

Size-selective adsorption of anionic dyes induced by layer space on layered double hydroxides hollow microspheres

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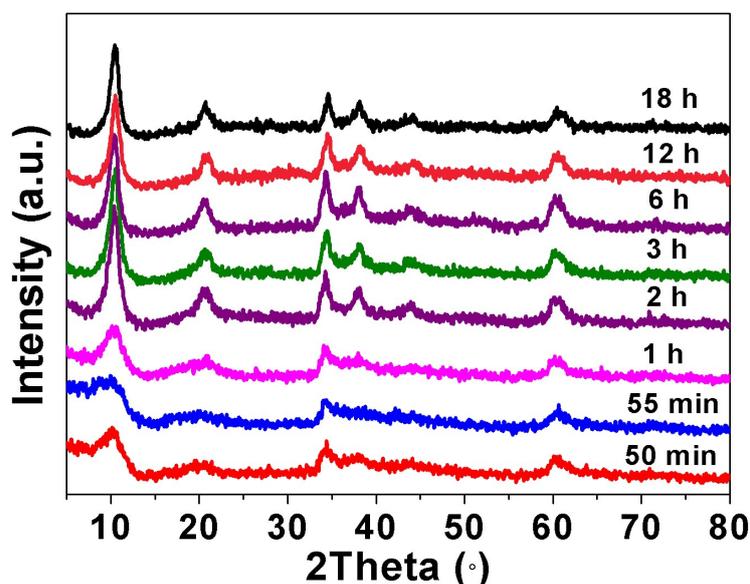


Fig. S1 XRD patterns of samples derived from the time dependent experiments

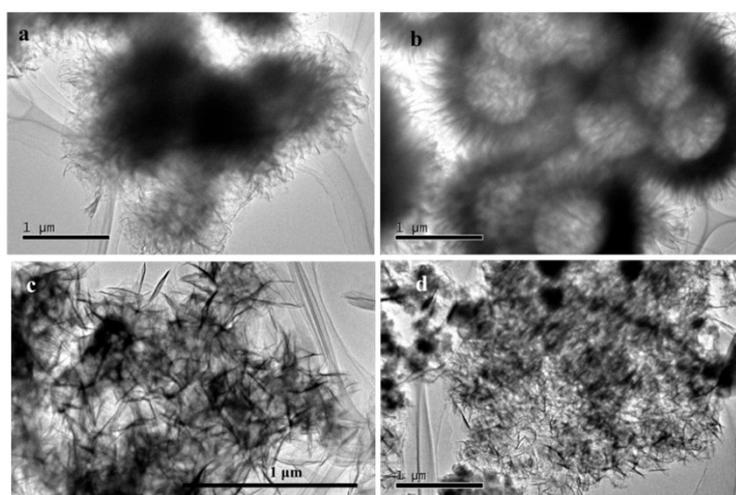


Fig. S2 TEM images of samples derived from different amounts of NH_4F added, (a) 0 g, (b) 0.2 g, (c) 0.4 g and (d) 0.8 g

