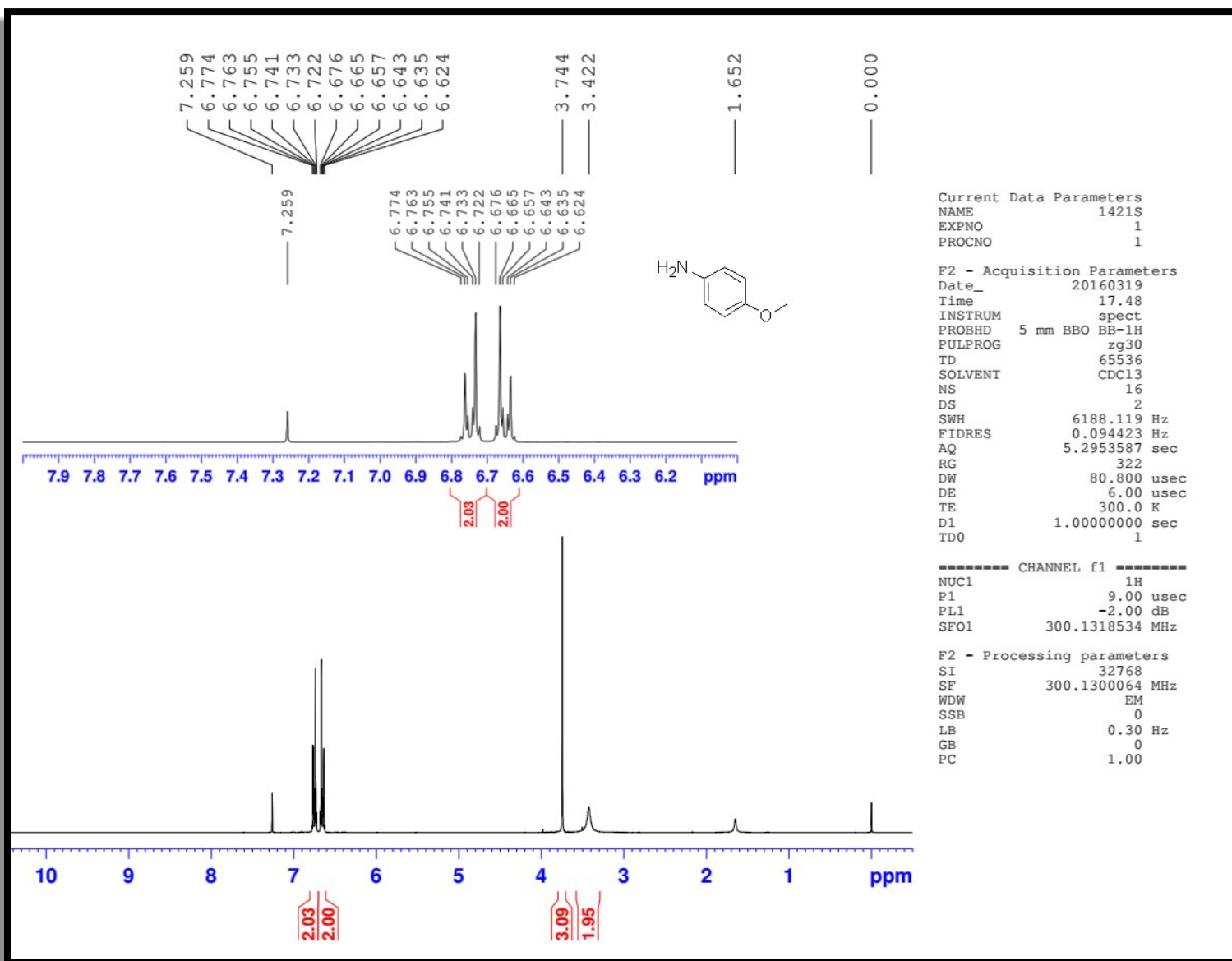
Figure S16.  $^1\text{H}$ -NMR spectra of products obtained using CuONPs-4 catalyst in the nitroarene reduction.



