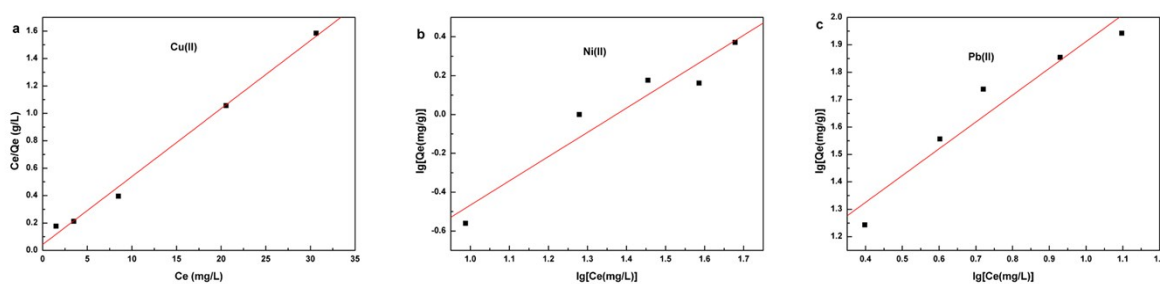


**Fig. S1.** FESEM images of ZnO samples prepared without the addition of ethanolamine: Zn<sup>2+</sup>/NaOH 1:2 (a), Zn<sup>2+</sup>/NaOH 1:8 (b), Zn<sup>2+</sup>/NaOH 1:2 and ethanol as solvent (c).



**Fig. S2.** Adsorption isotherms of commercial ZnO for removal of Cu(II) ions (a), Ni(II) ions (b), Pb(II) ions (c).

**Table S1.** Adsorption capacity of commercial ZnO for the adsorption of divalent heavy metals.

Sorbent	Sorbate		
	Cu <sup>2+</sup>	Ni <sup>2+</sup>	Pb <sup>2+</sup>
	Q <sub>m</sub> (mg/g)	K <sub>F</sub> (mg/g)	K <sub>F</sub> (mg/g)
commercial ZnO	20.0	0.02	8.6