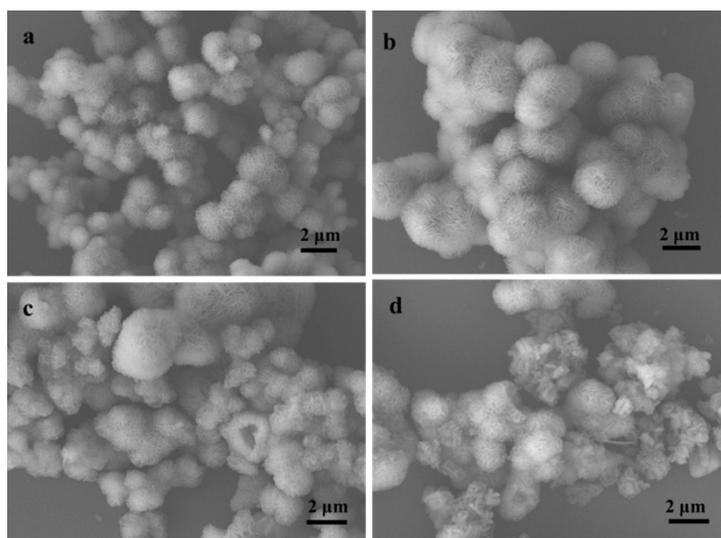


Fig.S3 SEM images of samples derived from different amounts of NH_4F added, (a) 0 g, (b) 0.2 g, (c) 0.4 g and (d) 0.8 g

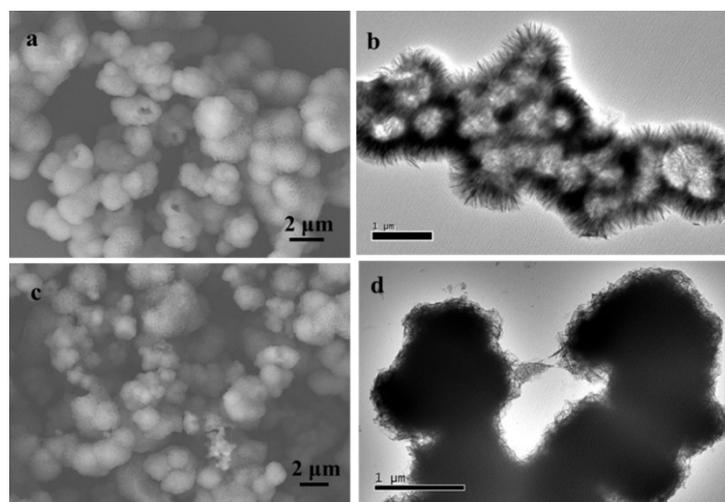


Fig. S4 SEM and TEM images of samples derived from NaF (a) and (b), NH_4Cl (c) and (d), respectively

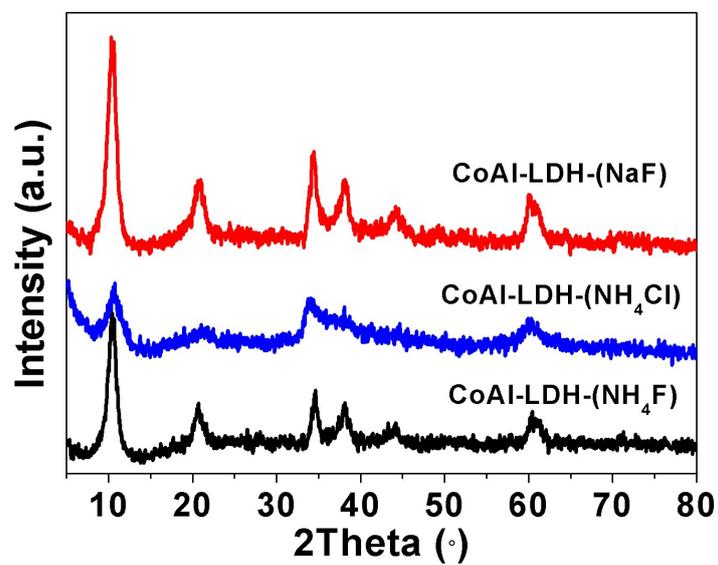


Fig. S5 XRD patterns of the samples prepared in the presence of NaF, NH₄Cl and NH₄F