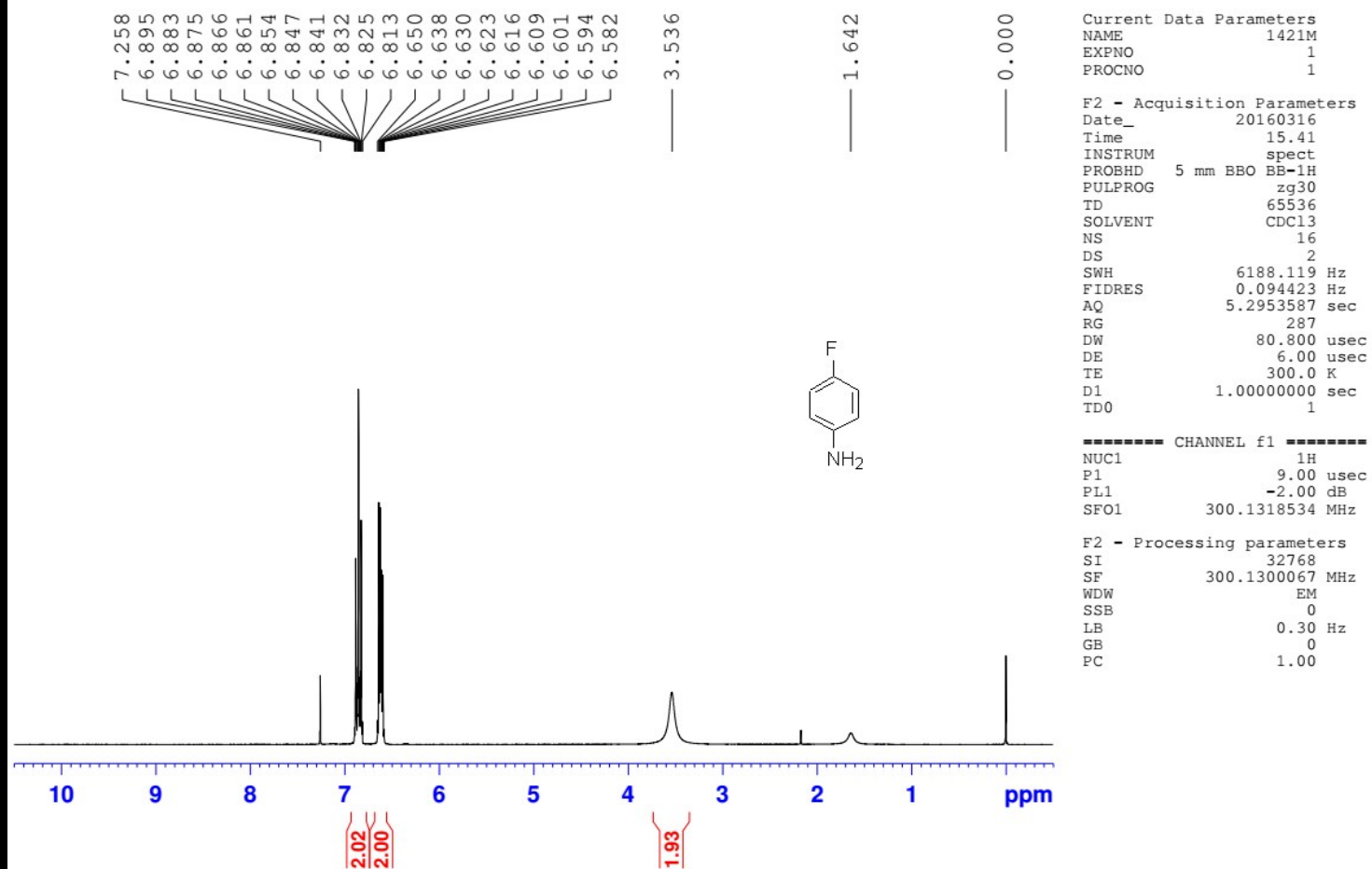
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