

SUPPORTING INFORMATION FOR

Comparison of the Ca²⁺ complexing properties of isosaccharinate and gluconate – is gluconate a reliable structural and functional model of isosaccharinate?

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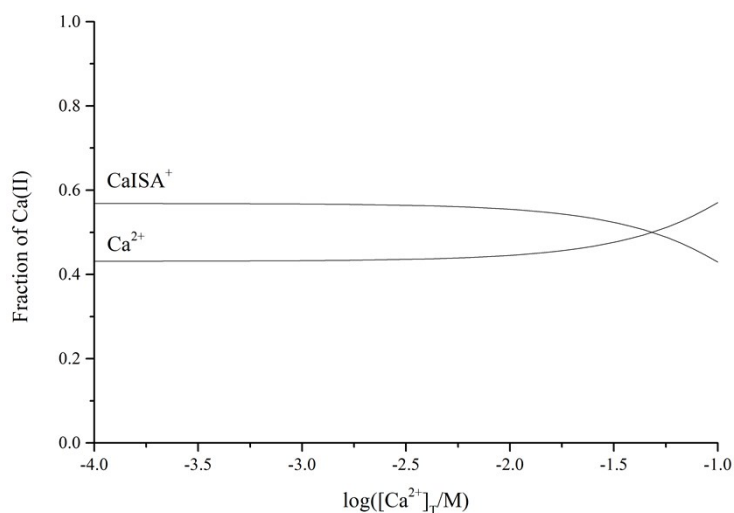


Figure S1. Distribution of Ca(II) among the aqueous species in a solution containing $[Isa^-]_T = 0.10$ M as a function of $\log([Ca^{2+}]_T/M)$ using the formation constant $\log K_{1,1} = 1.12$.

Table S1. The average coupling constants for the Isa^- protons in Hz, determined from the 1H NMR spectrum of aqueous solutions containing $[NaIsa]_T = 0.20$ M

J(Hz)/J(Hz)	H3	H3'	H4	H5	H5'	H6	H6'
H3		14.5	7.8				
H3'	14.5		4.8				
H4	7.8	4.8		4.2	6.9		
H5			4.2		11.5		
H5'			6.9	11.5			
H6							11.4
H6'						11.4	

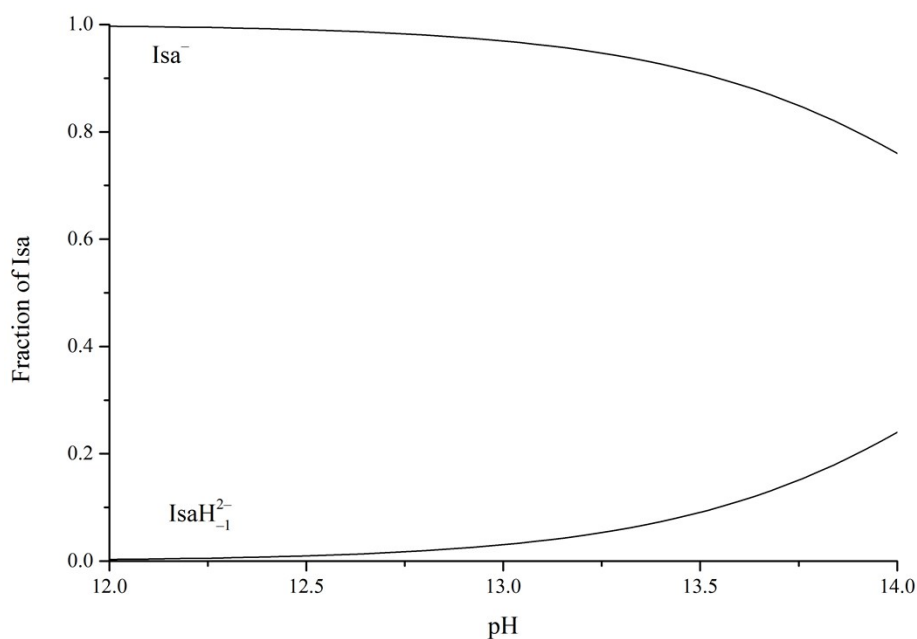


Figure S2. Distribution of Isa among the aqueous species in a solution containing $[\text{Isa}^-]_{\text{T}} = 0.10 \text{ M}$ as a function of pH using the deprotonation constant $\text{p}K_{\text{a}} = 14.5$.