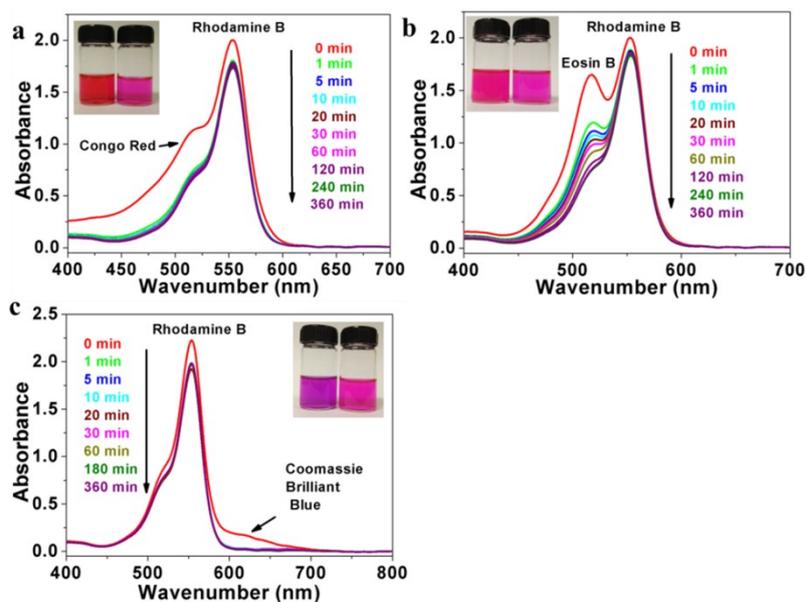


Fig. S6 UV-Vis spectral changes of 100 ml 10 ppm dye mixture in the presence of 20 mg flower-like CoAl-LDHs hollow spheres. (a) Congo Red & Rhodamine B; (b) Eosin B & Rhodamine B; (c) Coomassie Brilliant Blue & Rhodamine B. The inset photographs are before (left) and after adsorption (right).

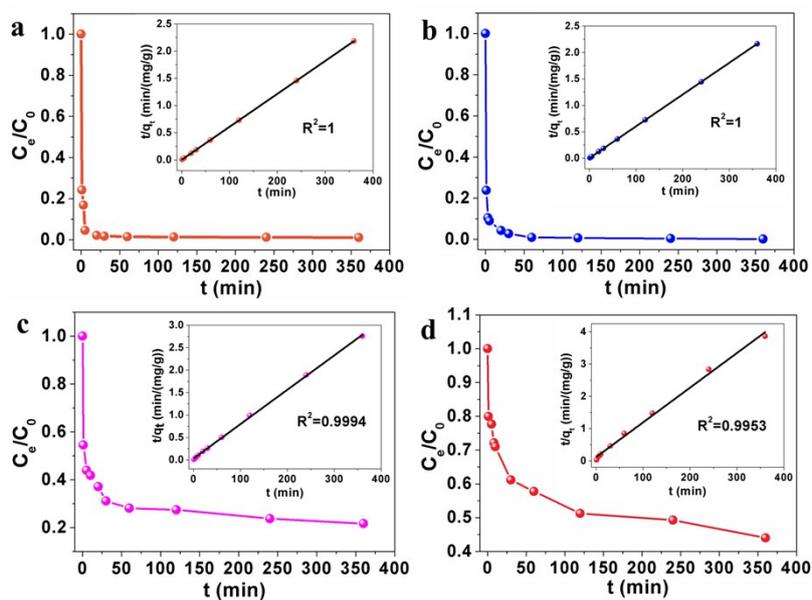


Fig. S7 Adsorption rates of Methyl Orange (a), Coomassie Brilliant Blue (b), Congo Red (c) and Eosin B (d) on the CoAl-LDHs hollow microspheres and all the initial concentration of dye molecules was 100 mg·L⁻¹. The insets show the corresponding pseudo-second-order kinetics model.