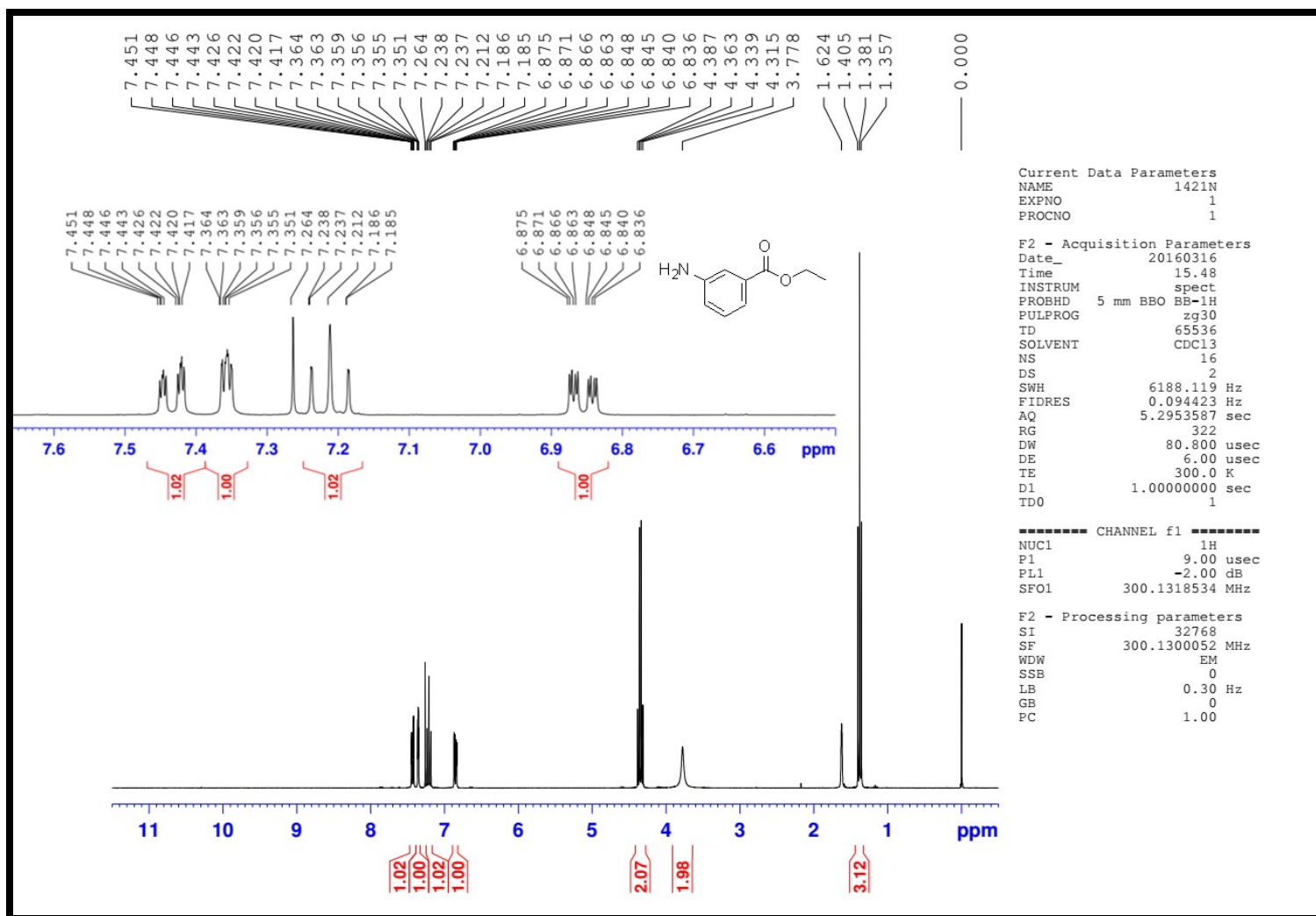
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Entry-3