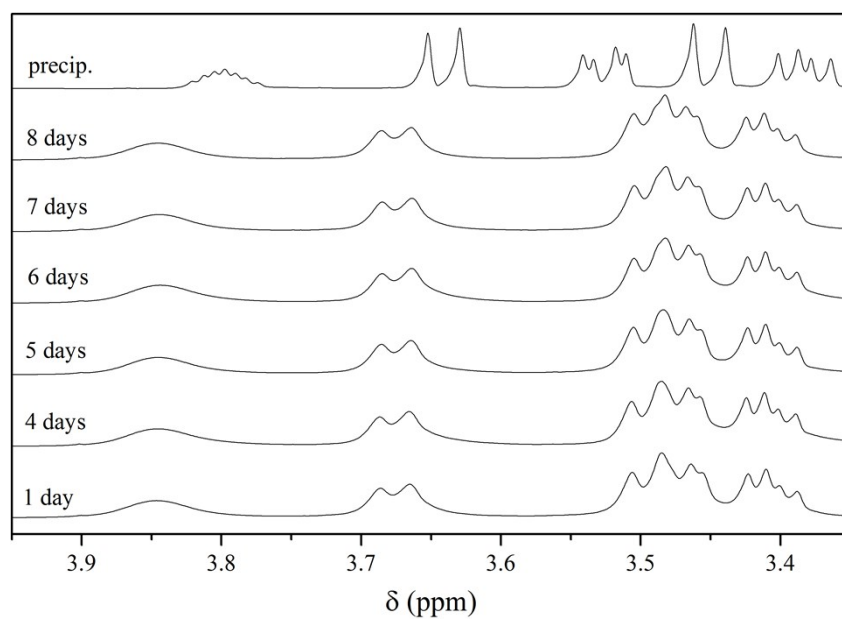


Figure S3. Time dependence of the ^1H NMR spectra of a solution containing $[\text{Isa}^-]_{\text{T}} = 0.20 \text{ M}$, $[\text{OH}^-]_{\text{T}} = 0.50 \text{ M}$ and $[\text{Ca}^{2+}]_{\text{T}} = 0.10 \text{ M}$, $T = (25 \pm 1) \text{ }^\circ\text{C}$. The spectrum on the top corresponds to the filtrate after removing the precipitate formed after 8 days.

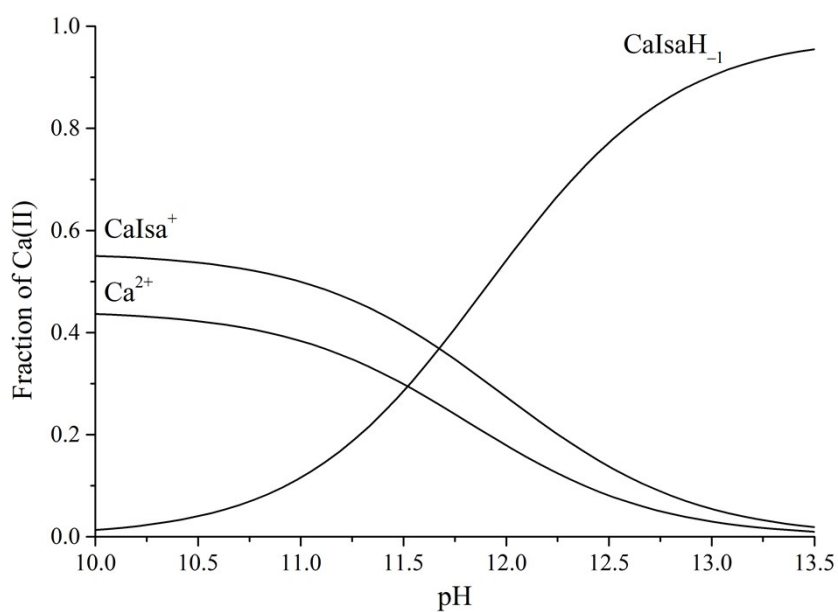


Figure S4. Distribution of Ca(II) among the various aqueous species in a solution containing $[\text{Isa}^-]_{\text{T}} = 0.10 \text{ M}$ and $[\text{Ca}^{2+}]_{\text{T}} = 0.05 \text{ M}$ as a function of pH.