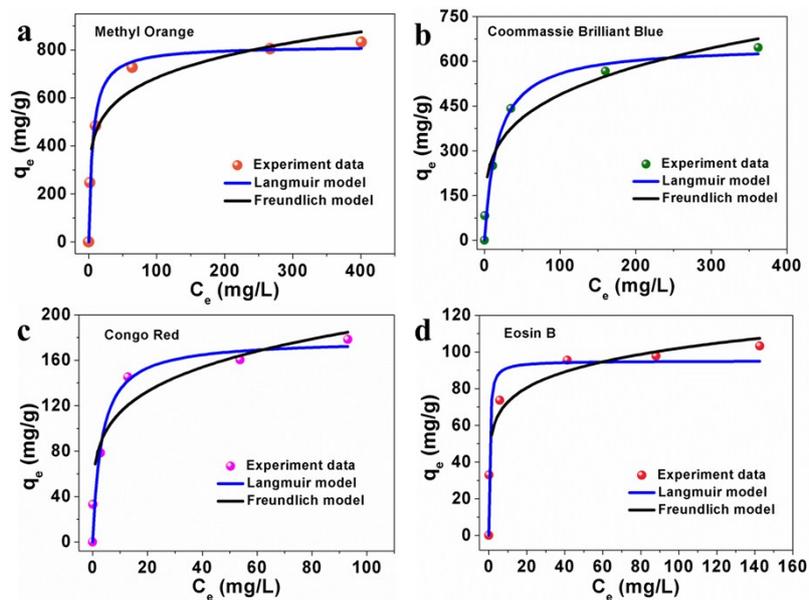


Fig. S8 Adsorption isotherm curves of Methyl Orange (a), Coomassie Brilliant Blue (b), Congo Red (c) and Eosin B (d) on the flower-like CoAl-LDHs hollow microspheres.

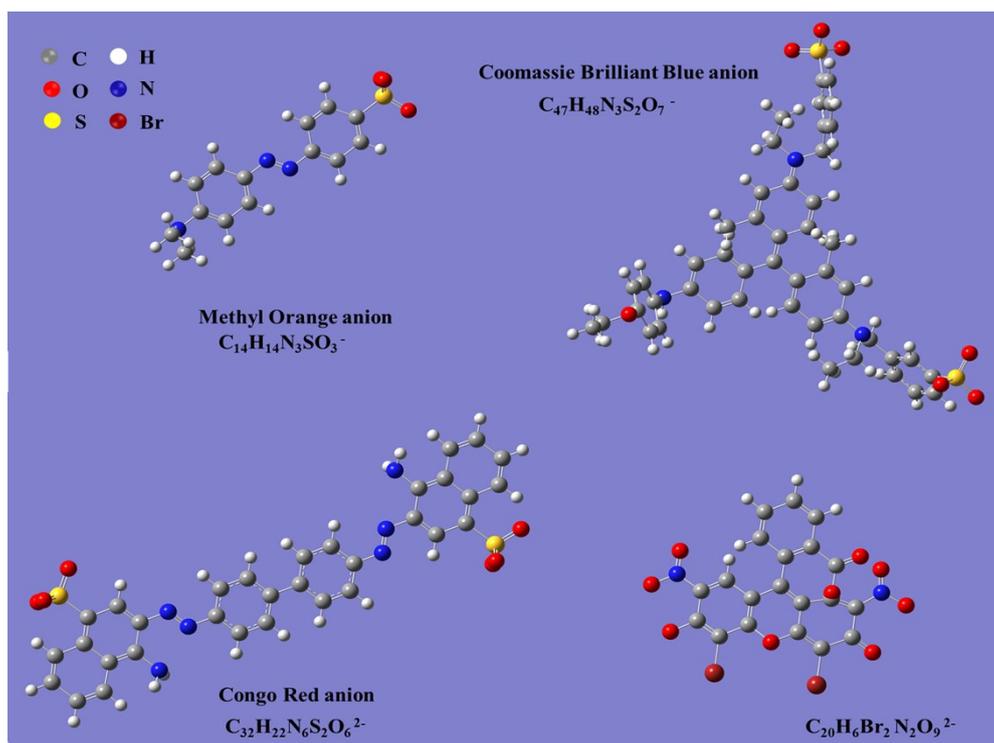


Fig. S9 3D molecular schematic illustrations (Drew by GaussView 5.0) of four anionic dyes