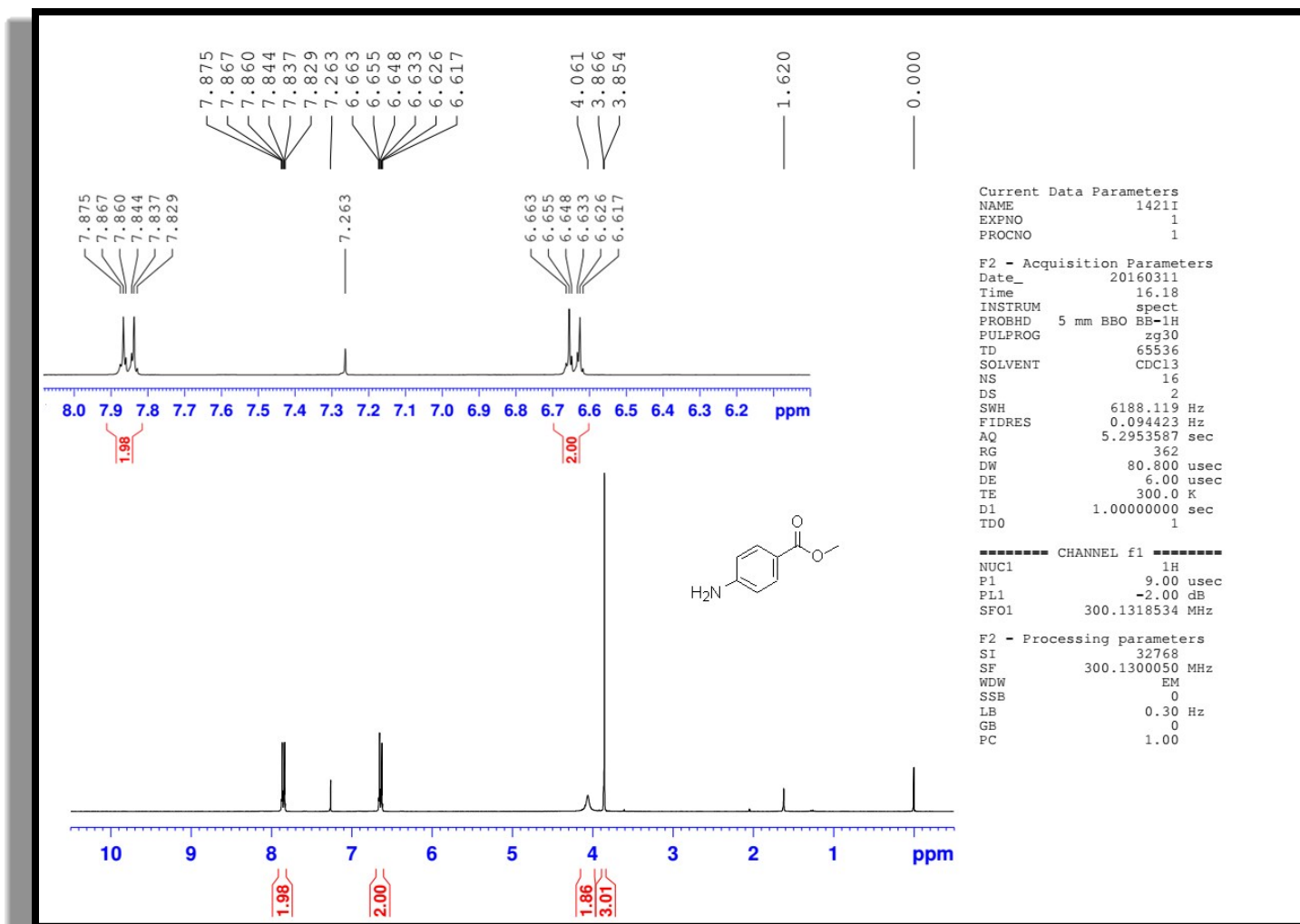


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