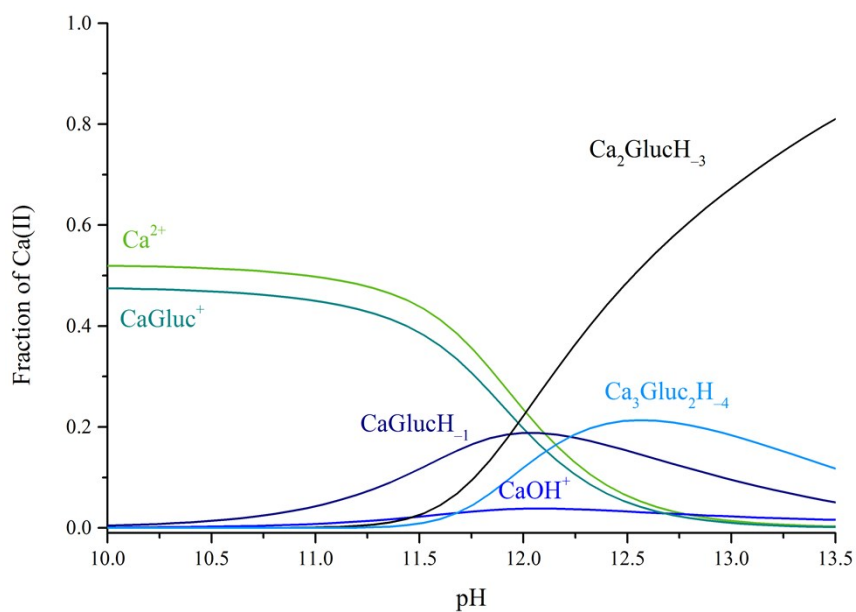


Figure S5. Distribution of Ca(II) among the various aqueous species in a solution containing $[\text{Gluc}^-]_{\text{T}} = 0.10 \text{ M}$ and $[\text{Ca}^{2+}]_{\text{T}} = 0.05 \text{ M}$ as a function of pH.

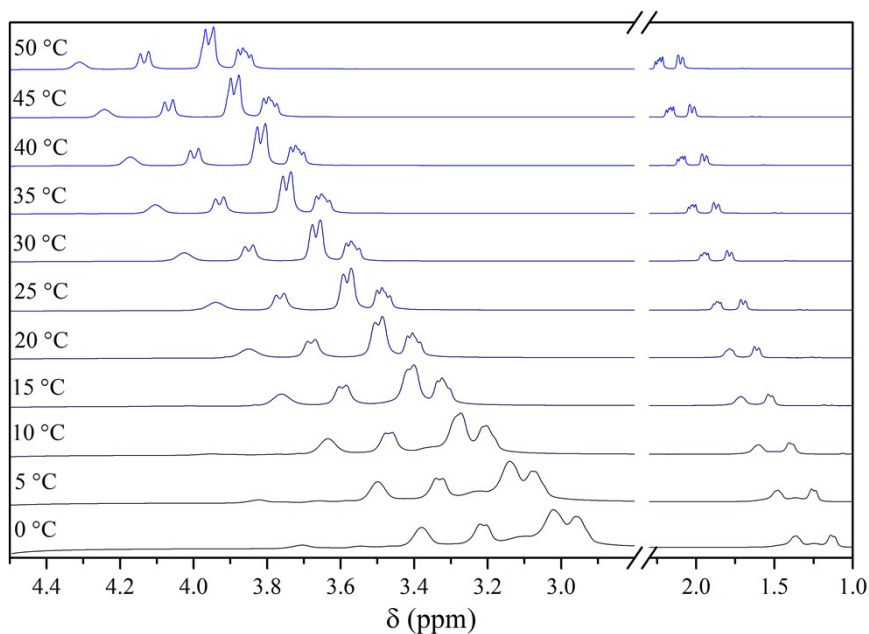


Figure S6. Temperature dependent ^1H NMR spectra in the temperature range of 0-50 $^{\circ}\text{C}$ of a solution containing $[\text{Isa}^-]_{\text{T}} = 0.20 \text{ M}$, $[\text{OH}^-]_{\text{T}} = 0.20 \text{ M}$ and $[\text{Ca}^{2+}]_{\text{T}} = 0.10 \text{ M}$.