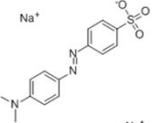
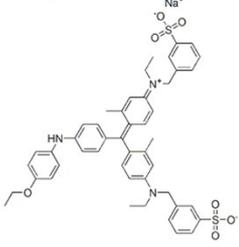
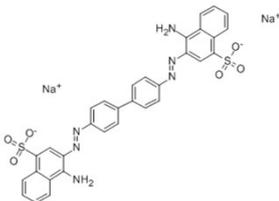
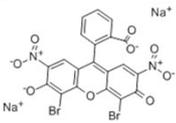
Table S1 Summary of the Langmuir and Freundlich isotherm model parameters for Methyl Orange, Coomassie Brilliant Blue, Congo Red and Eosin B uptake capacities on hierarchical flower-like CoAl-LDHs hollow microspheres

Sample name	Langmuir			Freundlich		
	q_m (mg/g)	b (L mg ⁻¹)	R ²	K_F (L mg ⁻¹)	1/n	R ²
CoAl-LDH-Methyl Orange	816.0	0.1776	0.9891	302.4	0.1772	0.9235
CoAl-LDH-Coomassie Brilliant Blue	654.8	0.05807	0.9849	152.9	0.2523	0.9595
CoAl-LDH-Congo Red	178.3	0.3050	0.9677	69.43	0.2158	0.9423
CoAl-LDH-Eosin B	95.14	2.411	0.9626	51.84	0.1469	0.9539

Table S2 Comparison of adsorption capacities for methyl orange on different adsorbents

Sample name	Surface area/m ² /g	Qm/mg/g	T/K	Reference
Co-Al-SO ₄ -LDH nanoscroll	63.49	1130	303	<i>J. Mater. Chem. A</i> , 2015, 3, 23395–23402
MnO ₂ -G-CNT	173.45	476.19	273	<i>New J. Chem.</i> , 2015, 39, 5484–5492
Mg-Al-LDH-NO ₃		1800.3	298	<i>J. Phys. Chem. C</i> , 2015, 119, 23388–23397
CNT@MCo ₂ O ₄ (M = Ni, Mn, Cu, Zn)	144.0, 103.64, 73.45, 120.58	1188.3, 790.5, 826.9, 935.4		<i>RSC Adv.</i> , 2015, 5, 79765–79773
calcined Zn/Al-CO ₃ -LDH		181.9	298	<i>J. Colloid Interf. Sci.</i> , 2007, 316, 284–291
rGO/Ni/MMO		210.8	298	<i>J. Colloid Interf. Sci.</i> , 2013, 408, 25–32
NH ₂ -MWCNTs	159.69	185.53	273	<i>RSC Adv.</i> , 2014, 4, 55162–55172
Mn ₃ O ₄ /KCC-1		746	273	<i>RSC Adv.</i> , 2015, 5, 106068–106076
H-d-MnO ₂ nanosheets	166.07	357	293	<i>J. Mater. Chem. A</i> , 2015, 3, 5674–5682
H-d-MnO ₂ nanoparticles	91.55	427		
CoAl-LDHs solid microspheres		770	298	this study
CoAl-LDHs hollow microspheres	167	816	298	this study

Table S3 The information of four anionic dyes

Dye name	Chemical formula	Structural formula	Width & Length*	Molecular weight/g/mol
Methyl Orange	$C_{14}H_{14}N_3SO_3Na$		5.5 Å & 15 Å	327.33
Coomassie Brilliant Blue	$C_{47}H_{48}N_3S_2O_7Na$		21.6 Å & 21.8 Å	854.02
Congo Red	$C_{32}H_{22}N_6S_2O_6Na_2$		7.4 Å & 26.7 Å	696.66
Eosin B	$C_{20}H_6Br_2N_2O_9Na_2$		12.4 Å & 10.3 Å	624.06

*: Calculated by GaussView 5.0