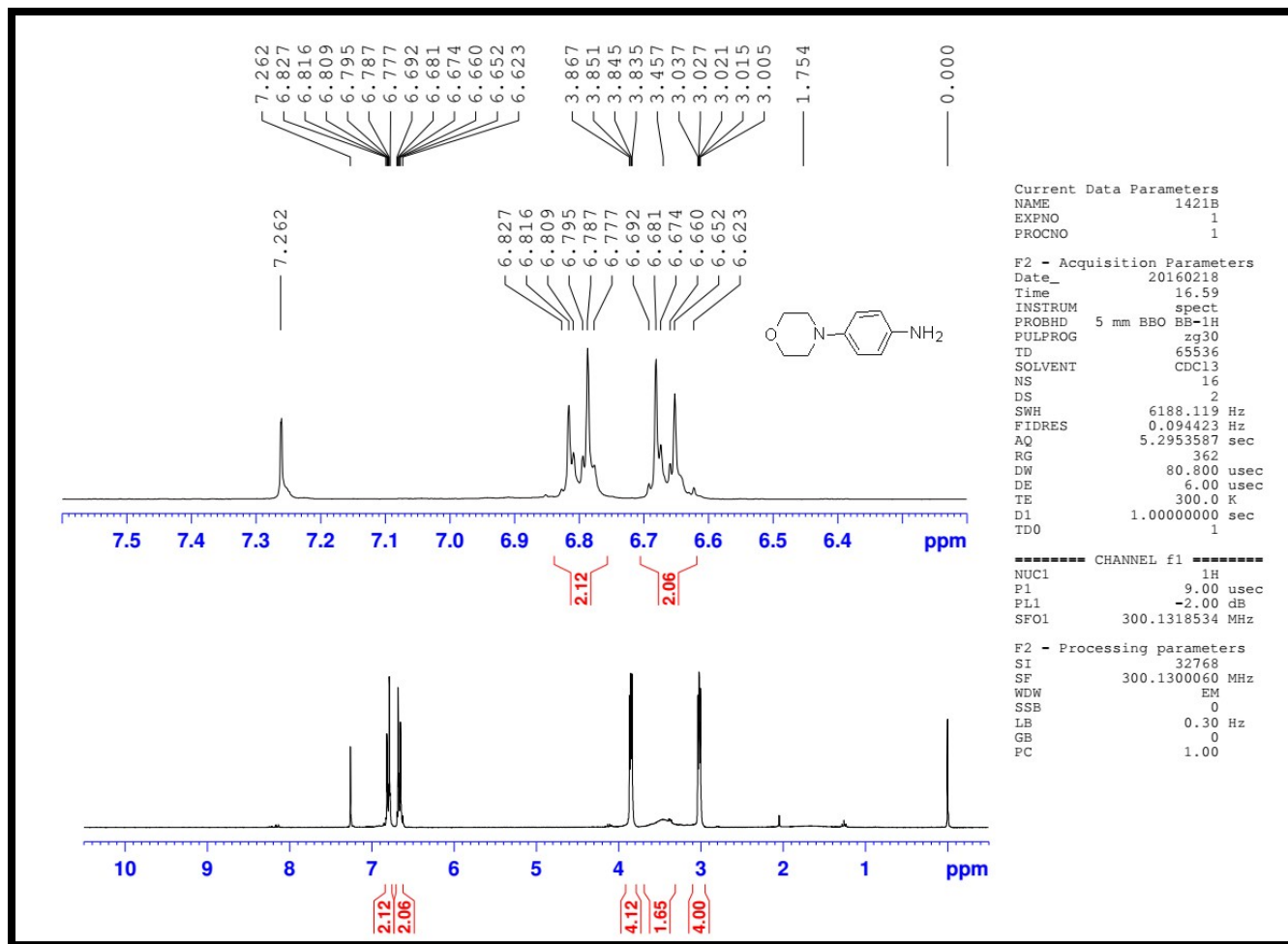


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