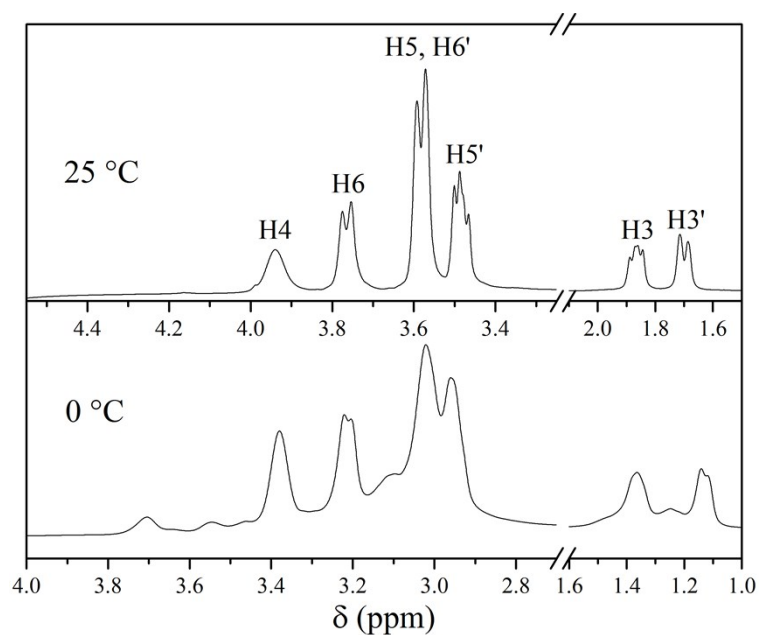


Figure S7. ^1H NMR spectra collected at 0 $^{\circ}\text{C}$ and 25 $^{\circ}\text{C}$ of a solution containing $[\text{Isa}^-]_{\text{T}} = 0.20 \text{ M}$, $[\text{OH}^-]_{\text{T}} = 0.20 \text{ M}$ and $[\text{Ca}^{2+}]_{\text{T}} = 0.10 \text{ M}$, $T = (25 \pm 1) ^{\circ}\text{C}$, $I = 1 \text{ M}$ (NaCl).

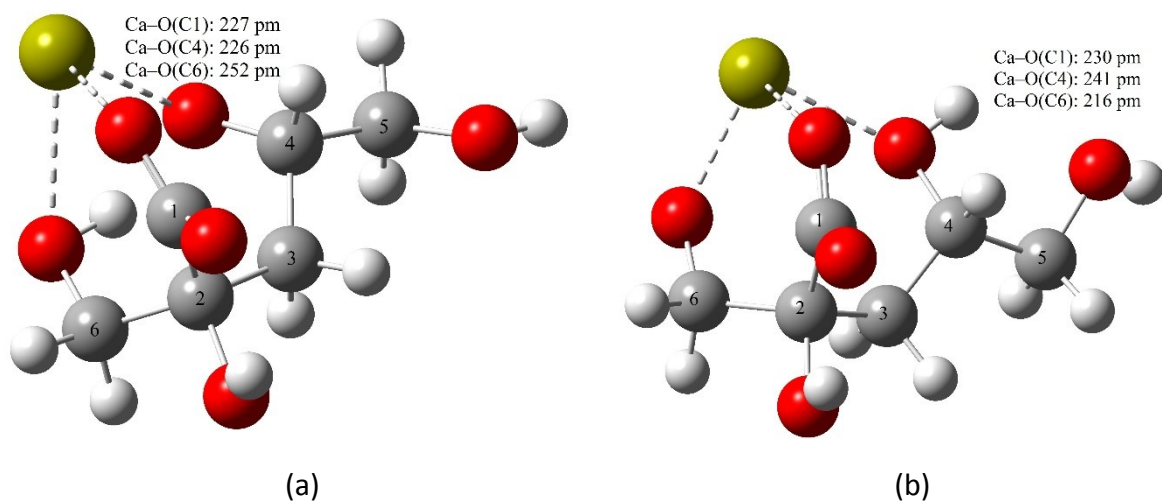


Figure S8. The optimum geometries of the CalsaH₁ complex with O(C6) as a binding site calculated at the B3LYP/6-311++G(d,p) level using the PCM model and explicit water molecules.