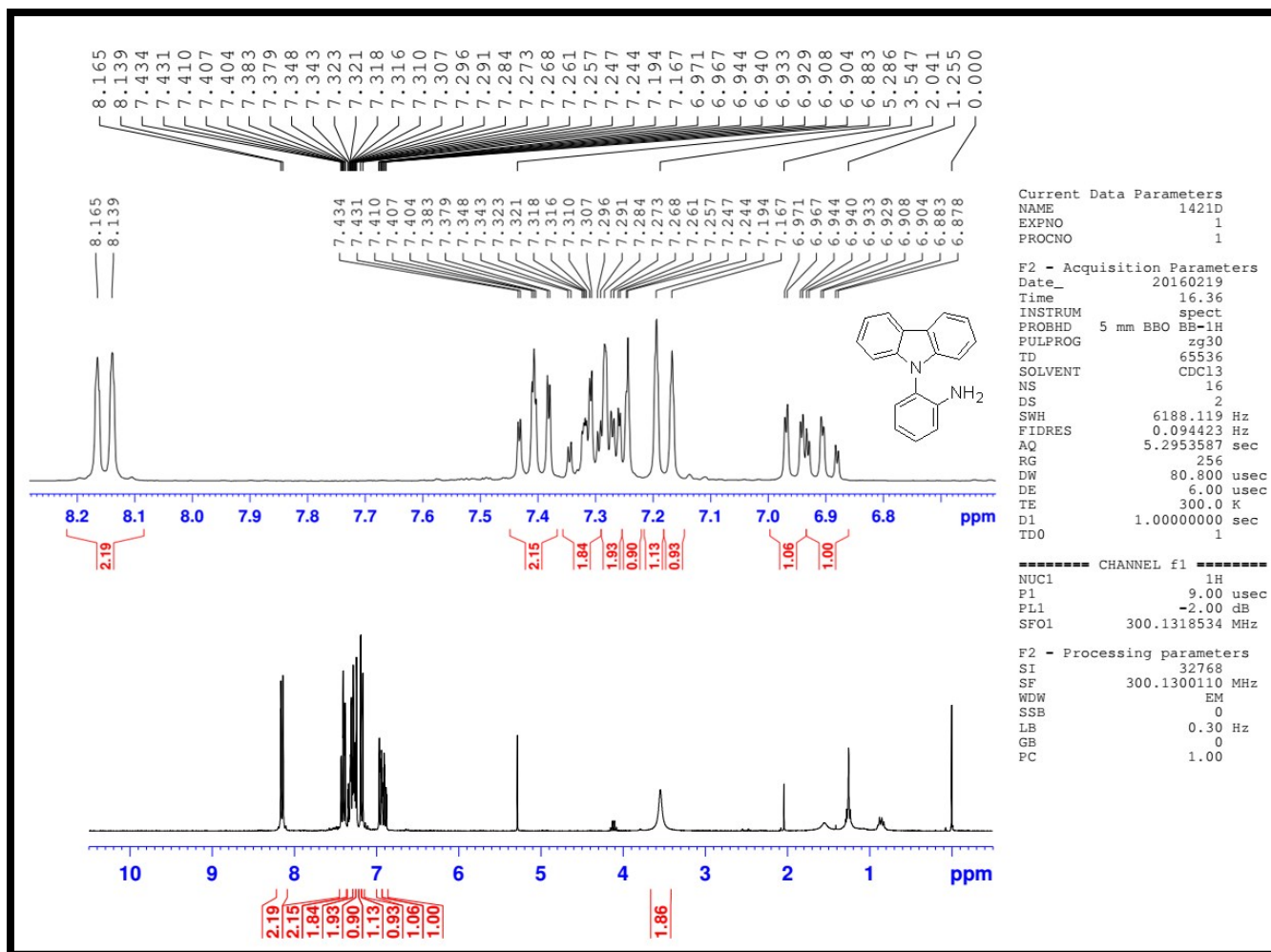




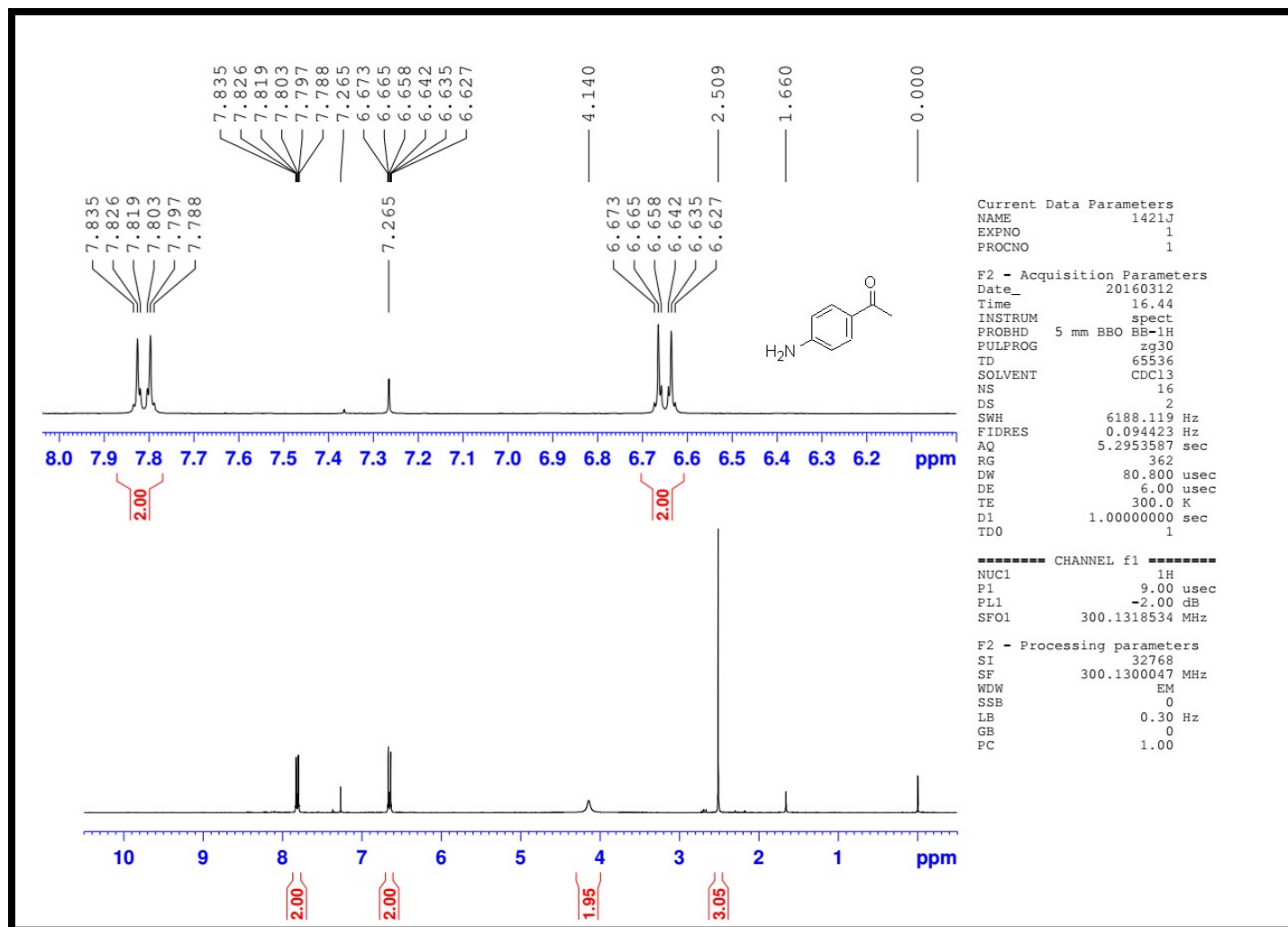
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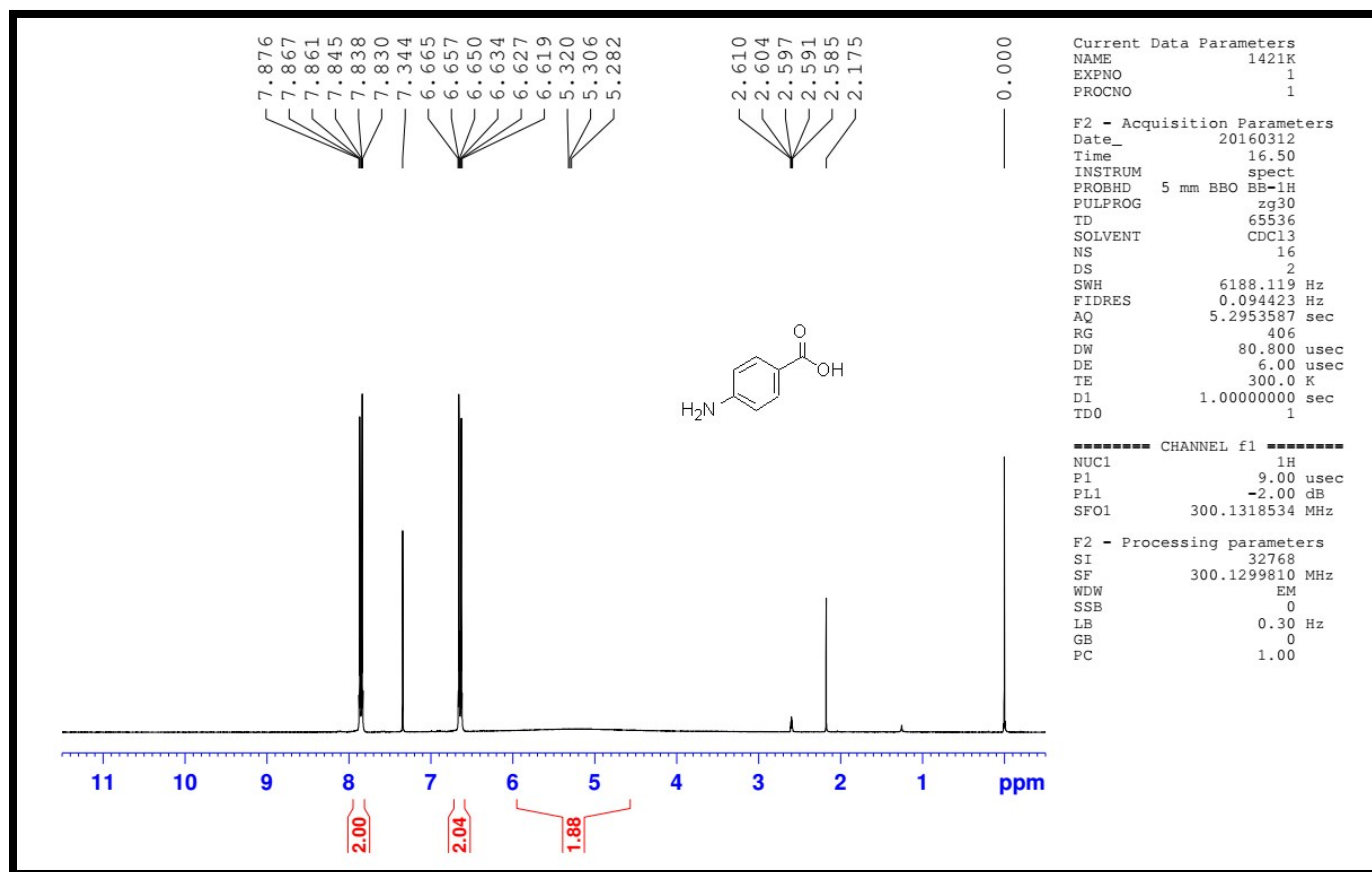
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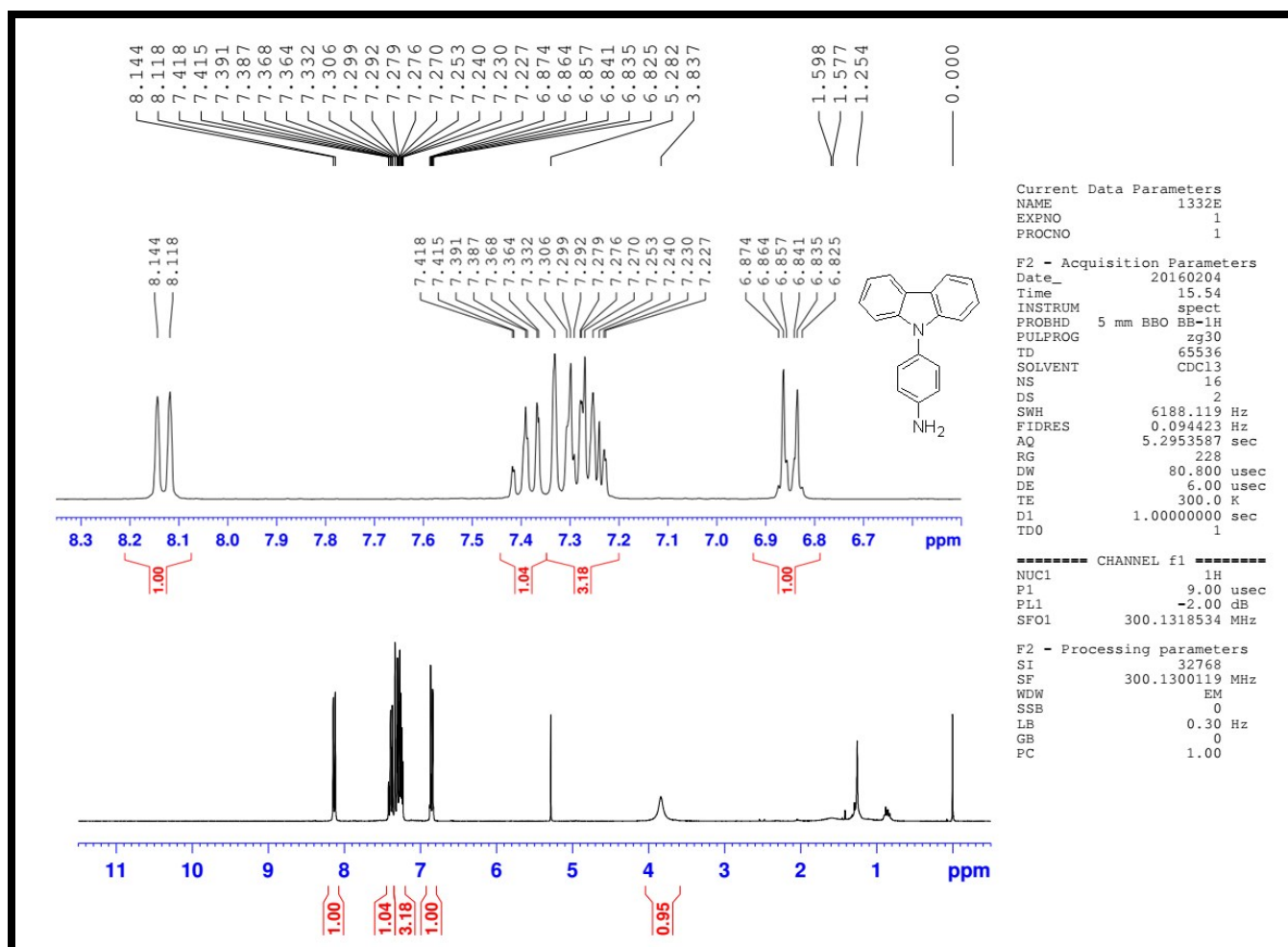
Entry-8



Entry-9



Entry-10



Entry-11